

Total questions in exam: 25 | Answered: 0

Question No. 2

Which of the following correctly matches the ion with the total number of electrons in the ion?

- Mg^{2+} , 14 electrons
- Br^- , 34 electrons
- Zn^{2+} , 28 electrons
- P^{3-} , 15 electrons

C

Question No. 17

The nucleus of an atom consists mainly of

- protons and electrons.
- protons, neutrons, and electrons.
- neutrons and electrons.
- protons and neutrons.

d

Total questions in exam: 25 | Answered: 15

Question No. 21

Iron has a density of 7.86 g/cm^3 . The volume occupied by 55.85 g of iron is

- 2.8 cm^3
- 0.141 cm^3
- 7.11 cm^3
- 439 cm^3

C

Total questions in exam: 25 | Answered: 0

Question No. 2

Which of the following correctly matches the ion with the total number of electrons in the ion?

- Mg^{2+} , 14 electrons
- Br^- , 34 electrons
- Zn^{2+} , 28 electrons
- P^{3-} , 15 electrons

C1.

Question No. 7

Which of the following statements is incorrect?

- Evaporation of water from a fish tank is evidence of a physical change.
- The condensation of steam on a mirror is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.
- The burning of a piece of charcoal to a white powder is an example of a physical change.

d

Question No. 7

Which of the following statements is incorrect?

- Evaporation of water from a fish tank is evidence of a physical change.
- The condensation of steam on a mirror is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.
- The burning of a piece of charcoal to a white powder is an example of a physical change.



Total questions in exam: 25 | Answered: 1

Question No. 17

The density of glycerin is 1.26 g/mL. Calculate the mass in grams of 800.0 mL glycerin.

- 1.008 Kg
- 1.01 g
- 100 g
- 1008 g

d

Total questions in exam: 26 | Answered: 1

Question No. 17

The density of glycerin is 1.26 g/mL . Calculate the mass in grams of 800.0 mL glycerin.

- 1.008 Kg
- 1.01 g
- 100 g
- 1008 g

Total questions in exam: 26 | Answered: 1

Question No. 17

The density of glycerin is 1.26 g/mL. Calculate the mass in grams of 800.0 mL glycerin.

- 1.008 Kg
- 1.01 g
- 100 g
- 1008 g

Question No. 9

Which of the following is NOT an example of a physical property?

- Solid ice can be very brittle
- Water can break into hydrogen and oxygen gases under electrolysis conditions
- Water can freeze solid at 32 degrees Fahrenheit
- Liquid water can turn into steam in a heated tea kettle

B

Total questions in exam: 25 | Answered: 1

Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Matter is rearranged.
- Atoms react with other atoms.
- Matter is conserved.

A

Total questions in exam: 25 | Answered: 1

Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Matter is rearranged.
- Atoms react with other atoms.
- Matter is conserved.

Total questions in exam: 25 | Answered: 1

Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Matter is rearranged.
- Atoms react with other atoms.
- Matter is conserved.

Total questions in exam: 25 | Answered: 1

Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Matter is rearranged.
- Atoms react with other atoms.
- Matter is conserved.

Total questions in exam: 25 | Answered: 1

Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Matter is rearranged.
- Atoms react with other atoms.
- Matter is conserved.

Total questions in exam: 25 | Answered: 1

Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Matter is rearranged.
- Atoms react with other atoms.
- Matter is conserved.

Question No. 5

The maximum number of electrons that may occupy the second energy level in the atom is _____

- 18
- 10
- 2
- 8

8

Total questions in exam: 25 | Answered: 1

Question No. 10

How many milliliters are in 0.005 L? [1 L = 1000 mL]

- 0.5 mL
- 0.50 mL
- 5 mL
- 0.000005 mL

C

Total questions in exam: 25 | Answered: 1

Question No. 10

How many milliliters are in 0.005 L? [1 L = 1000 mL]

- 0.5 mL
- 0.50 mL
- 5 mL
- 0.000005 mL

Question No. 7

A pure substance is:

- composed of two or more different types of atoms or molecules that has constant
- composed of two or more different types of atoms or molecules combined in vari
- composed of two or more regions with different compositions.
- composed of only one type of atoms or molecules.

Question No. 1

The elements cesium, potassium, and lithium _____.

- are isotopes of each other.
- have the same number of neutrons.
- are in the same period of elements.
- are of the same family.

d

Question No. 1

The elements cesium, potassium, and lithium _____.

- are isotopes of each other.
- have the same number of neutrons.
- are in the same period of elements.
- are of the same family.

Total questions in exam: 25 | Answered: 1

Question No. 7

Deposition is the process in which _____ changes into _____.

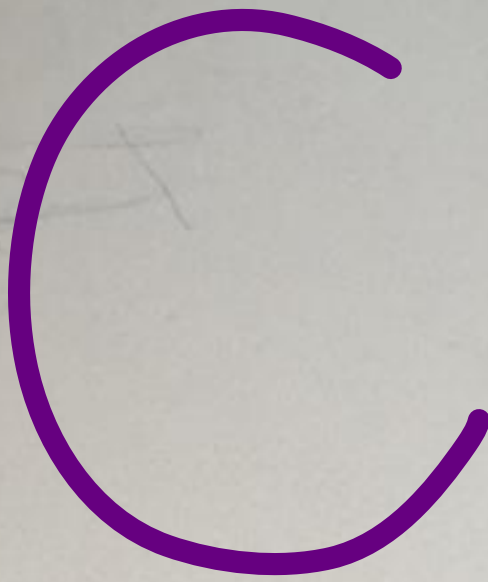
- a gas, a solid
- a solid, a gas
- a gas, a liquid
- a liquid, a solid

A

Question No. 3

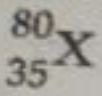
One element that has 4 valence electrons is _____

- lithium
- nitrogen
- carbon
- sulfur



Question No. 6

What does "X" represent in the following symbol?



- iron
- scandium
- chlorine
- bromine

d

Total questions in exam: 25 | Answered: 1

Question No. 5

Which of the following items is a mixture?

- table salt
- sulfur
- baking soda
- cereal and milk

d

Question No. 11

Which of the following is a compound?

- mercury (Hg) in a thermometer
- aluminum (Al) in a can
- salt (NaCl) in a shaker
- pure gold (Au) in ring

C

Question No. 12

Which of the following elements is not a metal?

- Ba
- Xe
- Mg
- Ag

b

Total questions in exam: 25 | Answered: 1

Question No. 4

In which of the following is the metric unit paired with its correct abbreviation?

- microgram / mg
- centimeter / km
- milliliter / mL
- kilogram / cg

C

Top 1

Questions in exam: 25 | Answered: 1

Chemistry_Quiz1_Sem2_2019

Question No. 13

What is the density of a solid sample in g/mL if its volume was found to be 3.05 mL and its mass was 7.03 g

- 2.30 g/mL
- 0.43 g/mL
- 21.4 g/mL
- 0.043 g/mL

A

Question No. 8

Which of the following statements is incorrect?

- The condensation of steam on a mirror is an example of a physical change.
- Evaporation of water from a fish tank is evidence of a physical change.
- The burning of a piece of charcoal to a white powder is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.

C

Question No. 2

Atoms of the same element with different mass numbers are called _____

- ions
- chemical families
- neutrons
- isotopes

d

Question No. 9

Which of the following is NOT an example of a physical property?

- Solid ice can be very brittle
- Water can break into hydrogen and oxygen gases under electrolysis conditions
- Water can freeze solid at 32 degrees Fahrenheit
- Liquid water can turn into steam in a heated tea kettle

D

Total questions in exam: 25 | Answered: 1

Question No. 14

Which one of the following does NOT occur in a chemical reaction?

- Matter is rearranged.
- Atoms are changed into other atoms.
- Matter is conserved.
- Atoms react with other atoms.

b

Question No. 16

A chemical change _____

- occurs when methane gas is flamed
- occurs when water is vaporized
- occurs when paper is shredded (cut)
- occurs when sugar is dissolved in water

A

Total questions in exam: 25 | Answered: 1

Question No. 15

Semiconductors are located in the periodic table on (or in) the _____.

- left side.
- line dividing metals and nonmetals.
- right side.
- first period.

b

Question No. 18

What is the charge on the ion formed by aluminum?

- 5-
- 13+
- 3-
- 3+

d

Question No. 20

The number of protons in the nucleus of an atom

- is called the atomic number.
- is the same for all isotopes of an element.
- identifies the atom as a particular element.
- All the answers are correct

d

QUESTION No. 20

The energy stored in the chemical bonds of a molecule is

- kinetic energy
- work
- specific heat
- potential energy

d



Total questions in exam: 25 | Answered: 8

Question No. 12

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The space occupied by the electrons is almost empty.
- The volume occupied by electrons is referred to as an electron cloud.

b

Save & Next حفظ والتالي



Total questions in exam: 25 | Answered: 8

Question No. 12

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The space occupied by the electrons is almost empty.
- The volume occupied by electrons is referred to as an electron cloud.

Save & Next حفظ والتالي

Question No. 16

Which of the following elements is a liquid at room temperature?

- helium
- sodium
- chlorine
- mercury

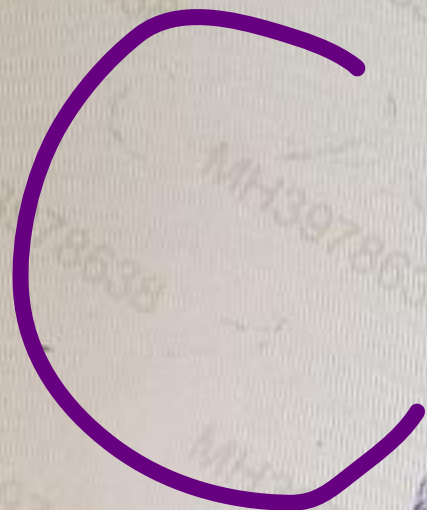
d

Save & Next حفظ والتالي

Question No. 4

Bromine is the only nonmetal that is a liquid at room temperature. Consider the isotope Br-81. Select the combination which lists the correct atomic number, neutron number, and mass number, respectively.

- 35, 81, 46
- 81, 46, 35
- 35, 46, 81
- 46, 81, 35



User: MH397

Number of m
Number of qu

2 Answered

20 Not Visited

1

2

8

9

15

16

22

23

Question No. 4

Bromine is the only nonmetal that is a liquid at room temperature. Consider the isotope Br-81. Select the combination which lists the correct atomic number, neutron number, and mass number, respectively.

- 35, 81, 46
- 81, 46, 35
- 35, 46, 81
- 46, 81, 35

User: MH397

Number of m
Number of qu

2 Answered

20 Not Visited

1	2	
8	9	
15	16	
22	23	

Atoms are neutral because _____

- neutrons neutralize the (+) and (-) charges in the atom
- equal numbers of protons and neutrons are present
- all atoms contain neutrons
- the atomic number equals the number of electrons

d

Question No. 9

Which one of the following is a homogeneous mixture?

- Alloys
- Vegetable salad
- water
- Sand in water

A

Question No. 9

Which one of the following is a homogeneous mixture?

- Alloys
- Vegetable salad
- water
- Sand in water

Question No. 9

Which one of the following is a homogeneous mixture?

- Alloys
- Vegetable salad
- water
- Sand in water

Question No. 9

Which one of the following is a homogeneous mixture?

- Alloys
- Vegetable salad
- water
- Sand in water

Question No. 17

The nucleus of an atom consists mainly of

- protons and electrons.
- protons, neutrons, and electrons.
- neutrons and electrons.
- protons and neutrons.

d

Question No. 17

The nucleus of an atom consists mainly of

- protons and electrons.
- protons, neutrons, and electrons.
- neutrons and electrons.
- protons and neutrons.

Total questions in exam: 25 | Answered: 8

Question No. 11

One decimeter is equal to _____.

- 1 m
- 10 cm
- 1000 mm
- 10 m

A

Save & Next حفظ و التالي

Question No. 8

Which of the following is NOT a correct (name, symbol) combination?

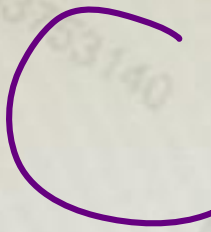
- beryllium, Be
- calcium, Ca
- magnesium, Mg
- manganese, Mn

Save & Next حفظ والتالي

Question No. 10

According to Hund's Rule _____.

- electrons pair first before filling other orbitals with same energy
- no two electrons in the same atom can have the same four quantum numbers.
- electrons fill orbitals of the same energy singly and pair up only after the orbitals are all half filled.
- electrons found in the same orbital must have the same spin.



Save & Next حفظ و التالي

Question No. 10

According to Hund's Rule _____.

- electrons pair first before filling other orbitals with same energy
- no two electrons in the same atom can have the same four quantum numbers.
- electrons fill orbitals of the same energy singly and pair up only after the orbitals are all half filled.
- electrons found in the same orbital must have the same spin.

Save & Next حفظ و التالي

Question No. 3

Scandium belongs to _____.

- transition metals
- alkaline earth metals
- noble gases
- halogens

A

Question No. 5

Sodium is an example of

- metalloids
- nonmetals
- metals
- noble gases

Save & Next حفظ التالي

HP Compaq LE1711

Question No. 4

Which one of these represents a physical change?

- water, when heated, forms water vapor
- sugar, when heated, becomes brown
- bleach turns hair yellow
- milk turns sour

A

Save & Next حفظ و التالي

HP Compaq LE1711

Question No. 4

Which one of these represents a physical change?

- water, when heated, forms water vapor
- sugar, when heated, becomes brown
- bleach turns hair yellow
- milk turns sour

Save & Next حفظ و التالي

HP Compaq LE1711

Total questions in exam: 25 | Answered: 8

Question No. 12

Consider the atoms of Co-59 and Co-60, both of these atoms have the same _____

- number of protons
- atomic mass number
- number of neutrons
- none of these is true

Total questions in exam: 25 | Answered: 8

Question No. 12

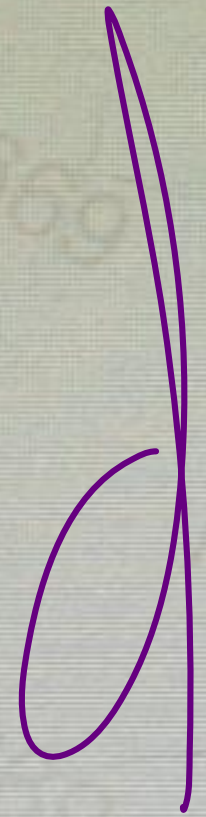
Consider the atoms of Co-59 and Co-60, both of these atoms have the same _____

- number of protons
- atomic mass number
- number of neutrons
- non of these is true

Question No. 4

What is the atomic symbol for manganese?

- Mo
- Mg
- Md
- Mn



Question No. 7

Which of the following statements is incorrect?

- Evaporation of water from a fish tank is evidence of a physical change.
- The condensation of steam on a mirror is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.
- The burning of a piece of charcoal to a white powder is an example of a physical change.

D

Question No. 3

Which one of the following does NOT occur in a chemical reaction?

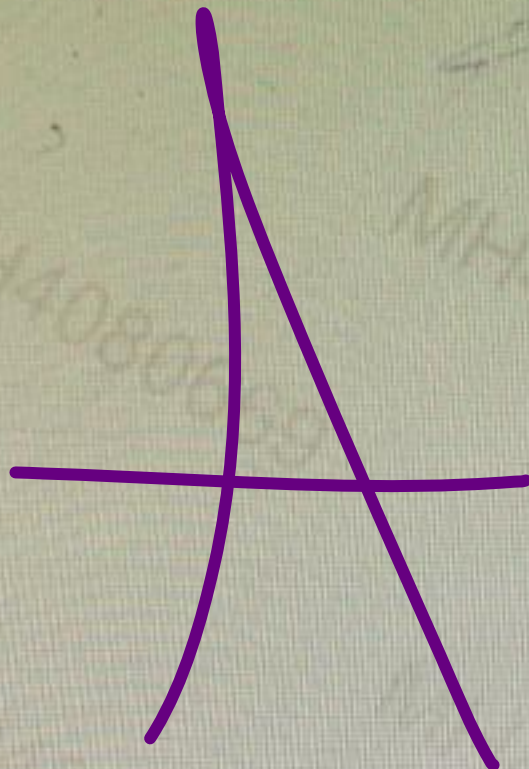
- Matter is conserved
- Matter is rearranged
- Atoms are changed into other atoms.
- Atoms react with other atoms.

C

Question No. 6

Condensation refers to which conversion?

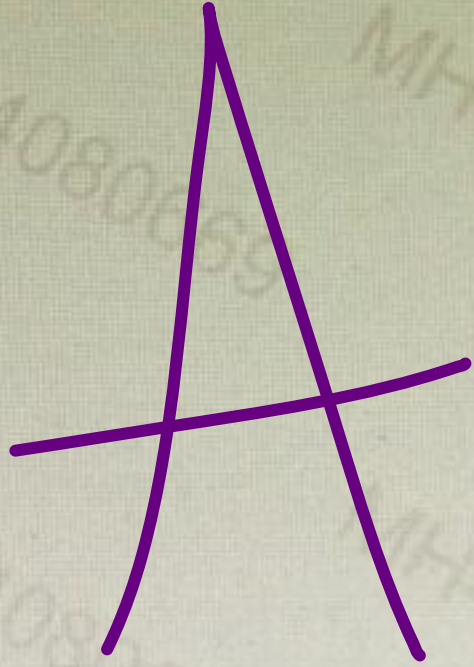
- gas -> liquid
- solid -> gas
- liquid -> gas
- solid -> liquid



Question No. 8

How many protons and electrons are present in O^{2-} ?

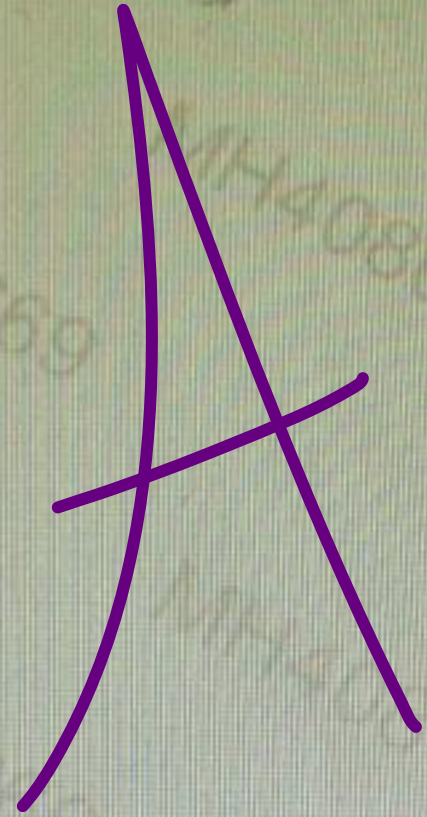
- 8 protons and 10 electrons
- 8 protons and 8 electrons
- 16 protons and 8 electrons
- 10 protons and 8 electrons



Question No. 3

Which of the following is an example of chemical property?

- flammability
- viscosity
- density
- color



Question No. 3

Which of the following is an example of chemical property?

- flammability
- viscosity
- density
- color

Question No. 2

Which of the following gives the correct numbers of protons, neutrons, and electrons in a neutral atom of $^{118}_{50}\text{Sn}$?

- 50 protons, 68 neutrons, 50 electrons
- 118 protons, 50 neutrons, 118 electrons
- 68 protons, 68 neutrons, 50 electrons
- 118 protons, 118 neutrons, 50 electrons

A

Question No. 2

Which of the following gives the correct numbers of protons, neutrons, and electrons in a neutral atom of $^{118}_{50}\text{Sn}$?

- 50 protons, 68 neutrons, 50 electrons
- 118 protons, 50 neutrons, 118 electrons
- 68 protons, 68 neutrons, 50 electrons
- 118 protons, 118 neutrons, 50 electrons

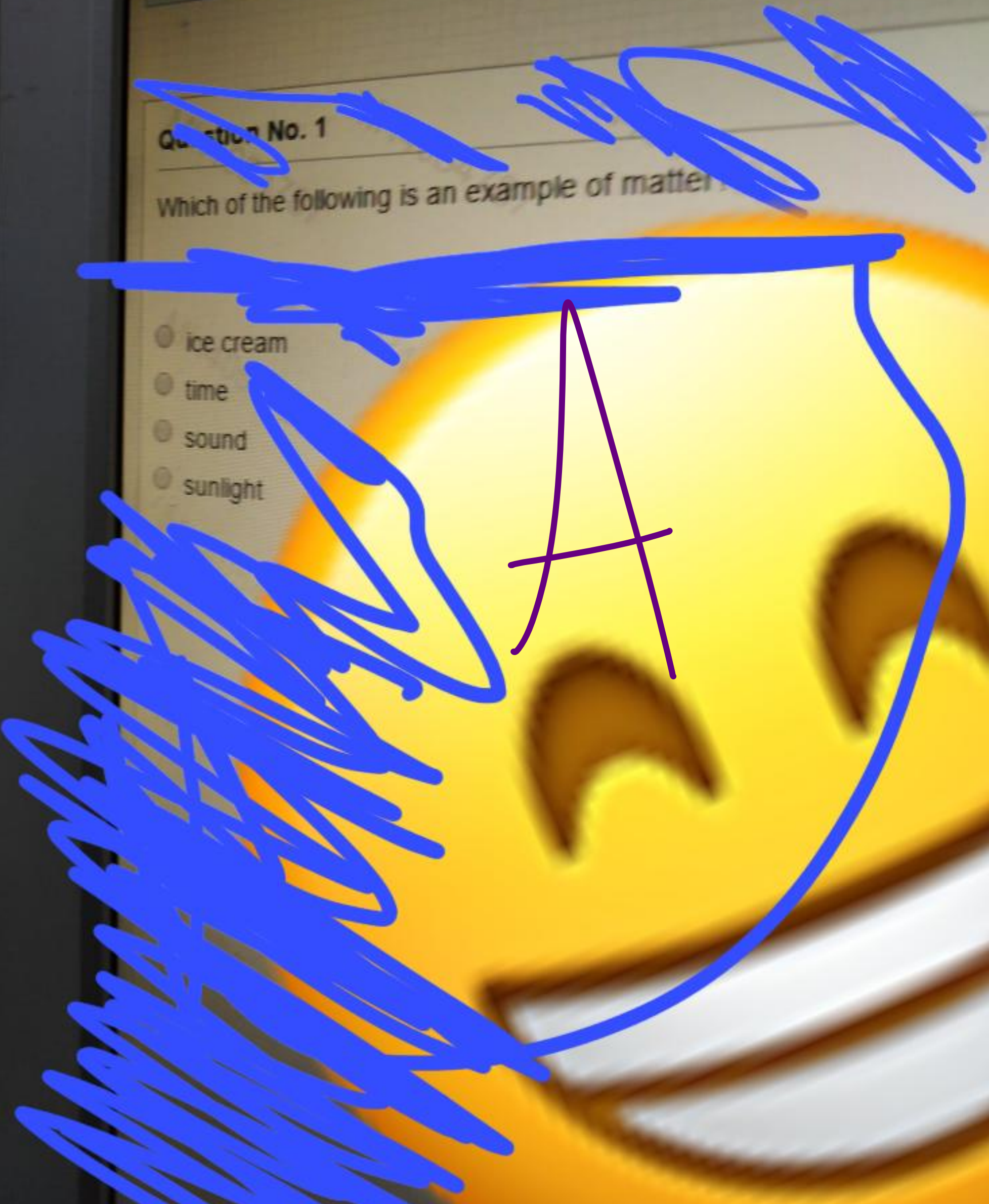
✓
1/21

Question No. 1

Which of the following is an example of matter?

- ice cream
- time
- sound
- sunlight

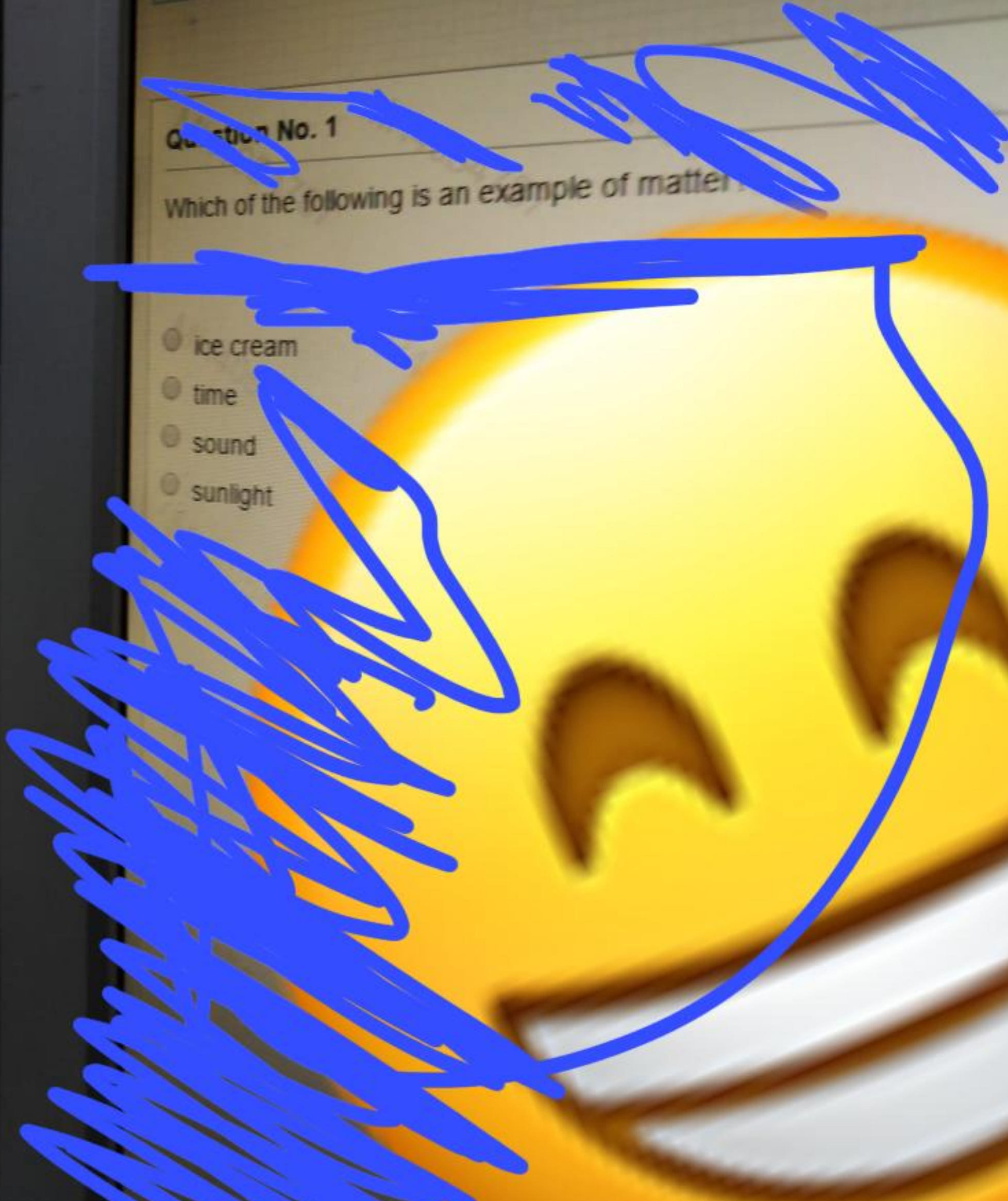
A



Question No. 1

Which of the following is an example of matter?

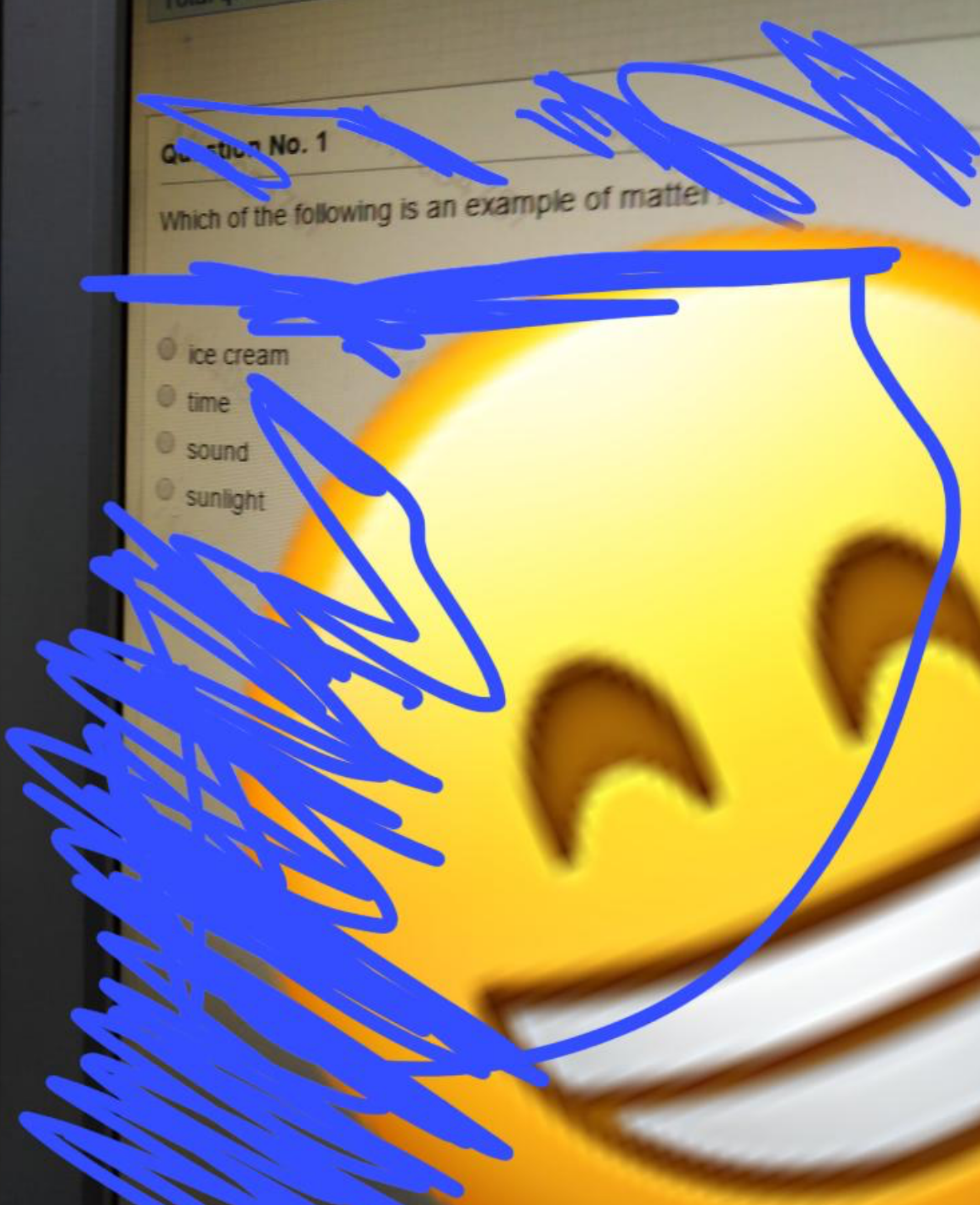
- ice cream
- time
- sound
- sunlight



Question No. 1

Which of the following is an example of matter?

- ice cream
- time
- sound
- sunlight



Total questions in exam: 25 | Answered: 1

Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Matter is rearranged.
- Atoms react with other atoms.
- Matter is conserved.

A

Question No. 3

Which of the following is an example of chemical property?

- flammability
- viscosity
- density
- color

A

Question No. 2

Which of the following gives the correct numbers of protons, neutrons, and electrons in a neutral atom of $^{118}_{50}\text{Sn}$?

- 50 protons, 68 neutrons, 50 electrons
- 118 protons, 50 neutrons, 118 electrons
- 68 protons, 68 neutrons, 50 electrons
- 118 protons, 118 neutrons, 50 electrons

A

Question No. 3

Which of the following is an example of chemical property?

- flammability
- viscosity
- density
- color

Question No. 9

Which one of the following is a homogeneous mixture?

- Alloys
- Vegetable salad
- water
- Sand in water

A

✓
28
1

Question No. 7

Which of the following statements is incorrect?

- Evaporation of water from a fish tank is evidence of a physical change.
- The condensation of steam on a mirror is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.
- The burning of a piece of charcoal to a white powder is an example of a physical change.

d

Question No. 8

How many protons and electrons are present in O^{2-} ?

- 8 protons and 10 electrons
- 8 protons and 8 electrons
- 16 protons and 8 electrons
- 10 protons and 8 electrons



7/8

Total questions in exam: 25 | Answered: 1

Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Matter is rearranged.
- Atoms react with other atoms.
- Matter is conserved.

A

11/11/2023

Total questions in exam 25 | Answered 2

Question No. 3

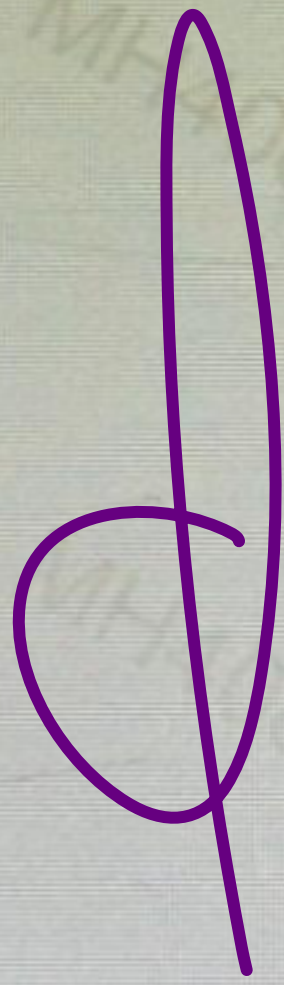
Which one of the following does NOT occur in a chemical reaction?

- Matter is conserved
- Matter is rearranged
- Atoms are changed into other atoms.
- Atoms react with other atoms.

Question No. 4

What is the atomic symbol for manganese?

- Mo
- Mg
- Md
- Mn



Total questions in exam: 25 | Answered: 0

Question No. 2

Which of the following correctly matches the ion with the total number of electrons in the ion?

- Mg^{2+} , 14 electrons
- Br^- , 34 electrons
- Zn^{2+} , 28 electrons
- P^{3-} , 15 electrons

Question No. 4

Bromine is the only nonmetal that is a liquid at room temperature. Consider the isotope Br-81. Select the combination which lists the correct atomic number, neutron number, and mass number, respectively.

- 35, 81, 46
- 81, 46, 35
- 35, 46, 81
- 46, 81, 35



User: MH397

Number of m
Number of qu

2 Answered

20 Not Visited

1

2

8

9

15

16

22

23

Question No. 4

Bromine is the only nonmetal that is a liquid at room temperature. Consider the isotope Br-81. Select the combination which lists the correct atomic number, neutron number, and mass number, respectively.

- 35, 81, 46
- 81, 46, 35
- 35, 46, 81
- 46, 81, 35

User: MH397

Number of m
Number of qu

2 Answered

20 Not Visited

1	2	
8	9	
15	16	
22	23	

Question No. 4

Bromine is the only nonmetal that is a liquid at room temperature. Consider the isotope Br-81. Select the combination which lists the correct atomic number, neutron number, and mass number, respectively.

- 35, 81, 46
- 81, 46, 35
- 35, 46, 81
- 46, 81, 35

No Image

User MH397

Number of m

Number of qu

2 Answered

20 Not Visited

1	2
8	9
15	16
22	23

Question No. 4

Bromine is the only nonmetal that is a liquid at room temperature. Consider the isotope Br-81. Select the combination which lists the correct atomic number, neutron number, and mass number, respectively.

- 35, 81, 46
- 81, 46, 35
- 35, 46, 81
- 46, 81, 35

No Image

User MH397

Number of m

Number of qu

2 Answered

20 Not Visited

1	2
8	9
15	16
22	23

Total questions in exam: 25 | Answered: 8

Question No. 12

Consider the atoms of Co-59 and Co-60, both of these atoms have the same _____

- number of protons
- atomic mass number
- number of neutrons
- non of these is true



Total questions in exam: 25 | Answered: 8

Question No. 12

Consider the atoms of Co-59 and Co-60, both of these atoms have the same _____

- number of protons
- atomic mass number
- number of neutrons
- non of these is true

Question No. 4

Which one of these represents a physical change?

- water, when heated, forms water vapor
- sugar, when heated, becomes brown
- bleach turns hair yellow
- milk turns sour

A

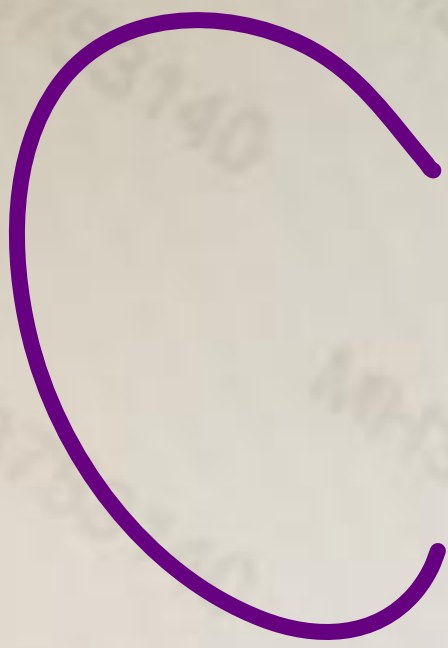
Save & Next حفظ و التالي

HP Compaq LE1711

Question No. 10

According to Hund's Rule _____.

- electrons pair first before filling other orbitals with same energy
- no two electrons in the same atom can have the same four quantum numbers.
- electrons fill orbitals of the same energy singly and pair up only after the orbitals are all half filled.
- electrons found in the same orbital must have the same spin.



Save & Next حفظ و التالي

Question No. 17

The nucleus of an atom consists mainly of

- protons and electrons.
- protons, neutrons, and electrons.
- neutrons and electrons.
- protons and neutrons.



Total questions in exam: 25 | Answered: 8

Question No. 12

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The space occupied by the electrons is almost empty.
- The volume occupied by electrons is referred to as an electron cloud.

b

Save & Next حفظ والتالي

QUESTION No. 20

The energy stored in the chemical bonds of a molecule is

- kinetic energy
- work
- specific heat
- potential energy

potential energy

D

Question No. 7

A pure substance is:

- composed of two or more different types of atoms or molecules that has constant
- composed of two or more different types of atoms or molecules combined in vari
- composed of two or more regions with different compositions.
- composed of only one type of atoms or molecules.

Question No. 14

Express 7.5 nm as picometers.

- 750 pm
- 7.5×10^6 pm
- 75.0 pm
- 7.5×10^3 pm



Question No. 21

Melting is a process in which _____ changes into _____, and it requires _____ temperature.

- a solid, a liquid, increasing
- a liquid, a gas, decreasing
- a liquid, a solid, increasing
- a solid, a gas, increasing

A

Question No. 21

Melting is a process in which _____ changes into _____, and it requires _____ temperature.

- a solid, a liquid, increasing
- a liquid, a gas, decreasing
- a liquid, a solid, increasing
- a solid, a gas, increasing

Total questions in exam: 25 | Answered: 8

Question No. 12

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The space occupied by the electrons is almost empty.
- The volume occupied by electrons is referred to as an electron cloud.

b

+966 56 617 3771

١٩ من الصور

تحديد



MKCL OES

Total questions in exam: 25 | Answered: 1

Question No. 16

A chemical change _____

- occurs when methane gas is flamed
- occurs when water is vaporized
- occurs when paper is shredded (cut)
- occurs when sugar is dissolved in water

A