Question No. 10

Condensation refers to which conversion?

- solid -> gas
- solid -> liquid
- liquid -> gas
- gas -> liquid

Save & Next

Question No. 4

Which of the following is a compound?

- mercury (Hg) in a thermometer
- aluminum (Al) in a can
- pure gold (Au) in ring
- salt (NaCl) in a shaker

Question No. 8

The correct prefix for the multiplier 0.000001 is _____.

- mega
- milli
- nano
- micro

Save & Next حفظ والتالي

Question No. 3

Which of the following items is a pure substance?

- air
- brass (zinc and copper)
- ice
- sea water

Question No. 2

What is the mass of 2.00 L of a solution with a density of 1.15 g/mL?

- [redacted]
- 1.15 kg
- 0.015 kg
- 2.30 kg

D

Question No. 1

What is the density of a solid sample in g/mL if its volume was found to be 3.05 mL and its mass was 7.03 g

- 2.30 g/mL
- 0.43 g/mL
- 0.043 g/mL
- 21.4 g/mL

Question No. 6

_____is the metric unit of the temperature of an object.

- Celsius
- Fahrenheit
- Quart
- Kelvin

Save & Next

Question No. 7

Convert 4 micrometers to meters. *[1 m = 1000,000 μm]*

- 4×10^{-6} m
- 4×10^{-9} m
- 4×10^{6} m
- 4×10^{-3} m

Save & Next

Question No. 8

Which of the following quantities represents the largest mass?

- 0.0010 kg
- 2.0×10^2 mg
- 2.0×10^2 g
- 1.0×10^5 μg

MKCL OES
Online Evaluation System

Chemis

Question No. 4

One of the following is an example of a pure substance:

- sea water
- oil and water
- diamond
- lemon juice and tea

Question No. 5

Which of the following is an example of a homogeneous mixture?

- salt dissolved in water
- sand mixed with oil
- gasoline mixed with water
- salad dressing

Save & Next

Question No. 3

Choose the pure substance from the list below.

- sugar
- milk
- lemonade
- air

Save & Next

Question No. 9

Identify the crystalline solid from the following.

- plastic
- water
- table salt
- glass

Question No. 1

What is the mass of 36.0 mL of a liquid if its density is 2.70 g/mL?

- 0.0750 g
- 13.3 g
- 0.748 g
- 97.2 g

Save & Next

Question No. 5

A piece of metal with a mass of 114 g was placed into a graduated cylinder that contained 25.00 mL of water, raising the water level to 42.50 mL. What is the density of the metal?

- 7.552 g/cm^3
- 7.154 g/cm^3
- 6.51 g/cm^3
- 2.68 g/cm^3

Question No. 10

Sublimation refers to which conversion?

- liquid -> gas
- gas -> liquid
- solid -> gas
- solid -> liquid

Save & Next حفظ وتالي

Question No. 11

Which one of these represents a chemical change?

- mixing carbon and oxygen (at room temperature)
- melting butter
- boiling water to form water vapor
- grilling a steak

Save & Next

Question No. 5

How would you classify sugar?

- mixture-heterogeneous
- mixture-homogeneous
- pure substance-element
- pure substance-compound

Question No. 9

Which of the following is NOT an example of matter?

- heat from a burning candle
- a dust particle
- a pencil eraser
- a balloon full of helium

Save & Next حفظ والتالي

Question No: 8

One decimeter is equal to

- 10 m
- 10 cm
- 1 m
- 1000 mm

Save & Next محفظ و التالي

Question No. 25

The smallest subatomic particles is _____.

- electron
- proton
- nucleus
- neutron

Question No. 22

What is the charge on the ion formed by sodium?

- 2-
- 2+
- 1+
- 1-

Question No. 23

A cation of +1 indicates that an element has _____.

- lost one proton.
- gained one proton.
- gained one electron.
- lost one electron.

Question No. 24

A neutron has an electrical charge of _____.

- -1
- -2
- +1
- 0

Question No. 21

The elements lithium, beryllium, and boron _____.

- are in the same group
- are in the same period
- have the same number of neutrons
- are isotopes of each other

Question No. 20

The elements beryllium, carbon, lithium _____.

- have the same mass number.
- are in the same group.
- have the same number of neutrons.
- are in the same period of elements.

Question No. 19

Valence electrons are electrons located _____.

- in the nucleus of an atom.
- in the first three shells of an atom.
- in the innermost energy level of an atom.
- in the outermost energy level of an atom.

Question No. 16

Who formulated the modern atomic theory?

- Antoine Lavoisier
- Joseph Thomson
- Nivaldo Tro
- John Dalton

Question No. 17

What element has the electron configuration $1s^2 2s^2 2p^5$?

- Na
- S
- F
- Mg

Chem

Question No. 18

The "p" orbital can hold up to _____ electrons.

- 8
- 2
- 18
- 6

Question No. 1

Determine the volume of an object that has a mass of 455.6 g and a density of 19.3 g/cm^3.

- 42.4 mL
- 18.5 mL
- 87.9 mL
- 23.6 mL

Question No. 8

In which of the following is the metric unit paired with its correct abbreviation?

- microgram / mg
- centimeter / km
- kilogram / cg
- milliliter / mL

Question No. 16

In the winter, if the outdoor temperature is -12 °F, What is the temperature on the Kelvin scale?

- 262 K
- 273 K
- 284 K
- 249 K

Question No. 17

"Every orbital in a subshell is singly occupied with one electron before any one orbital is doubly occupied" is known as ____

- Hund's rule
- Pauli exclusion principle
- Heisenberg uncertainty principle.
- Aufbau principle.

Question No. 18

What element has the electron configuration $1s^2 2s^2 2p^5$?

- Na
- Mg
- S
- F

MACL OLS

Question No. 3

Choose the pure substance from the list below

- sugar
- air
- milk
- lemonade

Question No. 15

Which of the following is the lowest temperature?

- 35 °F
- 0 °C
- -10 °C
- 100 K

[redacted]

Save & Next

Question No. 16

Which of the following is NOT a correct (name, symbol) combination?

- magnesium, Mg
- beryllium, Be
- manganese, Mn
- calcium, Ka

Save & Next حفظ وتالی

Question No. 11

A chemical change _____.

- occurs when paper is shredded (cut)
- occurs when methane gas is flamed
- occurs when sugar is dissolved in water
- occurs when water is vaporized

Save & Next حفظ والتالي

Question No. 12

All of the following can be considered physical properties EXCEPT _____.

- color
- taste
- boiling point
- flammability

ve & Next

Question No. 13

Which one of the following choices includes only chemical changes?

- rusting, evaporation, and deposition
- decomposition, digestion, and burning
- melting, evaporation, and sublimation
- melting, freezing, and digestion

Next

Question No. 9

The _____ is the building block of matter.

- atom
- element
- compound
- molecule

Save & Next

Question No. 10

Identify a solid.

- definite volume and no definite shape
- definite volume and definite shape
- no definite shape and no definite volume
- definite shape and compressible

Next