

تجمیعات

کیمیاء کوئیز 1

Names and Works of Scientists

(Capter 2 - Chemistry 101):

1- A. Lavoisier: law of conservation of mass.

لافوازييه: قانون حفظ الكتلة.

2- J. Proust: the law of definite proportion

بروست: قانون النسب المحددة او الثابتة.

3- J. Dalton: modren atomic theory & the law of multiple proportions.

دالتون: النظرية الذرية الحديثة وقانون النسب المتعددة.

4- J.J. Thomson: discovered the electron & suggested the "plum-pudding" model of the atom.

طومسون: اكتشف الالكترن واقترح نموذج "قطيرة الفاكهة" لشكل الذرة.

5- Rutherford: gold foil experiment & discovery of nucleus (or the positive protons).

رذرفورد: تجربة رقاقة الذهب، واكتشاف وجود النواة داخل الذرة (وتحديدا البروتونات الموجبة).

6- R. Millikan: oil drop experiment & claculated the charge of electron.

ميليكان: تجربة قطرة الزيت، وحساب شحنة الإلكترون.

7- J. Chadwick: discovered the neutrons (the neutral particles inside the nucleus).

شادويك: اكتشف النيوترونات (الجسيمات المتعادلة بداخل نواة الذرة).

8- N. Bohr: orbitals and quantum energy levels.

بور: مدارات ومستويات الطاقة الكمومية.

9- E. Schrodinger: electron-cloud model of the atom.

شرودينجر: نموذج السحابة الإلكترونية للذرة.

10- D. Mendeleev: periodic law & first periodic table (based on atomic masses of elements).

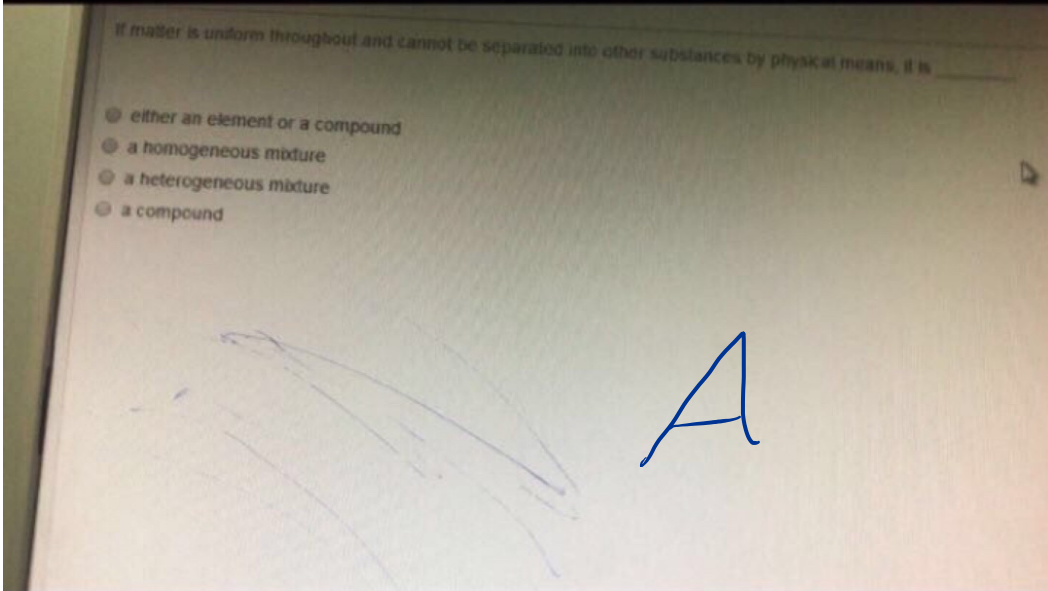
مندليف: القانون الدوري والجدول الدوري الأول للعناصر، المبني علي الكتل الذرية للعناصر.

11- H. Moseley: the modren periodic table of elements (based on atomic numbers of elements).

موزلي: الجدول الدوري الحديث للعناصر، المبني علي الأعداد الذرية للعناصر.

رجوع <

199 من 200



تجميعات 2018 A

اليوم عند ١:١٥ ص



MKCL OES
Online Evaluation System

Question No. 3

Compounds are pure substances that by definition consist of _____

- a single element
- two or more combined elements
- two or more mixed elements
- solids

حفظ و التالي Save & Next

B



A تجميعات 2018

اليوم عند 1:15 ص



Question No. 3

Which of the following statements about energy is FALSE?

- Energy cannot be transformed from one form to another.
- Energy is the capacity to do work.
- Energy can neither be created nor destroyed.
- An object possessing energy can do work on another object.

A

Question No. 24

Which of these elements has the highest electron affinity?

- Rubidium
- Cesium
- Lithium
- Potassium

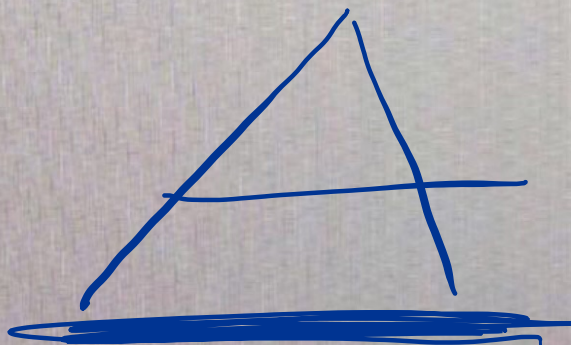
هڪڙو

Question No. 17

One element that has 8 valence electrons is _____

- Argon
- Helium
- Carbon
- Oxygen

~~D~~



Question No. 15

Helium atom has _____ valence electrons.

- 2
- 4
- 8
- 6

A

If matter is uniform throughout and cannot be separated into other substances by physical means, it is _____

- either an element or a compound
- a homogeneous mixture
- a heterogeneous mixture
- a compound

A

Question No. 1

What is the density of a solid sample in g/ml, if its volume was found to be 3.05 ml, and its mass was 7.03 g

- A 0.043 g/ml
- B 2.30 g/ml
- C 21.4 g/ml
- D 0.43 g/ml

* معلو ب الكثافة لزيادة لدرجة
لدرجة 3.05 ml ، كتلتها 7.03

B

Save & Next

Question No. 29

Which of the following is a general trend for the electronegativity of elements in the periodic table?

- increases from left to right, increases from bottom to top
- increases from left to right, decreases from bottom to top
- decreases from left to right, increases from bottom to top
- decreases from left to right, decreases from bottom to top

سکروف

Question No. 21

Which of the following statements is incorrect?

- All atoms of an element have the same atomic number.
- All atoms of an element must have the same mass.
- Atoms of an element may have different numbers of neutrons.
- All atoms of an element have the same number of electrons.

B

Question No. 23

The atomic radius of potassium is smaller than the atomic radius of _____

- cesium
- sodium
- lithium
- hydrogen

سکندر وف
A

Question No. 19

How many electrons are in Br^- ?

- 4
- 34
- 7
- 36

36

D

1	2	3
8	9	1
15	16	2
22	23	3

Question No. 4

Which of the following is an example of a homogeneous mixture?

- sand mixed with oil
- salad dressing
- salt dissolved in water
- gasoline mixed with water



جاني نفس السؤال بس كان غير متجنس و حطية

١:٥٢ م

a

Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 0.023 kg
- 4.03 kg
- 0.015 kg

C

Save & Next حفظ والتالي

Question No. 20

The smallest subatomic particles is _____.

- neutron
- nucleus
- electron
- proton

C

Question No. 16

Which of these elements is chemically similar to Sulfur?

- S
- O
- Cl
- P

B

Number of qu

15 Answered
6 Not Verified

1	2
8	9
15	16
22	

Question No. 17

Semiconductors are located in the periodic table on (or in) the _____.

- left side.
- line dividing metals and nonmetals.
- right side.
- first period.

B

Question No. 21

How many neutrons are there in an atom of lead (Pb) whose mass number is 207?

- 82
- 207
- 125
- 290

125

Question No. 18

An atom of ^{14}C contains _____ protons.

- 6
- 8
- 14
- 12

4

Question No. 22

Which of the following elements has the lowest electronegativity?

- Ba
- Mg
- Be
- Sr

انگیزه

Question No. 1

Determine the mass of an object that has a volume of 88.6 mL and a density of 12.0 g/mL.

- 866 g
- 1100 g
- 568 g
- 298 g

* ضرب الحجم بالكتافة

حاصلها كتلة

—————

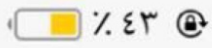
Question No. 12

Which of the following is the highest temperature?

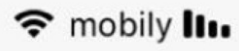
- 40 °C
- 200 °F
- 323 K
- freezing point of water



B

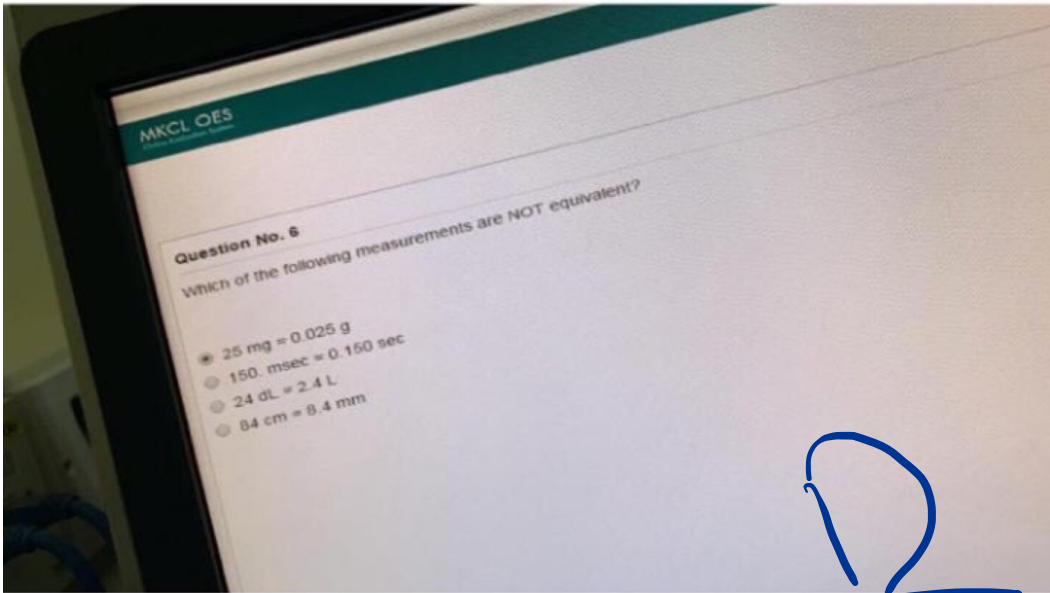


٤:١٠ م



أنت

٢٠١٨/١٠/٩، ١١:٤٧ ص



Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 0.023 kg
- 4.03 kg
- 0.015 kg

C

Save & Next حفظ والتالي

Question No. 25

Atoms X, Y, Z, and R have the following nuclear compositions:



Which two are isotopes?

- X & Y
- X & Z
- Y & R
- Z & R

B

Question No. 24

Express 7.5 nm in micrometers

- 750 μm
- 75.0 μm
- $7.5 \times 10^6 \mu\text{m}$
- $7.5 \times 10^3 \mu\text{m}$

Question No. 21

Which of the following statements is incorrect?

- All atoms of an element have the same atomic number.
- All atoms of an element must have the same mass.
- Atoms of an element may have different numbers of neutrons.
- All atoms of an element have the same number of electrons.

B

Question No. 21

The mass number of an atom can be calculated from the _____.

- number of electrons
- number of protons + neutrons
- number of electrons + protons
- number of protons

Save & Next حفظ و التالي

B

Question No. 14

What element has the electron configuration $1s^2 2s^2 2p^6 3s^2$?

- Be
- Mg
- Ca
- Na

B

Save & Next حفظ والتالي

Question No. 3

Compounds are pure substances that by definition consist of _____.

- a single element
- two or more combined elements
- two or more mixed elements
- solids

Save & Next حفظ و التالي

B

Question No. 20

A neutron has an electrical charge of _____.

- 0
- 1
- 2
- +1

Save & Next حفظ والتالي

A

Question No. 15

A chloride ion (Cl^-) has the same electron configuration as a(n) _____.

- Argon atom
- Neon atom
- Fluorine atom
- Sulfur atom

A

Question No. 7

Which of the following can NOT be considered as matter?

- heat
- sand
- wood
- paper

A

Save & Next حفظ و التالي

HP L1710

Question No. 24

Which of the following correctly matches the ion with the total number of electrons in the ion?

- Br⁻, 34 electrons
- Mg²⁺, 14 electrons
- P³⁻, 15 electrons
- Zn²⁺, 28 electrons

Save & Next حفظ و التالي





Question No. 14

When placing the electrons in orbitals of the same energy, the electrons are first placed singly, the rule is known as:

- Lavoisier's principle
- Mendeleev's rule
- Dalton's atomic theory
- Hund's rule



Save & Next حفظ والتالي

Question No. 6

In which of the following is the metric unit paired with its correct abbreviation?

- microgram / mg
- milliliter / mL
- kilogram / cg
- centimeter / km

Save & Next حفظ والتالي

B

Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 4.03 kg
- 0.015 kg
- 0.023 kg

B

Question No. 10

A physical change _____.

- occurs when glucose is converted into energy inside your cells
- occurs when iron rusts
- occurs when sugar is burnt
- occurs when water is evaporated

Save & Next حذركم

D

Question No. 10

When an egg is fried, what type of process is happening?

- melting
- a physical change
- evaporation
- a chemical change

D

User: AA400232

Number of ques

9 Answered

15 Not Visited

1	2	3
8	9	10
15	16	17
22	23	24

Question No. 3

One of the following is an example of a pure substance

- lemon juice and tea
- sea water
- diamond
- oil and water

11

Question No. 4

One of the following is an example of a mixture

- lemon juice and tea
- sodium
- table salt
- water

A

Question No. 6

How many kg are in 12.5 pounds? (1 kg = 2.21 lb)

- 5.65 kg
- 27.62 kg
- 2.76 kg
- 0.565 kg

A

Question No. 1

Determine the mass of an object that has a volume of 88.6 mL and a density of 12.5 g/mL.

- 866 g
- 1100 g
- 568 g
- 298 g

Question No. 4

One of the following is an example of a mixture.

- lemon juice and tea
- sodium
- table salt
- water

A

Question No. 18

How many electrons are in Rb^+ ?

- 4
- 34
- 7
- 35

34



Question No. 22

Which of the following elements has the lowest electronegativity?

- Ba
- Mg
- Be
- Sr

ہگڈون

Question No. 23

The atomic radius of potassium is smaller than the atomic radius of _____

- cesium
- sodium
- lithium
- hydrogen

سکون

Question No. 10

Which of the following is a physical change?

- rusting of iron
- Sublimation of dry ice
- flammability of propane gas
- baking the flour to make a cake

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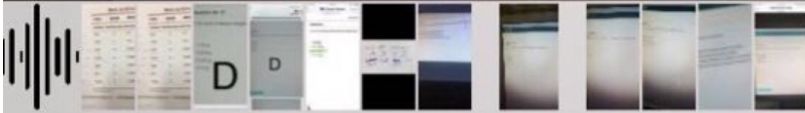
B

Question No. 23

The atomic radius of potassium is smaller than the atomic radius of _____

- cesium
- sodium
- lithium
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سکندرون



Question No. 26

Which of the following statements is incorrect?

- The condensation of steam on a mirror is an example of a physical change.
- Evaporation of water from a fish tank is evidence of a physical change.
- The burning of a piece of charcoal to a white powder is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.

C

Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 4.03 kg
- 0.015 kg
- 0.023 kg

B

Question No. 6

The correct prefix for the multiplier 0.000001 is _____

- micro
- mega
- mill
- nano

A

Save & Next حفظ و التالي

HP LI710

Question No. 12

Which of the following is the highest temperature?

- 40 °C
- 200 °F
- 323 K
- freezing point of water

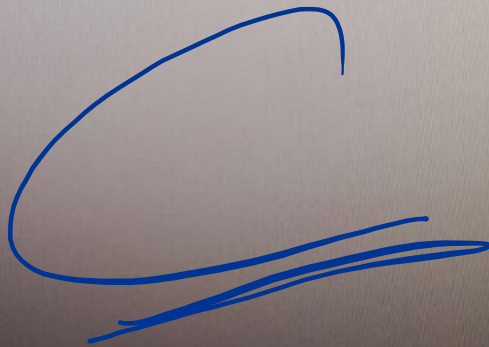
B

Save & Next حفظ والتالي

Question No. 20

The smallest subatomic particles is _____.

- neutron
- nucleus
- electron
- proton



Question No. 5

_____ is the metric unit of the temperature of an object.

- Fahrenheit
- Celsius
- Quart
- Kelvin

Save & Next حفظ واقتلي

B

Question No. 1

What is the density of a solid sample in g/mL if its volume was found to be 3.05 mL and its mass was 7.03 g

- 2.30 g/mL
- 21.4 g/mL
- 0.043 g/mL
- 0.43 g/mL

Save & Next حفظ و التالي

A

Question No. 8

Gases and liquids share the property of _____ .

- compressibility
- fixed volume
- rigid shape
- unfixed shape

Save & Next حفظ والتالي

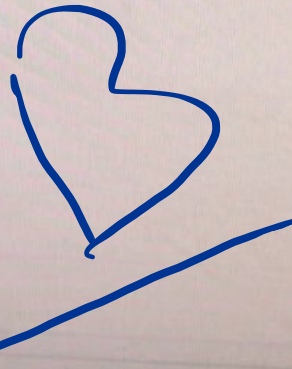
D

Question No. 10

Which of the following is a physical change?

- rusting of iron
- Sublimation of dry ice
- flammability of propane gas
- baking the flour to make a cake

Save & Next حفظ والتالي



Question No. 16

Which one of the following elements is a poor conductor of heat and electricity?

- copper
- fluorine
- lead
- iron

Save & Next حفظ و التالي

B

Question No. 7

The state of matter that has fixed volume and unfixed shape is called?

- liquid
- solid
- plasma
- gas

A

Save & Next حفظ و التالي

Question No. 6

In which of the following is the metric unit paired with its correct abbreviation?

- microgram / mg
- milliliter / mL
- kilogram / cg
- centimeter / km

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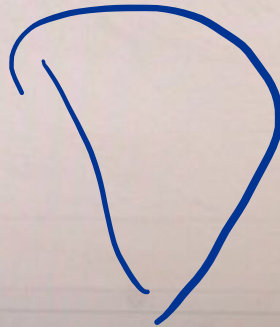
B

Question No. 8

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Save & Next حفظ والتالي



Question No. 22

Which of the following elements has the lowest electronegativity?

- Ba
- Mg
- Be
- Sr

A

کونو

Question No. 23

The atomic radius of potassium is smaller than the atomic radius of _____

- cesium
- sodium
- lithium
- hydrogen

A

سؤال 23

Question No. 9

Which of the following is a chemical property?

- Density
- Viscosity
- Melting point
- Acidity

Save & Next حفظ والتالي

2

Question No. 21

How many neutrons are there in an atom of lead (Pb) whose mass number is 207?

- 82
- 207
- 125
- 290

C

Question No. 18

An atom of ^{14}C contains _____ protons.

- 6
- 8
- 14
- 12

A

Question No. 6

9.31 g is the same mass as _____

- 93.1 cg
- 931 kg
- 9310 mg
- 931 μ g

931

Question No. 26

Which of the following statements is incorrect?

- The condensation of steam on a mirror is an example of a physical change.
- Evaporation of water from a fish tank is evidence of a physical change.
- The burning of a piece of charcoal to a white powder is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.

C

Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?


- 1.15 kg
- 0.023 kg
- 4.03 kg
- 0.015 kg

Save & Next حفظ والتالي

Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 0.023 kg
- 4.03 kg
- 0.015 kg



Save & Next حفظ والتالي

Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 4.03 kg
- 0.015 kg
- 0.023 kg

B

Question No. 4

Which of the following is an example of a homogeneous mixture?

- sand mixed with oil
- salad dressing
- salt dissolved in water
- gasoline mixed with water



Question No. 24

Which of the following correctly matches the ion with the total number of electrons in the ion?

- Br⁻, 34 electrons
- Mg²⁺, 14 electrons
- P³⁻, 15 electrons
- Zn²⁺, 28 electrons

Save & Next حفظ و التالي

D

Question No. 3

One of the following is an example of a pure substance

- lemon juice and tea
- sea water
- diamond
- oil and water

C

Question No. 4

Which of the following is an example of a homogeneous mixture?

- sand mixed with oil
- salad dressing
- salt dissolved in water
- gasoline mixed with water



Question No. 8

Melting is a process in which _____ changes into _____, and it requires _____ temperature.

- a solid, a liquid, increasing
- a liquid, a gas, decreasing
- a liquid, a solid, increasing
- a solid, a gas, increasing



User :AA40023

Number of ques

7 Answered

17 Not Visited

1	2	3
8	9	10
15	16	17
22	23	24

Question No. 4

One of the following is an example of a mixture:

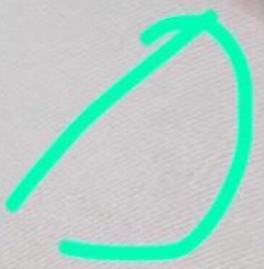
- lemon juice and tea
- sodium
- table salt
- water

X

Question No. 7

Which of the following items is an example of matter?

- light
- heat
- time
- air



Question No. 18

When an atom gains an electron, the

- a cation.
- an anion.
- an isotope.
- a proton.

5

Question No. 1

Determine the mass of an object that has a volume of 88.6 mL and a density of 12.5 g/mL.

- 866 g
- 1100 g
- 568 g
- 298 g

Question No. 5

For which of the following can the composition vary?

- both homogeneous and heterogeneous mixtures
- heterogeneous mixture
- homogeneous mixture
- pure substance

B

Question No. 4

Which of the following is an example of a homogeneous mixture?

- sand mixed with oil
- salad dressing
- salt dissolved in water
- gasoline mixed with water

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Question No. 9

Which of the following is a chemical property?

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Save & Next حفظ والتالي

D

Question No. 16

Which one of the following elements is a poor conductor of heat and electricity?

- copper
- fluorine
- lead
- iron

Save & Next حفظ و التالي

B

Question No. 8

Melting is a process in which _____ changes into _____, and it requires _____ temperature.

- a solid, a liquid, increasing
- a liquid, a gas, decreasing
- a liquid, a solid, increasing
- a solid, a gas, increasing

A

User :AA40023

Number of ques

7 Answered
17 Not Visited

1	2	3
8	9	10
15	16	17
22	23	24

Question No. 9

All of the following can be considered physical properties EXCEPT _____.

- mass
- flammability
- color
- boiling point

B

Question No. 17

Semiconductors are located in the periodic table on (or in) the _____.

- left side.
- line dividing metals and nonmetals.
- right side.
- first period.

B

Question No. 15

The maximum number of electrons that may occupy the second energy level in the atom is _____.

- 2
- 10
- 8
- 18

10

Number of ques

14 Answered

7 Not Visited

1	2
8	9
15	16
22	23

Question No. 14

Each electron occupies the lowest energy orbital available is known as the _____.

- Pauli exclusion principle.
- Aufbau principle.
- Hund's rule.
- Heisenberg uncertainty principle.

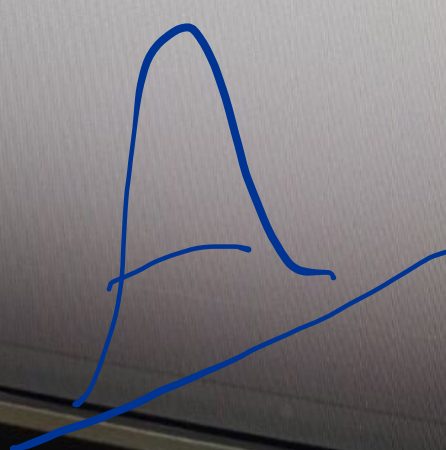
B

Question No. 13

Who formulated the modern atomic theory?

- John Dalton
- Nivaldo Tro
- Antoine Lavoisier
- Joseph Thomson

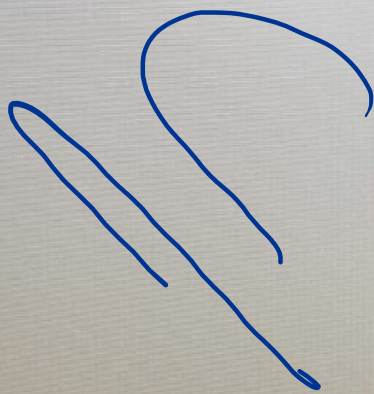
Solve & Next حل و التالي



Question No. 26

Which of the following statements is incorrect?

- The condensation of steam on a mirror is an example of a physical change.
- Evaporation of water from a fish tank is evidence of a physical change.
- The burning of a piece of charcoal to a white powder is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.



Question No. 6

Which of the following measurements are NOT equivalent?

- 25 mg = 0.025 g
- 150. msec = 0.150 sec
- 24 dL = 2.4 L
- 84 cm = 8.4 mm

2

Question No. 18

An atom of ^{14}C contains _____ protons.

- 6
- 8
- 14
- 12

A

Question No. 9

All of the following can be considered physical properties EXCEPT _____.

- mass
- flammability
- color
- boiling point

B

Question No. 10

When an egg is fried, what type of process is happening?

- melting
- a physical change
- evaporation
- a chemical change

D

User: AAA00232

Number of ques

9 Answered

15 Not Visited

1	2	3
8	9	10
15	16	17
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Question No. 13

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- Nivaldo Tro
- Antoine Lavoisier
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Semiconductors are located in the periodic table on (or in) the _____.

- left side.
- line dividing metals and nonmetals.
- right side.
- first period.

B

Question No. 16

Which of these elements is chemically similar to Sulfur?

- S
- O
- Cl
- P

B

Number of qu

15 Answered
6 Not Valued

1	2
8	9
15	16
22	

Question No. 15

The maximum number of electrons that may occupy the second energy level in the atom is _____.

- 2
- 10
- 8
- 18

8

Number of ques

14 Answered

7 Not Visited

1	2
8	9
15	16
22	23

Question No. 19

How many electrons are in Br^- ?

- 4
- 34
- 7
- 36

36

1	2	3
8	9	1
15	16	2
22	23	3

Question No. 20

The smallest subatomic particles is _____.

- neutron
- nucleus
- electron
- proton

C

Question No. 14

Each electron occupies the lowest energy orbital available is known as the _____.

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- Aufbau principle.
- Hund's rule.
- Heisenberg uncertainty principle.

A

Question No. 18

An atom of ^{14}C contains _____ protons.

- 6
- 8
- 14
- 12

A large handwritten number '7' in red ink is written across the center of the page, crossing the question text.

Question No. 22

Which of the following elements has the *lowest* electronegativity?

- Ba
- Mg
- Be
- Sr

سکریف

Question No. 22

Which of the following elements has the lowest electronegativity?

- Ba
- Mg
- Be
- Sr

Question No. 20

A neutron has an electrical charge of _____.

- 0
- 1
- 2
- +1

Save & Next حفظ و التالي

A

Question No. 15

A chloride ion (Cl^-) has the same electron configuration as a(n) _____.

- Argon atom
- Neon atom
- Fluorine atom
- Sulfur atom

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A

Question No. 23

The atomic radius of potassium is smaller than the atomic radius of _____

- cesium
- sodium
- lithium
- hydrogen

سکندر عرفان

Question No. 4

How would you classify sugar?

- pure substance-compound
- mixture-heterogeneous
- mixture-homogeneous
- pure substance-element

A

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Question No. 9

Which of the following is a chemical property?

- Density
- Viscosity
- Melting point
- Acidity

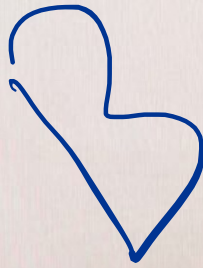
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Question No. 16

Which one of the following elements is a poor conductor of heat and electricity?

- copper
- fluorine
- lead
- iron



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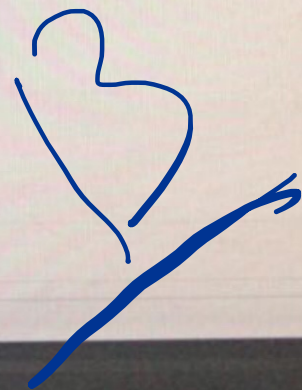


Question No. 21

The mass number of an atom can be calculated from the _____.

- number of electrons
- number of protons + neutrons
- number of electrons + protons
- number of protons

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Question No. 5

_____ is the metric unit of the temperature of an object.

- Fahrenheit
- Celsius
- Quart
- Kelvin

B

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If matter is uniform throughout and cannot be separated into other substances by physical means, it is _____

- either an element or a compound
- a homogeneous mixture
- a heterogeneous mixture
- a compound

A

Question No. 5

_____ is the metric unit of the temperature of an object.

- Fahrenheit
- Celsius
- Quart
- Kelvin

3

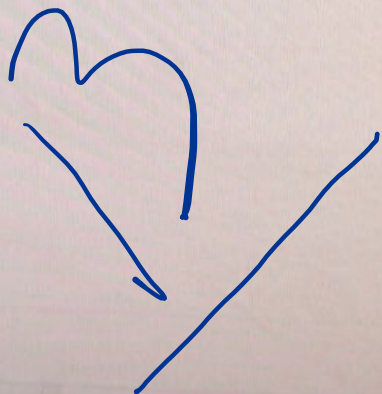
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Question No. 10

Which of the following is a physical change?

- rusting of iron
- Sublimation of dry ice
- flammability of propane gas
- baking the flour to make a cake

3



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Question No. 15

Helium atom has _____ valence electrons.

- 2
- 4
- 8
- 6

2

Question No. 9

Which of the following is a chemical property?

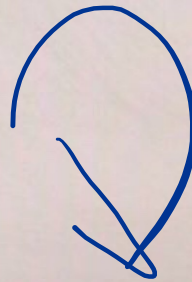
- Density
- Viscosity
- Melting point
- Acidity

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Question No. 8

Gases and liquids share the property of _____ .

- compressibility
- fixed volume
- rigid shape
- unfixed shape



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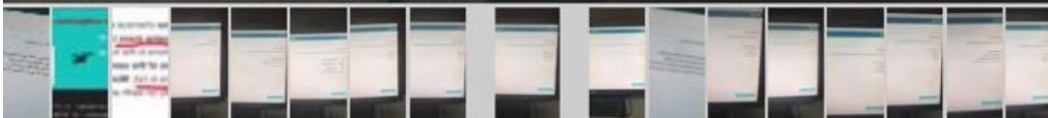
Question No. 19

In which of the following sets do all species have the same number of electrons?

- F⁻, Ne, Mg²⁺
- Ge, Se²⁻, Br⁻
- Br, Br⁻, Br⁺
- K⁺, Rb⁺, Cs⁺

A

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Question No. 7

The state of matter that has fixed volume and unfixed shape is called?

- liquid
- solid
- plasma
- gas

A

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Question No. 4

Which of the following is a homogeneous mixture?

- gasoline mixed with water
- sugar in tea
- salad dressing
- orange juice with pulp

B

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Question No. 12

Which of the following is the highest temperature?

- 40 °C
- 200 °F
- 323 K
- freezing point of water

B

Question No. 6

In which of the following is the metric unit paired with its correct abbreviation?

- microgram / mg
- milliliter / mL
- kilogram / cg
- centimeter / km

3



Question No. 25

Which of the following elements would you expect to have very similar properties to those of the element Radon, Rn?

- Ra
- At
- Ar
- Rb

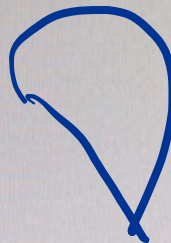
C

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Question No. 3

Choose the heterogeneous mixture from the list below.

- air
- carbon (graphite)
- chlorine gas
- water in gasoline



Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 0.023 kg
- 4.03 kg
- 0.015 kg

4

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Question No. 18

Isotopes are _____

- atoms of the same element that have different number of electrons.
- atoms of the same element that have different number of neutrons.
- atoms of the same element that have different number of protons.
- atoms of the same element that have the same number of neutrons.

B

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Question No. 25

Which of the following elements would you expect to have very similar properties to those of the element Radon, Rn?

- Ra
- Al
- Ar
- Rb

✓



Question No. 14

When placing the electrons in orbitals of the same energy, the electrons are first placed singly, the rule is known as:

- Lavoisier's principal
- Mendeleev's rule
- Dalton's atomic theory
- Hund's rule



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Question No. 16

The elements xenon, helium and argon _____

- are in the same group.
- are isotopes of each other.
- are in the same period of elements.
- have the same number of neutrons.

A

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Question No. 17

Which of the following elements has chemical properties similar to oxygen?

- nitrogen
- fluorine
- sulfur
- hydrogen



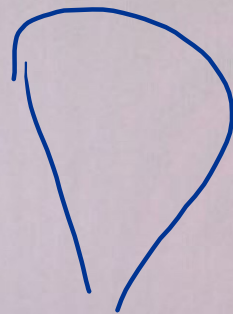
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Question No. 10

A physical change _____.

- occurs when glucose is converted into energy inside your cells
- occurs when iron rusts
- occurs when sugar is burnt
- occurs when water is evaporated



Question No. 9

All of the following can be considered physical properties EXCEPT _____.

- flammability
- taste
- color
- boiling point

A

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Question No. 8

Nitrogen gas has a(n) _____ volume and a(n) _____ shape.

- definite; indefinite
- definite; definite
- indefinite; definite
- indefinite; indefinite



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Question No. 24

Which of the following correctly matches the ion with the total number of electrons in the ion?

- Br⁻, 34 electrons
- Mg²⁺, 14 electrons
- P³⁻, 15 electrons
- Zn²⁺, 28 electrons

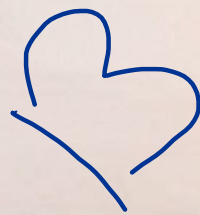
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D

Question No. 3

Compounds are pure substances that by definition consist of _____

- a single element
- two or more combined elements
- two or more mixed elements
- solids



Question No. 2

Helium has _____ valence electrons.

- 4
- 2
- 6
- 8

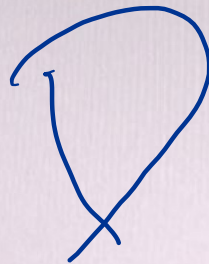
B

Question No. 8

Nitrogen gas has a(n) _____ volume and a(n) _____ shape.



- definite; indefinite
- definite; definite
- indefinite; definite
- indefinite; indefinite



Question No. 9

All of the following can be considered physical properties EXCEPT _____.

- flammability
- taste
- color
- boiling point

A

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Question No. 7

Which of the following can NOT be considered as matter?

- heat
- sand
- wood
- paper

A

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HP L1710

Question No. 4

How would you classify sugar?

- pure substance-compound
- mixture-heterogeneous
- mixture-homogeneous
- pure substance-element

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A

Question No. 15

The maximum number of electrons that may occupy the second energy level in the atom is _____.

- 2
- 10
- 8
- 18



Number of ques

14 Answered

7 Not Visited

1	2
8	9
15	16
22	23

Question No. 4

Which of the following is an example of a homogeneous mixture?

- sand mixed with oil
- salad dressing
- salt dissolved in water
- gasoline mixed with water



Question No. 13

Who formulated the modern atomic theory?

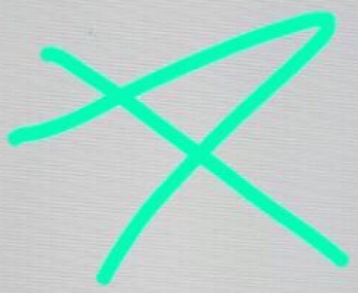
- John Dalton
- Nivaldo Tro
- Antoine Lavoisier
- Joseph Thomson

A

Question No. 8

Melting is a process in which _____ changes into _____, and it requires _____ temperature.

- a solid, a liquid, increasing
- a liquid, a gas, decreasing
- a liquid, a solid, increasing
- a solid, a gas, increasing



User :AA40023

Number of ques

7 Answered

17 Not Visited

1	2	3
8	9	10
15	16	17
22	23	24

Question No. 14

Each electron occupies the lowest energy orbital available is known as the _____.

- Pauli exclusion principle.
- Aufbau principle.
- Hund's rule.
- Heisenberg uncertainty principle.

B

Question No. 3

One of the following is an example of a pure substance

- lemon juice and tea
- sea water
- diamond
- oil and water

C

Question No. 6

Which of the following measurements are NOT equivalent?

- 25 mg = 0.025 g
- 150. msec = 0.150 sec
- 24 dL = 2.4 L
- 84 cm = 8.4 mm

Question No. 18

When an atom gains an electron, the

- a cation.
- an anion.
- an isotope.
- a proton.

5

Question No. 10

When an egg is fried, what type of process is happening?

- melting
- a physical change
- evaporation
- a chemical change

D

User: AAA00232

Number of ques

9 Answered

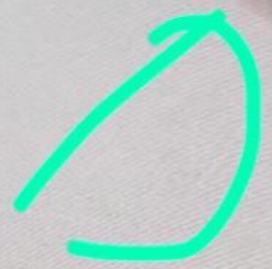
15 Not Visited

1	2	3
8	9	10
15	16	17
22	23	24

Question No. 7

Which of the following items is an example of matter?

- light
- heat
- time
- air



Question No. 9

All of the following can be considered physical properties EXCEPT _____.

- mass
- flammability
- color
- boiling point

B

Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 4.03 kg
- 0.015 kg
- 0.023 kg

B

Question No. 4

One of the following is an example of a mixture:

- lemon juice and tea
- sodium
- table salt
- water

X