

### **Question No. 1**

**Which of the following is the lowest temperature?**

- 273 Kelvin
- 100 Fahrenheit
- 100 Kelvin
- 100 Celsius

Total questions in exam: **20** | Answered: **6**

### **Question No. 7**

Which one of the following does NOT have the same electron configuration as the  $O^{2-}$  anion?

- Mg<sup>2+</sup> ion
  - Al<sup>3+</sup> ion
  - Na<sup>1-</sup> ion
  - Oxygen atom

X

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Online Evaluation System

Total questions in exam: 20 | Answered: 8

**Question No. 9**

Which of the following name/symbol pairs is incorrect?

- Lead, Pb
- Palladium, Pd
- Platinum, Pt
- Phosphorus, Pr



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Total questions in exam: 20 | Answered: 8

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Total questions in exam: 20 | Answered: 9

**Question No. 10**

The number of electrons in the shell with  $n = 4$  is ....

- 32 electrons
- 18 electrons
- 16 electrons
- 8 electrons

Total questions in exam: 20 | Answered: 14

Question No. 15

Which of the following elements has chemical properties similar to magnesium, Mg?

- Aluminum, Al
- Calcium, Ca
- Manganese, Mn
- Sodium, Na

X

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Total questions in exam: 20 | Answered: 16

Question No. 17

Which of the following elements has the lowest ability to gain electrons?

- Fluorine, F
- Carbon, C
- Oxygen, O
- Nitrogen, N



Total questions in exam: 20 | Answered: 18

## Question No. 19

Germanium is an example of .....

- Alkali metals
- Conductors
- Halogens
- Semiconductors

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Calculator

Instructions

End Test

Total questions in exam: 20 | Answered: 4

Question No. 6

Which one of the following does NOT have the same electron configuration as the F<sup>-</sup> anion?

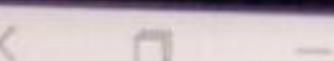
- Ne
- Al<sup>3+</sup>
- Na<sup>+</sup>
- S<sup>2-</sup>

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Which element has the following ground state electron configuration  $1S^2 2S^2 2P^6 3S^2$ ?

- Na
- Al
- Si
- Mg



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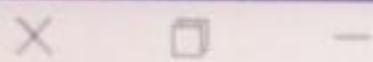
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The number of electrons in the shell with  $n = 3$  is .

- 18 electrons
- 8 electrons
- 32 electrons
- 2 electrons

Save & Next



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Total questions : 1

Which of the following elements has chemical properties similar to oxygen, O?

- Fluorine, F
- Nitrogen, N
- Sulfur, S
- Hydrogen, H

Save & Next



Print



X

?

What is the isotopic symbol of the aluminum isotope containing 14 neutrons?

- $^{14}_{13}\text{Al}$
- $^{27}_{14}\text{Al}$
- $^{14}_{14}\text{Al}$
- $^{27}_{13}\text{Al}$

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The potential energy is .....

- the stored energy, and it is based on the position or composition.
- the energy of motion.
- a form of kinetic energy.
- the stored energy, and it is based on the motion.

Save & Next



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How many nanometers are there in 16 decimeters?

- $1.6 \times 10^9$  nanometers
- $16 \times 10^{10}$  nanometers
- $16 \times 10^{-9}$  nanometers
- $1.6 \times 10^{-9}$  nanometers



Digit 1



Which one of these elements is a nonmetal?

Sr

Fe

H

Pb



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Which of the following statements about gases is false?

- Gases expand to fill the container they are placed in.
- Gases are highly compressible.
- Gas particles do not move around or past each other.
- Distances between particles of gases are very large.



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Group 8A elements are also called .....

- Noble gases
- Alkaline earth metals
- Alkali metals
- Halogens

X

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## MKCL QES

Which of the following is the highest temperature?

- 1 Celsius
- 270 Kelvin
- 0 Kelvin
- 0 Fahrenheit



Which of the following is NOT a correct name/symbol combination?

- Beryllium, Be
- Iron, Fe
- Manganese, Ma
- Phosphorus, P



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What Kelvin temperature is equal to 80 Fahrenheit?

- 299.8 K
- 27.7 K
- 30 K
- 305 K

Which one of the following does NOT have the same electron configuration as an oxygen atom?

- Oxygen atom
- Mg<sup>2+</sup> ion
- Na<sup>1+</sup> ion
- Al<sup>3+</sup> ion



Which of the following is an example of a chemical change?

- Grinding of pepper
- Melting of ice
- Digesting of an egg
- Freezing of water

X

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The ion which has 35 protons and 36 electrons is .....

- Br<sup>+</sup>
- Kr<sup>+</sup>
- Kr<sup>-</sup>
- Br<sup>-</sup>



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