

**Don't look back,
you're not going
that way.**

تجمیعات مید اول

کیمیاء (محلول)

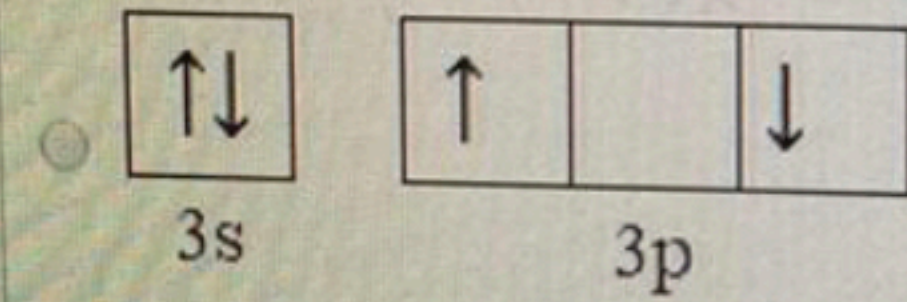
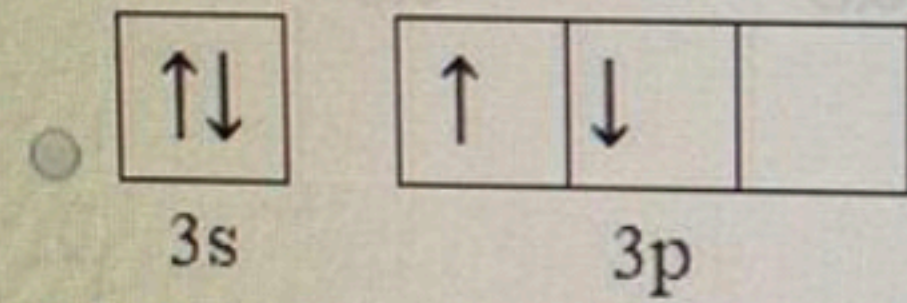
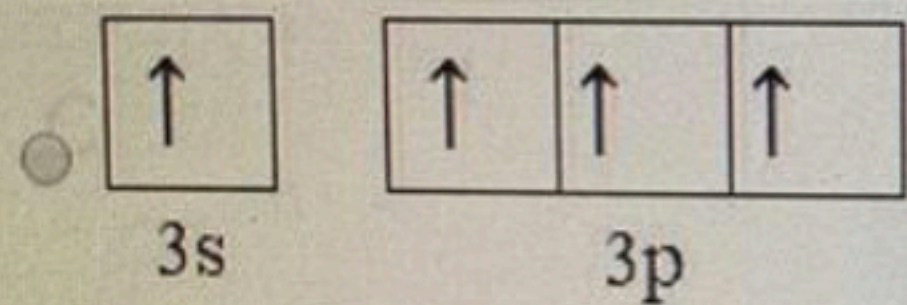
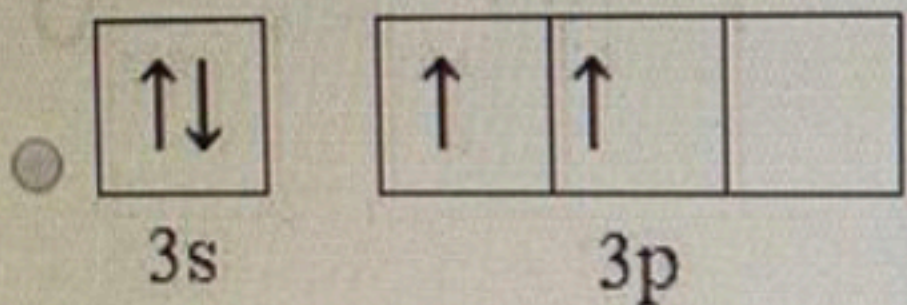
2020-1441

Eng.dhoom

Total questions in exam: 25 | Answered: 0

Question No. 13

The orbital diagram of Silicon that **violates** Aufbau principal



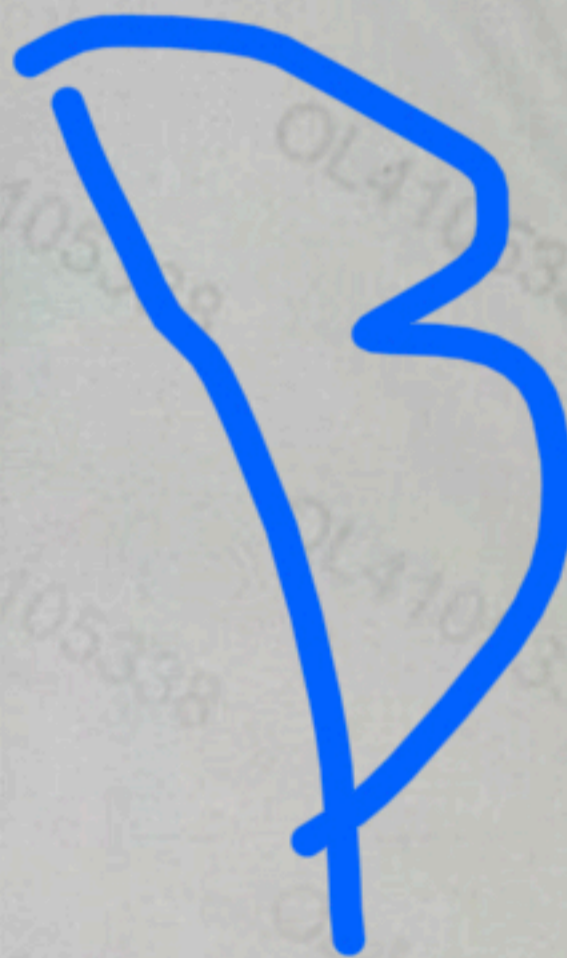
B

Total questions in exam: 25 | Answered: 0

Question No. 2

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The volume occupied by electrons is referred to as an electron cloud.
- The space occupied by the electrons is almost empty.



Total questions in exam: 25 | Answered: 0

Question No. 15

All of the following statements about different elements are true, EXCEPT

- Barium is an alkaline earth metal.
- Bromine is a metalloid.
- Krypton is one of the noble gases.
- Manganese is a transition metal.

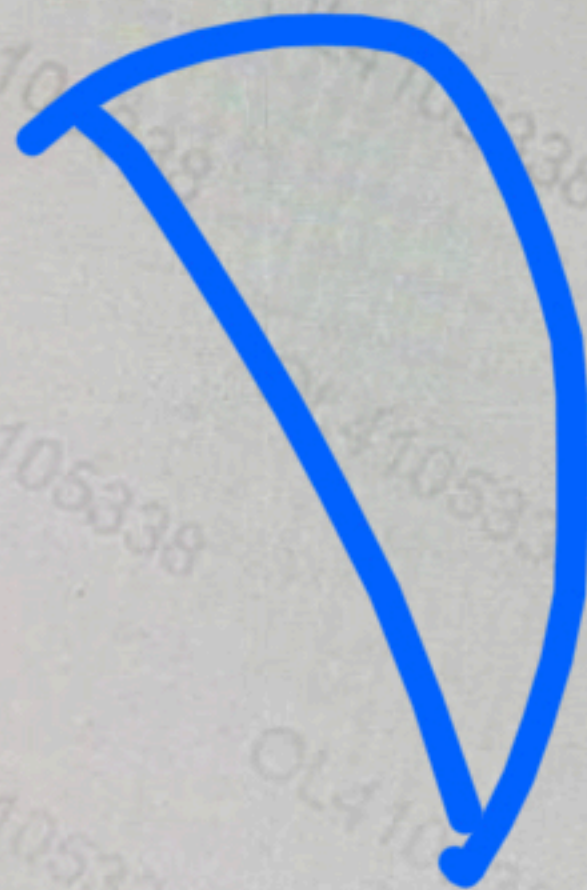
B

Total questions in exam: 25 | Answered: 0

Question No. 4

What volume of a liquid with a density of 13.53 g/mL is needed to provide 155 grams of this liquid?

- 2097.155 mL
- 141.47 mL
- 0.087 mL
- 11.45 mL



Save & Next

Total questions in exam: 25 | Answered: 0

Question No. 8

Which of the following is NOT a way of changing the state of matter?

- Melting
- Freezing
- Mixing
- sublimation

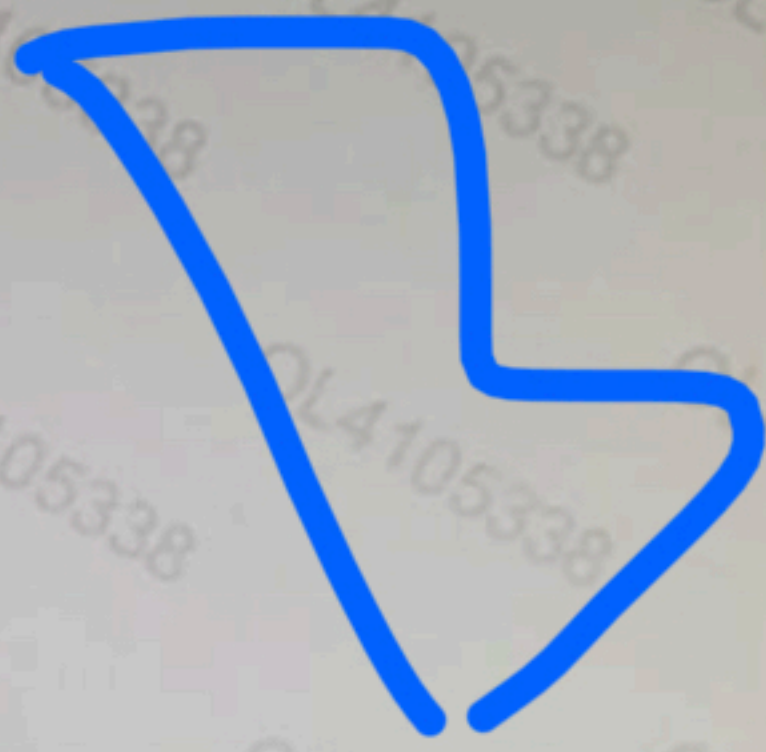
Save & Next

Total questions in exam: 25 | Answered: 0

Question No. 3

Which of these elements is chemically similar to Chlorine (Cl)?

- Sulfur
- Bromine
- Krypton
- Germanium



Save & Next

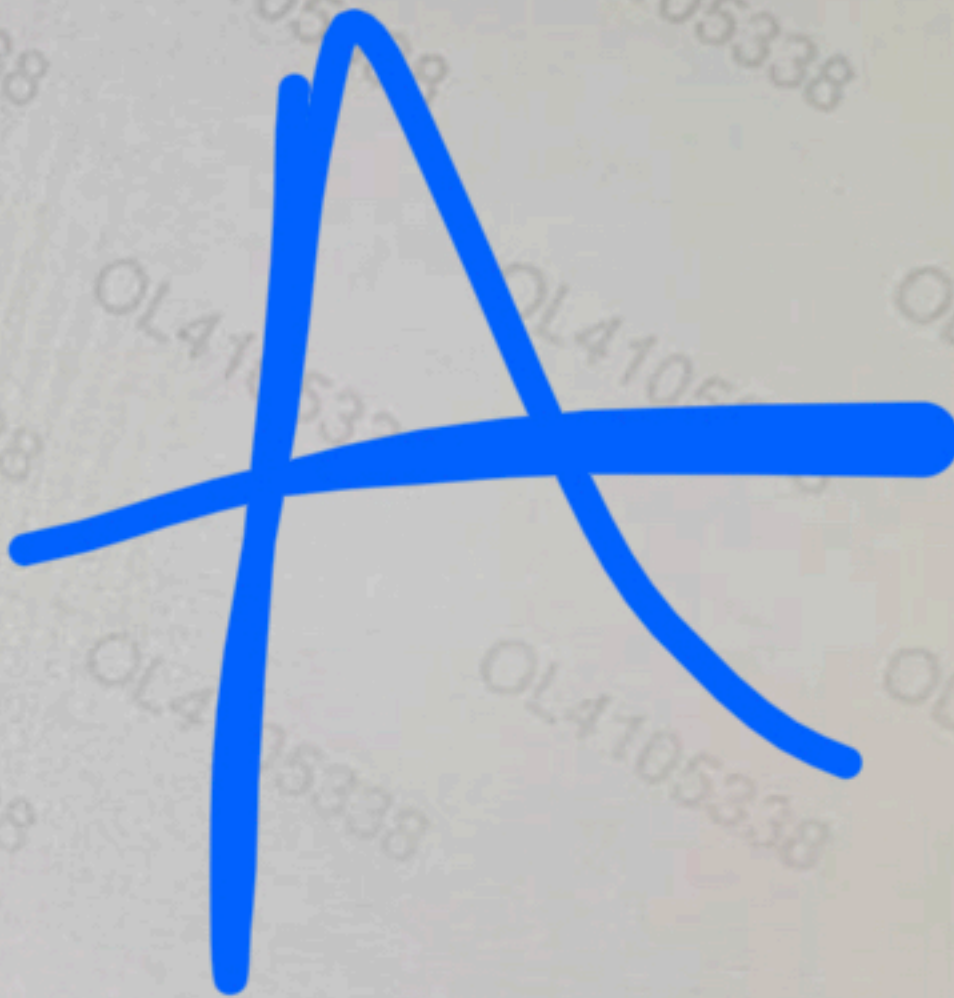


Total questions in exam: 25 | Answered: 0

Question No. 7

An object has a mass of 30.0 g and a volume of 5.0 cm³. What is its density?

- 6.00 g/cm³
- 17.00 g/cm³
- 6.00 g
- 7.00 g



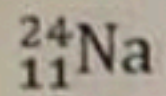
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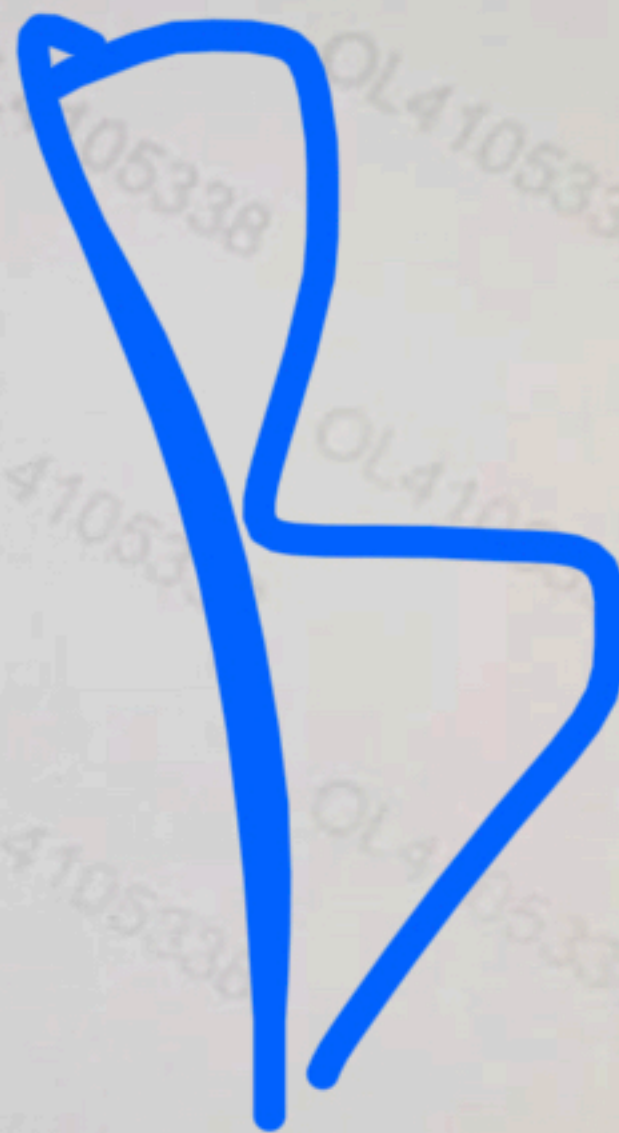
Total questions in exam: 25 | Answered: 0

Question No. 6

How many protons, neutrons, and electrons respectively are there in an atom of Sodium?



- 24, 11, 12
- 11, 13, 11
- 13, 12, 24
- 11, 11, 24



Save & Next

Total questions in exam: 25 | Answered: 0

Question No. 7

An object has a mass of 30.0 g and a volume of 5.0 cm³. What is its density?

- 6.00 g/cm³
- 17.00 g/cm³
- 6.00 g
- 7.00 g



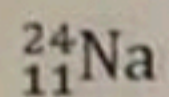
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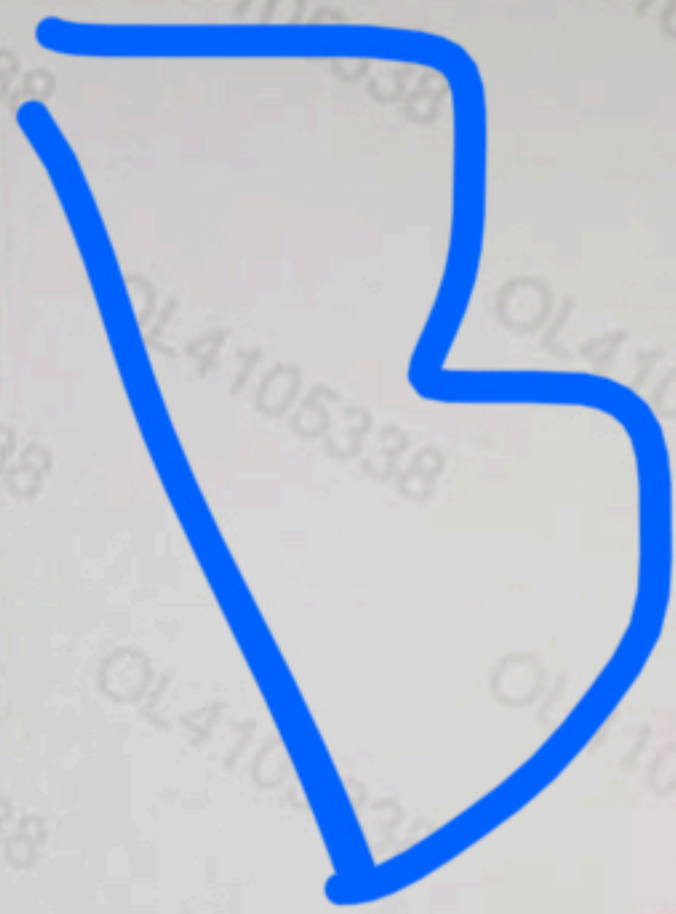
Total questions in exam: 25 | Answered: 0

Question No. 6

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- 11, 11, 24

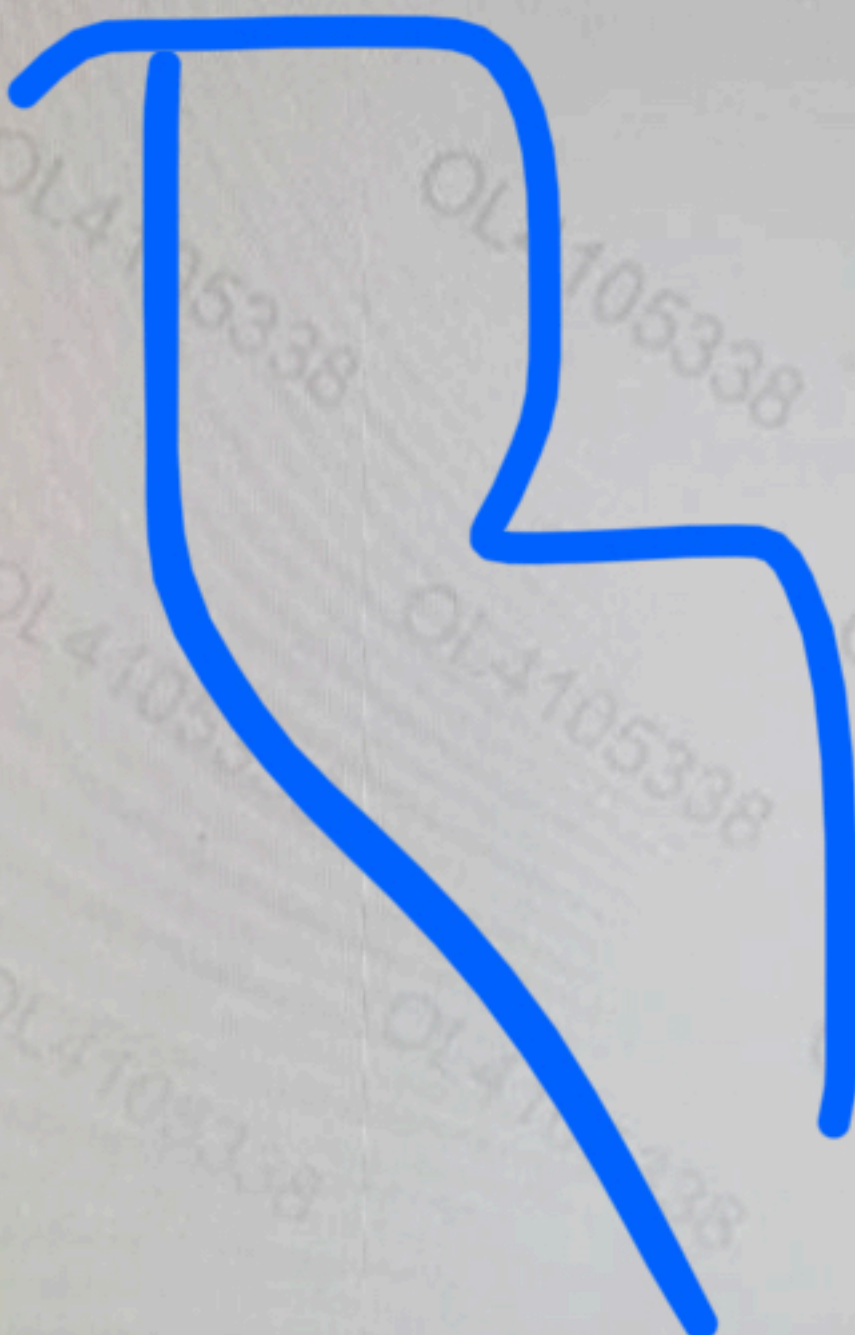


Save & Next

Question No. 9

Neon (Ne) has valence electrons.

- 2
- 8
- 6
- 4

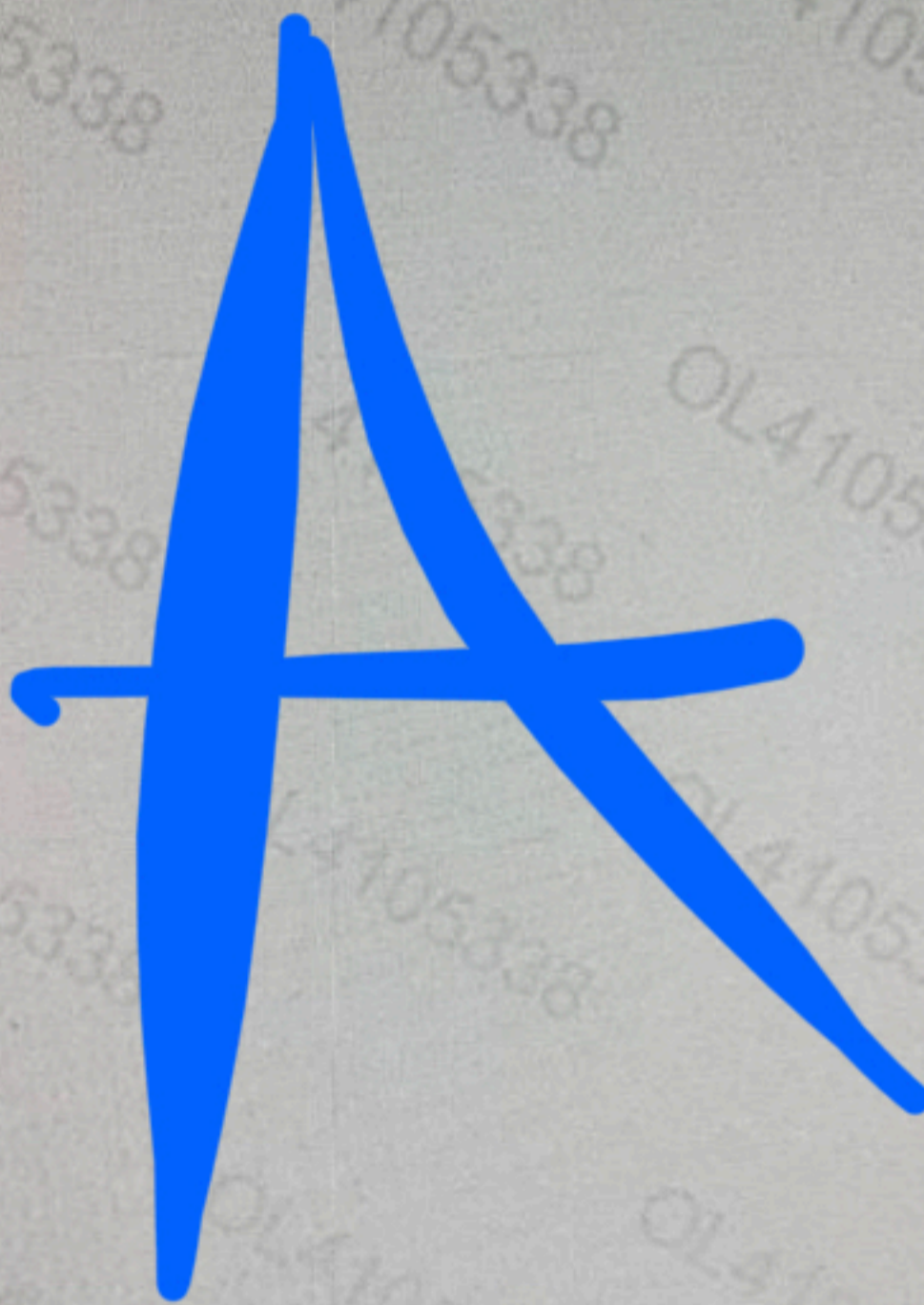


Save & Next

1 questions in exam: 25 | Answered: 0

Question No. 10

Express 7.5 nm as picometers.

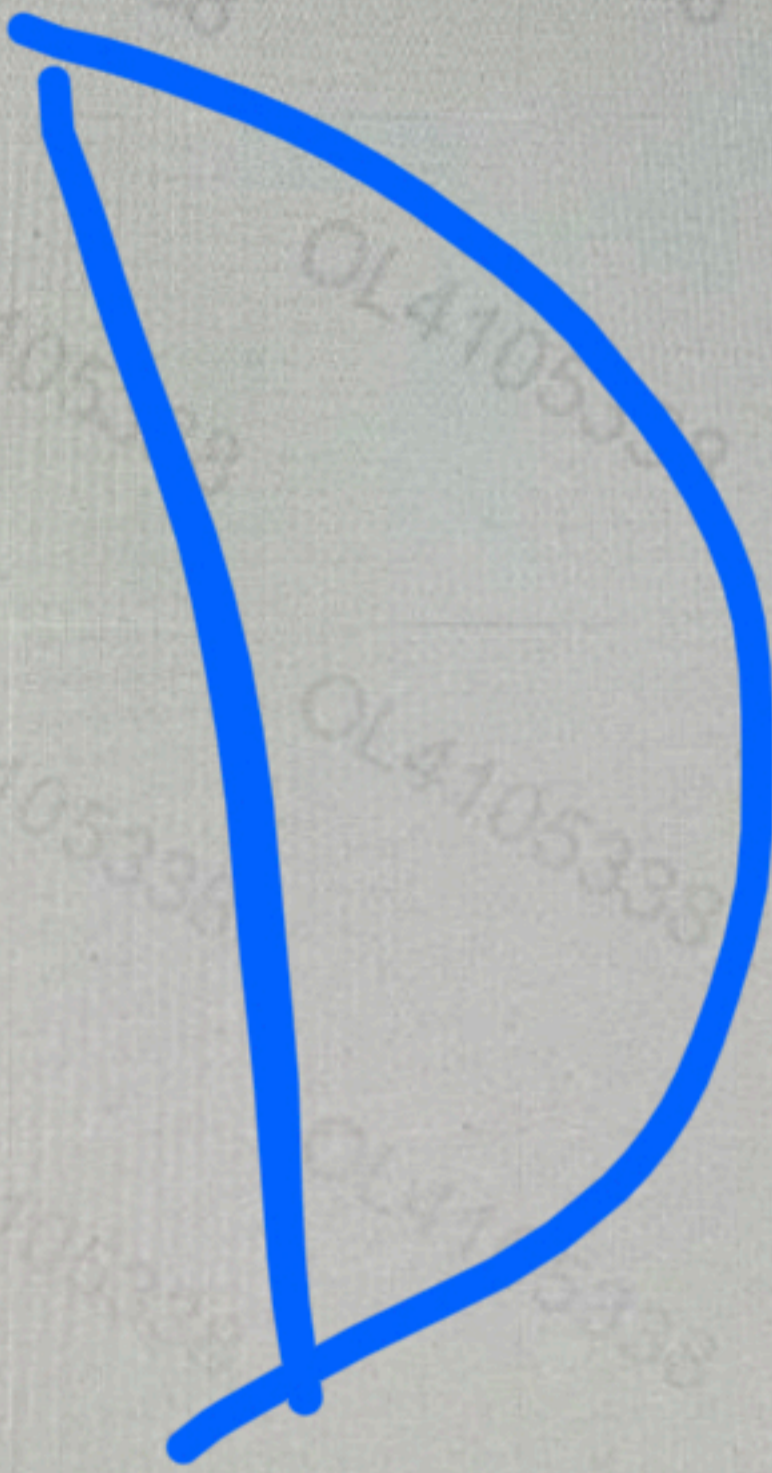
 7.5×10^3 pm 750 pm 7.5×10^{-6} pm 75.0 pm

Total questions in exam: 25 | Answered: 0

Question No. 14

Which of the following quantities represents the largest mass?

- 2.0×10^2 mg
- 0.0010 kg
- 1.0×10^5 mg
- 2.0×10^2 g.

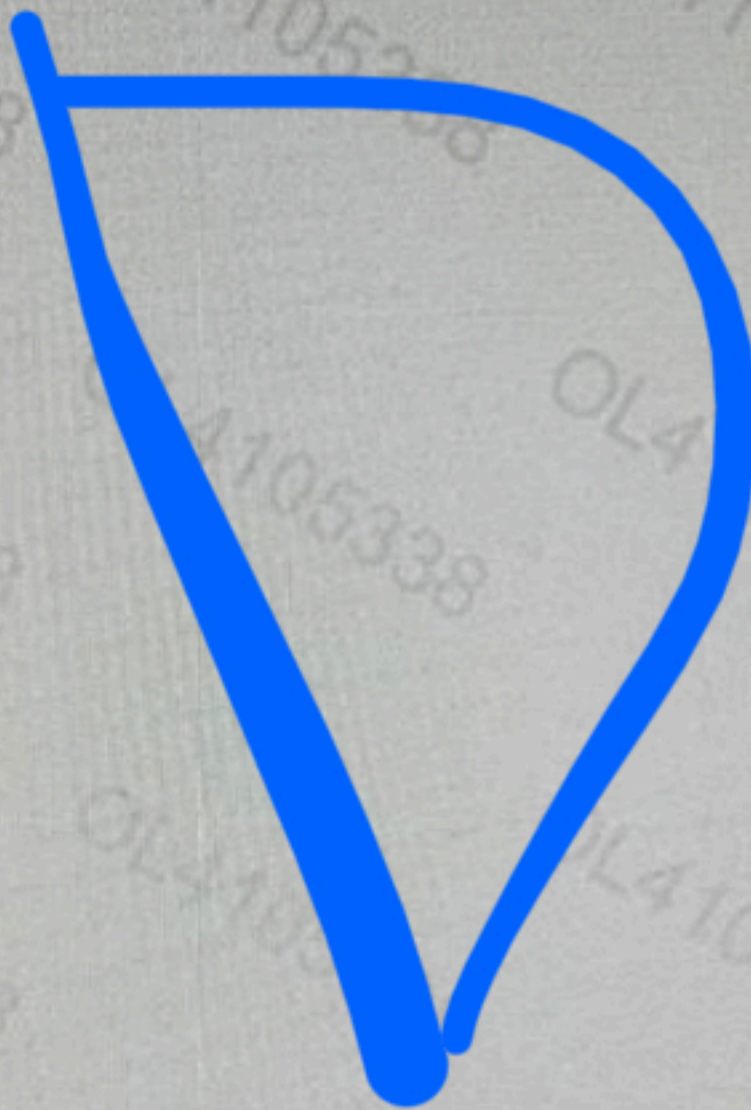


Total questions in exam: 25 | Answered: 0

Question No. 11

The charge of the Indium (In) ion is

- 1+
- 2+
- 4+
- 3+



Valence electrons are located at the

- lower energy levels.
- highest energy levels.
- outside the atom.
- nucleus.

B

Save & Next

HP L3710



Total questions in exam: 25 | Answered: 11

Question No. 5

Which of these is a physical property?

- Lead (Pb) becomes a liquid when heated to 601 degree Celsius
- Acidity of sulfuric acid
- Flammability of gasoline
- Toxicity of cyanide

A

Save & Next

10.85.7.215

P LE1851w

MKCL OES Exam Client



Total questions in exam: 25 | Answered: 11

Question No. 2

A⁻

A

A pure substance is

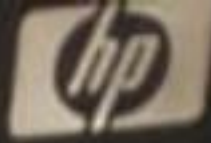
- composed of two or more different types of atoms or molecules combined in variable proportions.
- composed of only one type of atoms or molecules.
- composed of two or more regions with different compositions.
- composed of two or more different types of atoms or molecules homogenously mixed together.

B

Save & Next

06-7-2019

E1851w



Total questions in exam: 25 | Answered: 11

Question No. 20

A⁻

Which of the following statements is **NOT** correct?

- Atoms are made of subatomic particles.
- An atom has mass.
- Atoms are invisible.
- Each element is formed of different atoms.

D

Save & Next

10657215

Total questions in exam: 25 | Answered: 11

Question No. 11

Which orbital-filling diagram does NOT obey Hund's rule?

- \uparrow $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$
 $4s$ $3d$
- $\uparrow\downarrow$ $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ $\underline{\hspace{1cm}}$
 $4s$ $3d$
- $\uparrow\downarrow$ $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ $\underline{\hspace{1cm}}$ $\underline{\hspace{1cm}}$
 $4s$ $3d$
- $\uparrow\uparrow$ $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ \uparrow $\underline{\hspace{1cm}}$ $\underline{\hspace{1cm}}$
 $4s$ $3d$



Save & Next

10/05/2015

HP LE1851w

Total questions in exam: 25 | Answered: 11

Question No. 16

..... has one kind of atoms while has more than one kind

- A compound, an element
- An element, a compound
- A mixture, an element
- A mixture, a compound

B

Save & Next

Total questions in exam: 25 | Answered: 11

Question No. 6

How would you classify sugar?

- A heterogeneous mixture.
- A homogeneous mixture.
- An element.
- A compound.

D

Save & Next

10.05.7.215

Total questions in exam: 25 | Answered: 11

Question No. 7

When an atom gains electron, the resulting particle is called

- an isotope.
- a proton.
- a cation.
- an anion.



Save & Next

18/09/2015

HP LE1851w

Total questions in exam: 25 | Answered: 11

Question No. 10

Isotopes are

B

- atoms of the same element that have the same number of neutrons.
- atoms of the same element that have different number of neutrons.
- atoms of the same element that have different number of electrons.
- atoms of the same element that have different number of protons.

Save & Next

Total questions in exam: 25 | Answered: 1

Question No. 3

The elements of group 1A and 7A are, and the elements of group 8A are.....

- reactive, unreactive
- unreactive, unreactive
- reactive, reactive
- unreactive, reactive

[Save & Next](#)



User : A

Number
Number

15 Ans

0 Not

1 2

8 9

15 16

22 23

Question No. 19

The fact that Carbon and Oxygen can combine together in different ratio of 1:1 and 1:2 is a representation to the Law of

- definite proportions.
- conservation of matter.
- conservation of energy.
- multiple proportions.

D

Save & Next

10/05/2020

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HP Compaq L61711

The building-block of a compound, whose properties are the same as those of the compound is the

- element.
- atom.
- molecule.
- mixture.

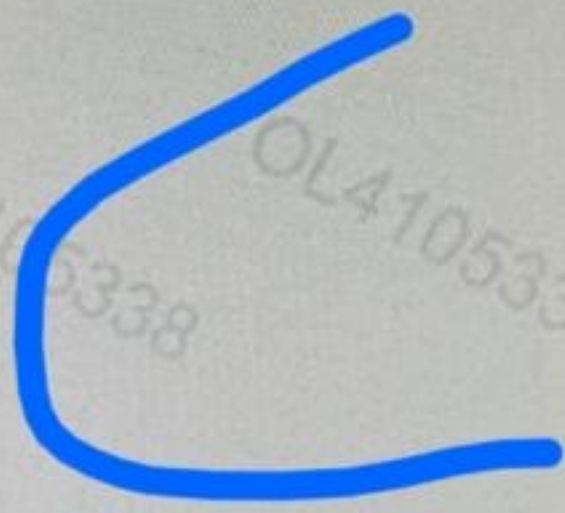


Save & Next

Total questions in exam: 25 | Answered: 0

Question No. 16

Which of these represents a chemical change?

- Mixing sugar and water at room temperature.
 - Boiling water to form steam.
 - Digestion of food in the stomach.
 - Melting butter at room temperature.
- 

Total questions in exam: 25 | Answered: 11

Question No. 22

Which of the following can NOT be considered as matter?

- wood
- sand
- paper
- light



Save & Next

08/05/2015

HP LE1851w

MKCL OES Exam Client Version 2.0.0.1

Total questions in exam: 25 | Answered: 0

Question No. 2

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The volume occupied by electrons is referred to as an electron cloud.
- The space occupied by the electrons is almost empty.

B

Save & Next

Question No. 19

The fact that Carbon and Oxygen can combine together in different ratio of 1:1 and 1:2 is a representation to the Law of

- definite proportions.
- conservation of matter.
- conservation of energy.
- multiple proportions.

Save & Next

User : A

Number
Number

15 Ans

0 Not

1 2

8 9

15 16

22 23

Thestate has lowest (K.E) while
thehas the highest(K.E)

Solid,gas

Solid,liquid

liquid,gas

Solid ,liqud

A



Chemistry_Quiz1_Sem2_2020

CL OES

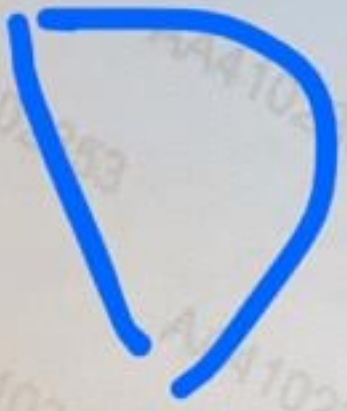
Total questions in exam: 25 | Answered: 15

A⁻ A A⁺

Question No. 17

Combustion of fuel in car engine is considered as a

- physical change.
- physical property.
- chemical property.
- chemical change.



Save & Next

10.66.7.215

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HP Compaq (E171)

Question No. 12

An isotope represented as Fe-58

- is an isotope of fluorine.
- has 32 protons.
- contains 32 neutrons.
- must have 58 electrons.



User: AA4102383

Number of main questions: 28

Number of questions: 28

7 Answered

11 Not Answered

12 Not Marked

0 Partially Answered

1	2	3	4	5	6	7
8	9	10	11	12	13	14
15	16	17	18	19	20	21
22	23	24	25			

Calculator

Instructions

Help

End Test

Table of the elements

9		10		11		12		13	14	15	16	17	18
8B				8B		8B		3A	4A	5A	6A	7A	8A
27 Co Cobalt 58.93	28 Ni Nickel 58.69	29 Cu Copper 63.55	30 Zn Zinc 65.39	31 Ga Gallium 69.72	32 Ge Germanium 72.61	33 As Arsenic 74.92	34 Se Selenium 78.96	35 Br Bromine 79.90	36 Kr Krypton 83.80	37 Rb Rubidium 85.47	38 Sr Strontium 87.62	39 Y Yttrium 88.91	40 Zr Zirconium 91.22
45 Rh Rhodium 102.91	46 Pd Palladium 106.42	47 Ag Silver 107.87	48 Cd Cadmium 112.41	49 In Indium 114.82	50 Sn Tin 118.71	51 Sb Antimony 121.76	52 Te Tellurium 127.60	53 I Iodine 126.90	54 Xe Xenon 131.29	55 Cs Cesium 132.91	56 Ba Barium 137.33	57 La Lanthanum 138.91	58 Ce Cerium 140.12
77 Ir Iridium 223.83	78 Pt Platinum 195.08	79 Au Gold 196.97	80 Hg Mercury 200.59	81 Tl Thallium 204.38	82 Pb Lead 207.2	83 Bi Bismuth 208.98	84 Po Polonium (209)	85 At Astatine (210)	86 Rn Radon (222)	87 Fr Francium (223)	88 Ra Radium (226)	89 Ac Actinium (227)	90 Th Thorium (232)

All of the following SI unit of measurement.except

- Secand
- meter
- Kelvin
- yard


D

Deposition the process in
which.....changed into.....?

- A. gas,solid
- B.liquid,solid
- C.solid,gas
- D.gas ,liquid

A

Nh₄ amonia class as

- Molecular elements
- atomic elements 
- mixture of emelents
- molecular compound

Total questions in exam: 28 | Answered: 3



User: AA4102363

Number of main questions: 28
Number of questions answered: 3

3 Correct
18 Not Answered

Question No. 1

A A A

The charge of the nucleus of a Calcium atom is

- 2+
- 40+
- 20-
- 20+



Save & Next

1	2	3	4
5	6	10	11
15	16	17	18
22	23	24	25

Calculator
Help

Periodic table of the elements

Key

- Atomic number
- Element symbol
- Element name
- Average atomic mass*

8		9		10		11		12		13	14	15	16	17
26		27		28		29		30		31	32	33	34	35
Fe		Co		Ni		Cu		Zn		Al	Si	P	S	Cl
Iron		Cobalt		Nickel		Copper		Zinc		Aluminum	Silicon	Phosphorus	Sulfur	Chlorine
55.85		58.93		58.69		63.55		65.39		26.98	28.09	30.97	32.07	35.45
44		45		46		47		48		49	50	51	52	53
Ru		Rh		Pd		Ag		Cd		In	Sn	Sb	Te	I
Ruthenium		Rhodium		Palladium		Silver		Cadmium		Indium	Tin	Antimony	Tellurium	Iodine
101.07		102.91		106.42		107.87		112.41		114.82	118.71	121.76	127.4	126.9
76		77		78		79		80		81	82	83	84	85
Os		Ir		Pt		Au		Hg		Tl	Pb	Bi	Po	At
Osmium		Iridium		Platinum		Gold		Mercury		Thallium	Lead	Bismuth	Polonium	Astatine
190.23		192.22		195.08		196.97		200.59		204.38	207.2	208.98	209	210
109		110		111		112		113		114	115	116	117	118
Mt		Ds		Nh		Fl		Mc		Lv	Ts	Og	Uu	Uub
Meitnerium		Darmstadtium		Nihonium		Flerovium		Moscovium		Livermorium	Tennessine	Oganesson	Ununseptium	Ununoctium
(288)		(285)		(284)		(289)		(288)		(286)	(289)	(286)	(285)	(284)

Total questions in exam: 25 | Answered: 0

Question No. 1

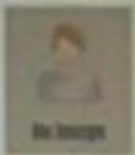
A⁻ A A⁺A chloride ion (Cl⁻) has the same electron configuration as

.....

- an argon atom.
- a sulfur atom.
- a neon atom.
- a fluorine atom.



Save & Next



User: AA4102353

Number of marks: 1
Number of questions: 250 Answered
24 Not Visited

1	2	3
8	9	10
15	16	17
22	23	24

Calculator

Notepad

Total questions in exam: 28 | Answered: 1

Question No. 1

A A A

Which of the following measurements are NOT equivalent?

- 25 mg = 0.025 g
- 24 dL = 2.4 L
- 150 msec = 0.150 sec
- 84 cm = 8.4 mm



Save & Next

User: AA

Number of Questions: 28

Number of Correct Answers: 1

Score: 17

Calculator

Stop

Periodic table of the elements

Key
 - Atomic number
 - Element symbol
 - Element name
 - Average atomic mass*

8		9		10		11		12		13	14	15	16
3A		4A		5A		6A		7A		8A		9A	
5	6	7	8	9	10	11	12	13	14	15	16	17	18
B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar
Boron 10.81	Carbon 12.01	Nitrogen 14.01	Oxygen 16.00	Fluorine 18.99	Neon 20.18	Sodium 22.99	Magnesium 24.31	Aluminum 26.98	Silicon 28.09	Phosphorus 30.97	Sulfur 32.07	Chlorine 35.45	Argon 39.95
19	20	21	22	23	24	25	26	27	28	29	30	31	32
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Cu	Zn	Ga	Ge	As	Se
Potassium 39.10	Calcium 40.08	Scandium 44.96	Titanium 47.88	Vanadium 50.94	Chromium 52.00	Manganese 54.94	Iron 55.85	Copper 63.55	Zinc 65.39	Gallium 69.72	Germanium 72.61	Arsenic 74.92	Selenium 78.96
39	40	41	42	43	44	45	46	47	48	49	50	51	52
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Ag	Cd	In	Sn	Sb	Te
Rubidium 85.47	Strontium 87.62	Yttrium 88.91	Zirconium 91.22	Niobium 92.91	Molybdenum 95.94	Technetium 98.91	Ruthenium 101.07	Silver 107.87	Cadmium 112.41	Indium 114.82	Tin 118.71	Antimony 121.76	Tellurium 127.60



Total questions in exam: 25 | Answered: 0

Question No. 4



What is the prefix multiplier used to represent the factor 10^{-6} ?

- Mega
- nano
- kilo
- micro



Save & Next

Total questions in exam: 25 | Answered: 0

Question No. 3



Which of the following properties is **NOT** a characteristic of the group 1A (1) metallic elements?

- They are shiny.
- They are good conductors of heat.
- Most of them are liquids at room temperature.
- They are good conductors of electricity.

[Save & Next](#)

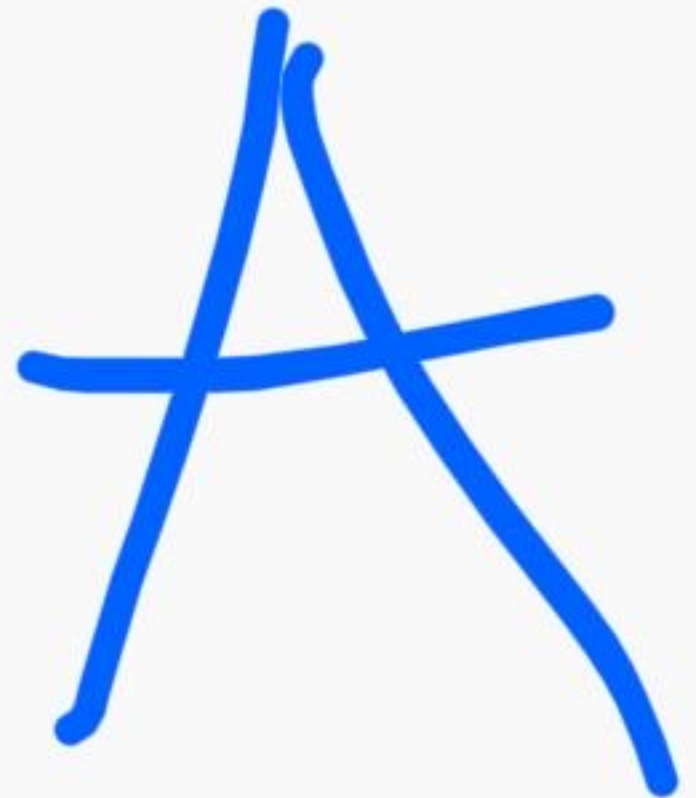
How would you classify the color
of
the rose

Physical properties

Physical change

Chemical properties

Chemical change



Total questions in exam: 25 | Answered: 23

Question No. 6

Valence electrons are located in

- the first three shells of an atom
- the outermost energy level of an atom
- the nucleus of an atom
- the innermost energy level of an atom

B

Save & Next

HP L1710

Valence electrons are located at the

- lower energy levels.
- highest energy levels.
- outside the atom.
- nucleus.

B

Save & Next

HP L3710



According to Aufbau principle,
which one is correct:

A- 3s 3p 3d

B- 4p 4d 4f

C- 3p 4s 3d



هذا الي جاني

Question No. 28

Which of the following properties is **NOT** a characteristic of the group 8A (18) in the periodic table?

- They are poor conductors of electricity.
- They are poor conductors of heat.
- They are gases at room temperature.
- They are shiny.



Save & Next

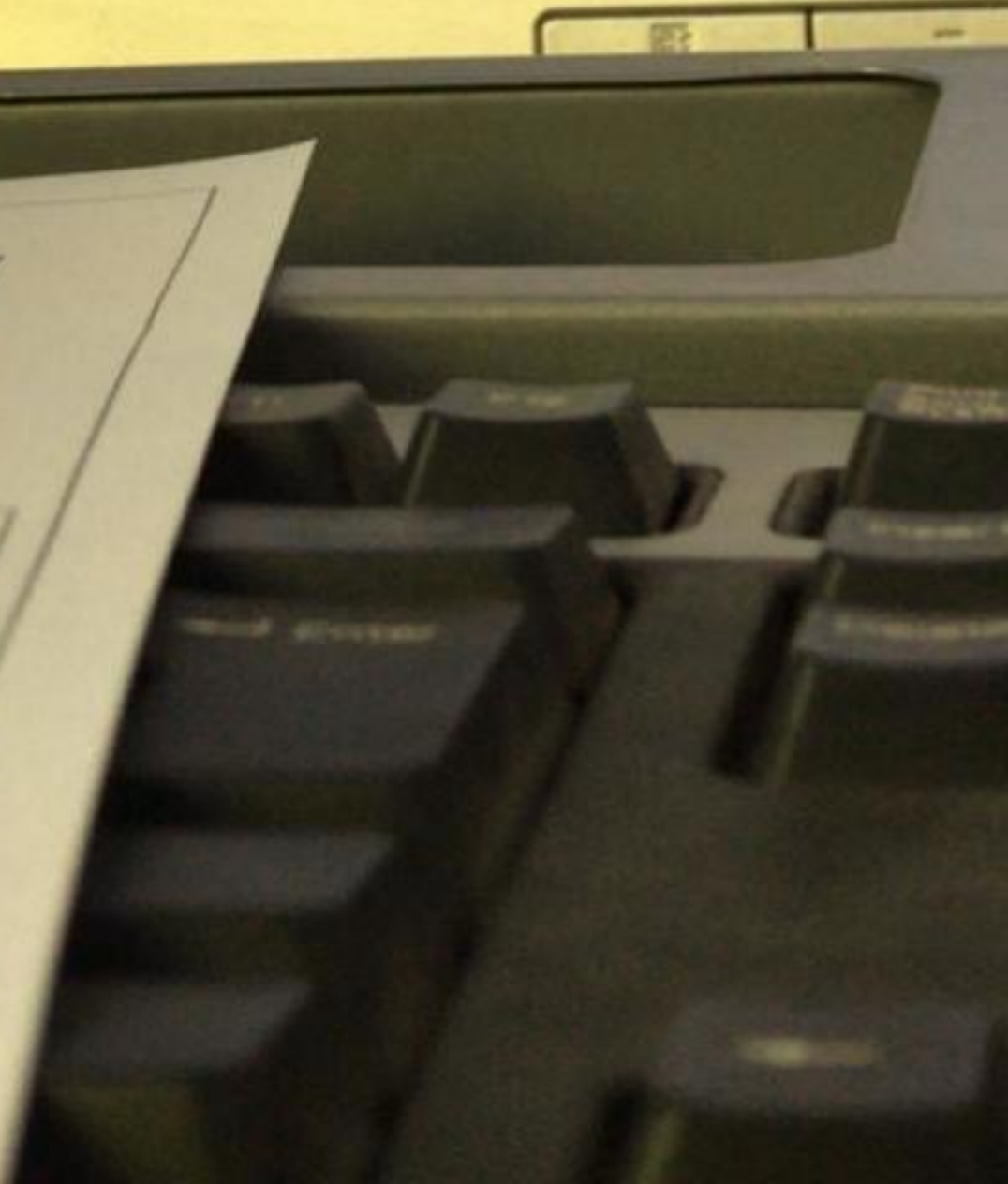
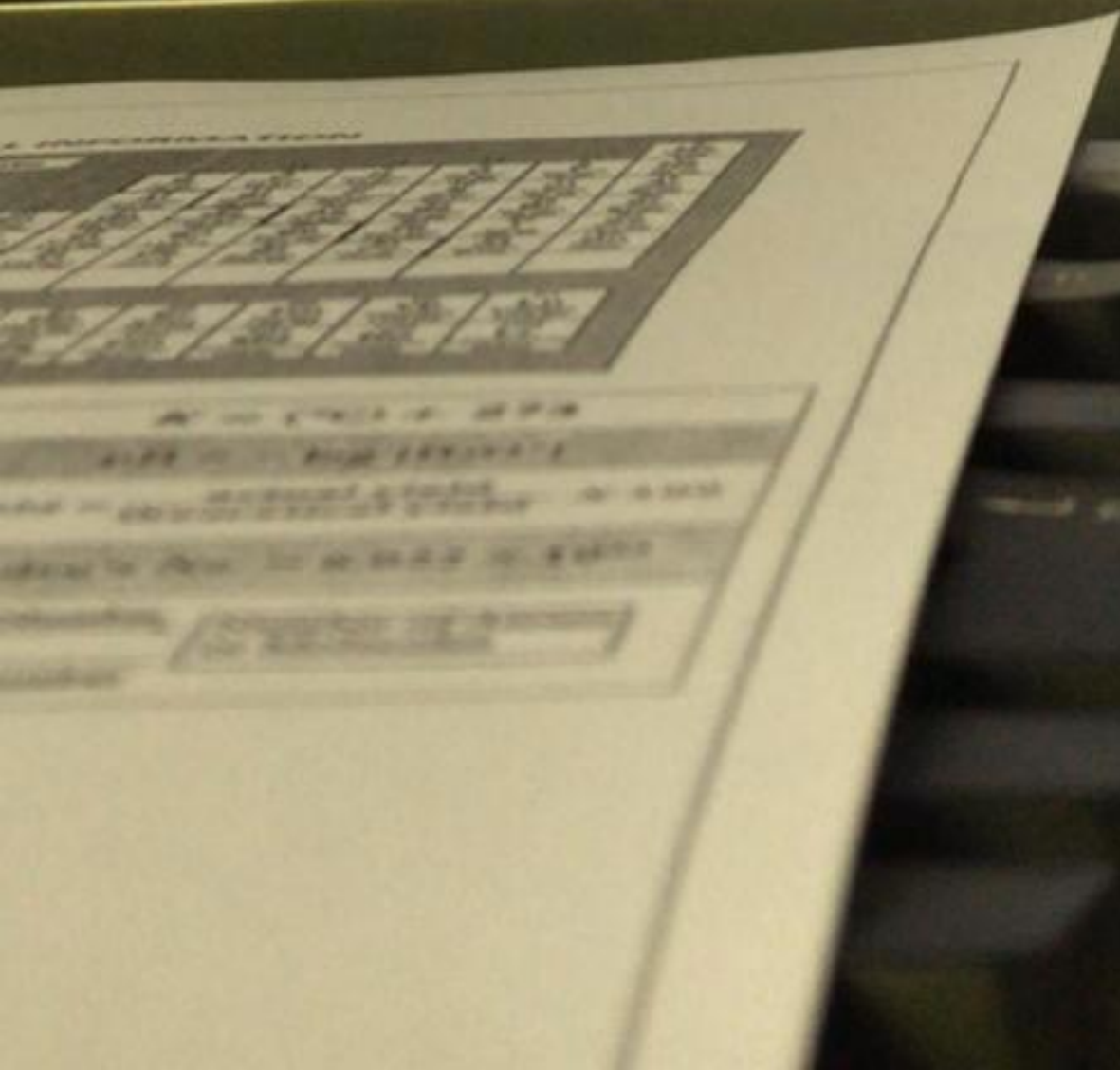
The Rutherford gold foil experiment demonstrated that atoms



- are visible to the naked eye
- are homogeneous
- consist of an almost empty nucleus surrounded by a dense cloud of electrons.
- consist of a dense nucleus surrounded by mostly empty space.

Save & Next

Compaq LE1711



Total questions in exam: 25 | Answered: 2

Question No. 14

Compounds are composed of

- bonded atoms of two or more elements.
- only one type of atoms.
- bonded atoms of the same element.
- two or more mixed substances.

A

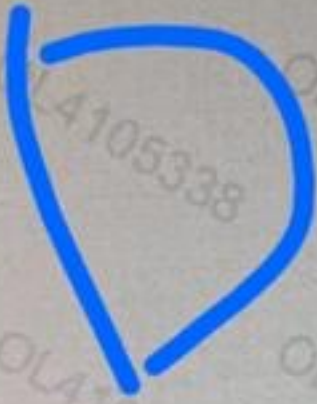
Save & Next

10/10/2015

Question No. 1

The drops of water that appear on the outside of a glass of cold juice are an example of

- sublimation.
- freezing.
- evaporation.
- condensation.



Save & Next

Question No. 23

Which of the following measurements are NOT equivalent?

- 24 dL = 2.4 L
- 150 msec = 0.150 sec
- 84 cm = 8.4 mm
- 25 mg = 0.025 g

[Save & Next](#)



The value of the electrical charge of electron was measured by

- J. Chadwick.
- R. Millikan.
- E. Rutherford.
- J. Dalton.

B

Save & Next

Question No. 24

The gaseous state is characterized by

- hardness
- fixed shape
- compressibility
- fixed volume



Save & Next

Total questions in exam: 25 | Answered: 0

Question No. 18

The maximum number of electrons that may occupy the second energy level ($n = 2$) in the atom

- 2
- 8
- 10
- 18

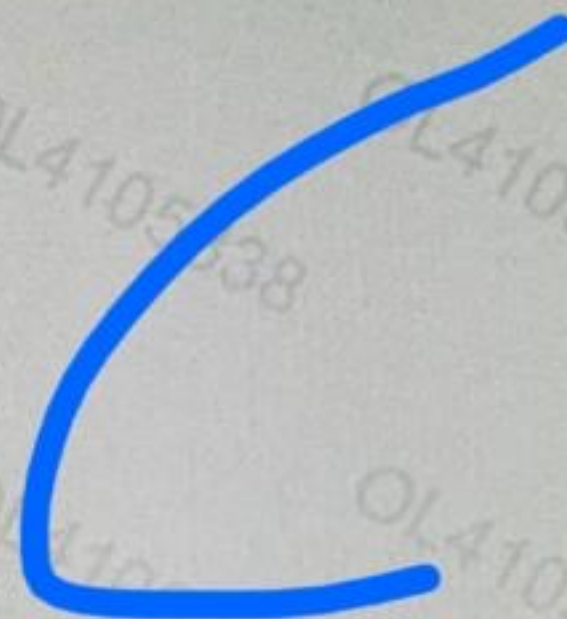
B

Total questions in exam: 25 | Answered: 0

Question No. 16

Which of these represents a chemical change?

- Mixing sugar and water at room temperature.
- Boiling water to form steam.
- Digestion of food in the stomach.
- Melting butter at room temperature.



Questions in exam: 25 | Answered: 0

Question No. 12

The element vanadium (V) is one of the

 noble gases. halogens. transition metals. alkaline earth metals.

Total questions in exam: 25 | Answered: 7

Question No. 3

Which one of the following species has same number of electrons as Mg^{2+} cation?

- K⁺
- Cl⁻
- S²⁻
- F⁻





Question No. 1

How many micrometers are there in 6.0 km?

- 6.0×10^6 micrometers
- 1.7×10^{-9} micrometers
- 6.0×10^9 micrometers
- 1.7×10^{-4} micrometers

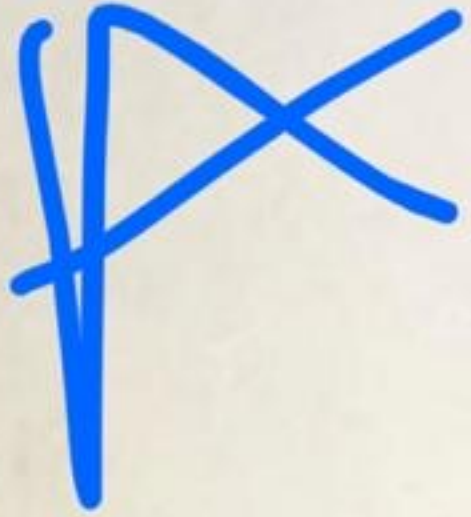
C

Save & Next

Question No. 11

The composition is NOT the same in all parts of the

- Heterogeneous mixture
- pure substance
- Homogeneous and heterogeneous mixtures
- Homogeneous mixture



Save & Next

Question No. 11

The composition is NOT the same in all parts of the

- Heterogeneous mixture
- pure substance
- Homogeneous and heterogeneous mixtures
- Homogeneous mixture

13

Save & Next

Question No. 22

Which of these elements is a good conductor of electricity?

- He
- Ag
- S
- N

B

Save & Next

Question No. 12

..... is an example of physical changes.

- Dry ice subliming
- Burning wood
- Combustion of gasoline
- Rusting of iron

A

Save & Next

Question No. 8

How many protons are there in an isotope of Sodium with a mass number of 25?

- 11
- 25
- 14
- 15

A

Save & Next

Total questions in exam: 25 | Answered: 2

Question No. 2

How many inches are there in 0.5 km? (1 inch = 2.54 cm)

- 127000 inch
- 19685 inch
- 1968.5 inch
- 12700 inch

B

Total questions in exam: 25 | Answered: 0

Question No. 7

An object has a mass of 30.0 g and a volume of 5.0 cm³. What is its density?

- 6.00 g/cm³
- 17.00 g/cm³
- 6.00 g
- 7.00 g

Question No. 21

What is the charge on the Barium ion?

- 1+
- 2+
- 2-
- 1-

3

Save & Next

Total questions in exam: 25 | Answered: 0

Question No. 1

Which of the following pairs are isotopes?

- H-2 and He-2
- Pa-141 and Po-209
- U-238 and U-240
- C-12 and Co-59



Save & Next

Total questions in exam: 25 | Answered: 12

Question No. 15

An anion is defined as _____

- a group of stable atoms.
- a charged atom or group of atoms with a net negative charge.
- an atom or group of atoms with a net positive charge.
- a stable atom.

Save & Next

Total questions in exam: 25 | Answered: 12

Question No. 20

How many unpaired electrons are there in an Iron atom (Fe) in its ground state?

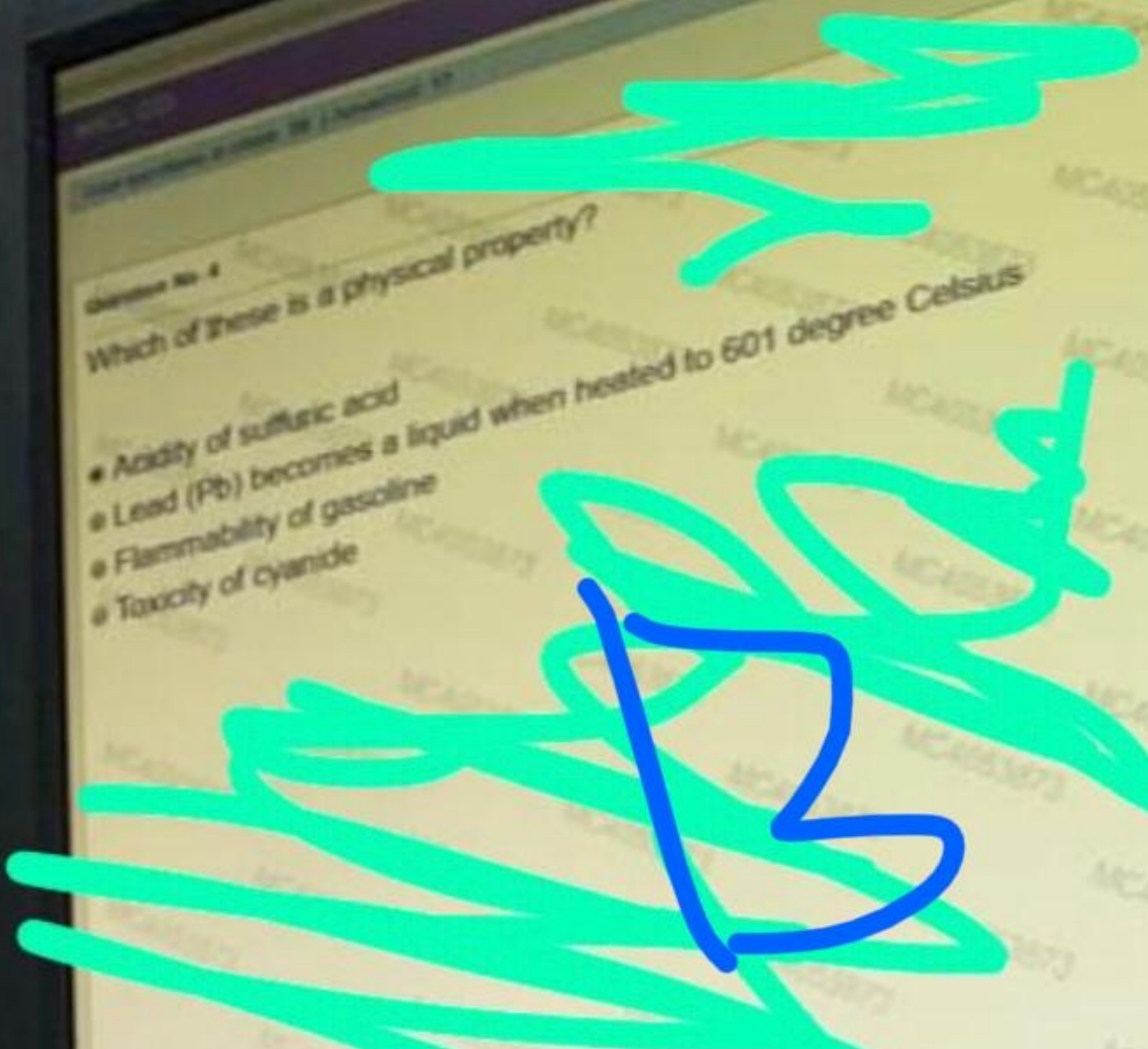
- 2
- 4
- 6
- 0

B

Save & Next

Question No. 4
Which of these is a physical property?

- Acidity of sulfuric acid
- Lead (Pb) becomes a liquid when heated to 601 degree Celsius
- Flammability of gasoline
- Toxicity of cyanide



رسالة محوّلة →

جانبي S8

Hetro

Homo

Molecular element

Compound

ص 9:28



Total questions in exam: 25 | Answered: 12

Question No. 24

What is the class of P_4 ?

- Compound
- Homogeneous mixture
- Molecular element
- Heterogeneous mixture



Save & Next

10/10/2018 11:17 (Answered: 17)

Which of the following statements is true?

- Energy cannot be converted between potential and kinetic energies.
- As the temperature of a substance increases, its kinetic energy decreases.
- Liquids have both potential and kinetic energies, regardless to their physical states.
- Solids have the greatest kinetic energy and gases have the lowest kinetic energy.



Question No. 1

All of the following are examples of compounds, EXCEPT

- Co
- CO₂
- FeO
- CO



The building-block of a compound, whose properties are the same as those of the compound is the

- element.
- atom.
- molecule.
- mixture.



Question No. 12

An element cannot

- combine with other elements to form compound.
- be part of a heterogeneous mixture.
- be part of a homogeneous mixture.
- be broken down into simpler substances in a chemical process.

b

Save & Next

Question No. 14

Consider the Bromine isotope Br-81, select the combination which lists respectively the atomic number, neutrons number, and mass number.

- 46, 81, 35
- 35, 46, 81
- 35, 81, 46
- 81, 46, 35

B

Save & Next

Question No. 19

What decimal power does the abbreviation "pico" represent?

- 1×10^{-1}
- 1×10^{-12}
- 1×10^{-9}
- 1×10^6

B

Save & Next

Question No. 13

The electronic distribution of the electrons on the principal energy levels of ^{17}Cl is

- 2, 8, 7
- 2, 2, 8, 5
- 2, 8, 2, 5
- 2, 2, 5, 8



Save & Next

Question No. 7

In the periodic table, group 7A (17) elements are also called

.....

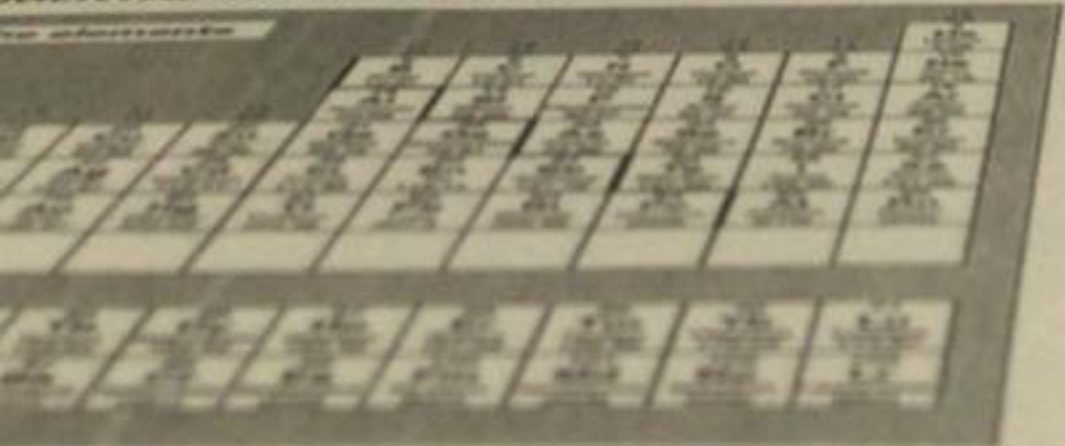
- alkaline earth metals.
- alkali metals.
- noble gases.
- halogens.

D

Save & Next

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EXPERIMENTAL INFORMATION



$R = (T) + 273$
 $\ln P = - \ln P_0 + \frac{U}{RT}$
Percent yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$
Avogadro's No. = 6.022×10^{23}

Question No. 1

The "f" sublevel can hold up to electrons.

- 18
- 8
- 14
- 6



Save & Next

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EXAM FOR SUPPLEMENTAL INFORMATION

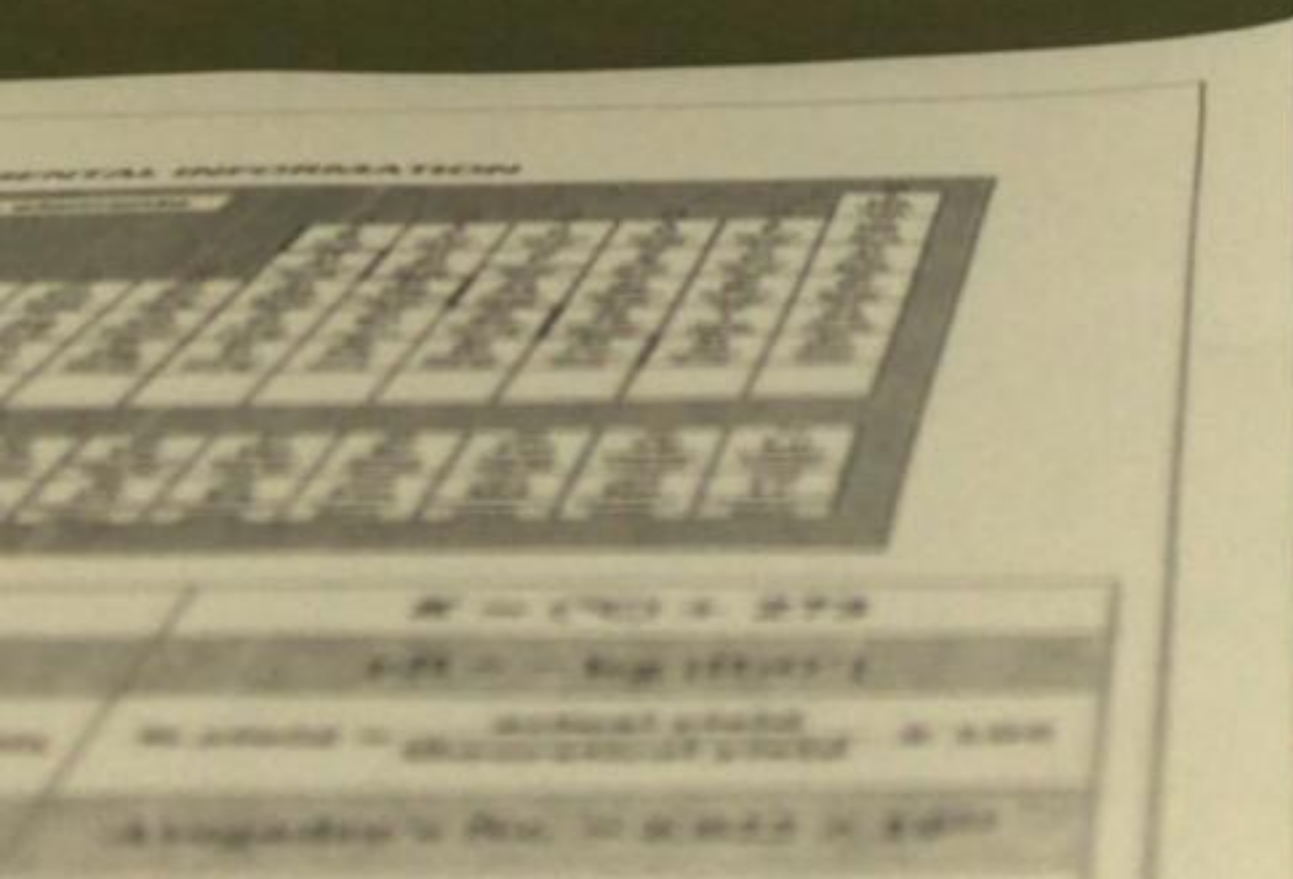
Question No.	Answer	Mark
1	14	4
2		
3		
4		
5		
6		
7		
8		
9		
10		

An iodine (I) atom has a mass number, $A = 131$.
How many protons, neutrons, and electrons does it have?

- 53 protons, 131 neutrons, 54 electrons.
- 53 protons, 78 neutrons, 53 electrons.
- 39 protons, 78 neutrons, 39 electrons.
- 53 protons, 78 neutrons, 54 electrons.



Save & Next



Liquids have a mass, a volume, and a shape.

- constant, fixed, variable
- constant, fixed, definite
- constant, variable, definite
- variable, variable, definite

A

Save & Next

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TABLE 202 SUPPLEMENTAL INFORMATION

$^{\circ}\text{C} = \text{K} - 273$	$\text{K} = ^{\circ}\text{C} + 273$
$\text{pH} = -\log [\text{H}^+]$	$\text{pH} = -\log [\text{H}^+]$
$\text{m} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$	$\text{m} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$
Avogadro's No. = 6.022×10^{23}	Avogadro's No. = 6.022×10^{23}

$\text{m} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

Avogadro's No. = 6.022×10^{23}

If the density chloroform is 1.492 g/mL, what is the mass of 225 mL?

- 226.49 g
- 150.80 g
- 335.70 g
- 6.63×10^{-3} g



Save & Next

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TABLE FOR SUPPLEMENTAL INFORMATION

A FOR ANSWERING EACH OF THE QUESTIONS

Q. No.	1	2	3	4	5	6	7	8	9	10
Answer										

$^{\circ}\text{C} = \text{K} - 273$ $\text{K} = (^{\circ}\text{C}) + 273$

Total questions in exam: 25 | Answered: 6

Question No. 3

If the density chloroform is 1.492 g/mL, what is the mass of 225 mL?

- 6.63×10^{-3} g
- 150.80 g
- 335.70 g
- 226.49 g

Total questions in exam: 25 | Answered: 6

Question No. 9

A specific isotope of an element is known to have 15 protons and 16 neutrons. Which symbol would properly represent this isotope?

- $^{31}_{15}\text{Ga}$
- $^{31}_{15}\text{P}$
- $^{31}_{16}\text{S}$
- $^{31}_{15}\text{P}$

B

Save & Next



30

Question No. 8

Which of the following statements is TRUE about the state of matter?

- Solids and liquids have fixed volume and rigid shape.
- Gases are not compressible.
- Matter changes from one state to another by changing the temperature.
- Table salt and diamond are examples of amorphous solids.

Save & Next

Question No. 13

Which of the following are isotopes for each other?

 ${}_{19}^{41}\text{A}$, ${}_{19}^{39}\text{B}$, ${}_{20}^{41}\text{C}$ and ${}_{18}^{39}\text{D}$

- A and C
- A and B
- B and D
- C and D

B

Question No. 16

"Only two electrons, with opposing spins, are allowed in each orbital" is known as the

- Heisenberg uncertainty principle.
- Hund's rule.
- Aufbau principle.
- Pauli exclusion principle.

D

Question No. 17

All of the following statements about different elements are true, EXCEPT

- Chlorine is a metalloid.
- Iron is a transition metal.
- Neon is a noble gas.
- Calcium is an alkaline earth metal.



Save & Next

ETON

Question No. 18

Which of the following can be considered as a physical change?

- Milk sours
- Hardness
- Ice melts
- Toxicity

B

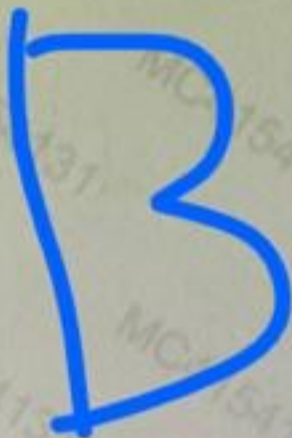
Save & Next

Total questions in exam: 25 | Answered: 17

Question No. 20

The temperature unit in the metric system is

- Fahrenheit
- Celsius
- Pound
- Kelvin



Save & Next

Question No. 19

Which of the following statements is true?

- Energy cannot be converted between potential and kinetic energies.
- Solids have the greatest kinetic energy and gases have the lowest kinetic energy.
- As we increase the temperature of a substance, its kinetic energy decreases.
- All substances have both potential and kinetic energies, regardless to their physical states.

Save & Next

Question No. 19

Which of the following statements is true?

D

- Energy cannot be converted between potential and kinetic energies.
- Solids have the greatest kinetic energy and gases have the lowest kinetic energy.
- As we increase the temperature of a substance, its kinetic energy decreases.
- All substances have both potential and kinetic energies, regardless to their physical states.

Save & Next

ETON

hp

Which of the following determines the identity of an atom?

- The number of neutrons.
- The number of electrons.
- The number of protons.
- The total number of protons and neutrons.



Save & Next

Total questions in exam: 25 | Answered: 7

Question No. 3

Which one of the following species has same number of electrons as Mg^{2+} cation?

- K⁺
- Cl⁻
- S²⁻
- F⁻

D

Save & Next

Question No. 22

Toluene boils at 231.08 °F. This temperature equals _____ K

- 110.6
- 383.75
- 417.94
- 504.23

B

Question No. 14

The density of carbon tetrachloride is 1.59 g/mL. What is the mass of 3.65 mL of carbon tetrachloride?

- 0.43 g
- 2.29 g
- 5.80 g
- 5.24 g



Question No. 17

All of the following statements about different elements are true, EXCEPT

- Chlorine is a metalloid.
- Iron is a transition metal.
- Neon is a noble gas.
- Calcium is an alkaline earth metal.

A

Save & Next

ETON