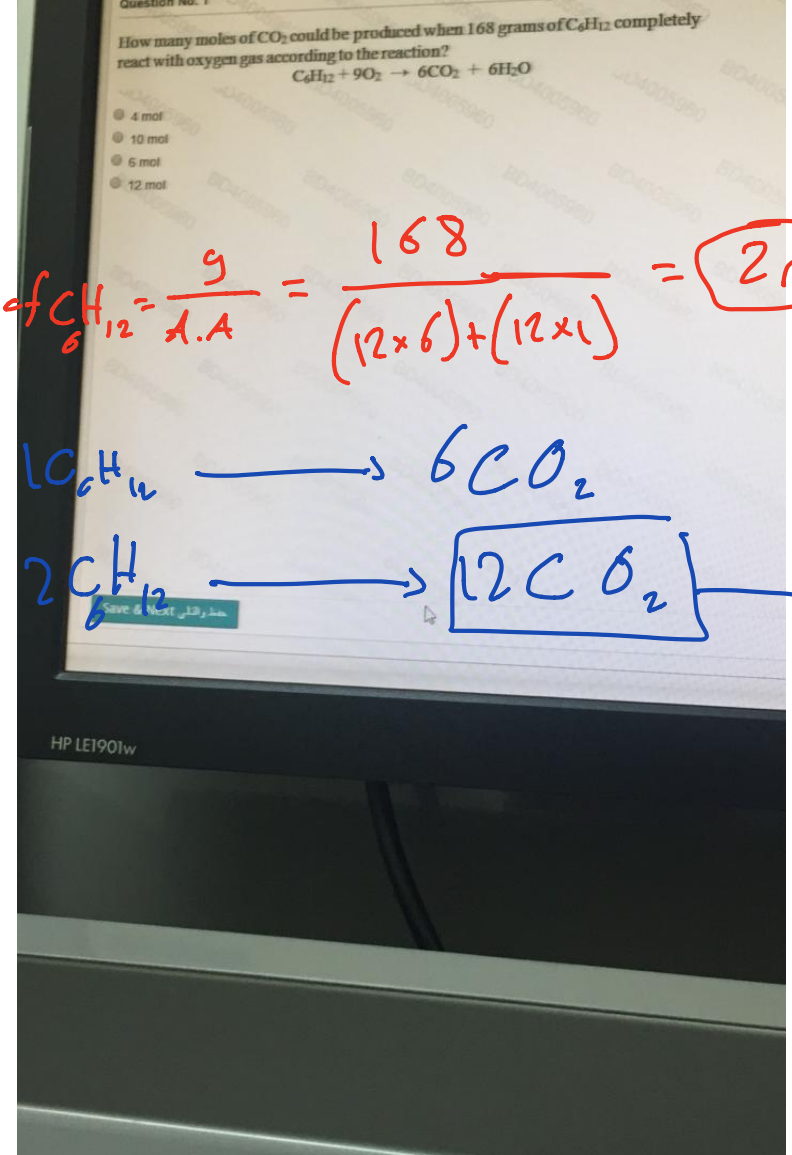


mole of C_6H_{12} = $\frac{g}{A.A}$ = $\frac{168}{(12\times6)+(12\times1)}$ = 2m

$1C_6H_{12} \longrightarrow 6CO_2$

$2C_6H_{12} \longrightarrow 12CO_2$ → D

Total questions in exam: 40 | Answered: 0

Question No. 2

What is the general term for a substance dissolved in water?

- covalent substance
- aqueous solution
- ionic salt
- water solution

Save & Next حفظ والتالي

LE1901w

Total questions in exam: 40 | Answered: 0

Question No. 1

In an experiment, 50.0 g of silicon tetrachloride ($SiCl_4$) is treated with 20.0 g of water to produce silicon dioxide (SiO_2) according to the following balanced equation:

$$SiCl_{4(S)} + 2H_2O_{(l)} \rightarrow SiO_{2(s)} + 4HCl_{(g)}$$

The limiting reactant for the above experiment is:

- $SiCl_4$
- SiO_2
- H_2O
- HCl

Save & Next حفظ والتالي

A

(1)

$$\text{mole of } SiCl_4 = \frac{g}{A.A}$$

$$= \frac{50}{(28\times1)+(35.45\times4)} = \boxed{0.29\,m}$$

$$1\,SiCl_4 \longrightarrow 1\,SiO_2$$

$$0.29\,SiCl_4 \longrightarrow 0.29\,SiCl_4$$

(2)

$$\text{mole of } H_2O = \frac{g}{A.A} = \frac{20}{16+2}$$

$$= 1.11\,m$$

$$2\,H_2O \longrightarrow 1\,SiO_2$$

$$1.11\,H_2O \longrightarrow 0.55$$

MKCL OES

Total questions in exam: 40 | Answered: 0

Chemistry_FT_Sem1_2018

Question No. 1

When the substances in the equation below are at equilibrium, at pressure P and temperature T, the equilibrium can be shifted to favor the products by

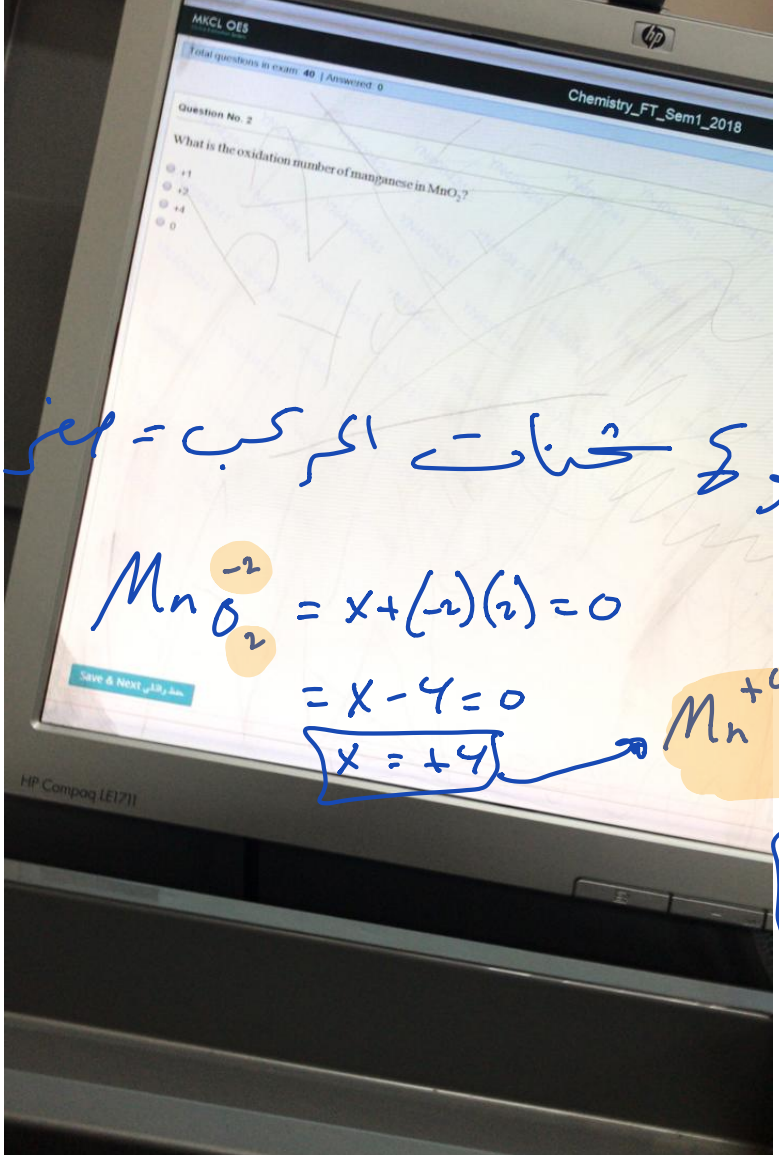
$$CuO(s) + H_2(g) \rightleftharpoons H_2O(g) + Cu(s) + Heat$$

- decreasing the pressure
- decreasing the temperature
- adding more CuO
- increasing the pressure

B

& Next حفظ والتالي

LE1711



* مجموعی حصہ کے مرکب = صفر

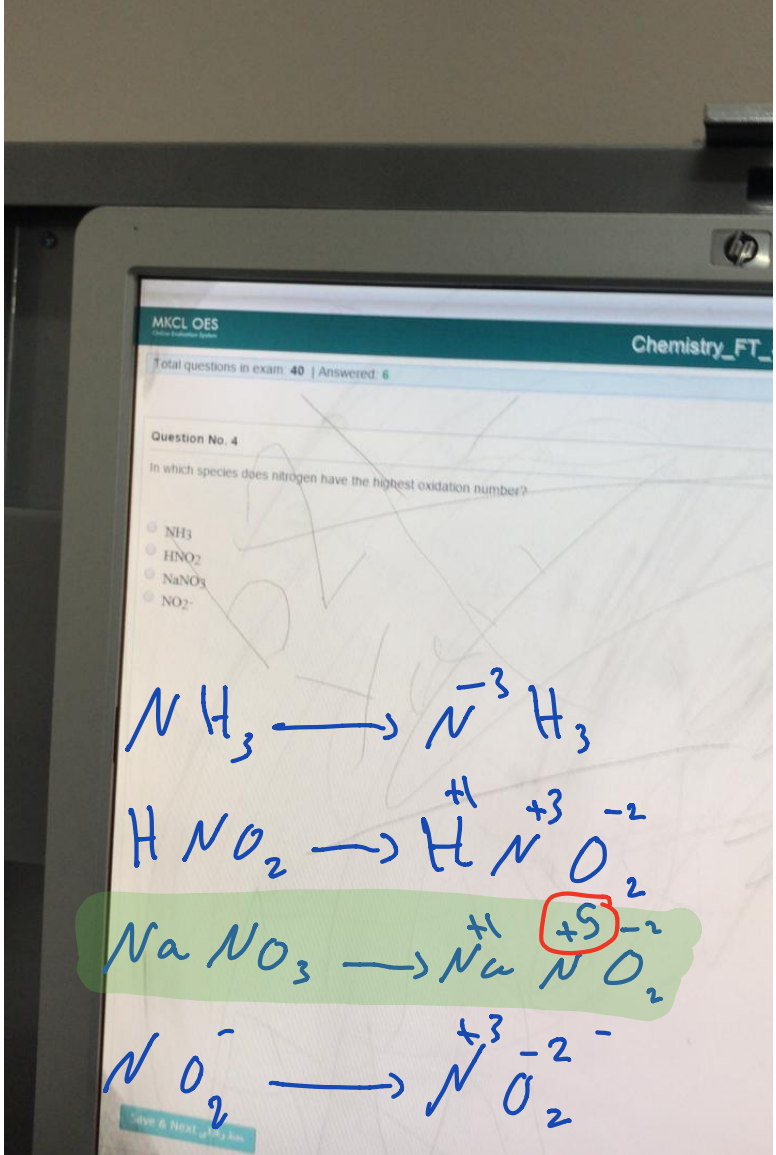
$Mn^{-2}O_2 = x + (-2)(2) = 0$

$= x - 4 = 0$

$x = +4$

$Mn^{+4}O_2^{-2}$

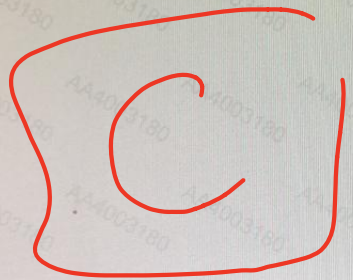
C


MKCL OES
Chemistry_FT_
Total questions in exam: 40 | Answered: 6
Question No. 4
In which species does nitrogen have the highest oxidation number?
NH3
HNO2
NaNO3
NO2-
NH3 → N-3 H3
HNO2 → H+1 N+3 O2-2
NaNO3 → Na+1 N+5 O3-2
NO2- → N+3 O2-2-
C

Question No. 5

For the reaction: $C_{(s)} + H_2O_{(g)} \rightleftharpoons H_{2(g)} + CO_{(g)}$ ΔH is positive (endothermic)
What would be the effect of removing H_2 gas from the reaction vessel?

- More water will be formed.
- The reaction will shift to the left.
- The reaction will shift to the right.
- The reaction will not be affected.



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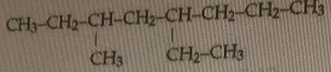
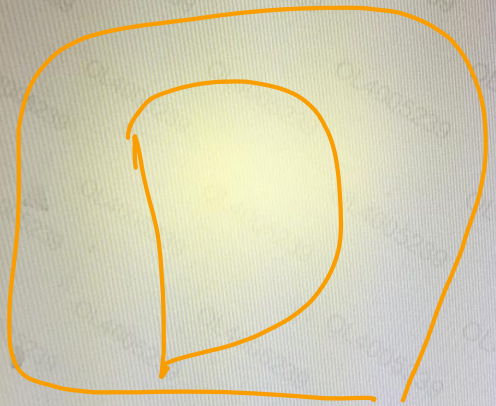
HP L1710

Question No. 2

What is the IUPAC name for the following?



- 4-ethyl-6-methyloctane
- 3-ethyl-5-methyloctane
- isooctane
- 5-ethyl-3-methyloctane



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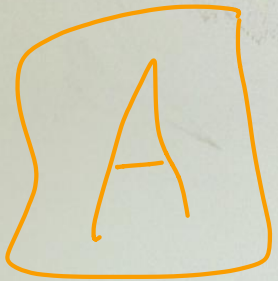
HP Compaq LE1711

Question No. 6

Which of the following solutions is the most basic?

- $[H_3O^+] = 1.0 \times 10^{-10}$ M
- $[OH^-] = 1.0 \times 10^{-10}$ M
- $[OH^-] < 1.0 \times 10^{-10}$ M
- $[H_3O^+] > 1.0 \times 10^{-7}$ M



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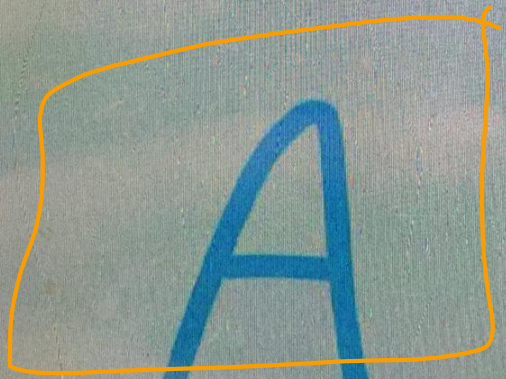
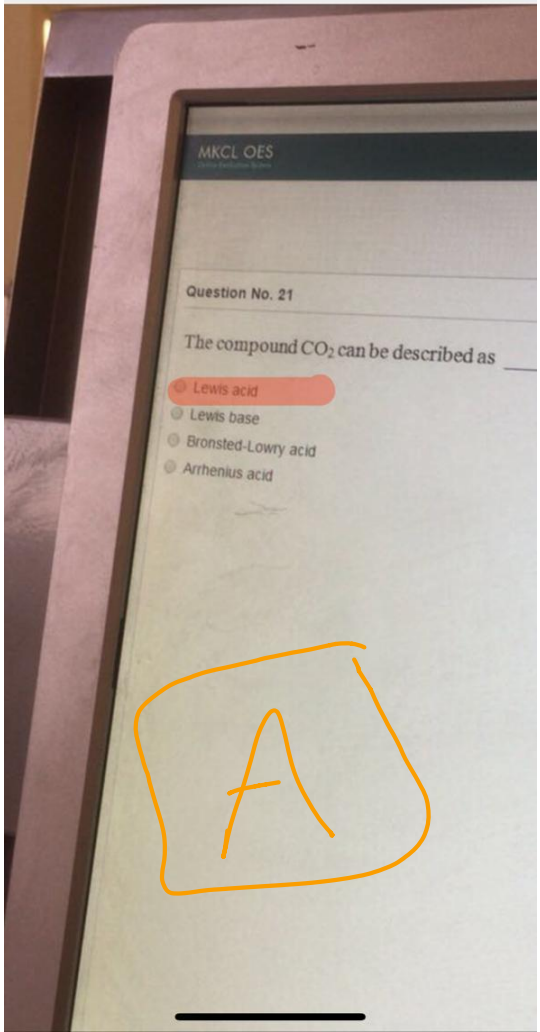
A chemical reaction has reached equilibrium when _____

- the rate of the forward reaction equals the rate of the reverse reaction
- all products have been removed from the reaction mixture
- all reactants have been converted to products
- the concentrations of reactants and products are equal



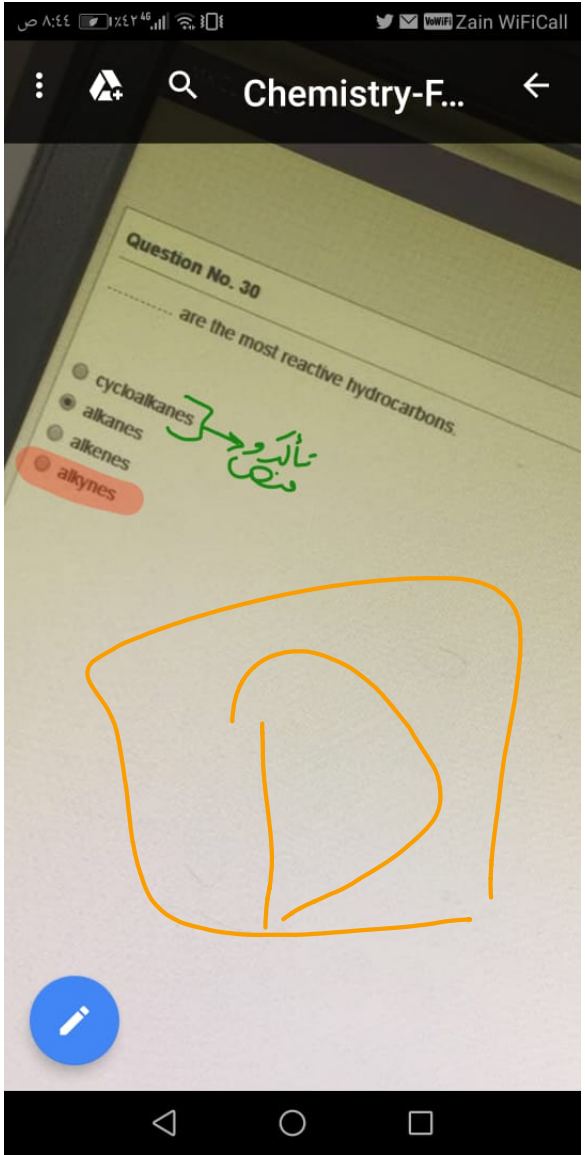


Question No. 9

In the following reaction, what is the effect of adding more NO_2 to the starting reaction mixture?

$$2NO_2\,(g) \rightleftarrows N_2O_4\,(g)$$

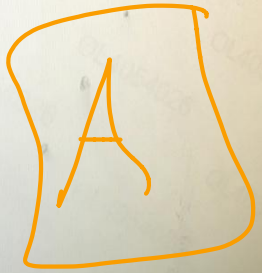
- It would make the reaction more endothermic.
- It would increase the final quantity of products.
- It would make the reaction more exothermic.
- It would decrease the final quantity of products.


Chemistry-F...
Question No. 30
..... are the most reactive hydrocarbons.
cycloalkanes
alkanes
alkenes
alkynes

Question No. 3

What is the name of this compound?

- cyclohexane
- cyclopentane
- cyclooctane
- cycloheptane



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Online Evaluation System

Chemistry_FT_Sem1_

Total questions in exam: **40** | Answered: **0**

Question No. 1

What is the molecular formula of a compound that has a molar mass of 116 g/mol and its empirical formula is C_2H_5?

- C_6H_{15}
- C_2H_5
- C_8H_{20}
- C_6H_{20}

Molecular formula = e · empirical

$$e = \frac{\text{molar mass of molecular}}{\text{molar mass of empirical}}$$

$$e = \frac{116}{29} = 4$$

Molecular formula = $4 \cdot C_2H_5$

$= C_8H_{20}$

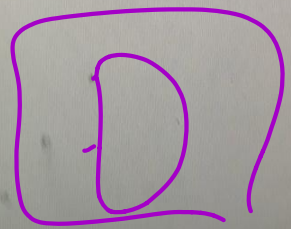
C

Total questions in exam: 40 | Answered: 0

Question No. 4

The main characteristic of all weak electrolyte solutions is that they _____

- do not conduct electricity
- completely ionize in aqueous solutions
- do not dissolve in water
- partially ionize in aqueous solutions



MKCL OES
Online Evaluation System

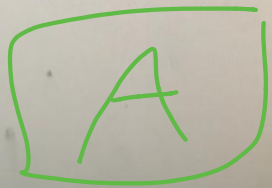
Chemist

Total questions in exam: **40** | Answered: **0**

Question No. 5

The molarity (M) of an aqueous solution containing 22.5 g of sucrose ($C_{12}H_{22}O_{11}$) in 35.5 mL of solution is _____.

- 1.85
- 0.0657
- 0.104
- 3.52



$$\text{Molarity} = \frac{\text{amount of solution}}{\text{Volume}}$$

$$\text{mole} = \frac{22.5}{342} = 0.06\,m$$

$$\text{Volume} = \frac{35.5}{1000} = 0.0355$$

$$\text{molarity} = \frac{0,06}{0,0355} = \boxed{1.69\,M}$$

A

Total questions in exam: **40** | Answered: **0**

Question No. 6

How many grams of CO_2 could be produced when 44 grams of C_3H_7COOH completely react with oxygen gas according to the reaction?

$$C_3H_7COOH + 5O_2 \rightarrow 4CO_2 + 4H_2O$$

- 44 g
- 22 g
- 133 g
- 88 g

① $C_3H_5COOH = \frac{44}{88} = \frac{1}{2}$ mole

② $1\ C_3H_5COOH \longrightarrow 4\ CO_2$

$\frac{1}{2}\ C_3H_5COOH \longrightarrow 2\ CO_2$

③ grams $= 2 \cdot (12\times1) + (16\times2)$

$= 2 \times 44$

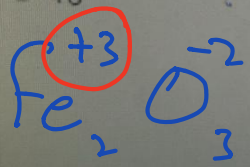
$= 88$

D

Question No. 8

What is the oxidation number of iron in Fe_2O_3?

- +3
- -3
- -6
- +6

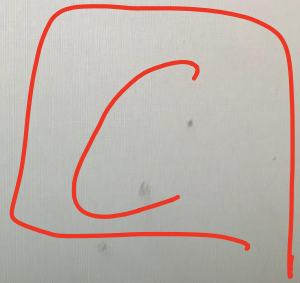


A

Question No. 7

The compound NH_3 can be described as _____.

- Bronsted-Lowry acid
- Arrhenius acid
- Lewis base
- Lewis acid



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Online Evaluation System

Chemistry_FT_Sem1_2018

Total questions in exam: **40** | Answered: **0**

Question No. 9

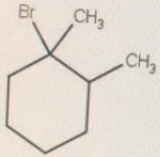
Calculate the volume (in liter) of a solution that contains 3.12 moles of NaCl if the molarity of this solution is 6.67 M NaCl.

- 2.823 L
- 2.141 L
- 0.208 L
- 0.468 L

$$\text{Volume} = \frac{n}{M} = \frac{3.12}{6.67} = 0.468\ \text{L} \rightarrow \text{D}$$

Question No. 10

Choose the correct name for the following compound:



- 1-bromo-1,2-dimethylbenzene
- 1-bromo-1,2-dimethylcyclohexane
- 2-bromo-2-methyltoluene
- 2-bromo-1,2-dimethylcyclohexane

B

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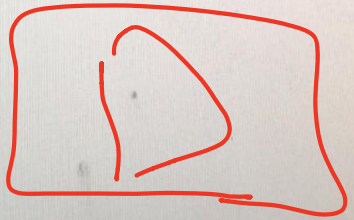
Question No. 11

In the "Basic" solutions, _____

- $pH < 7$ and $[H_3O^+] > 10^{-7}$ M
- $pH = 7$ and $[H_3O^+] = 10^{-7}$ M
- $pH > 7$ and $[H_3O^+] > 10^{-7}$ M
- $pH > 7$ and $[H_3O^+] < 10^{-7}$ M



MKCL OES

Chemistry_F

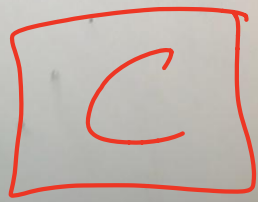
Total questions in exam: **40** | Answered: **0**

Question No. 17

In the reaction below, what is the theoretical yield in moles for NO when 5 moles of NH_3 react with 7 moles of O_2?

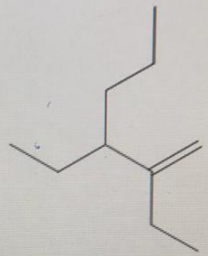
$$4\ NH_3 + 5\ O_2 \rightarrow 4\ NO + 6\ H_2O$$

- 3.6 mol
- 2.4 mol
- 5.0 mol
- 4.8 mol



Question No. 12

Name the following organic compound:



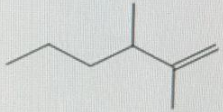
- 4-ethyl-3-methyleneheptane
- 2,3-diethyl-1-hexene
- 2,3-diethyl-1-hexyne
- 2-ethyl-3-propyl-1-pentene

B

Total questions in exam: **40** | Answered: **0**

Question No. 15

Provide the name of the compound below.



- 2,3-dimethyl-1-hexene
- 2,3-dimethyl-2-hexene
- 4,5-dimethyl-5-hexene
- 4,5-dimethyl-6-hexene

A

Total questions in exam: 40 | Answered: 13

Question No. 12

The reaction for the decomposition of PCl_5 to chlorine and PCl_3 is shown below.

$$PCl_5\,(s) \rightleftharpoons PCl_3\,(g) + Cl_2(g)$$

If the equilibrium concentrations are $[PCl_5] = 1.0\ M$, $[PCl_3] = 1.0\ M$, $[Cl_2] = 0.10\ M$, what is the value of the equilibrium constant?

- $K_c = 1.0 \times 10^{2}$
- $K_c = 1.0 \times 10^{-2}$
- $K_c = 1.0 \times 10^{-4}$
- $K_c = 1.0 \times 10^{-1}$

$K_c = [1][0.1]$

$K_c = 0.1 = 1 \times 10^{-1}$ → D

Question No. 13

Consider the reaction: $2\ SO_2(g) + O_2(g) \leftrightarrow 2\ SO_3(g)$

If, at equilibrium at a certain temperature, $[SO_2] = 1.50$ M, $[O_2] = 0.120$ M, and $[SO_3] = 1.25$ M, what is the value of the equilibrium constant K_{eq}?

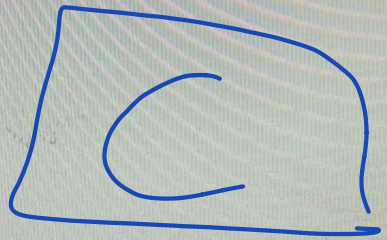
- 0.14
- 6.94
- 5.79
- 8.68

$$K_{eq} = \frac{[1.25]^2}{[1.5]^2[0.12]} = 5.79 \rightarrow \text{C}$$

Question No. 15

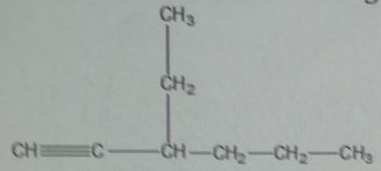
A compound that has a molar mass of 60 g/mol and an empirical formula of CH_2O , its molecular formula is:

- CH_2O.
- C_2H_4O.
- $C_2H_4O_2$.
- $C_3H_6O_3$.

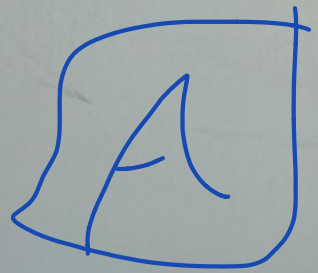


Question No. 18

What is the correct systematic name of the following compound?



- 3-ethyl-1-heptyne
- 4-ethyl-5-hexyne
- 3-ethyl-1-hexyne
- 4-ethyl-1-hexyne



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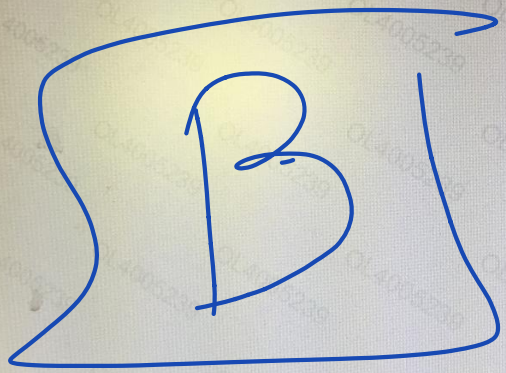
Question No. 4

What is the family of this organic compound?

CH_3-CH_2-NH_2

- phenol
- amine
- amide
- ether



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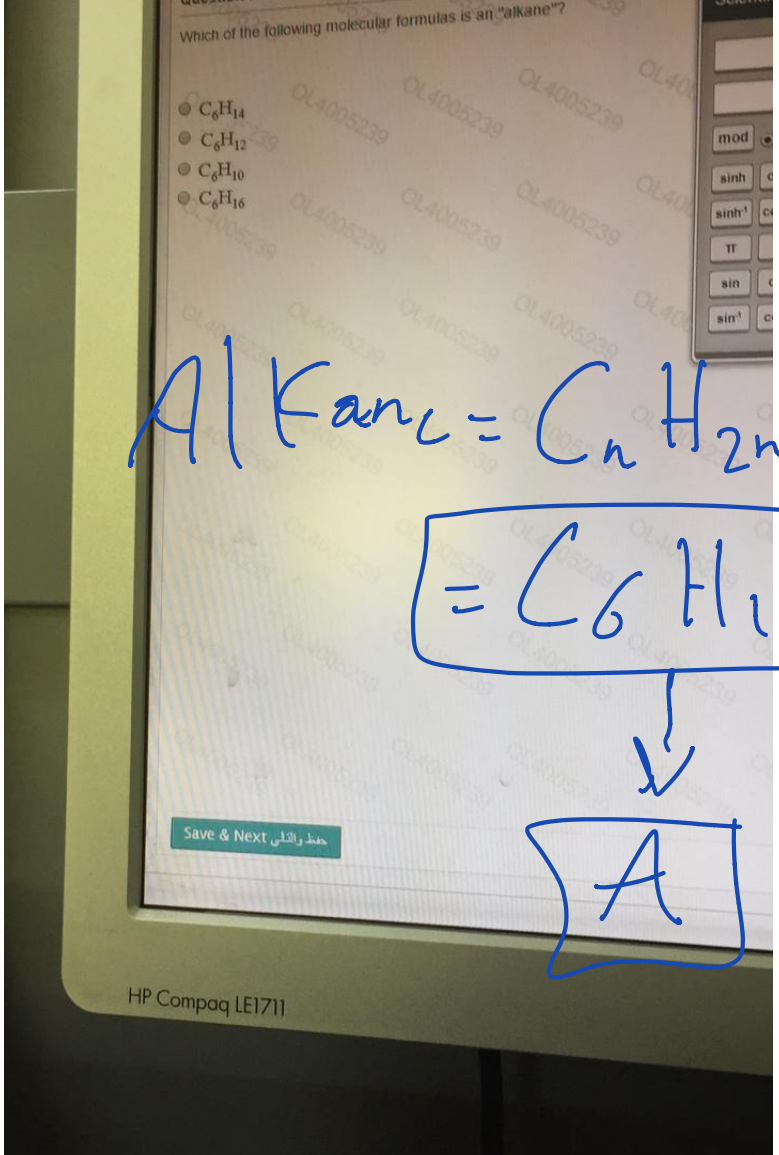
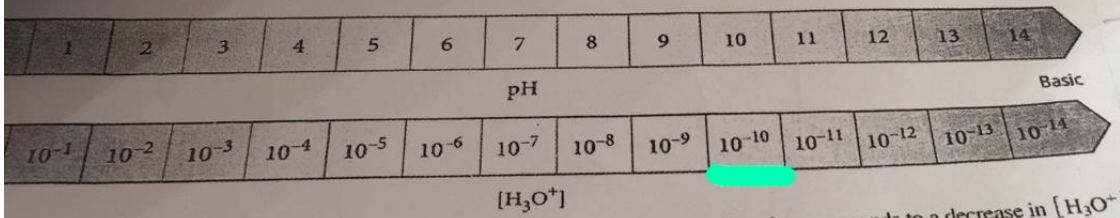
HP Compaq LE1711


Which of the following molecular formulas is an "alkane"?
C_6H_{14}
C_6H_{12}
C_6H_{10}
C_6H_{16}
Save & Next حفظ والتالي
HP Compaq LE1711
Alkane $= C_nH_{2n+2}$
$= C_6H_{14}$
A

pH, start with the equation that defines pH. e of pH and then solve for $[H_3O^+]$. Since the given to two decimal places, the $[H_3O^+]$ is written to two (Remember that $10^{\log x} = x$. Some calculators use an inv ent this function.).

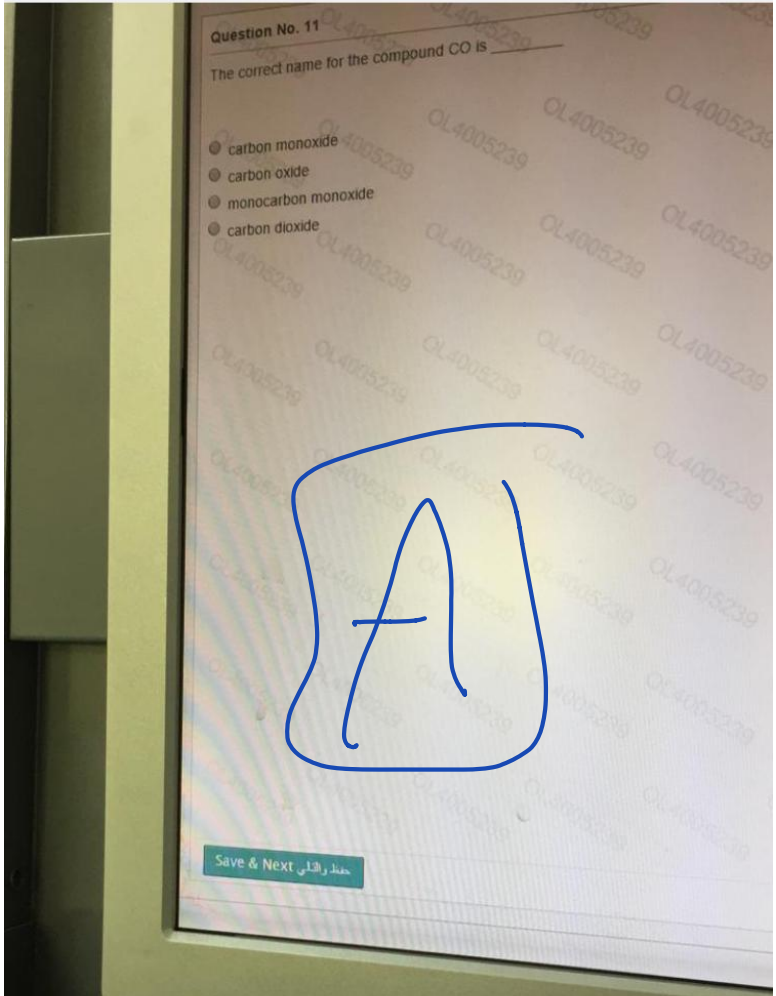
$$\begin{aligned} pH &= -\log[H_3O^+] \\ 4.80 &= -\log[H_3O^+] \\ -4.80 &= \log[H_3O^+] \\ 10^{-4.80} &= 10^{\log[H_3O^+]} \\ 10^{-4.80} &= [H_3O^+] \\ [H_3O^+] &= 1.6 \times 10^{-5}\ M \end{aligned}$$

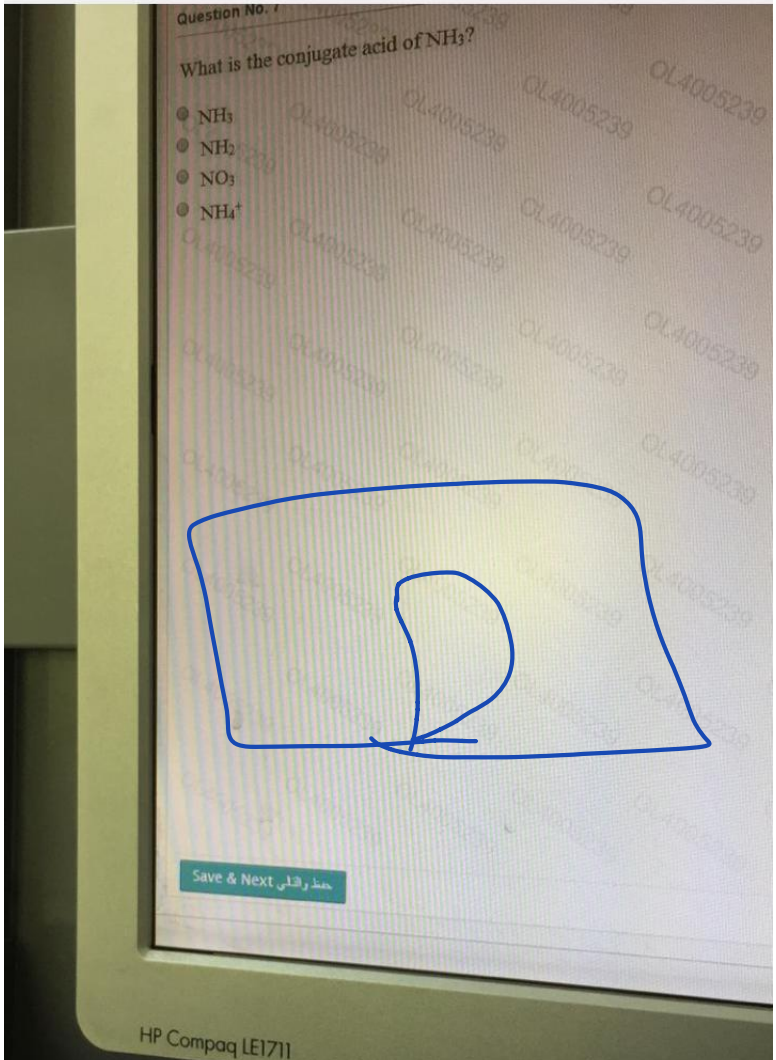
The pH Scale



▲ **Figure 5.11 The pH Scale** An increase of 1 on the pH scale corresponds to a decrease in $[H_3O^+]$ by a factor of 10.


ص 9:00


Question No. 11
The correct name for the compound CO is _____
carbon monoxide
carbon oxide
monocarbon monoxide
carbon dioxide
A
Save & Next حفظ والتالي

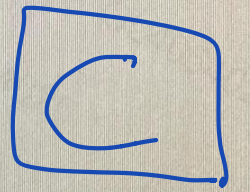

Question No: 7
What is the conjugate acid of NH_3?
NH_3
NH_2^-
NO_3^-
NH_4^+
Save & Next حفظ والتالي
HP Compaq LE1711

Total questions in exam: **40** | Answered: **0**

Question No. 27

After a chemical reaction reaches equilibrium, _____.

- The amount of products is decreasing.
- The amount of products is increasing.
- The amount of reactants and products are constant.
- The amount of reactants and products are equal.

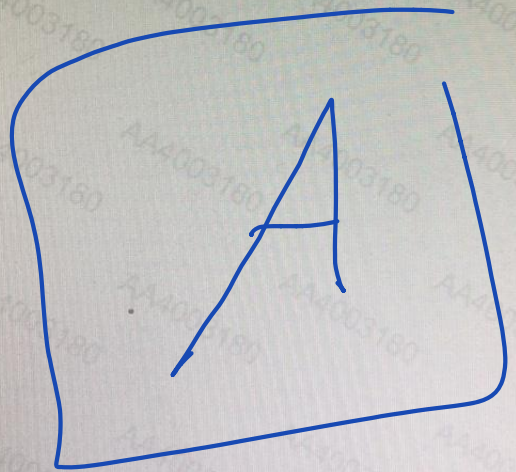


Question No. 6

In the reaction:

$$Cu_{(s)} + 2Ag^{+}_{(aq)} \rightarrow Cu^{2+}_{(aq)} + 2Ag_{(s)}$$

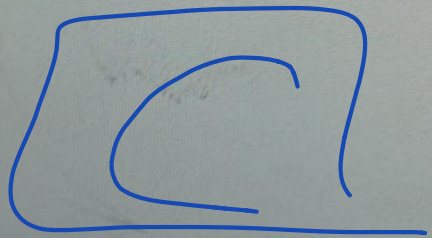
- $Cu_{(s)}$ is the reducing agent and $Ag^{+}_{(aq)}$ is reduced.
- $Ag^{+}_{(aq)}$ is the reducing agent and $Cu_{(s)}$ is reduced.
- $Ag^{+}_{(aq)}$ is oxidizing agent and $Cu_{(s)}$ is reduced
- $Cu_{(s)}$ is the oxidizing agent and $Ag^{+}_{(aq)}$ is oxidized.



Question No. 23

The chemical formula of the compound formed between sodium and fluorine is _____.

- NaF_2
- Na_2F
- NaF
- NaF_3



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Total questions in exam: **40** | Answered: **0**

Question No. 25

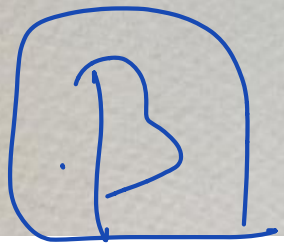
What is the term for the pairs of valence electrons that are not shared in a molecule?

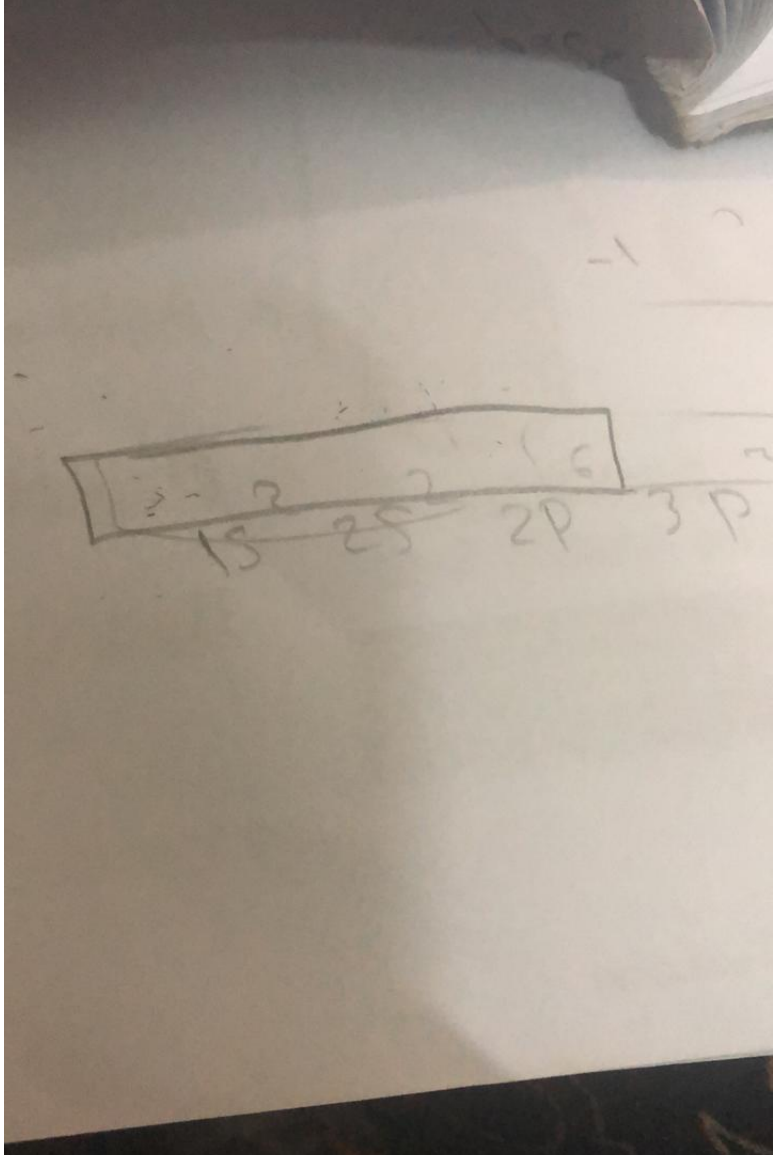
- core electrons
- bonding electrons
- lone pairs of electrons
- sharing electrons

Question No. 12

The substance that causes the reduction of another substance is called:

- anode
- reducing agent
- oxidizing agent
- cathode




1S
2S
2P
3P

Total questions in exam: 40 | Answered: 26

70:59

User: BD4005957

Number of main questions: 40
Number of questions: 40

26 Answered | 14 Not Answered | 0 Not Visited | 0 Partially Answered

Question No. 5

Consider the following reaction at equilibrium, decreasing the pressure will _____.

$$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$$

- shift the reaction to the right
- shift the reaction to the left
- have no effect
- cannot be determined, since the temperature is unknown

1 2 3 4 5 6 7
8 9 10 11 12 13 14
15 16 17 18 19 20 21
22 23 24 25 26 27 28
29 30 31 32 33 34 35
36 37 38 39 40

Calculator | Instructions

End Test

Save & Next حفظ والتالي

A

Question No. 26

The mass percent composition of sulfur in H_2S is:

- 32.7%
- 22.7%
- 94.1%
- 5.9%

$$\frac{32}{34} \times 100 = 94.1\%$$

C

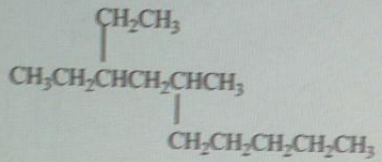
Question No. 20

Calculate the mass of 500 atoms of iron (Fe).

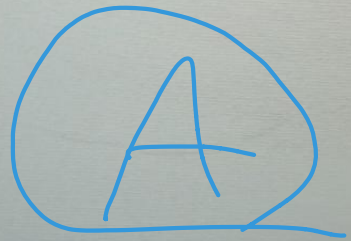
- 56 g Fe
- 6.02×10^{23} g Fe
- 1.22 g Fe
- 4.64×10^{-20} g Fe

Question No. 40

What is the IUPAC name of this compound?



- 3-ethyl-5-methyldecane
- 1-octylpentane
- 3-ethyl-2-pentylhexane
- 8-ethyl-6-methyldecane



Save & Next

Total questions in exam: **40** | Answered: **11**

Question No. 34

What is the correct equilibrium constant expression for the following reaction?

$$2\ Cu(s) + O_2(g) \rightarrow 2\ CuO(s)$$

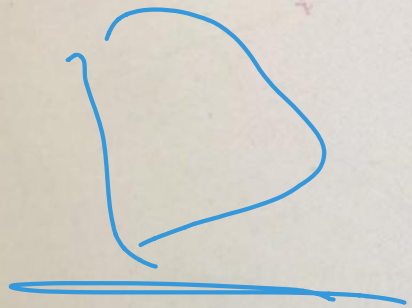
- $K_{eq} = [CuO]^2 / [O_2]$
- $K_{eq} = [CuO]^2 / [Cu]^2[O_2]$
- $K_{eq} = 1 / [O_2]$
- $K_{eq} = [O_2]$

DES

uestion No. 42

When heat (q) has positive value, this means that _____.

- The system loses thermal energy.
- The work (w) = 0.
- The surrounding gains thermal energy.
- The system gains thermal energy.



MKCL OES

Chemistry_FT_Sem1_2016

Total questions in exam: 40 | Answered: 0

Question No. 3

The compound below is an

- acid
- amide
- ester
- amine

Save & Next حفظ و التالي

MKCL OES
Online Evaluation System

Chem

Total questions in exam: **40** | Answered: **5**

Question No. 12

If a drain cleaning solution has a pH = 13, this solution is _____.

- strongly acidic
- strongly basic
- weakly basic
- weakly acidic

Save & Next सेव्ह व पुढील

MKCL OES

Chemistry_F

Total questions in exam: 40 | Answered: 11

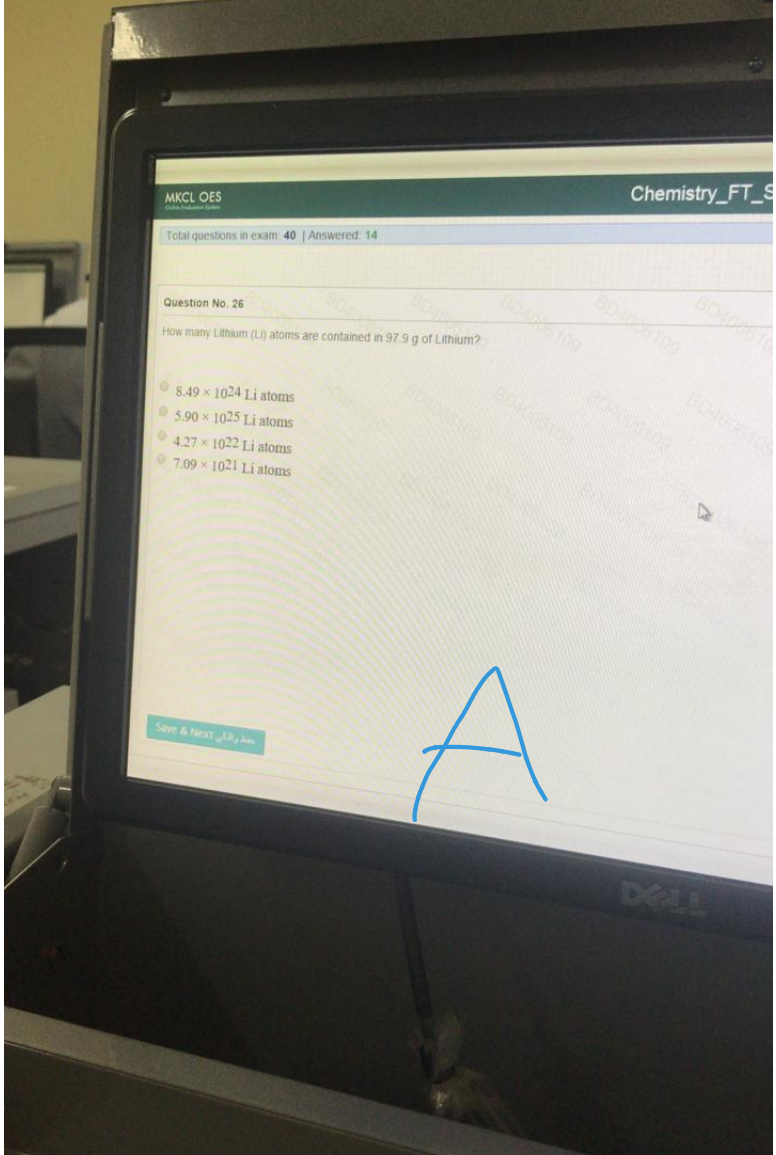
Question No. 14

The following structure corresponds to a ..... alcohol.

- Primary
- Tertiary
- Quaternary
- Secondary

B

Save & Next


MKCL OES
Chemistry_FT_S
Total questions in exam: 40 | Answered: 14
Question No. 26
How many Lithium (Li) atoms are contained in 97.9 g of Lithium?
8.49 × 10^24 Li atoms
5.90 × 10^25 Li atoms
4.27 × 10^22 Li atoms
7.09 × 10^21 Li atoms
Save & Next
A

Total questions in exam: **40** | Answered: **11**

Question No. 32

Which of the following pairs is NOT a conjugate acid-base pair according to the concept of Bronsted-Lowry?

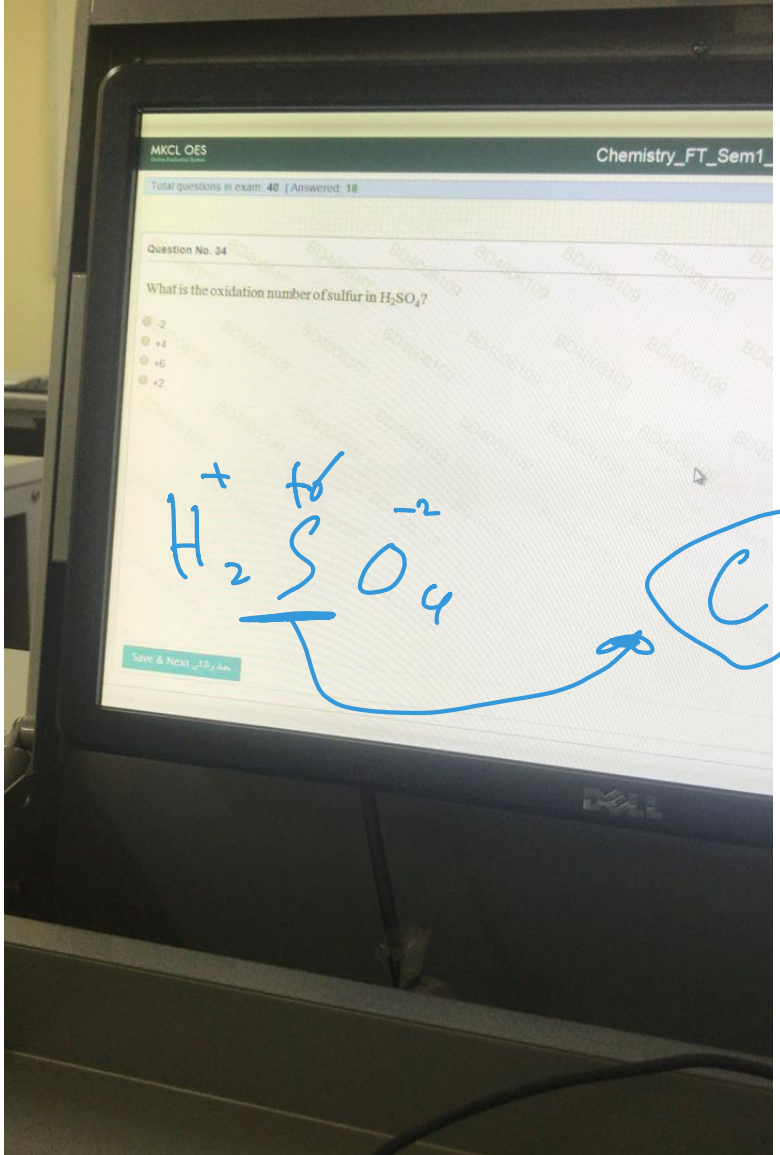
- H_3PO_4 and HPO_4^{2-}
- H_3PO_4 and $H_2PO_4^-$
- $H_2PO_4^-$ and HPO_4^{2-}
- HPO_4^{2-} and PO_4^{3-}

Total questions in exam: **40** | Answered: **0**

Question No. 22

The name of the chemical compound Cu_2CO_3 is:

- copper(II) carbonate
- copper(III) carbonate
- copper(I) carbonate
- copper carbonate


MKCL OES
Chemistry_FT_Sem1_
Total questions in exam: 40 | Answered: 18
Question No. 34
What is the oxidation number of sulfur in H_2SO_4?
-2
+4
+6
+2
Save & Next
+
+6
-2
H_2SO_4
C

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Online Evaluation System

Chemi

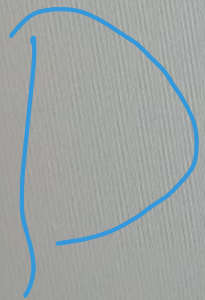
Total questions in exam: **40** | Answered: **0**

Question No. 5

Which of the changes listed below will shift the equilibrium position to the *ri*
the following reversible reaction?

$$CH_4(g) + H_2O(g) + \text{heat} \rightleftarrows CO(g) + 3\ H_2(g)$$

- A decrease of volume
- A decrease of $[CH_4]$
- A decrease of temperature
- A decrease of $[CO]$



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Online Evaluation System

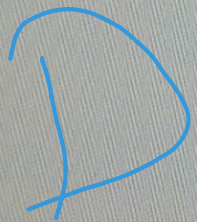
Chemi

Total questions in exam: **40** | Answered: **0**

Question No. 2

Which of the following expression symbols is used for quantifying acidity and basicity?

- aH
- bH
- eH
- pH



MKCL OES

Chemistry_FT_Sem

Total questions in exam: **40** | Answered: **0**

Question No. 7

Which of the following is the electron dot formula (Lewis structure) for an atom of strontium?

(a) Sr· (b) $\dot{S}r\cdot$ (c) $\cdot\dot{S}r:$ (d) $\cdot\ddot{S}r:$

- (a)
- (b)
- (c)
- (d)

Save & Next

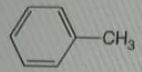
MKCL OES

Chemis

Total questions in exam: **40** | Answered: **0**

Question No. 1

What is the name of compound shown below?



- benzene
- phenol
- toluene
- aniline

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Chemistry_FT

Total questions in exam: **40** | Answered: 0

Question No. 8

What is the $[OH^-]$ in a solution that has a $[H_3O^+] = 1 \times 10^{-6}$ M?

- 1×10^{-2} M
- 1×10^{-6} M
- 1×10^{-10} M
- 1×10^{-8} M

$$[H_3O^+] \cdot [OH^-] = 1 \times 10^{-14}$$

$$[OH^-] = \frac{1 \times 10^{-14}}{1 \times 10^{-6}}$$

$$[OH^-] = 1 \times 10^{-8} \rightarrow (1)$$

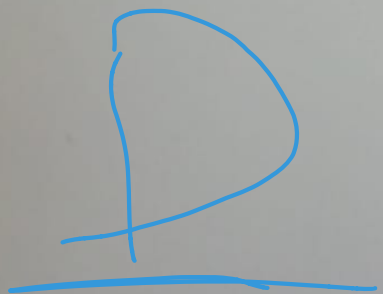
Save & Next

Total questions in exam: **40** | Answered: **0**

Question No. 3

Which one of the following is a Lewis base?

- BF_3
- $AlCl_3$
- NH_4^+
- NH_3



MKCL OES

Chemistry_FT_Sem1_

Total questions in exam: 40 | Answered: 27

Question No. 38

Consider the following reaction at equilibrium. What effect will increasing the temperature have on the system?

$$C_3H_8(g) + 5\,O_2(g) \rightleftharpoons 3\,CO_2(g) + 4\,H_2O(l) + \text{heat}$$

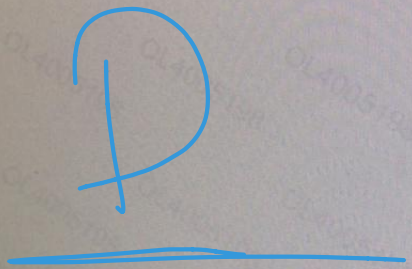
- The reaction will shift to the right in the direction of products.
- The equilibrium constant will increase.
- The reaction will shift to the left in the direction of reactants.
- More CO_2 will be formed.

Save & Next

Question No. 3

What is the oxidation number of nitrogen in NO_3^{-1}?

- 0
- -3
- -5
- +5



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MKCL OES
Online Evaluation System

Chemistry_FT_Sem1_

Total questions in exam: **40** | Answered: **7**

Question No. 5

What is the chemical formula of the product formed by the reaction between aluminum and oxygen?

- AlO
- Al_3O_2
- Al_2O_3
- Al_3O

Al^{+3} O^{-2} → Al_2O_3

(C)

Save & Next حفظ والتالي

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Online Evaluation System

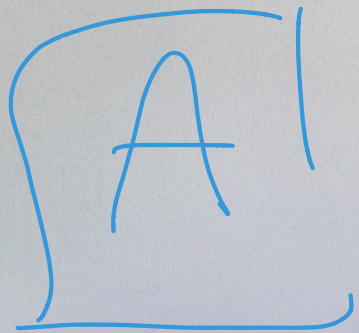
Chemistry_FT_

Total questions in exam: **40** | Answered: **2**

Question No. 5

What is the term for the concentration expression that relates the moles of solute dissolved in each liter

- molarity (M)
- mass/mass percent (m/m %)
- molality (m)
- parts per million (ppm)



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Question No. 39

What are the correct coefficients needed to balance the reaction below?

$$N_2H_4 + H_2O_2 \rightarrow N_2 + H_2O$$

- 2, 4, 2, 8
- 1, 1, 1, 1
- 1, 4, 1, 4
- 1, 2, 1, 4

Save & Next حفظ والتالي

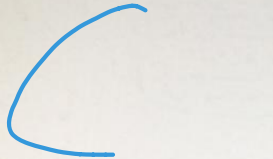
HP Compaq LE1711

Question No. 38

Which of the following is the electron dot formula (Lewis structure) for an atom of oxygen?

(a) O· (b) Ö· (c) ·Ö· (d) ·Ö:

(c)
(b)
(d)
(a)



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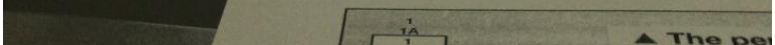
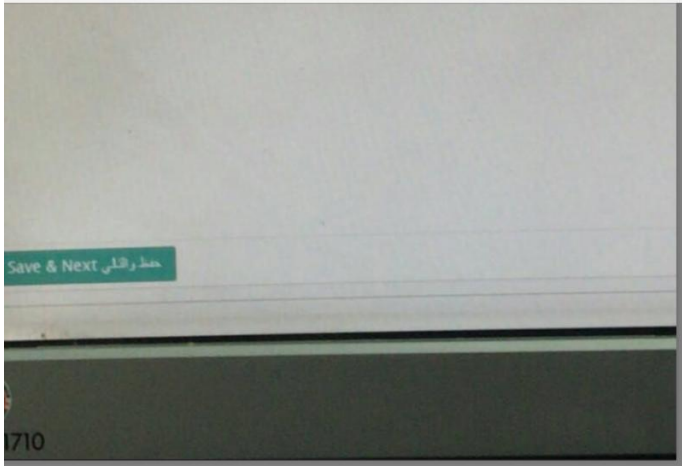
HP Compaq LE1711

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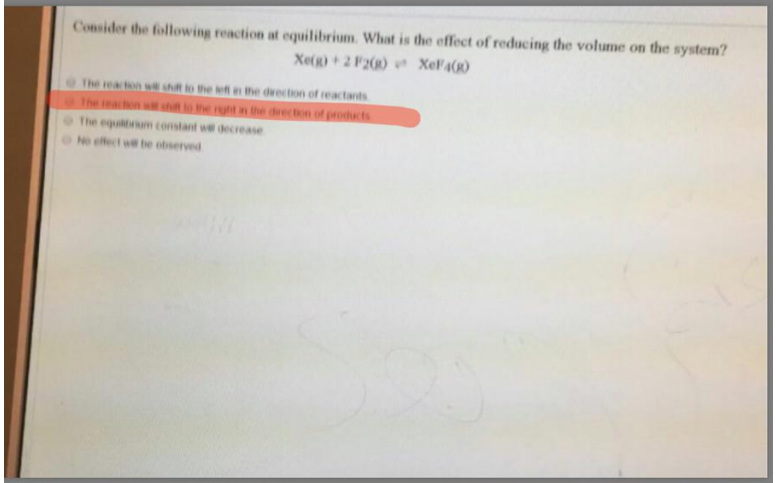
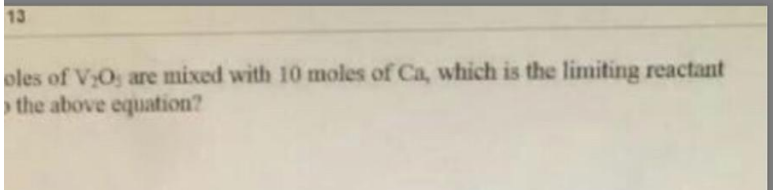
1710

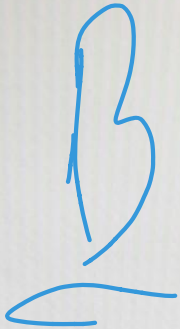

Consider the following reaction at equilibrium. What is the effect of reducing the volume on the system?
Xe(g) + 2 F2(g) ⇌ XeF4(g)
The reaction will shift to the left in the direction of reactants.
The reaction will shift to the right in the direction of products.
The equilibrium constant will decrease.
No effect will be observed.


13
oles of V2O5 are mixed with 10 moles of Ca, which is the limiting reactant
the above equation?

Question No. 13

Which of the following pairs of systematic names and common names is correctly matching?

- toluene = hydroxybenzene
- aniline = aminobenzene
- acetylene = ethene
- phenol = methylbenzene



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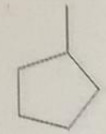
HP Compaq LE1711

Question No. 12

Provide the name of the compound below.



- methylcyclohexane
- ethylcyclopentane
- methylcyclopentane
- methylcyclopropane

Scientific Calculator

M

mod Deg Rad MC

sinh cosh tanh Exp ( )

$\sinh^{-1}$ $\cosh^{-1}$ $\tanh^{-1}$ $\log_2 x$ ln log 7

π e n! $\log_y x$ e^x 10^x 4

sin cos tan x^y x^3 x^2 1

Save &


Save & Next حفظ وتالي
HP Compaq [E171]
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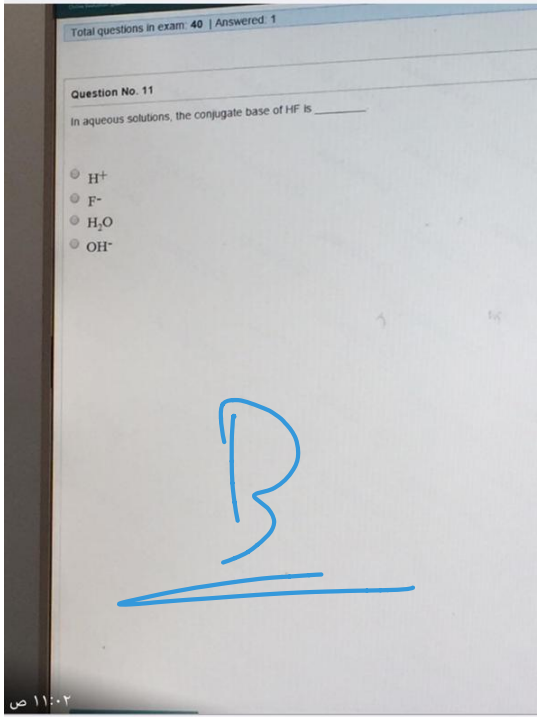
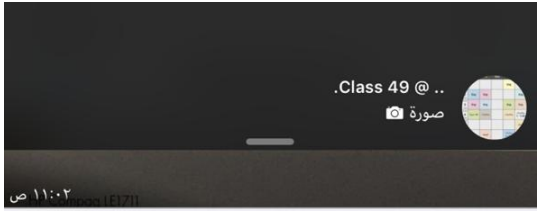

Total questions in exam: 40 | Answered: 1
Question No. 11
In aqueous solutions, the conjugate base of HF is _____.
H^+
F^-
H_2O
OH^-
B
١١:٠٢ ص


MKCL OES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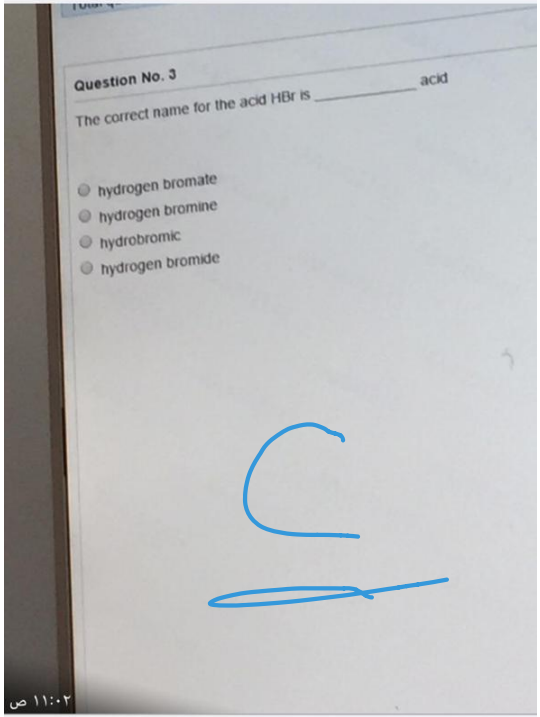
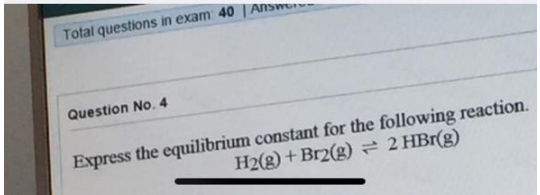
صورة
١١:٠٢ ص


Question No. 3
The correct name for the acid HBr is _____ acid
hydrogen bromate
hydrogen bromine
hydrobromic
hydrogen bromide
١١:٠٢ ص


Total questions in exam: 40
Question No. 4
Express the equilibrium constant for the following reaction.
$H_2(g) + Br_2(g) \rightleftharpoons 2\,HBr(g)$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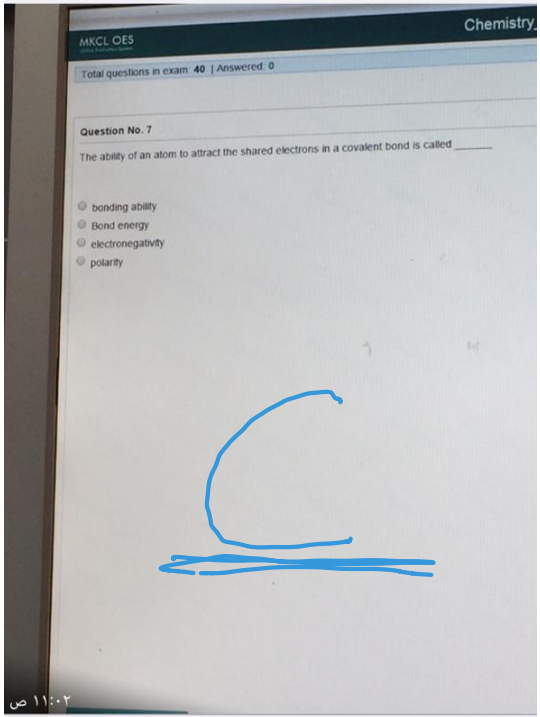
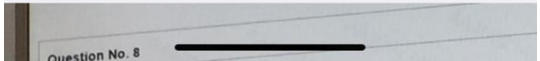
١١:٠٢ ص


MKCL OES
Chemistry
Total questions in exam: 40 | Answered: 0
Question No. 7
The ability of an atom to attract the shared electrons in a covalent bond is called _____
bonding ability
Bond energy
electronegativity
polarity
١١:٠٢ ص


Question No. 8

MKCL OES

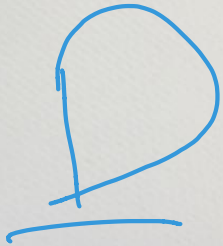
Chemistry_FT_Sem1_

Total questions in exam: **40** | Answered: **12**

Question No. 18

What is the molecular formula of a compound that has a molar mass of 180 g/mol and its empirical formula is CH_2O?

- $C_2H_2O_2$
- $C_4H_8O_4$
- $C_5H_{10}O_5$
- $C_6H_{12}O_6$



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MKCL OES

Chemistry_FT_Sem1_2018

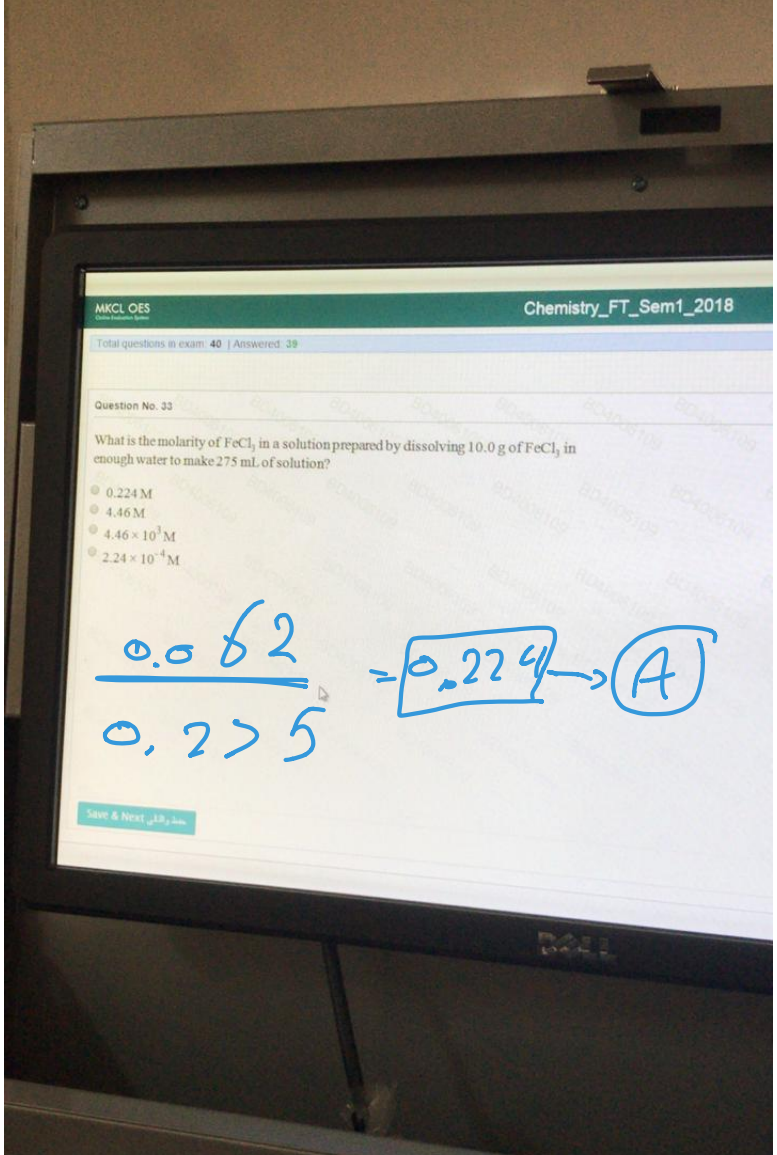
Total questions in exam: 40 | Answered: 35

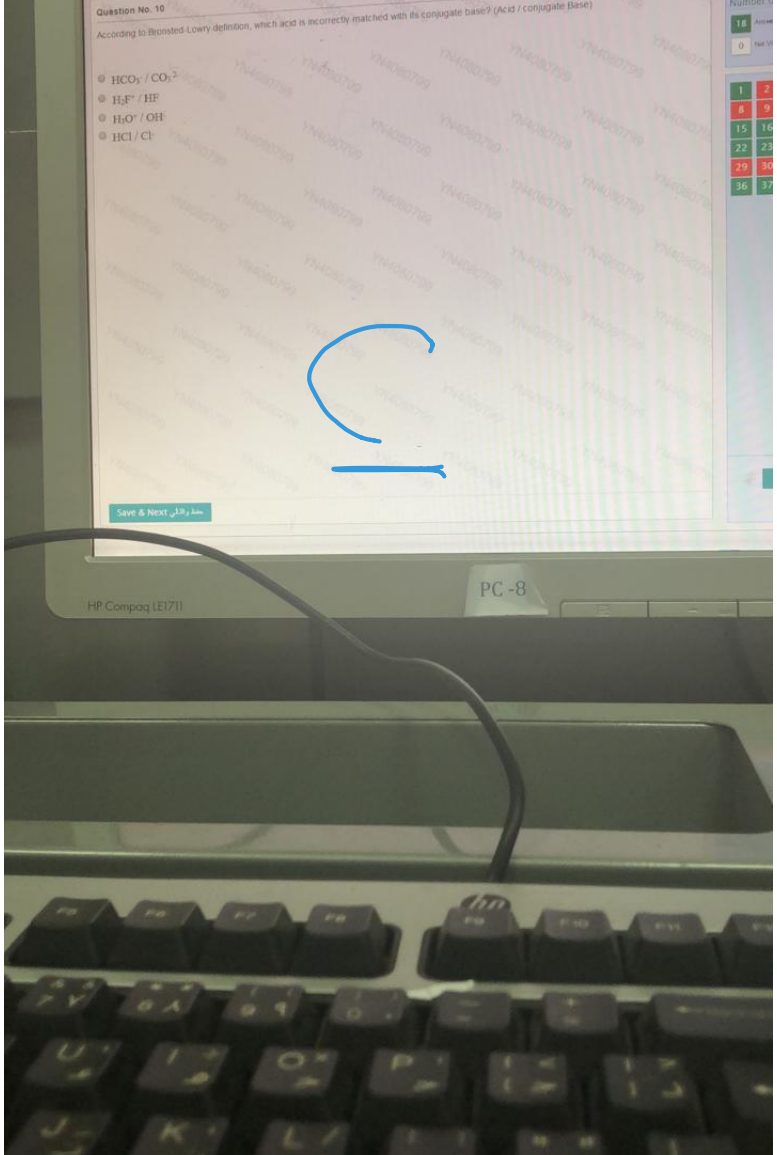
Question No. 21

What is the term for the concentration expression that relates the moles of solute dissolved in each liter of solution?

- mass/mass percent (m/m %)
- molality (m)
- parts per million (ppm)
- molarity (M)

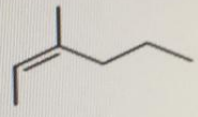
Save & Next


MKCL OES
Chemistry_FT_Sem1_2018
Total questions in exam: 40 | Answered: 39
Question No. 33
What is the molarity of $FeCl_3$ in a solution prepared by dissolving 10.0 g of $FeCl_3$ in enough water to make 275 mL of solution?
0.224 M
4.46 M
4.46×10^{3} M
2.24×10^{-4} M
0.062
0.275
= 0.224 → (A)
Save & Next

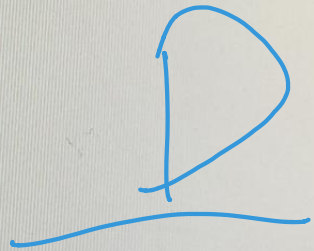

Question No. 10
According to Bronsted-Lowry definition, which acid is incorrectly matched with its conjugate base? (Acid / conjugate Base)
HCO_3^- / CO_3^{2-}
H_2F^+ / HF
H_3O^+ / OH^-
HCl / Cl^-
Save & Next
HP Compaq LE1711
PC -8

Question No. 4

What is the name of the following compound?



- 4-ethyl-4-hexene
- 3-methylenehexane
- 3-methyl-3-hexene
- 3-methyl-2-hexene



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Total questions in exam: 40 | Answered: 16

Question No. 1

An aqueous solution of _____ is considered as strong electrolyte, thus, it can conduct electricity.

- CO_2
- $C_6H_{12}O_6$
- LiCl
- C_8H_{18}

Save & Next حفظ والتالي

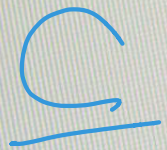
Total questions in exam: 40 | Answered: 16

Question No. 7

Identify the Bronsted-Lowry conjugate acid in the following reaction.

$$H_2O + CO_3^{2-} \rightarrow HCO_3^- + OH^-$$

- H_2O
- OH^-
- HCO_3^-
- CO_3^{2-}



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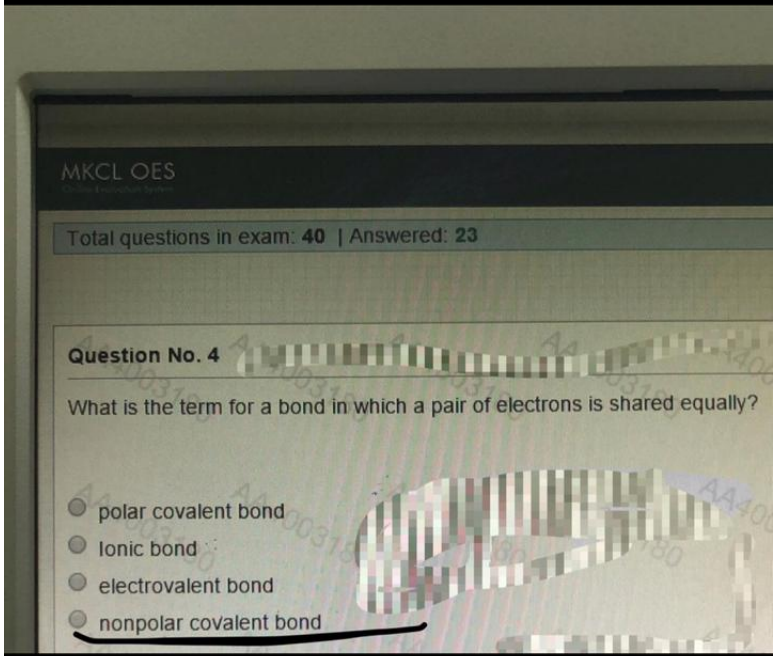
MKCL OES
Total questions in exam: 40 | Answered: 23
Question No. 4
What is the term for a bond in which a pair of electrons is shared equally?
polar covalent bond
Ionic bond
electrovalent bond
nonpolar covalent bond

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Chemistry_FT_Se

Total questions in exam: **40** | Answered: **17**

Question No. 16

How many grams of NaCl are there in 55.0 mL of a 1.90 M aqueous solution of NaCl?

- 0.105
- 3.21
- 12.2
- 6.11

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Question No. 30

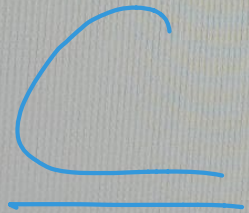
Which of the following is a general property of an acidic solution?

- tastes sour
- turns litmus paper to red
- pH is less than 7
- all are correct

Question No. 25

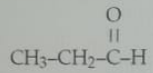
What is the molecular formula of a compound that has a molar mass of 68 g/mol and its empirical formula is HO?

- H_2O
- H_2O_3
- H_4O_4
- H_2O_4



Question No. 28

What is the family of this organic compound?



- aldehyde
- ketone
- carboxylic acid
- ester

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Total questions in exam: 40 | Answered: 12

Question No. 20

Which of the following substances contains a nonpolar covalent bond?

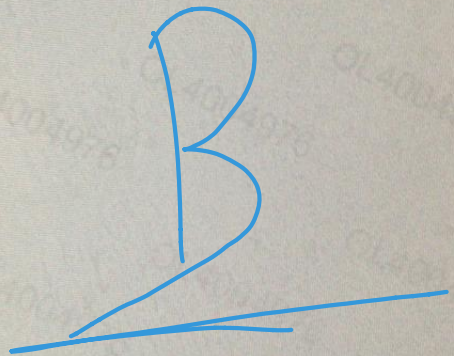
- H_3O^+
- NaCl
- NH_3
- N_2

Total questions in exam: 40 | Answered: 6

Question No. 4

What is the IUPAC name for CH_3-CH_2-C≡CH?

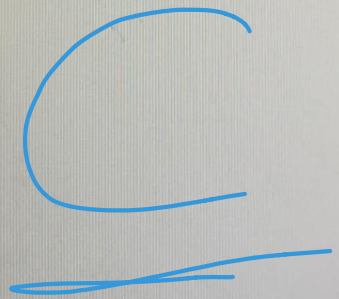
- 3-butyne
- 1-butyne
- 2-butyne
- butyne



Question No. 28

Which of the following pairs of species is NOT a conjugate acid-base pair?

- H_2O and OH^-
- H_2SO_4 and HSO_4^-
- NH_3 and NH_2
- HSO_4^- and SO_4^{2-}



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Online Evaluation System

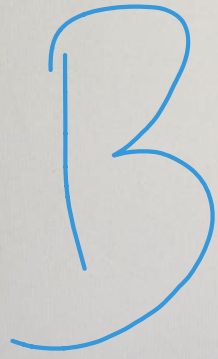
Chem

Total questions in exam: **40** | Answered: **24**

Question No. 36

If the reaction is endothermic, which of the following is always true?

- the reaction rate is fast
- the reaction takes in heat
- the reaction gives out heat
- the reaction rate is slow

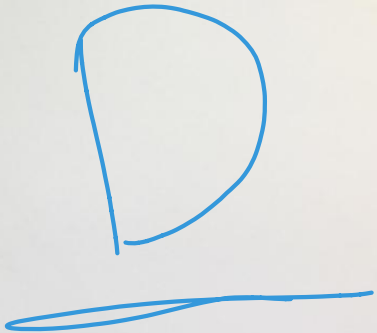


Total questions in exam: 40 | Answered: 25

Question No. 22

Which of these organic compounds is "unsaturated"?

- C_3H_8
- CH_3OH
- CH_4
- C_5H_{10}



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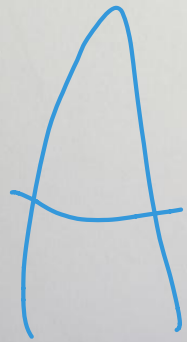
Question No. 10

The mass% of H in Eethane (C_2H_6) is _____.

- 20.1
- 74.9
- 79.9
- 4.0

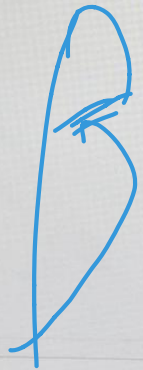


Total questions in exam: **40** | Answered: **12**

Question No. 11

What mass (g) of NaBr is contained in 0.25 L of a sodium bromide solution that has a molarity of 1.20 M?

- 2.32 g
- 30.9 g
- 37.3 g
- 4.93 g



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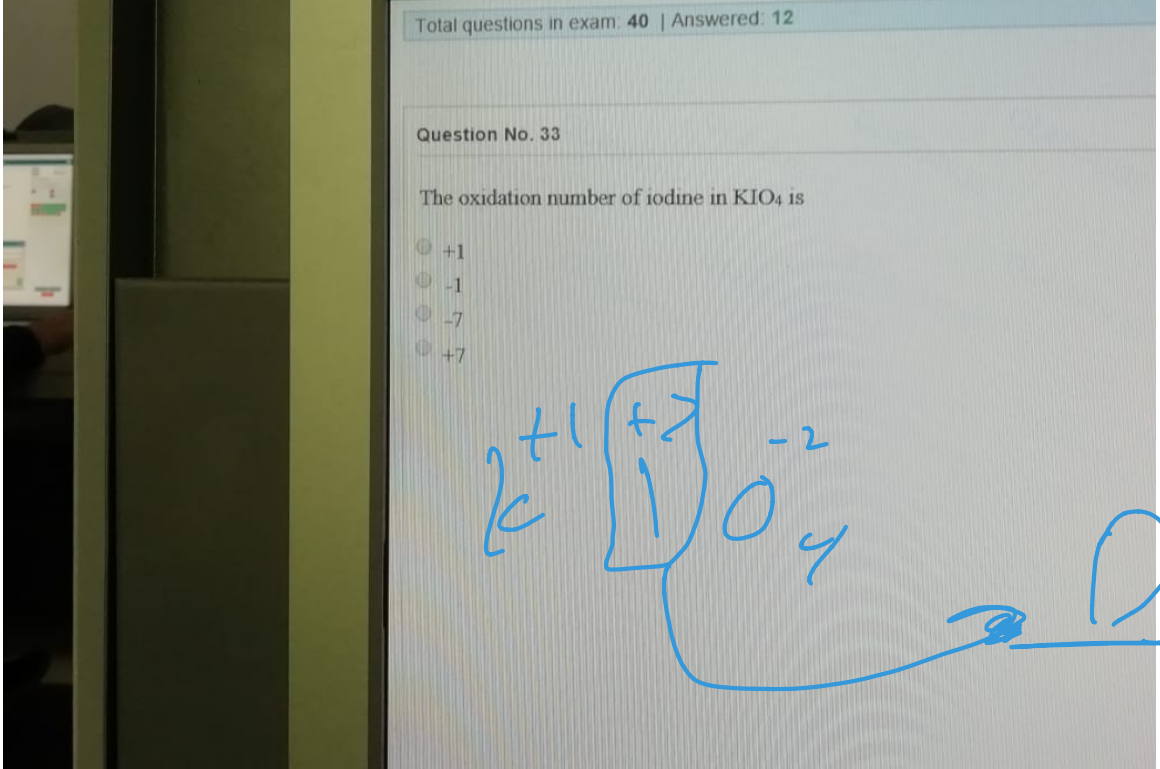
P Compaq LE1711

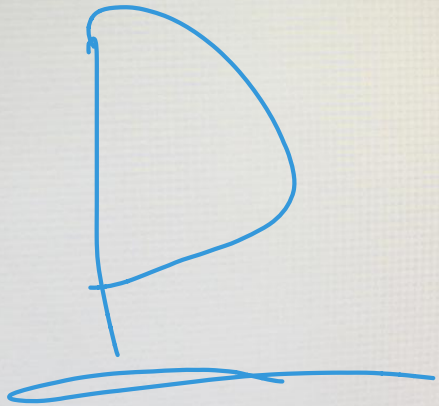

Total questions in exam: 40 | Answered: 12
Question No. 33
The oxidation number of iodine in KIO_4 is
+1
-1
-7
+7

Question No. 12

In the following reaction, what is the effect on the direction of the reaction if more SO_3 is added to the reaction mixture?

$$2SO_2\,(g) + O_2\,(g) \rightleftharpoons 2SO_3\,(g)$$

- The equilibrium shifts to produce more products.
- The rate of formation of products is increased.
- The position of the equilibrium remains unchanged.
- The equilibrium shifts to produce more reactants.



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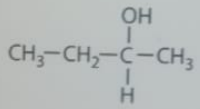
Total questions in exam: **40** | Answered: **31**

Question No. 15

What is the type of the following alcohol?



- Primary
- Secondary
- Tertiary
- Quaternary

Question No. 27

Refer to the equilibrium shown below. Which of the following will shift the reaction to the right?

$$CH_4\,(g) + 2O_2\,(g) \rightleftharpoons CO_2\,(g) + 2H_2O\,(g)$$

- adding excess oxygen
- increasing the pressure
- removing carbon dioxide as soon as it is formed
- adding O_2 and removing CO_2

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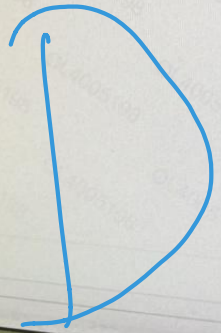

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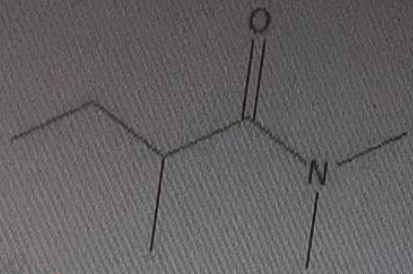
Question No. 38

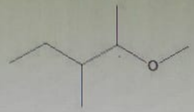
Which family does the following organic compound belong to?



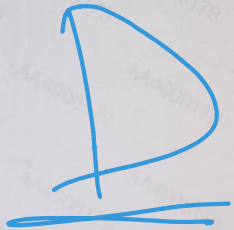
- amine
- ether
- carboxylic acid
- amide

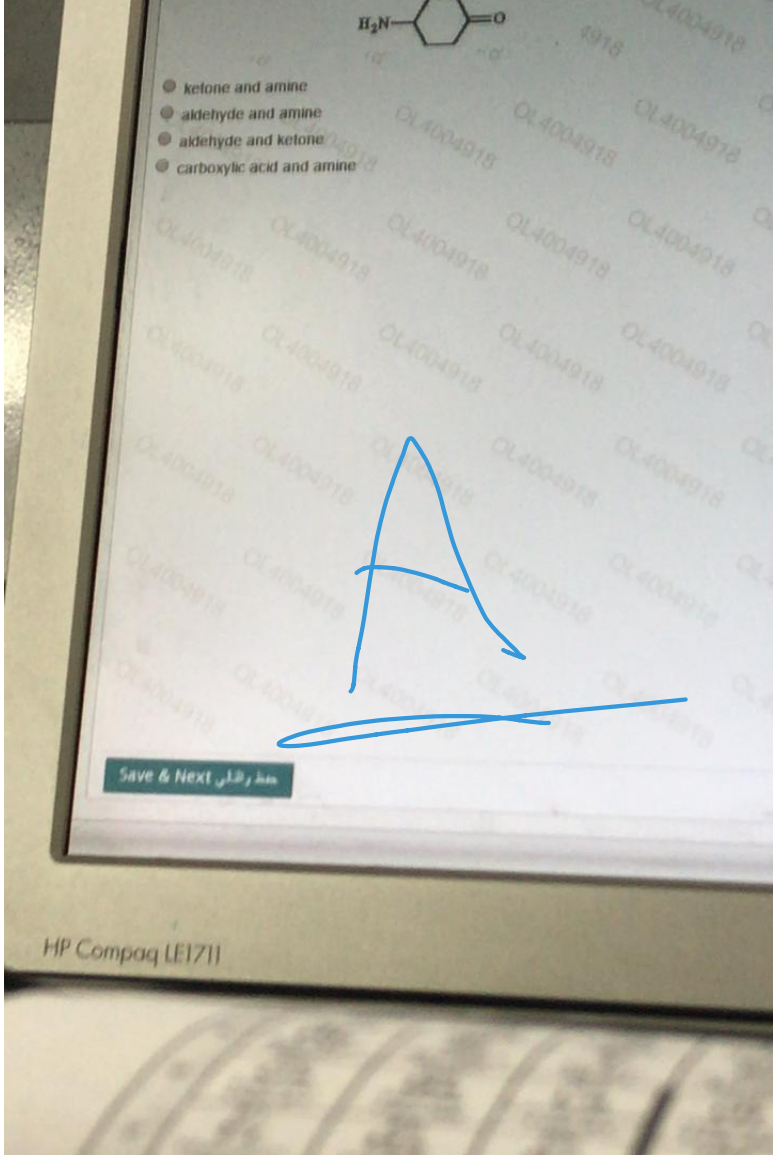
Question No. 2

To which family does the following organic compound belong?



- alcohol
- aldehyde
- carboxylic acid
- ether




H_2N
O
ketone and amine
aldehyde and amine
aldehyde and ketone
carboxylic acid and amine
Save & Next
HP Compaq LE1711

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Chemistry_FT_Sem1_2018

Total questions in exam: 40 | Answered: 27

Question No. 12

Calculate the mass percent composition of phosphorous in K_3PO_4.

- 30.2 %
- 26.75%
- 55.3 %
- 14.6 %

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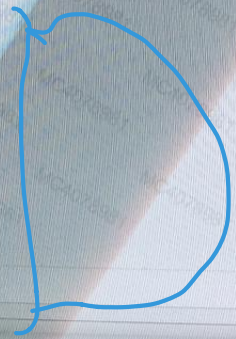

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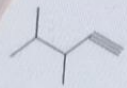
Chemistry_FT_Sem1_2018

Total questions in exam: 40 | Answered: 27

Question No. 15

Provide the name of the compound below.



- 2,3-dimethyl-5-pentyne
- 3,4-dimethyl-1-pentyne
- 2,3-dimethyl-4-pentyne
- 3,4-dimethyl-2-pentyne

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1 8 15 22 29 36

DELL

Chemistry_FT_Sem1_20

Total questions in exam: 40 | Answered: 27

Question No. 21

How many moles of cesium (Cs) are contained in 595 kg of cesium?

- 1.26×10^3 moles Cs
- 4.48×10^3 moles Cs
- 7.91×10^4 moles Cs
- 2.23×10^2 moles Cs

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B

DELL

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Chemistry_FT_Sem1_2018

Total questions in exam: 40 | Answered: 27

Question No. 23

What is the coefficient of chlorine gas after balancing the following equation?

$$\_\_Fe(s) + \_\_Cl_2(g) \rightarrow \_\_FeCl_3(s)$$

- 4
- 1
- 3
- 2

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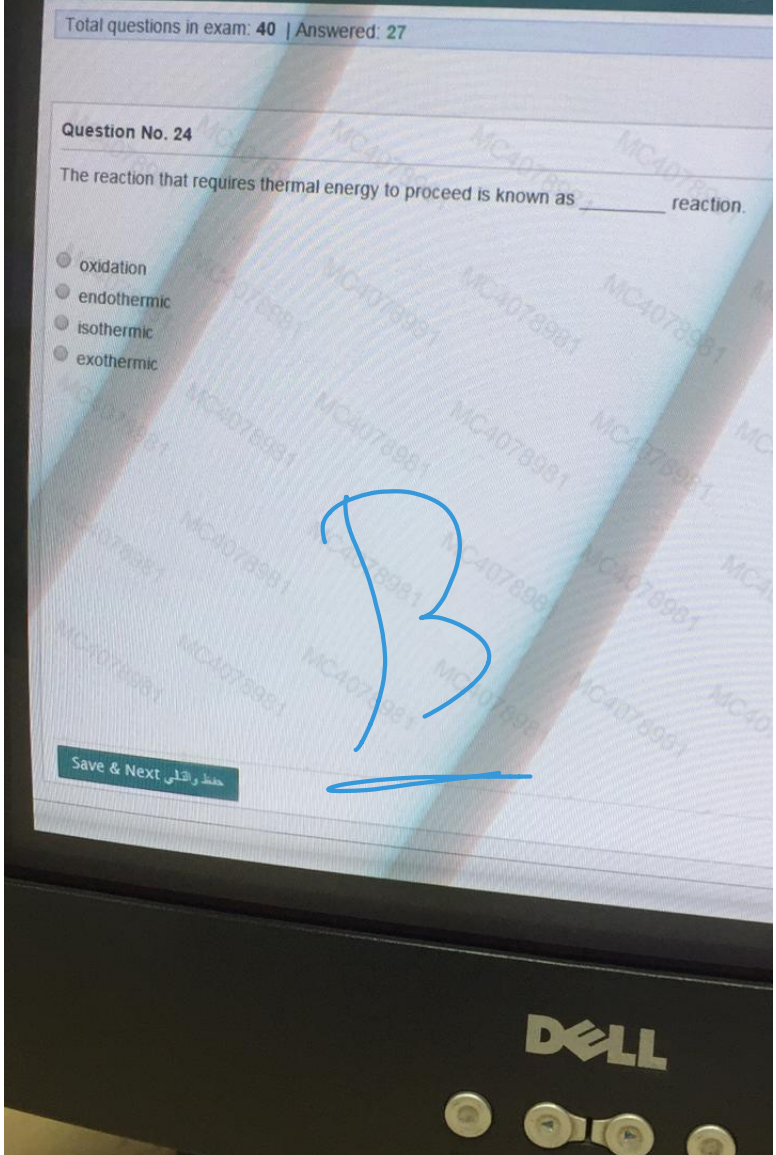
Total questions in exam: 40 | Answered: 27
Question No. 24
The reaction that requires thermal energy to proceed is known as _____ reaction.
oxidation
endothermic
isothermic
exothermic
B
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DELL

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Chemistry_FT_Sem1_201

Total questions in exam: **40** | Answered: **27**

Question No. 35

The combustion of ammonia in the presence of excess oxygen yields NO_2 and H_2O:

$$4\,NH_3\,(g) + 7\,O_2\,(g) \rightarrow 4\,NO_2\,(g) + 6\,H_2O\,(g)$$

The combustion of 43.9 g of ammonia produces _____ g of NO_2.

- 43.9
- 2.58
- 119
- 178

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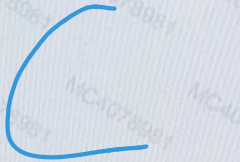

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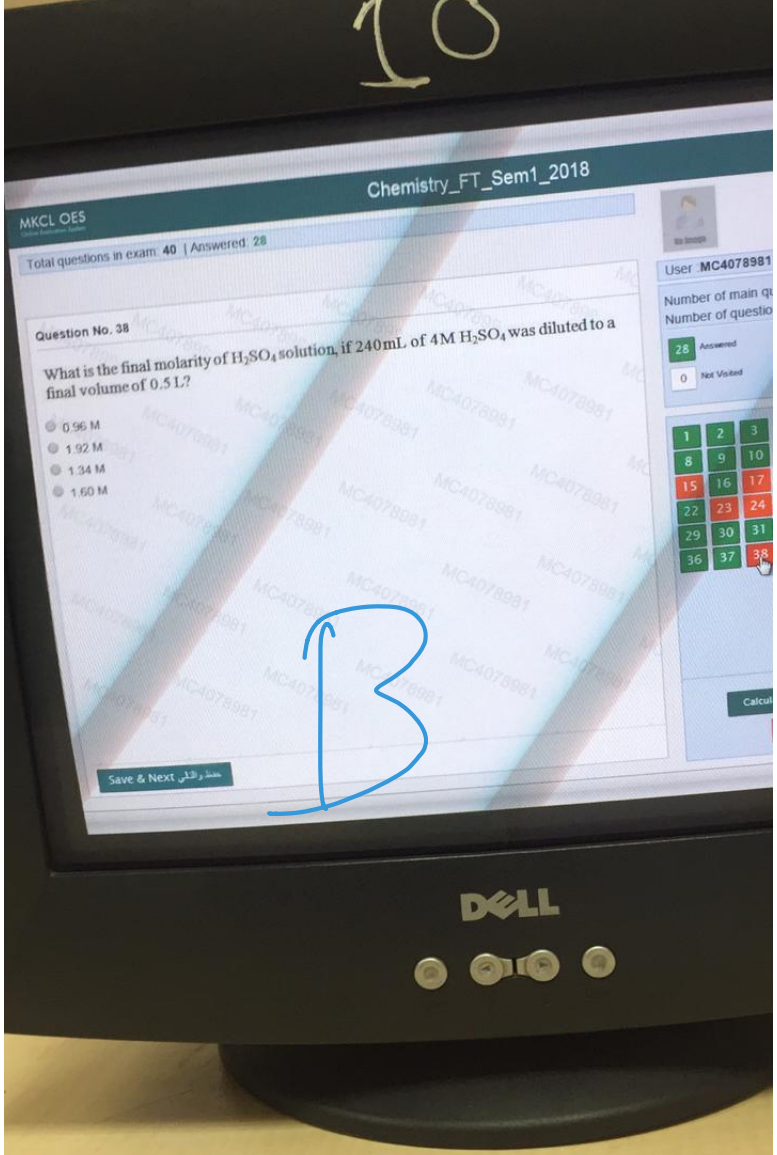
MKCL OES
Chemistry_FT_Sem1_2018
Total questions in exam: 40 | Answered: 28
User: MC4078981
Question No. 38
What is the final molarity of H_2SO_4 solution, if 240 mL of 4 M H_2SO_4 was diluted to a final volume of 0.5 L?
0.96 M
1.92 M
1.34 M
1.60 M
Number of main q
Number of questio
28 Answered
0 Not Visited
1 2 3
8 9 10
15 16 17
22 23 24
29 30 31
36 37 38
Save & Next
DELL

MKCL OES
Online Evaluation System

Chemistry_FT_Sem1_2

Total questions in exam: 40 | Answered: 27

Question No. 20

In an oxidation-reduction reaction, the reduced substance always _____

- shows gain of electrons.
- takes on oxygen atoms.
- gives up hydrogen atoms.
- shows loss of electrons.

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DELL

What is the oxidation number of carbon in CO_3^{2-} ?

- +2
- +6
- +4
- -2

$[CO_3^{-2}]^{-2}$

$\longrightarrow x + (-2 \times 3) = -2$

$x - 6 = -2$

$x = -2 + 6$

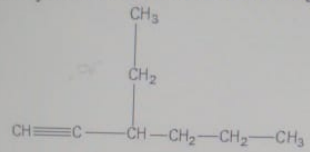
$x = +4$

C

Total questions in exam 40 | Answered

Question No. 8

What is the correct systematic name of the following compound?



- 4-ethyl-1-hexyne
- 4-ethyl-5-hexyne
- 3-ethyl-1-heptyne
- 3-ethyl-1-hexyne

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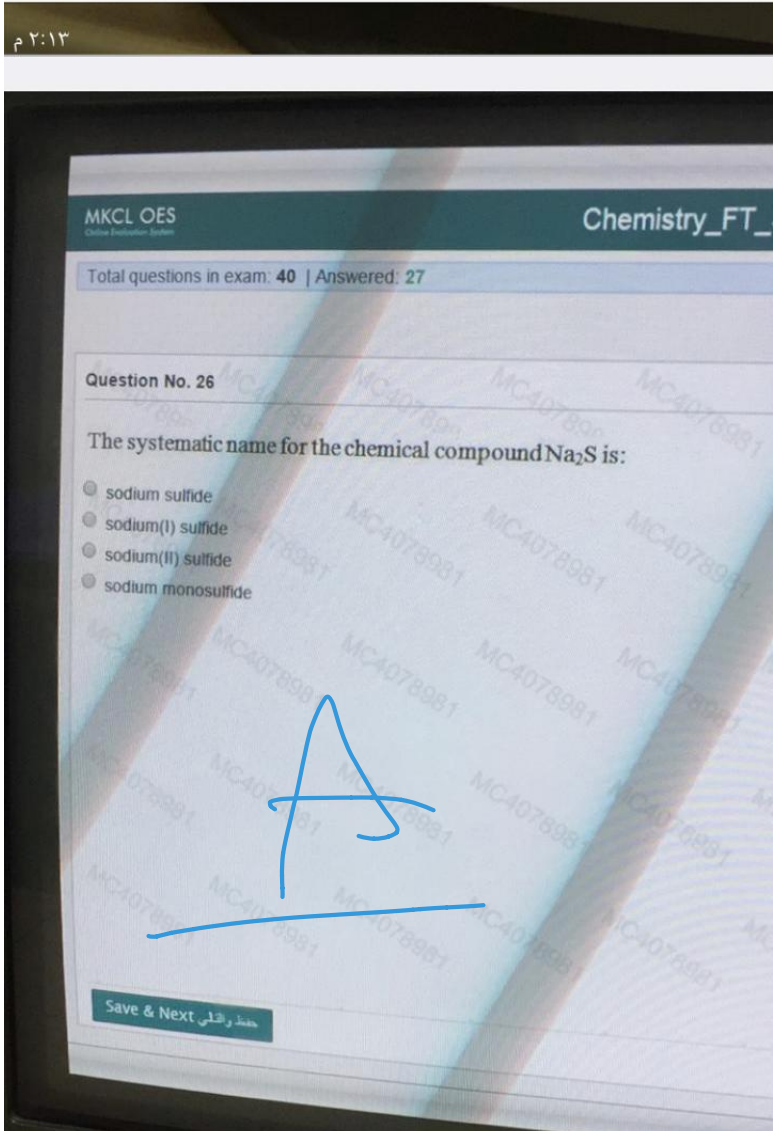
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MKCL OES
Chemistry_FT_
Total questions in exam: 40 | Answered: 27
Question No. 26
The systematic name for the chemical compound Na_2S is:
sodium sulfide
sodium(I) sulfide
sodium(II) sulfide
sodium monosulfide
A
Save & Next حفظ والتالي