

Total questions in exam: 25 | Answered: 1

Question No. 1

Of the elements: Li, B, C, and F, the element with the most metallic character is \_\_\_\_\_

- F
- Li
- C
- B

B

Save & Next

Total questions in exam: 25 | Answered: 1

## Question No. 2

In which of these substances are the atoms held together by polar covalent bonding?

- ClF
- S<sub>8</sub>
- CsCl
- SrCl<sub>2</sub>

A

Save &amp; Next

Question No. 3

A compound that has a molar mass of 60 g/mol and an empirical formula of  $\text{CH}_2\text{O}$ ,  
molecular formula is:

- $\text{C}_2\text{H}_4\text{O}$
- $\text{C}_3\text{H}_6\text{O}_3$
- $\text{C}_2\text{H}_4\text{O}_2$
- $\text{CH}_2\text{O}$



Save & ...



Total questions

25

Answered 0

Question No. 8

What is the mass% of carbon in  $(C_2H_4O)_n$ ?

- 38.7%
- 7.74%
- 20.8%
- 60%

Handwritten scribbles and a large arrow pointing towards the top right corner of the screen.

Total questions in exam: 25 | Answered: 4

Q. 5

The atomic radius of krypton is smaller than the atomic radius of \_\_\_\_\_

- xenon
- neon
- helium
- argon

A

Total questions in exam: 25 | Answered: 4

## Question No. 6

What is the percent yield for a reaction if its theoretical yield is 130 g and its actual yield is 74 g?

- 80%
- 64%
- 57%
- 53%



Total Questions in Exam: 25 | Answered: 4

Question No. 7

What is the coefficient of oxygen gas after balancing the following equation?



- 3
- 5
- 2
- 1

A

Total questions in exam: 25 | Answered: 4

What is the mass% of carbon in  $(C_2H_6O_2)$ ?

- 38.7%
- 7.74%
- 20.6%
- 60%

A

Save &amp; Next



Total questions in exam: 25 | Answered: 4

The chemical formula of the compound formed between magnesium and chlorine is

- $Mg_3Cl$
- $Mg_2Cl$
- $MgCl_2$
- $MgCl$



Total questions in exam: 25 | Answered: 4

What is the final molarity of  $\text{HNO}_3$  solution, if 300 mL of 2M  $\text{HNO}_3$  was diluted to a final volume of 0.2 L?

- 5.0 M
- 6.0 M
- 4.0 M
- 3.0 M



Next

Total questions in exam: 20 | Answered: 10

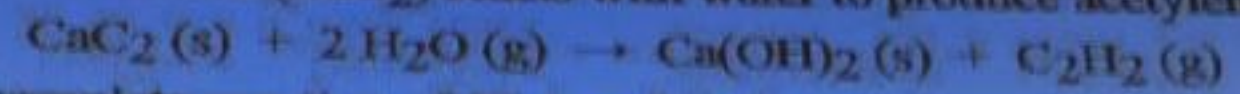
Question 10

A strong electrolyte is a substance that \_\_\_\_\_ in aqueous solutions.

- partially ionizes
- completely ionizes
- does not ion
- reacts

B

Calcium carbide ( $\text{CaC}_2$ ) reacts with water to produce acetylene ( $\text{C}_2\text{H}_2$ ):



The complete reaction of 57.4 g of  $\text{CaC}_2$  requires consumption of \_\_\_\_\_ g of  $\text{H}_2\text{O}$ .

- 32.0
- 1.79
- 64.1
- 0.895

32.0

A

10

What is the term for the shared valence electrons in a covalent molecule?

- lone pair electrons
- nonbonding electro
- core electrons
- bonding electrons



Question No. 1

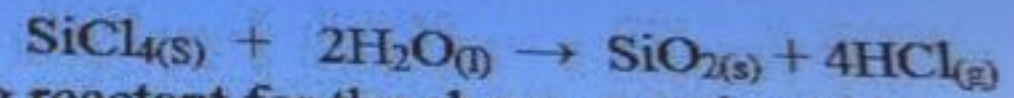
Select the element whose Lewis dot symbol is correct

- Te
- He
- Rn
- Fr



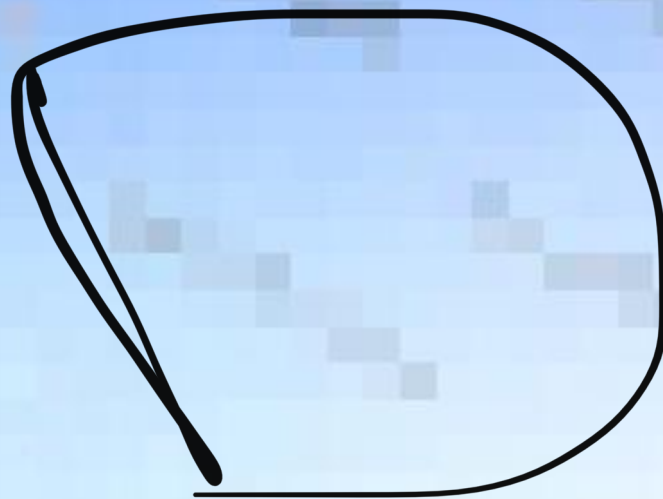
Question

In an experiment, 50.0 g of silicon tetrachloride ( $\text{SiCl}_4$ ) is treated with 20.0 g of water to produce silicon dioxide ( $\text{SiO}_2$ ) according to the following balanced equation:



The limiting reactant for the above experiment is:

- $\text{SiO}_2$
- $\text{HCl}$
- $\text{H}_2\text{O}$
- $\text{SiCl}_4$



Total questions in exam: 25 | Answered: 10

Question No. 16

What is the oxidation number of sulfur in  $\text{SO}_3^{2-}$ ?

- +4
- +6
- +2
- 2

D



Fluorine is considered \_\_\_\_\_

- a molecular compound
- an ionic compound
- a molecular element
- an atomic element



Q11

What is the molarity of  $\text{FeCl}_3$  in a solution prepared by dissolving 10.0 g of  $\text{FeCl}_3$  in enough water to make 275 mL of solution?

- 4.46 M
- $4.46 \times 10^3$  M
- $2.24 \times 10^{-4}$  M
- 0.224 M



If 0.274 moles of a substance weighs 62.5 g, what is the molar mass of the substance, in units of g/mol?

- $4.38 \times 10^{-3} \text{ g/mol}$
- $2.17 \times 10^2 \text{ g/mol}$
- $2.28 \times 10^2 \text{ g/mol}$
- $1.71 \times 10^1 \text{ g/mol}$

QV

If 0.274 moles of a substance weighs 62.5 g, what is the molar mass of the substance, in units of g/mol?

- $4.38 \times 10^{-3} \text{ g/mol}$
- $2.17 \times 10^2 \text{ g/mol}$
- $2.28 \times 10^2 \text{ g/mol}$
- $1.71 \times 10^1 \text{ g/mol}$



Total questions in exam: 25 | Answered: 13

Q

The coefficients (a,b,c,d) needed to balance the equation below are:



- (2,2,1,1)
- (1,1,2,2)
- (1,2,1,1)
- (2,1,2,1)

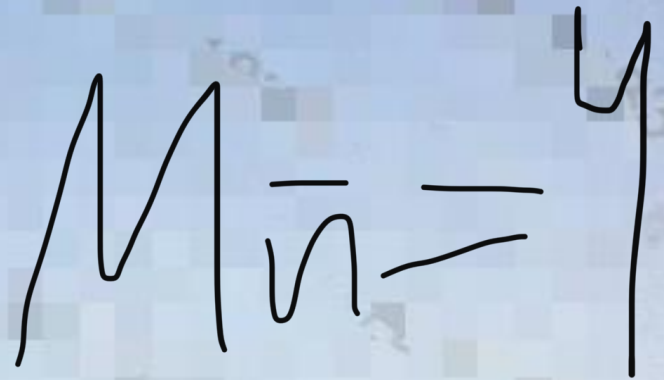
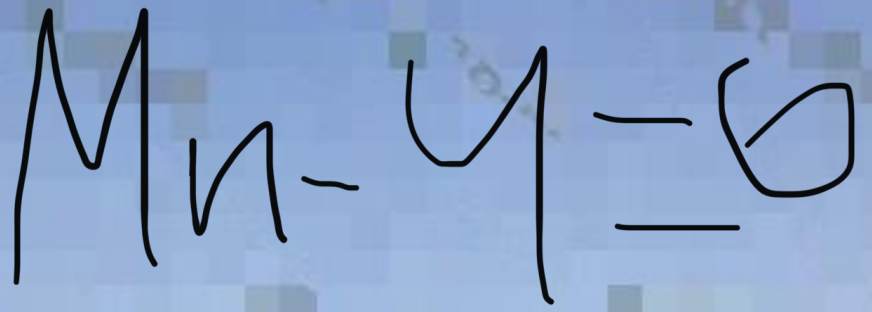


Q1

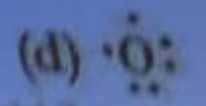
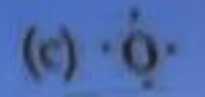
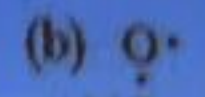
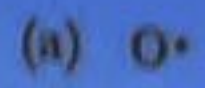
What is the oxidation number of manganese in  $\text{MnO}_2$ ?

 +2 +4 +1 0

B



Which of the following is the electron dot formula (Lewis structure) for an atom of oxygen?



- (c)
- (d)
- (b)
- (a)

B

Q. 23

Two mole of any substance contains \_\_\_\_\_ particles?

- $1.20 \times 10^{24}$
- $3.011 \times 10^{24}$
- $12.044 \times 10^{24}$
- $6.022 \times 10^{23}$

A

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2x



Total questions in exam: 25 | Answered: 18

Question

The name of the chemical compound  $\text{CuSO}_4$  is:

- copper(I) sulfate
- copper sulfate
- copper(II) sulfate
- copper(III) sulfate



Total questions in exam: 25 | Answered: 18

Question No. \_\_\_\_\_

The correct name for the compound  $\text{CCl}_4$  is:

- carbon dichloride
- carbon tetrachloride
- carbon trichloride
- carbon chloride

B

In an experiment, 50.0 g of silicon tetrachloride ( $\text{SiCl}_4$ ) is treated with 20.0 g of water to produce silicon dioxide ( $\text{SiO}_2$ ) according to the following balanced equation.



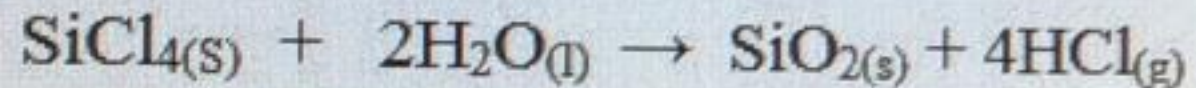
The limiting reactant for the above experiment is:

- $\text{SiO}_2$
- $\text{HCl}$
- $\text{H}_2\text{O}$
- $\text{SiCl}_4$

Total questions in exam: 25 | Answered: 23

Q1

In an experiment, 50.0 g of silicon tetrachloride ( $\text{SiCl}_4$ ) is treated with 20.0 g of water to produce silicon dioxide ( $\text{SiO}_2$ ) according to the following balanced equation:



The limiting reactant for the above experiment is:

- $\text{SiO}_2$
- $\text{HCl}$
- $\text{H}_2\text{O}$
- $\text{SiCl}_4$



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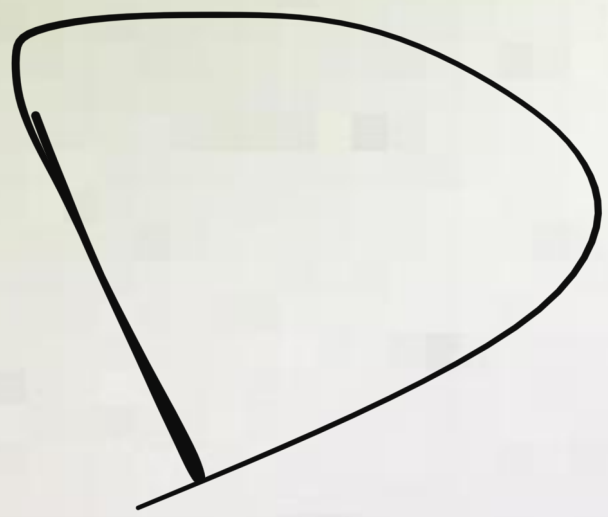


MKCL OES  
Total questions in exam: 25 | Answered: 3

Question

How many moles of  $(\text{NH}_4)_2\text{S}$  are there in 75 % of  $(\text{NH}_4)_2\text{S}$ ?

- 1.0
- 3.4
- 7.5
- 1.1





21

The coefficients (a,b,c,d) needed to balance the equation below are:  
 $a \text{FeS} + b \text{O}_2 \rightarrow c \text{Fe}_2\text{O}_3 + d \text{SO}_2$

- (4,7,4,2)
- (4,4,7,2)
- (7,4,2,4)
- (4,7,2,4)



MCQ QES

Total questions in exam: 25 | Answered: 2

Question No. 20

Select the element whose Lewis dot symbol is correct.

- $\cdot \cdot \cdot \cdot$
- $\cdot \cdot \cdot \cdot$
- $\cdot \cdot \cdot \cdot$
- $\cdot \cdot \cdot \cdot$



Que. p. 21

The coefficients (a,b,c,d) needed to balance the equation below are:

$$a \text{FeS} + b \text{O}_2 \rightarrow c \text{Fe}_2\text{O}_3 + d \text{SO}_2$$

- (4,7,4,2)
- (4,4,7,2)
- (7,4,2,4)
- (4,7,2,4)





Total questions: 25 / Answered: 0

Question No. 2

A chemical reaction produces 7.25 moles of barium sulfate,  $\text{BaSO}_4$ . What mass (in grams) of barium sulfate is produced?

- 185 g
- 31.1 g
- $1.69 \times 10^3$  g
- 233 g



NaCl

What is the molarity of a solution made by dissolving 25.00 g of NaCl in enough water to

- 0.479 M
- 0.526 M
- 0.684 M
- 0.308 M

A large, handwritten scribble in black ink, consisting of several overlapping loops and lines, is drawn over the right side of the page, partially obscuring the text and options.

Total questions in exam: 25 | Answered: 7

The most correct name for the compound  $P_4S_{10}$  is:

- tetraphosphorus nonasulfide
- triphosphorus decasulfide
- tetraphosphorus octasulfide
- tetraphosphorus decasulfide

D

Question No

The charge of hydroxide ion is

- 1-
- 2+
- 1+
- 2-



OH<sup>-</sup>

Total questions in exam: 25 | Answered: 5

Question No. 2

What is the final molarity of KOH solution, if 0.064 L of 3.4 M KOH was diluted to of 0.4 L?

- 0.61 M
- 0.54 M
- 0.68 M
- 0.27 M

A large, hand-drawn black outline of the letter 'B' is positioned in the center of the page, overlapping the multiple-choice options.



Total questions in exam: 25 | Answered: 7

Question No

Which of the following elements does NOT exist as a diatomic molecule?

- carbon
- nitrogen
- oxygen
- hydrogen

A



Chemistry\_Quiz2\_Sem1\_2019

MKCL OES

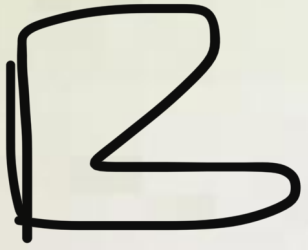
Total questions in exam: 25 | Answered: 0

A<sup>-</sup> A A<sup>+</sup>

Question No. 25

To prepare 25 mL of 0.25 M  $\text{MgCl}_2$  solution from a 0.75 M  $\text{MgCl}_2$  stock solution, you should mix

- 12 mL of stock solution mixed with enough water to make 25 mL
- 8.3 mL of stock solution mixed with enough water to make 25 mL
- 9.6 mL of stock solution mixed with any amount of water
- 11.3 mL of stock solution added to 25 mL water



Ques...

Which of the following elements has the greatest ability to lose electrons?

- Francium
- Sodium
- Lithium
- Potassium

Handwritten scribbles and a large curved line.



Question N

What is the final molarity of  $\text{KH}_2\text{PO}_4$  solution, if 25 mL of 4M  $\text{KH}_2\text{PO}_4$  was diluted to a final volume of 500 mL?

- 0.6 M
- 0.4 M
- 1.8 M
- 0.2 M



Number of main questions  
Number of questions 25



Answered

15

Not Visited

1

2

3

4

5

6

7

8

9

10

11

12

13

14

15

16

17

18

19

20

21

22

23

24

25

Calculate the mass percent composition of sulfur in  $\text{Cr}_2(\text{SO}_4)_3$ .

- 44.9%
- 5.4%
- 32.4%
- 22.1%

Number of math questions  
Number of questions

Answered  
 Not Answered

1	2	3
4	5	6
7	8	9
10	11	12

Question 1

Which of the following compounds gives a strong electrolyte aqueous solution?

- $C_6H_{12}O_6$
- $NH_3$
- $Na_2CO_3$
- $CCl_4$



Number of main q  
Number of question

6 Answered

15 Not Visited

1	2	3
6	9	10
15	16	17
22	23	24

What is the molecular formula of a compound that has a molar mass of 30 g/mol and its empirical formula is  $\text{CH}_2\text{O}$ ?

- $\text{C}_4\text{H}_8\text{O}_4$
- $\text{C}_3\text{H}_6\text{O}_3$
- $\text{C}_2\text{H}_4\text{O}_2$
- $\text{CH}_2\text{O}$

D

Number of m  
Number of q

23 Answered

0 Not Visited

1	2	
8	9	
15	16	
22	23	

Question

Calculate the mass percent composition of sulfur in  $\text{Ga}_2(\text{SO}_4)_3$ .

- 44.9 %
- 9.4 %
- 32.6 %
- 22.5 %

D

Number of  
Number of

21

Answer

0

Not Vis

1

2

8

9

15

16

22

23

Question No. 1

The atomic size (radius) of atoms \_\_\_\_\_

- decreases going down within a group.
- increases going across a period.
- does not change going down with in a group.
- increases going down within a group.

D

Which of the following substances contains a nonpolar covalent bond?

- $\text{NH}_3$
- $\text{NaCl}$
- $\text{N}_2$
- $\text{H}_3\text{O}^+$



Number of ma  
Number of que

21

Answered

0

Not Visited

1

2

3

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16

17

22

23

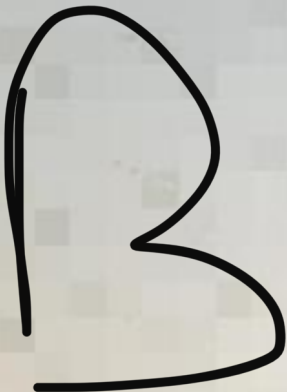
24

Total questions in exam: 25 (Answered: 1)

Question

If 4 moles of an element are weighing 92 g, this element is \_\_\_\_\_

- Ca
- Na
- Fe
- S





Total questions in exam: 25 / 10

Question

The most correct name for the compound  $\text{CS}_2$  is:

- carbon trisulfide
- monocarbon disulfide
- carbon disulfide
- carbon sulfide



Total questions in exam: 25 | Answered: 4

Ques...

A A A

Household sugar, sucrose, has the molecular formula  $C_{12}H_{22}O_{11}$ . What is the mass percent of carbon in sucrose?

- 51.4 %
- 6.5 %
- 62.8 %
- 42.1 %

?

.

Q.

What substance is the reducing agent in the following redox reaction?



- Cu<sup>2+</sup>
- Cu
- Zn
- Zn<sup>2+</sup>



HWCE OES

Total questions in exam: 28 | Answered: 4

Question 1

The molarity (M) of an aqueous solution containing 22.5 g of sucrose ( $C_{12}H_{22}O_{11}$ ) in 35.5 mL of solution is \_\_\_\_\_

- 1.85
- 0.0657
- 0.104
- 3.52

السؤال

$$M = \frac{m}{V(L)}$$

Question No. 3

The number which is located on the LEFT of a chemical formula that helps to balance a chemical equation is called \_\_\_\_\_

- exponent
- superscript
- coefficient
- subscript



Question No. 11

Fluorine is considered \_\_\_\_\_


- a molecular element
- an atomic element
- an ionic compound
- a molecular compound

A

Total questions in exam: 25 | Answered: 5

Question No. 14

What is the term for the pairs of valence electrons that are not shared in a molecule?

- core electrons
  - bonding electrons
  - lone pairs of electrons
  - sharing electrons
- 

What substance is the reducing agent in the following redox reaction?



- $\text{Cu}^{2+}$
- $\text{Cu}$
- $\text{Zn}$
- $\text{Zn}^{2+}$





The most correct name for the compound  $\text{CS}_2$  is:

- carbon trisulfide
- monocarbon disulfide
- carbon disulfide
- carbon sulfide

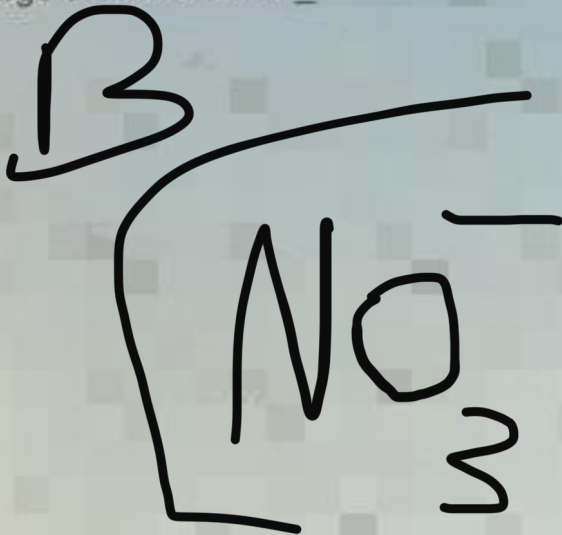


Total questions in exam: 25 | Answered: 5

Question ...

The charge of nitrate ion is \_

- 1+
- 1-
- 2+
- 2-



Total questions in exam: 26 | Answered: 5

Question No. \_\_\_\_\_

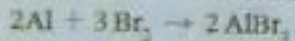
What is the term for the pairs of valence electrons that are not shared in a molecule?

- core electrons
- bonding electrons
- lone pairs of electrons
- sharing electrons



c

In the chemical reaction below: what is oxidized and what is reduced in the following reaction?



- AlBr<sub>3</sub> is reduced and Br<sub>2</sub> is oxidized.
- Al is reduced and Br<sub>2</sub> is oxidized.
- AlBr<sub>3</sub> is reduced and Al is oxidized.
- Al is oxidized and Br<sub>2</sub> is reduced.

D

## Question

What is the mass percentage of O in  $C_{14}H_{20}N_2SO_4$  if the molar mass of the compound is 312.4 g/mole.

- 6.45%
- 53.81%
- 20.49%
- 10.27%



Q

What is the mass percentage of O in  $C_{11}H_{20}N_2SO_4$  if the molar mass of the compound is 312.4 g/mole.

- 6.45%
- 53.81%
- 20.49%
- 10.27%

Question no.

In the chemical reaction below: what is oxidized and what is reduced in the following reaction?



- $\text{AlBr}_3$  is reduced and  $\text{Br}_2$  is oxidized.
- $\text{Al}$  is reduced and  $\text{Br}_2$  is oxidized.
- $\text{AlBr}_3$  is reduced and  $\text{Al}$  is oxidized.
- $\text{Al}$  is oxidized and  $\text{Br}_2$  is reduced.

D

Total questions in exam 28 | Answered 8

QUESTION 10

Which of the following substances gives a strong electrolyte?

- $N_2$
- $H_2O$
- $MgCl_2$
- $CH_4$







Total questions in exam 25 | Answered 5

Q

What is the final molarity of  $\text{HNO}_3$  solution, if 95 mL of 2M  $\text{HNO}_3$  was diluted to a final volume of 0.3 L?

- 0.43 M
- 0.42 M
- 0.63 M
- 0.53 M



Question

Which of the following represents the correct Lewis dot structure for oxygen molecule?



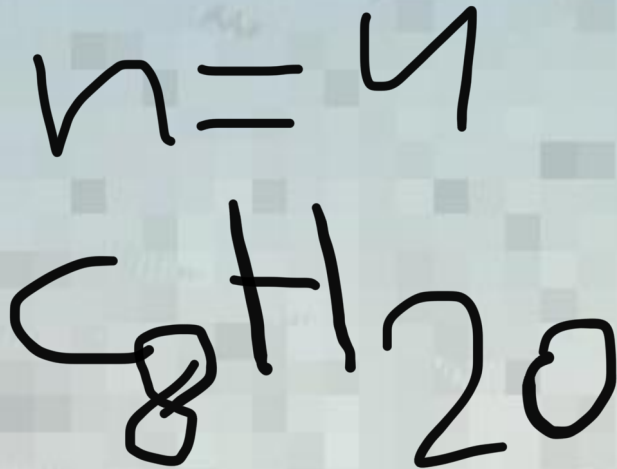
- (b)
- (a)
- (d)
- (c)



Question no. 77

What is the molecular formula of a compound that has a molar mass of 116 g/mol and its empirical formula is  $C_2H_5$ ?

- $C_4H_{10}$
- $C_2H_5$
- $C_3H_7.5$
- $C_8H_{20}$



Total questions in exam: 25 | Answered: 8

Question:

Which of these atoms is the most electronegative?

- Al
- Mg
- Na
- P

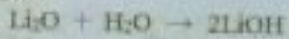


ink

Total questions in exam: 25 | Answered: 5

Q2

In the reaction below, what is the theoretical yield in moles for LiOH when 3 grams of  $\text{Li}_2\text{O}$  react with 4 grams of  $\text{H}_2\text{O}$ ?



- 0.4 mol
- 0.2 mol
- 0.8 mol
- 0.6 mol

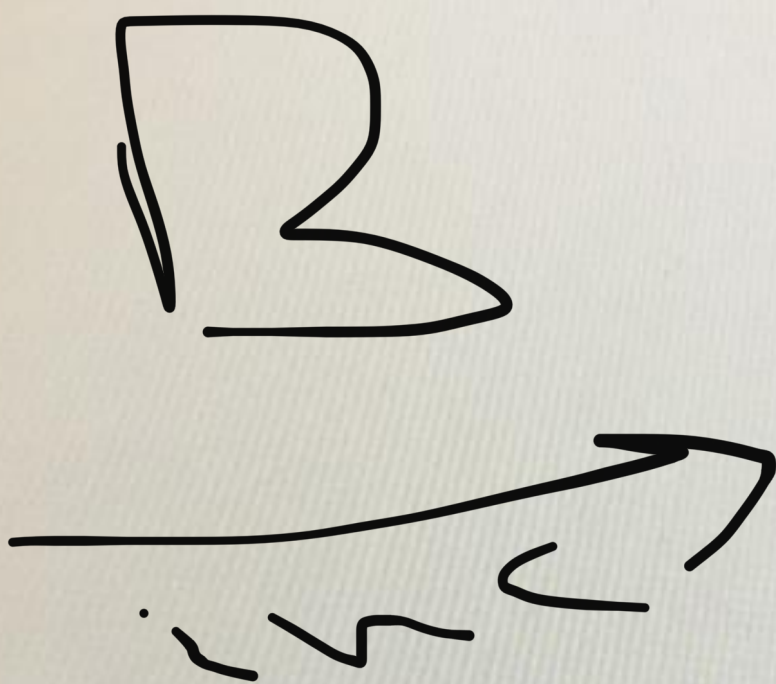
B

Total questions in exam: 25 | Answered: 20

## Question No. 18

Which of the following elements has the highest electronegativity?

- Li
- N
- Be
- C

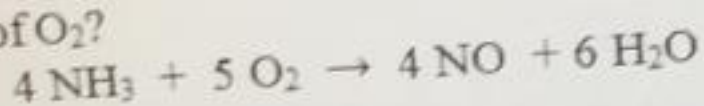


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Question No. 19

In the reaction below, what is the theoretical yield in moles for NO when 3 moles of  $\text{NH}_3$  react with 3 moles of  $\text{O}_2$ ?



- 2.6 mol
- 2.8 mol
- 2.4 mol
- 3.0 mol

[Save & Next](#)

Total questions in exam: 25 | Answered: 20

Question No. 20

Select the element whose Lewis dot symbol is correct.

- Ra•
- :Te:
- Fr•
- He•

A

Save & Next

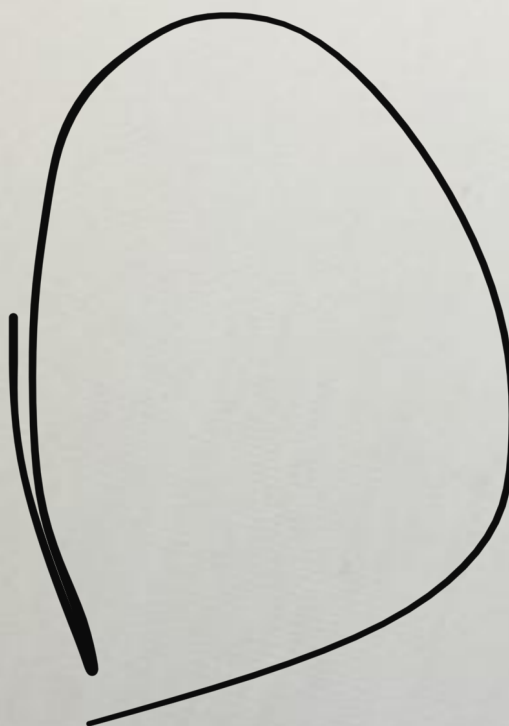


Total questions in exam: 25 | Answered: 20

Question No. 22

Oxidation cannot occur without \_\_\_\_\_

- air
- water
- acid
- reduction



Question No. 23

The correct name for the compound CO is \_\_\_\_\_

- carbon oxide
- carbon monoxide
- monocarbon monoxide
- carbon dioxide

B



Total questions in exam: 25 | Answered: 0

Question No. 7

A<sup>-</sup>

A

Which of the following compounds gives a nonelectrolyte aqueous solution?

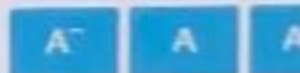
- $\text{Na}_2\text{CO}_3$
- $\text{CCl}_4$
- $\text{HI}$
- $\text{HBr}$

ب  
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All questions in exam: 25 | Answered: 0

Question No. 5



the reaction:



$\text{Cu}_{(s)}$  is the oxidizing agent and  $\text{Ag}^+_{(aq)}$  is oxidized.

$\text{Cu}_{(s)}$  is the reducing agent and  $\text{Ag}^+_{(aq)}$  is reduced.

$\text{Ag}^+_{(aq)}$  is oxidizing agent and  $\text{Cu}_{(s)}$  is reduced

$\text{Ag}^+_{(aq)}$  is the reducing agent and  $\text{Cu}_{(s)}$  is reduced.



Questions in exam: 25 | Answered: 0

Question No. 4

A<sup>-</sup> A A

Calculate the mass percent composition of oxygen in  $\text{Ga}_2(\text{SO}_4)_3$ .

- 1.9%
- 2.6%
- 37%
- 2.5%

A  
37.01%

Save & Next



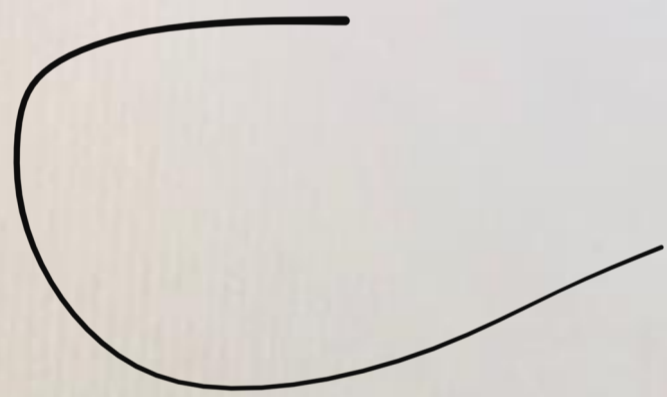
Total questions in exam: 25 | Answered: 0

Question No. 6



The ionization energy of silicon is lower than the ionization energy of \_\_\_\_\_.

- sodium
- aluminum
- phosphorus
- magnesium





Total questions in exam: 25 | Answered: 0

Question No. 7

A<sup>-</sup>

A

Which of the following compounds gives a nonelectrolyte aqueous solution?

- $\text{Na}_2\text{CO}_3$
- $\text{CCl}_4$
- $\text{HI}$
- $\text{HBr}$

MS



Total questions in exam: 25 | Answered: 0

Question No. 9

A<sup>-</sup>

A

The molarity (M) of an aqueous solution containing 52.5 g of sucrose ( $C_{12}H_{22}O_{11}$ ) in 35.5 mL of solution is \_\_\_\_\_.

- 0.104
- 4.32
- 1.48
- 5.46

B



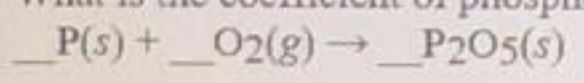


Total questions in exam: 25 | Answered: 0

Question No. 8

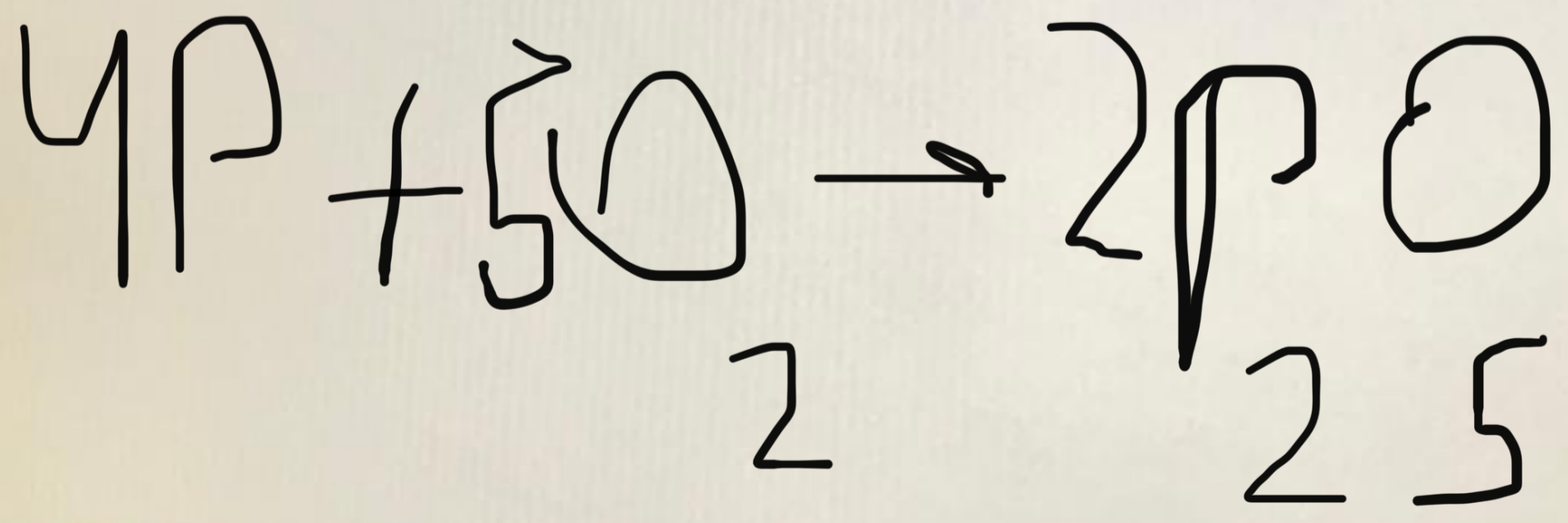
A<sup>-</sup> A

What is the coefficient of phosphorus after balancing the following equation?



- 4
- 5
- 2
- 1

A



Save & Next

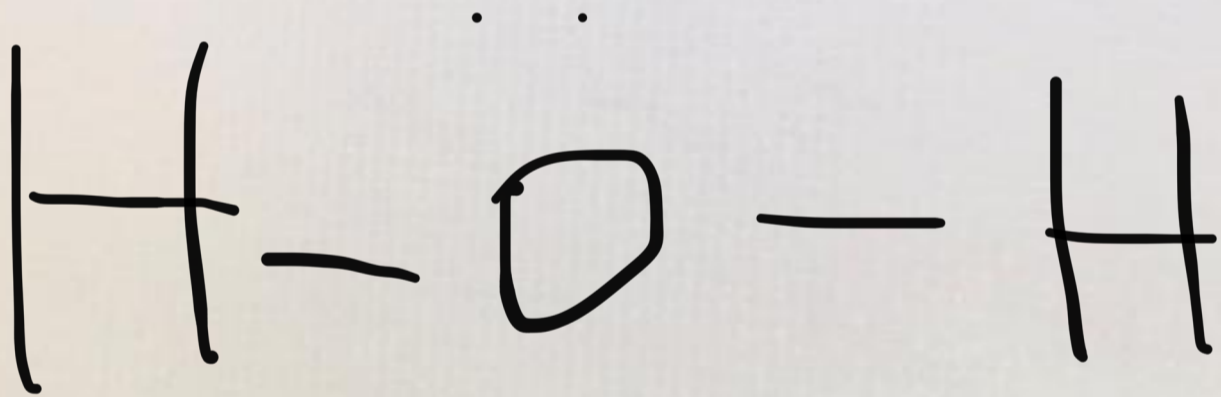
Total questions in exam: 25 | Answered: 0

Question No. 11

A<sup>-</sup>

Based on Lewis structures, the number of lone pairs of electrons in the water molecule is \_\_\_\_\_

- 3
- 2
- 4
- 8



B



Total questions in exam: 25 | Answered: 0

Question No. 12

In a solution, the solute is \_\_\_\_\_.

- the substance present in the greatest amount
- always water
- always a solid
- the substance present in the smallest amount

D

~~the substance present in the smallest amount~~

Save & Next

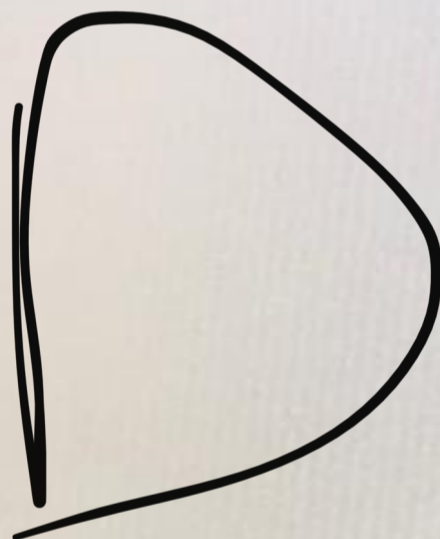
Total questions in exam: 25 | Answered: 0

Question No. 10

A

What is the percent yield for a reaction if its theoretical yield is 144 g and its actual yield is 46 g?

- 53%
- 22%
- 60%
- 32%



Total questions in exam: 25 | Answered: 0

Question No. 14

A<sup>-</sup>

A

Calculate the molar mass of  $\text{CaCO}_3$ .

- 100 g/mol
- 120 g/mol
- 50 g/mol
- 87 g/mol

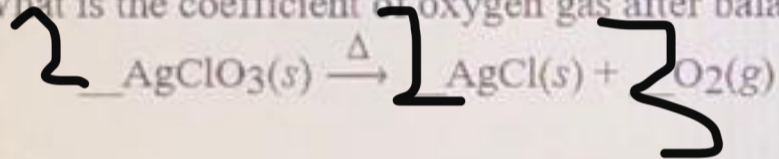
A

Total questions in exam: 25 | Answered: 0

Question No. 12

A

What is the coefficient of oxygen gas after balancing the following equation?

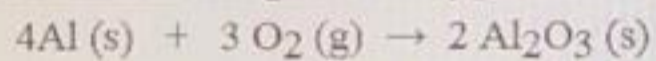


- 3
- 4
- 2
- 1

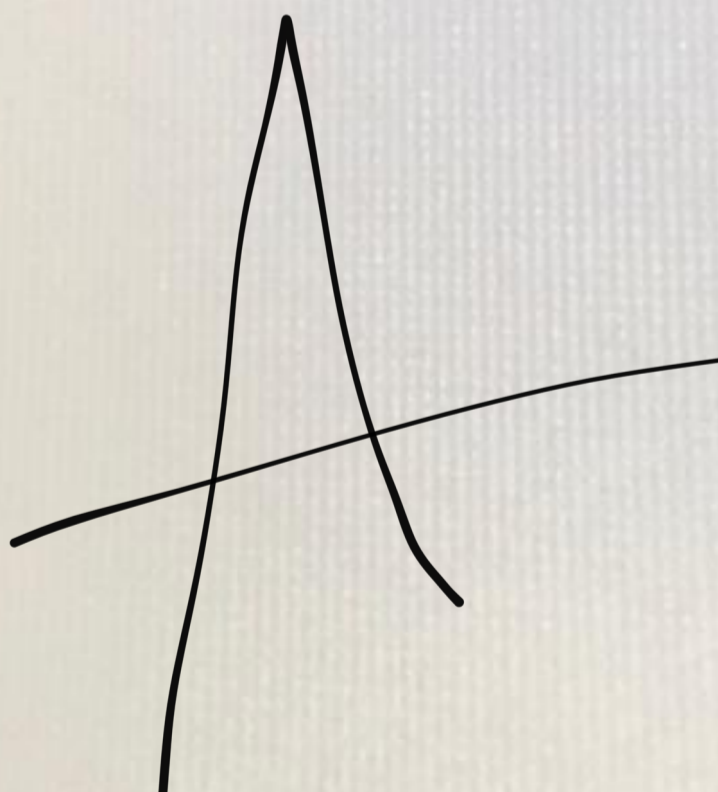
A

Total questions in exam: 25 | Answered: 0

Question No. 15

Solid aluminum and gaseous oxygen react in a combination reaction to produce  $\text{Al}_2\text{O}_3$ The maximum amount of  $\text{Al}_2\text{O}_3$  that can be produced from 2.5 g of Al and 2.5 g of  $\text{O}_2$  is \_\_\_\_\_ g.

- 4.7
- 5.3
- 7.4
- 9.4



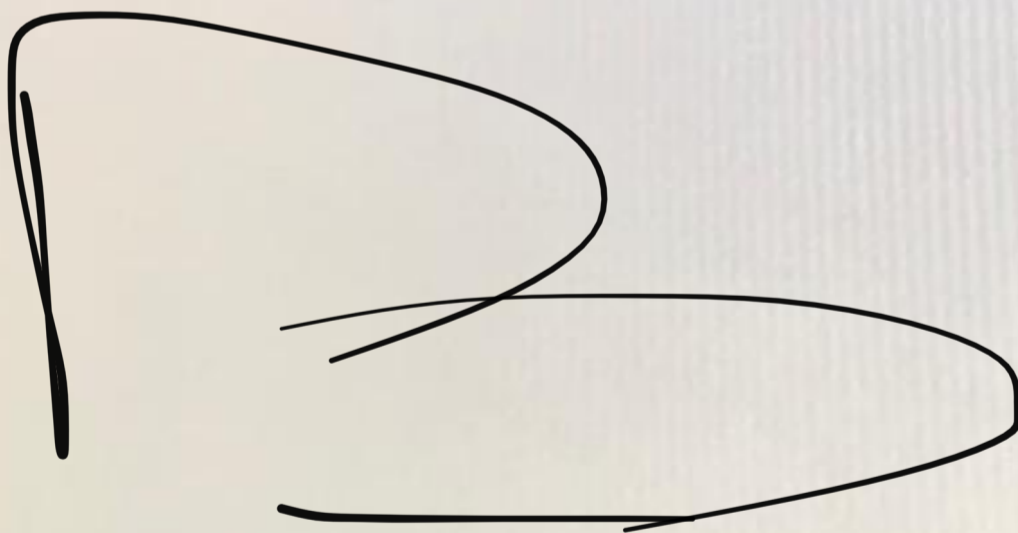
Total questions in exam: 25 | Answered: 0

Question No. 13

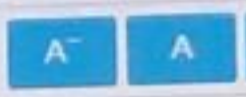
A<sup>-</sup>

Which of the following statements is FALSE regarding an ionic bond?

- Breaking an ionic bond between two ions requires energy.
- The ionic bond is formed between two nonmetal atoms.
- An ionic bond is the attraction between a metal ion and a nonmetal ion.
- Metal atoms lose electrons and nonmetal atoms gain electrons.







Question No. 19

A chemical reaction produces 7.25 moles of barium sulfate,  $\text{BaSO}_4$ . What mass (in grams) of barium sulfate is produced?

- 31.1 g
- $1.69 \times 10^3$  g
- 233 g
- 185 g

B

Total questions in exam: 25 | Answered: 0

Question No. 16

A

Analysis of an unknown substance showed that it has a high boiling point and is brittle. It is an insulator as a solid but conducts electricity when melted. Which of the following substances would have those characteristics?

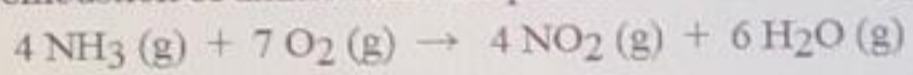
- CO
- KBr
- Al
- HCl

A large, handwritten number '16' in black ink, indicating the question number.

Total questions in exam: 25 | Answered: 0

Question No. 17

The combustion of ammonia in the presence of excess oxygen yields  $\text{NO}_2$  and  $\text{H}_2\text{O}$ :



The combustion of 43.9 g of ammonia produces \_\_\_\_\_ g of  $\text{NO}_2$ .

- 119
- 43.9
- 178
- 2.58

A large, handwritten black letter 'B' is drawn on the screen, positioned to the right of the multiple-choice options.

Total questions in exam: 25 | Answered: 0

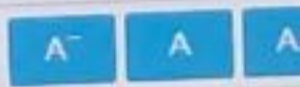
Question No. 18

A

Which of the following represents the trend of the electronegativity in the periodic table?

- decreases from left to right, increases from bottom to top
- increases from left to right, increases from bottom to top
- increases from left to right, decreases from bottom to top
- decreases from left to right, decreases from bottom to top

A



Question No. 22

The oxidation number of iodine in  $\text{KIO}_4$  is

- 7
- +7
- 1
- +1

B

$$\text{I} + 1 + 4(-2) = 0$$

$$\text{I} + 1 - 8 = 0$$

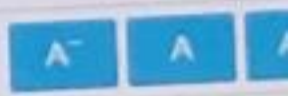
$$\text{I} = +7$$



Chemistry\_Quiz2\_Sem1\_2019

KCL OES

Total questions in exam: 25 | Answered: 0



Question No. 24

The name of the chemical compound CuOH is \_\_\_\_\_

- copper(III) hydroxide
- copper(II) hydroxide
- copper hydroxide
- copper(I) hydroxide

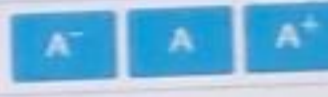
$Cu - 2 + 1$   
 $Cu = 2 - 1$   
 $Cu = 1$   
**D**



# Chemistry\_Quiz2\_Sem1\_2019

CL OES

al questions in exam: 25 | Answered: 0



Question No. 20

The correct name for the compound CO is \_\_\_\_\_

- carbon monoxide
- carbon dioxide
- carbon oxide
- monocarbon monoxide

A



## Chemistry\_Quiz2\_Sem1\_2019

MKCL OES  
Online Evaluation System

Total questions in exam: 25 | Answered: 0

A

A

Question No. 23

One of the following elements exists as an atomic element:

- oxygen
- neon
- hydrogen
- chlorine

B





Chemistry\_Quiz2\_Sem1\_2019

CL OES

al questions in exam: 25 | Answered: 0

A<sup>-</sup>

A

A<sup>+</sup>

Question No. 21

How many moles of  $(\text{NH}_4)_2\text{S}$  are there in 75 g of  $(\text{NH}_4)_2\text{S}$ ?

3.4

7.5

1.1

1.9

C

Total questions in exam: 25 | Answered: 0

## Question No. 25

What is the final molarity of NaOH solution, if 0.028 L of 4 M NaOH was diluted to a final volume of 0.32 L?

- 0.16 M
- 0.08 M
- 0.35 M
- 0.25 M



Total questions in exam: 25 | Answered: 0

Question No. 13

A<sup>-</sup>

Which of the following statements is FALSE regarding an ionic bond?

- Breaking an ionic bond between two ions requires energy.
- The ionic bond is formed between two nonmetal atoms.
- An ionic bond is the attraction between a metal ion and a nonmetal ion.
- Metal atoms lose electrons and nonmetal atoms gain electrons.

B



# Chemistry Quiz 2

TAKE TEST

Total questions in exam: 28 | Attempt 1

Question No: 3

The main characteristic of all electrolyte solutions is that they

- contain molecules
- react with other solutions
- conduct heat
- conduct electricity



SAVE & NEXT



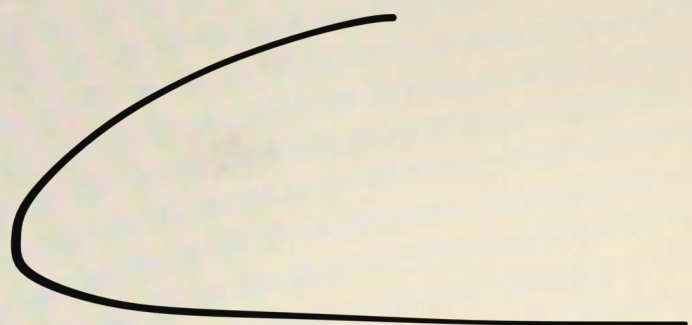
Total questions in exam: 25 | Answered: 1

## Question No. 2

When 22.0 g NaCl and 21.0 g H<sub>2</sub>SO<sub>4</sub> are mixed and react according to the equation below, which is the limiting reagent?



- HCl
- Na<sub>2</sub>SO<sub>4</sub>
- H<sub>2</sub>SO<sub>4</sub>
- NaCl

[Save & Next](#)

Question No. 25

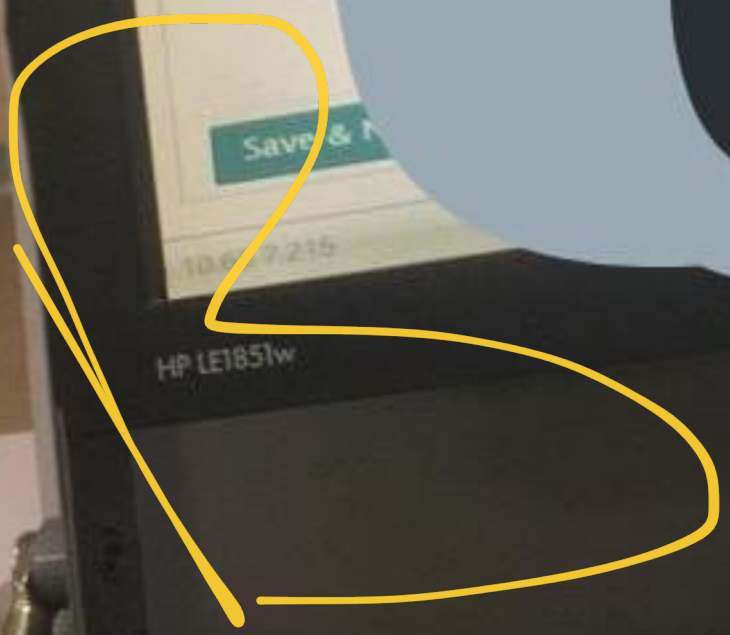
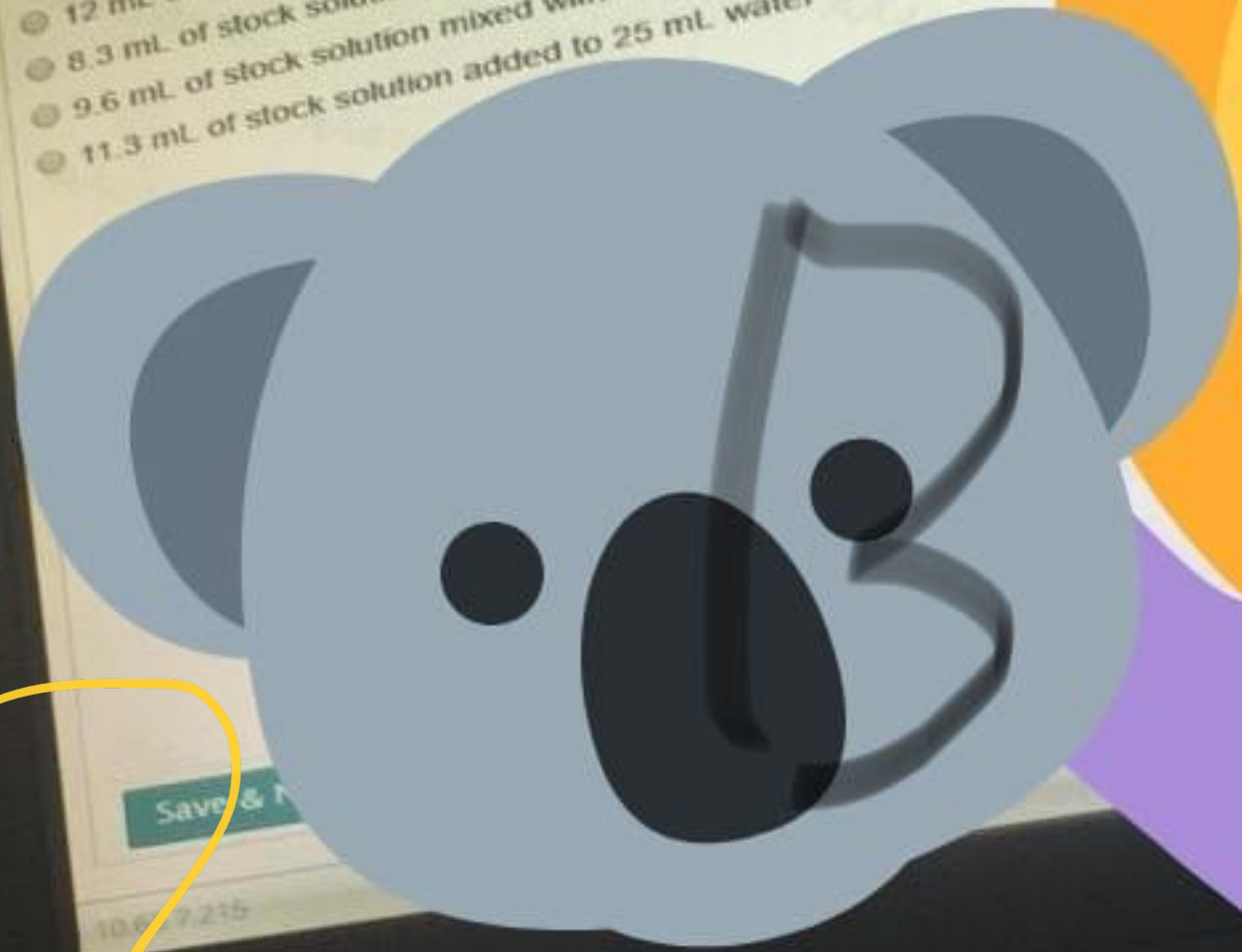
To prepare 25 mL of 0.25 M  $MgCl_2$  solution from a 0.75 M  $MgCl_2$  stock solution, you should mix

- 12 mL of stock solution mixed with enough water to make 25 mL
- 8.3 mL of stock solution mixed with enough water to make 25 mL
- 9.6 mL of stock solution mixed with any amount of water
- 11.3 mL of stock solution added to 25 mL water

Save & P

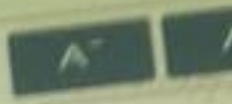
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HP LE1851w





Total questions in exam: 25 | Answered: 1

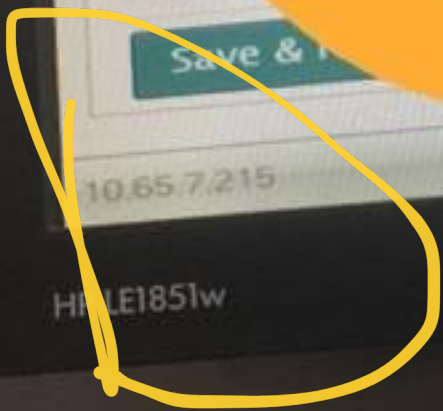
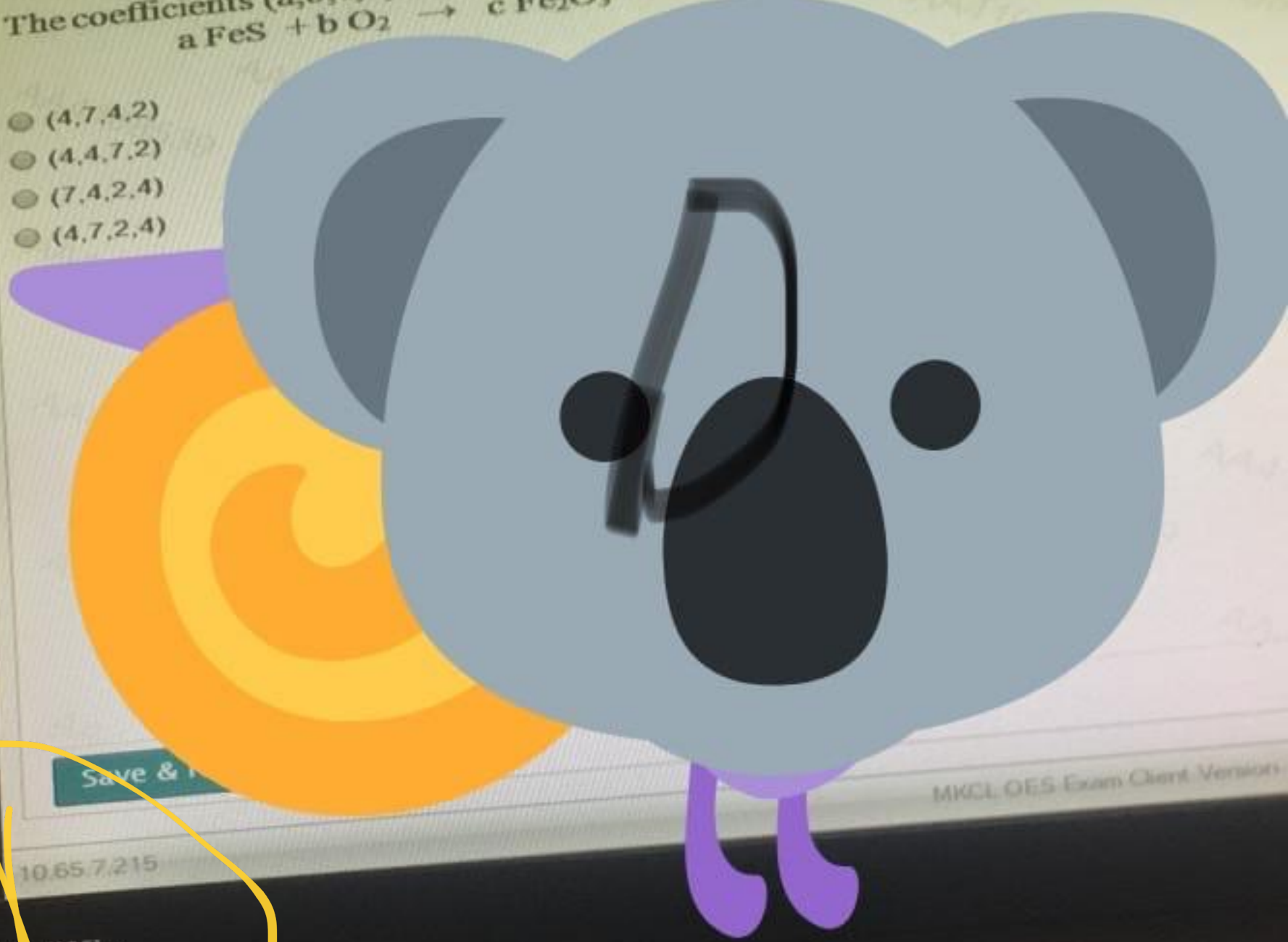


Question No. 21

The coefficients (a,b,c,d) needed to balance the equation below are:  
 $a \text{FeS} + b \text{O}_2 \rightarrow c \text{Fe}_2\text{O}_3 + d \text{SO}_2$

- (4,7,4,2)
- (4,4,7,2)
- (7,4,2,4)
- (4,7,2,4)

Save & ...



Total questions in exam: 25 | Answered: 7

Question No. 24

The most correct name for the compound  $P_4S_{10}$  is:

- tetraphosphorus nonasulfide
- triphosphorus decasulfide
- tetraphosphorus octasulfide
- tetraphosphorus decasulfide





# Chemistry Quiz

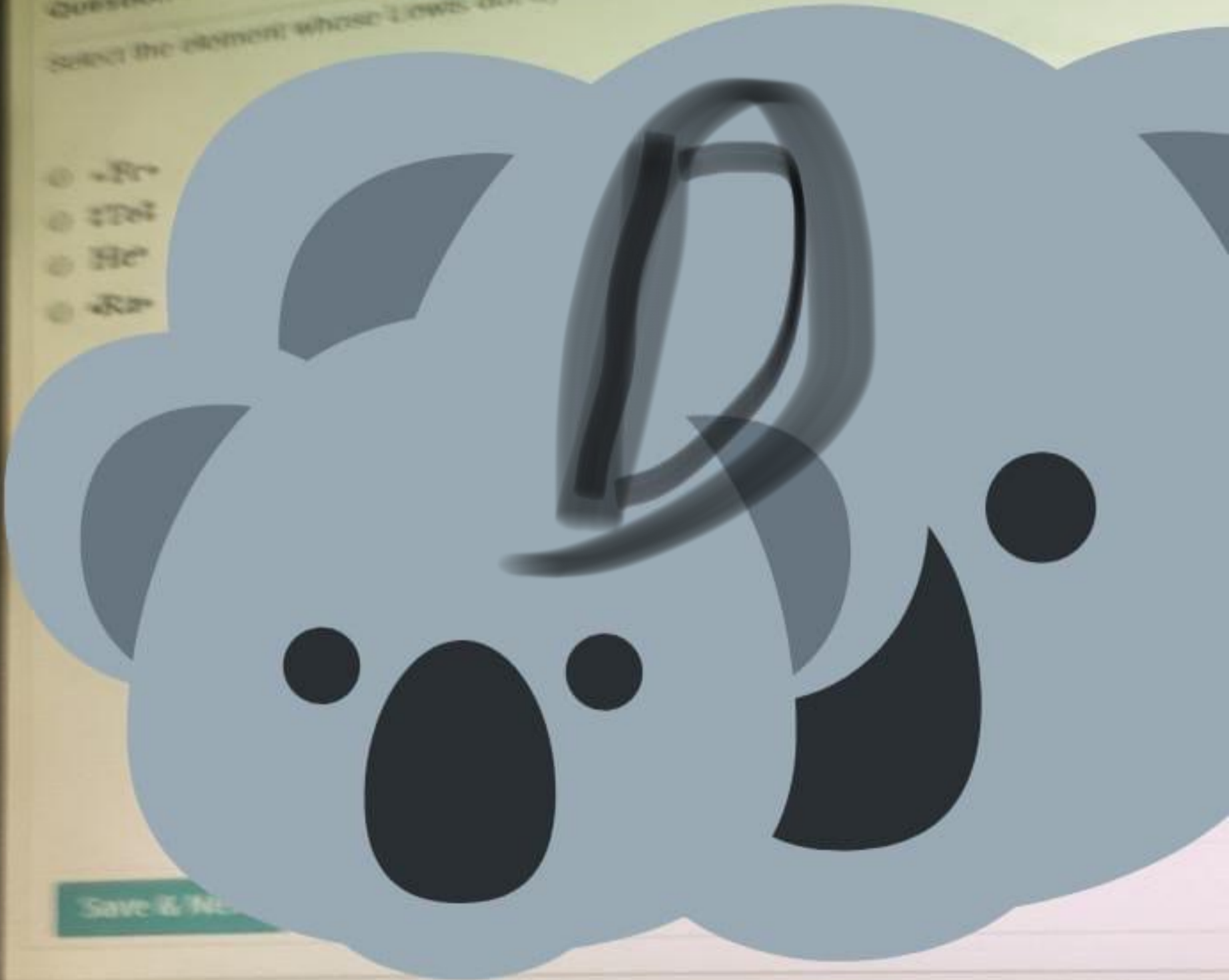
ANSWERS

Total Questions in Quiz: 20

Question No. 20

Select the element whose Lewis dot symbol is correct.

- $\text{Li} \cdot$
- $\cdot\text{Li}$
- $\cdot\text{Li}\cdot$
- $\cdot\text{Li}\cdot\cdot$



Save & Next

MKCL OES

Total questions in exam: 10

Question No. 16

How many moles of  $(\text{NH}_4)_2\text{S}$  are there in 75 g of  $(\text{NH}_4)_2\text{S}$ ?

- 1.9
- 3.4
- 7.5
- 1.1



Question No. 2

The ionization energy of rubidium is higher than the ionization energy of \_\_\_\_\_.

- potassium
- cesium
- sodium
- lithium

B

Question No. 20

The most correct name for the compound  $\text{SO}_3$  is:

- sulfur trioxide
- sulfur(II) oxide
- sulfur oxide
- mono sulfur trioxide

A

Question No. 23

The correct name for the acid  $\text{HBr}$  is \_\_\_\_\_ acid

- hydrogen bromide
- hydrogen bromine
- hydrogen bromate
- hydrobromic



Question No. 25

A<sup>-</sup>

A

A<sup>+</sup>

In which of these substances are the atoms held together by polar covalent bonding?

**B**

- ClF
- S<sub>8</sub>
- CsCl
- SrCl<sub>2</sub>

Question No. 24

Which of the following symbols indicates a solid substance in a chemical

- (g)
- (aq)
- (l)
- (s)



**Question No. 12**

In a solution, the solute is \_\_\_\_\_.

- the substance present in the greatest amount
- always water
- always a solid
- the substance present in the smallest amount





Question No. 13

Which of the following exists as a diatomic molecule?

- phosphorus
- calcium
- carbon
- hydrogen





The empirical formula of the molecular compound  $C_4H_{12}$  is:

- $C_2H_7$
- $C_3H_8$
- $C_4H_{14}$
- $CH_2$



Which of the following compounds gives a nonelectrolyte aqueous solution?

- NaOH
- C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>
- HCl
- Na<sub>2</sub>CO<sub>3</sub>

B

Question No. 4

What is the final molarity of  $\text{H}_2\text{SO}_4$  solution, if 120 mL of 4 M  $\text{H}_2\text{SO}_4$  is diluted to a final volume of 0.5 L?

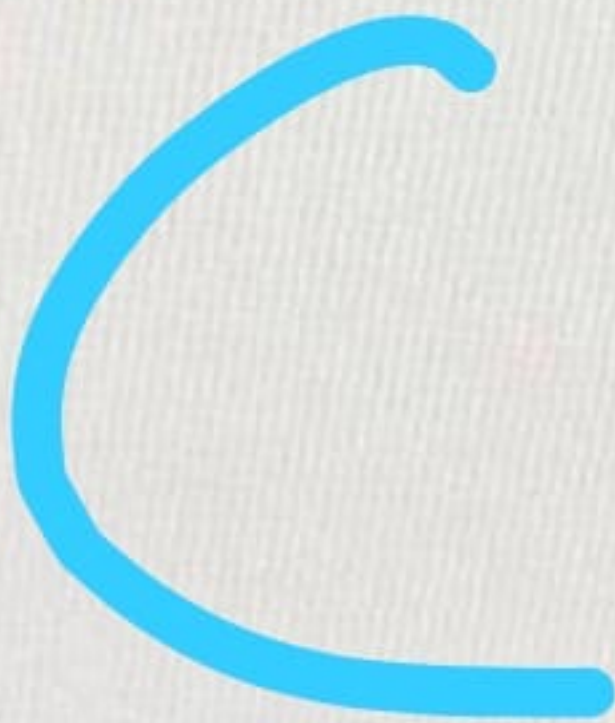
- 0.68 M
- 0.96 M
- 0.60 M
- 0.52 M

B

Question No. 10

What is the term for a bond composed of two electron pairs shared between two atoms?

- triple bond
- electrovalent bond
- double bond
- single bond





Total questions in exam: 25 | Answered: 0

Question No. 1

A

What is the oxidation number of carbon in  $\text{CO}_3^{2-}$  ?

- +2
- 2
- +6
- +4



Save & Next

Total questions in exam: 25 | Answered: 0

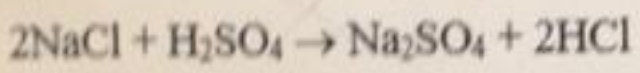
Question No. 2

A<sup>-</sup>

A

A<sup>+</sup>

When 22.0 g NaCl and 21.0 g H<sub>2</sub>SO<sub>4</sub> are mixed and react according to the equation below, which is the limiting reagent?



- H<sub>2</sub>SO<sub>4</sub>
- NaCl
- HCl
- Na<sub>2</sub>SO<sub>4</sub>

A

Save &amp; Next

Total questions in exam: 25 | Answered: 0

A<sup>-</sup>

A

A<sup>+</sup>

Question No. 3

The mass % of Al in aluminum sulfate ( $\text{Al}_2(\text{SO}_4)_3$ ) is \_\_\_\_\_.

- 7.88%
- 15.77%
- 45.7%
- 21.93%

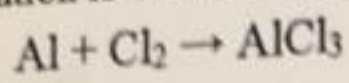
**B**

Save &amp; Next



Question No. 5

When the following equation is balanced, the coefficient for  $\text{AlCl}_3$  will be \_\_\_\_

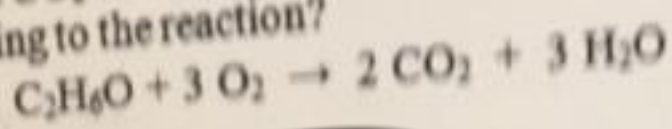



- 3
- 1
- 4
- 2

[Save & Next](#)

Question No. 4

How many molecules of  $\text{CO}_2$  could be produced when 2 moles of  $\text{C}_2\text{H}_6\text{O}$  completely react with oxygen gas according to the reaction?



- 24.08 x  $10^{23}$  molecules.
- 12.04 x  $10^{23}$  molecules.
- 4 molecules.
- 2 molecules. 

[Save & Next](#)

Total questions in exam: 25 | Answered: 1



## Question No. 7

Which of the following species has an oxidation number of zero?

- Al<sup>+3</sup>
- O<sub>2</sub>
- Na<sup>+1</sup>
- Mg<sup>+2</sup>

2

B

[Save & Next](#)

Total questions in exam: 25 | Answered: 1

A

Question No. 9

Which of the following elements has the highest electronegativity?

- Li
- Be
- C
- N

Save &amp; Next

Question No. 6

\_\_\_\_\_ gives a strong electrolyte aqueous solution.

- NH<sub>3</sub>
- CCl<sub>4</sub>
- CaCl<sub>2</sub>
- CS<sub>2</sub>

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ایسی

Save &amp; Next

Total questions in exam: 25 | Answered: 1

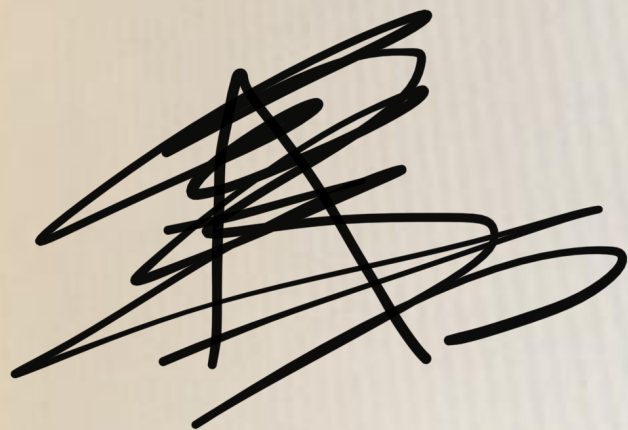
A<sup>-</sup>

A

## Question No. 8

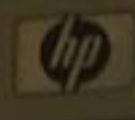
In a solution, the solute is \_\_\_\_\_

- always a solid
- the substance present in the smallest amount
- always water
- the substance present in the greatest amount



B

Save &amp; Next



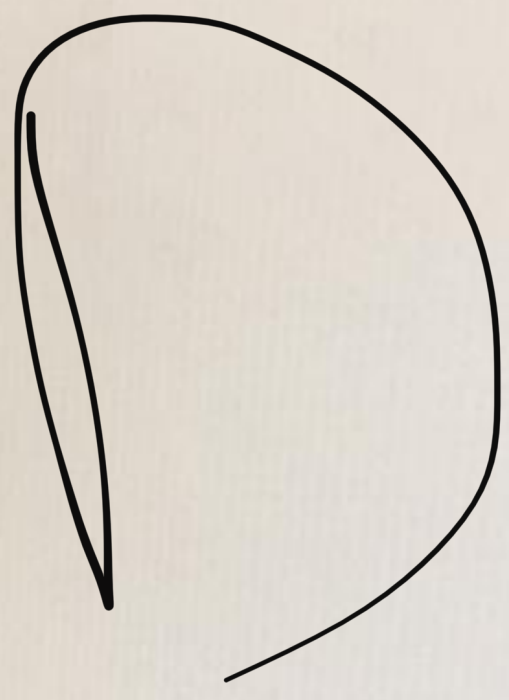
Total questions in exam: 25 | Answered: 1

A<sup>-</sup> A A<sup>+</sup>

Question No. 10

What is the molarity of HCl solution prepared by diluting 600 mL of 1.0 M HCl to a total volume of 1500 mL?

- 0.2 M
- 0.1 M
- 2.0 M
- 0.4 M



Save & Next

Total questions in exam: 25 | Answered: 8

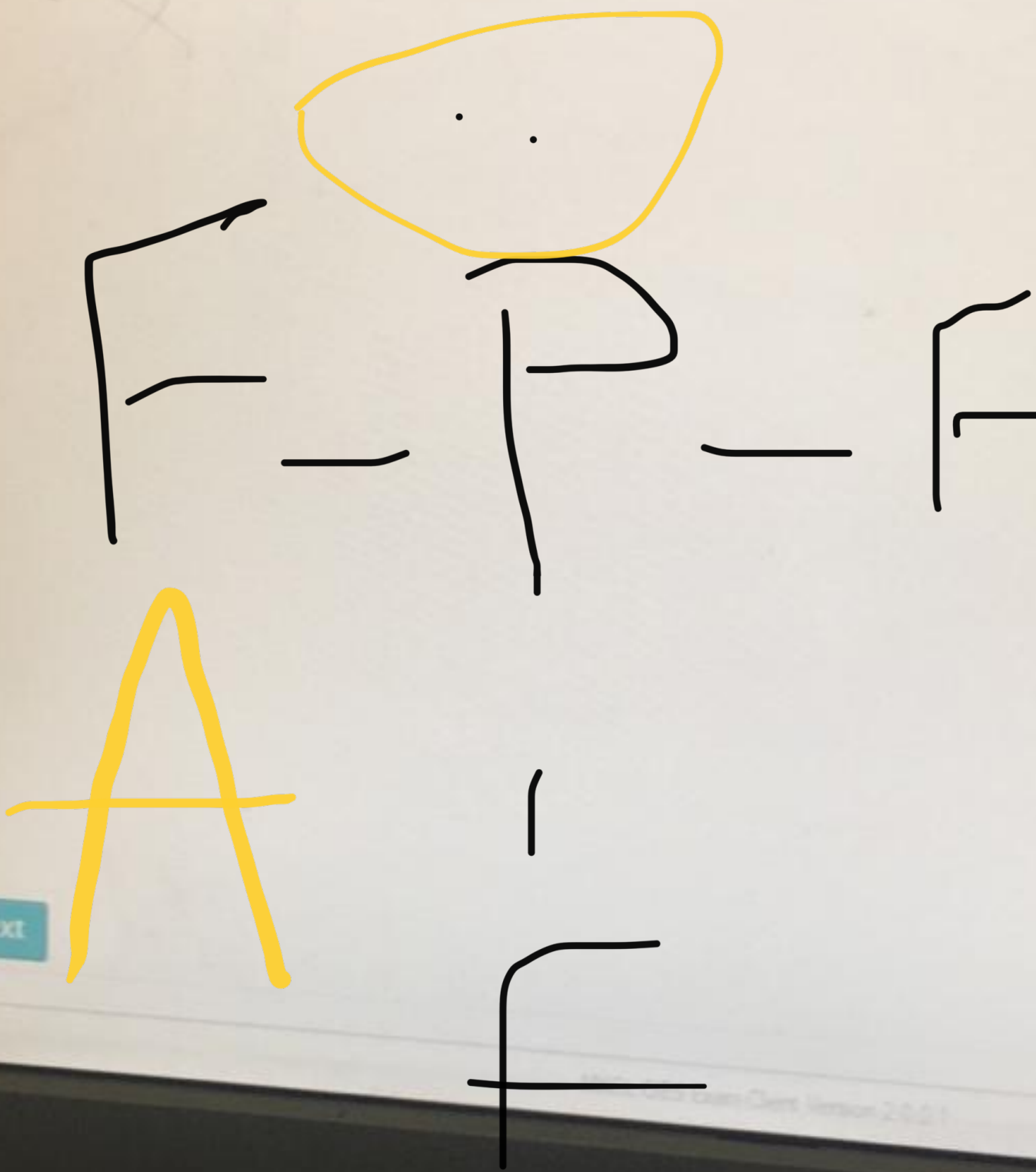
Question No. 13

A<sup>-</sup>

A

A<sup>+</sup>Using Lewis dot structure, find the number of lone pairs of electrons on the "P" atom in PF<sub>3</sub>.

- 1 pair
- 0 pairs
- 3 pairs
- 2 pairs



Save &amp; Next



Total questions in exam: 25 | Answered: 1

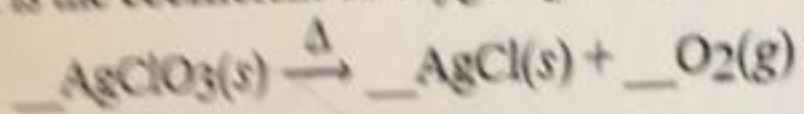
A<sup>-</sup>

A

A<sup>+</sup>

Question No. 12

What is the coefficient of oxygen gas after balancing the following equation?



- 2
- 1
- 4
- 3

Save &amp; Next

Question No. 14

A<sup>-</sup>

Predict which of the following compounds has an ionic bond.

- IBr
- HI
- LiH
- CCl<sub>4</sub>

Save & Next

Total questions in exam: 25 | Answered: 1

A<sup>-</sup>

A

A<sup>+</sup>

Question No. 11

Calculate the mass percent composition of phosphorous in  $K_3PO_4$ .

- 14.6 %
- 26.75%
- 30.2 %
- 55.3 %



Save &amp; Next

Total questions in exam: 25 | Answered: 8

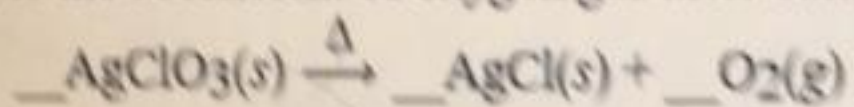
Question No. 12

A

A

A

What is the coefficient of oxygen gas after balancing the following equation?



- 2
- 1
- 4
- 3



Save &amp; Next

Total questions in exam: 25 | Answered: 10

Question No. 18

A<sup>-</sup> A A<sup>+</sup>

What volume of 8.25 M HCl solution must be diluted to prepare 2.40 L of 0.50 M HCl solution?

- 438.8 mL
- 3.96 L
- 0.256 L
- 145.4 mL

35.6

Save &amp; Next



User: AA

Number of

Number of

10 Answered

3 Next Questions

1

2

8

9

15

16

22

23

Cancel

Next

Total questions in exam: 25 | Answered: 8

Question No. 16

A<sup>-</sup>

A

The empirical formula of the molecular compound  $C_6H_{12}$  is:

- $C_6H_{14}$
- $CH_2$
- $C_3H_8$
- $C_2H_7$

B  
- 6

Save &amp; Next

10/05/2016

MKCL OES Exam Client Version 2.0.0.1

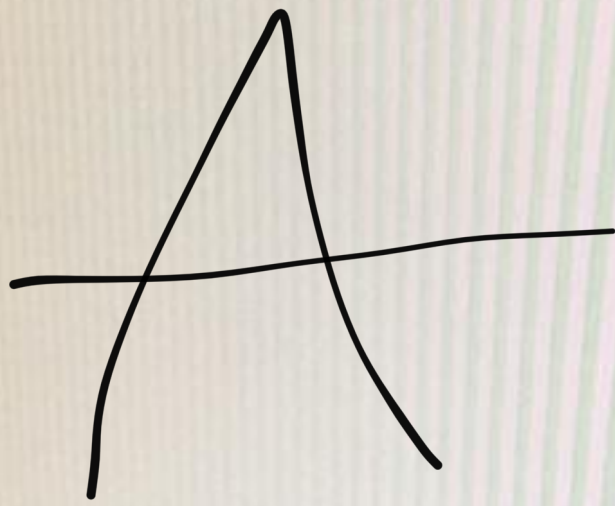
HP 11210

Total questions in exam: 25 | Answered: 10

## Question No. 19

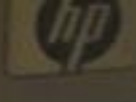
Ionization energy is \_\_\_\_\_

- the energy needed to remove the least tightly bound electron.
- the energy an ion gains from an electron.
- higher for potassium than for lithium.
- highest for metals in Group 1A(1).

A large, handwritten letter 'A' is drawn in black ink on the screen, positioned below the multiple-choice options.

Save &amp; Next

10/05/2015



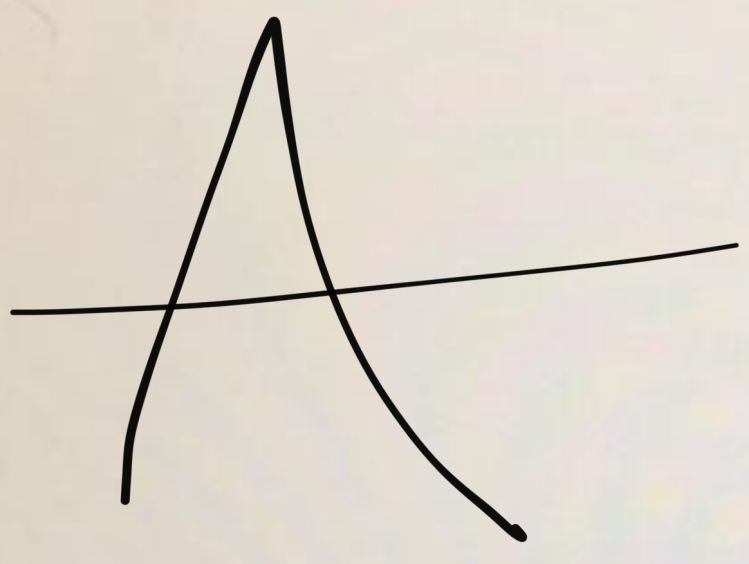
Total questions in exam: 25 | Answered: 10

Question No. 17

A-

How many moles of  $(\text{NH}_4)_2\text{S}$  are there in 150 g of  $(\text{NH}_4)_2\text{S}$ ?

- 2.21
- 1.04
- 1.56
- 1.5



Save & Next

10/05/2019



## Question No. 15

What is the charge on the Fe ions in  $\text{Fe}_2\text{O}_3$ ?

- 1+
- 2-
- 3+
- 2+

Save & Next



Total questions in exam: 25 | Answered: 13

Question No. 22

A<sup>-</sup> A A<sup>+</sup>

Analysis of an unknown substance showed that it has a high boiling point and is brittle. It is an insulator as a solid but conducts electricity when melted. Which of the following substances would have those characteristics?

- Al
- HCl
- KBr
- CO



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Total questions in exam: 25 | Answered: 13

## Question No. 24

Neon is an example of \_\_\_\_\_

- ionic compounds
- molecular compounds
- diatomic molecules
- atomic elements

  
Save & Next

Total questions in exam: 25 | Answered: 13

Question No. 25

A<sup>-</sup>

A

A<sup>+</sup>

A chemical reaction produces 7.25 moles of barium sulfate, BaSO<sub>4</sub>. What mass (in grams) of barium sulfate is produced?

- 233 g
- 185 g
- $1.69 \times 10^3$  g
- 31.1 g

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Total questions in exam: 25 | Answered: 13

A<sup>-</sup>

Question No. 23

The name of the chemical compound  $\text{CuSO}_4$  is:

- copper(II) sulfate
- copper(III) sulfate
- copper sulfate
- copper(I) sulfate



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Question No. 20

What is the percent yield for a reaction if its theoretical yield is 130 g and its actual yield is 74 g?

- 57%
- 53%
- 64%
- 80%

A  
57%

Total questions in exam: 25 | Answered: 23

Question No. 21

A<sup>-</sup>

A

A<sup>+</sup>The most correct name for the compound  $\text{SO}_3$  is:

- sulfur oxide
- mono sulfur trioxide
- sulfur(II) oxide
- sulfur trioxide



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