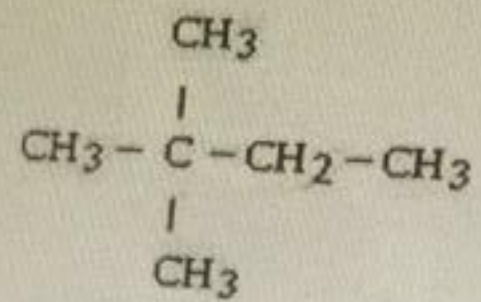


Question No. 32

What is the IUPAC name for the following?



- 2,2-dimethylbutane
- 3,3-dimethylbutane
- 2-dimethylbutane
- dimethylbutane

ASave & Next حفظ و التالي

All of the following is a general property of a basic solution Except?

- tastes sour
- neutralizes acids
- feels slippery
- turns litmus paper blue



A

Question No. 8

Which one of the following elements forms two or more ions with different ionic charges?

- O
- F
- K
- Fe

D

Question No. 7

What is the molecular formula of a compound that has a molar mass of 180 g/mol and its empirical formula is CH_2O ?

- $\text{C}_4\text{H}_8\text{O}_4$
- $\text{C}_6\text{H}_{12}\text{O}_6$
- $\text{C}_2\text{H}_2\text{O}_2$
- $\text{C}_5\text{H}_{10}\text{O}_5$

B

Question No. 28

CO_2 acts as a Lewis acid in the reaction $\text{CaO(s)} + \text{CO}_2 \rightarrow \text{CaCO}_3\text{(s)}$ because it _____.

- turns blue litmus to red
- reacts with a metal
- is a proton donor
- is an electron-pair acceptor

D

Question No. 39

What is the systematic name suffix of carboxylic acid?

- oic acid
- oate.
- al
- one

A

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Question No. 3

The elements in a column of the periodic table are known as

- a period.
- Metalloids.
- Noble gases.
- a group.

D

Question No. 40

Amino acids are the "building blocks" of

- proteins.
- vitamins
- fats.
- carbohydrates.

A

Save & Next حفظ والتالي

Question No. 27

Lewis base is defined as

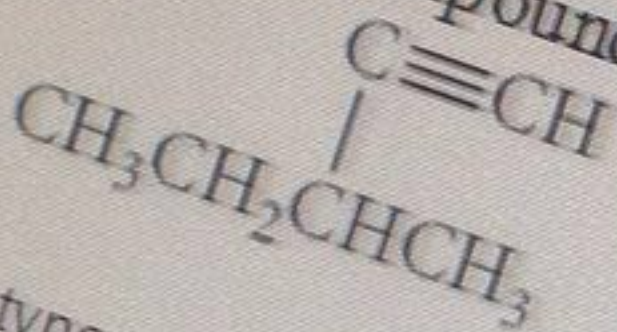
- an electron pair acceptor
- an electron pair donor
- Produces OH^- ions in an aqueous solution
- a proton acceptor

B



Question No. 35

Name the following compound

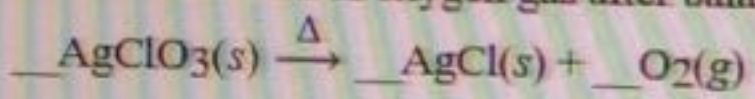


- 3-methyl-4-pentyne
- 3-methyl-1-pentyne
- 2-ethynebutane
- 3-ethyl-1-butyne

B

Question No. 5

What is the coefficient of oxygen gas after balancing the following equation?



- 2
- 1
- 3
- 4

C

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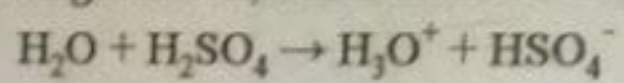
Question No. 2

The charge on the nucleus of a calcium atom is

- 20+
- 40+
- 20-
- 2+

A

According to the following reaction, which molecule is acting as a conjugate acid?



- HSO_4^-
- H_2SO_4
- H_3O^+
- H_2O

A

B

Question No. 38

Conversion of an aldehyde to an alcohol is known as reaction.

- Addition
- Reduction
- Esterification
- Oxidation

B



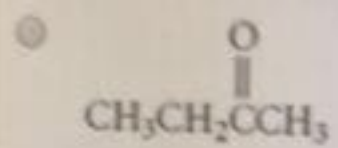
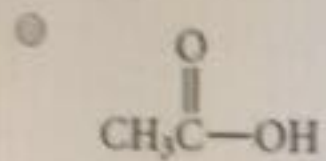
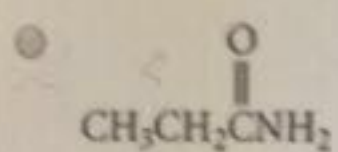
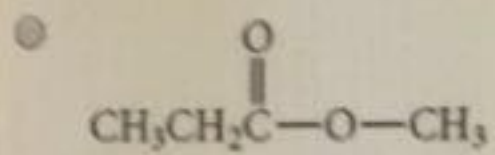
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All Media

Question No. 37

Which of the following compounds is an ester?



A

Save & Next حذرقای



Question No. 1

The number of electrons in the outer energy level of a neutral atom of bromine is

- 3
- 5
- 7
- 2

C

Save & Next حفظ واقلی

Question No. 8

The name of the chemical compound CuBr is:

- Copper(I) bromide
- Copper(II) bromide
- Copper(III) bromide
- Copper bromide

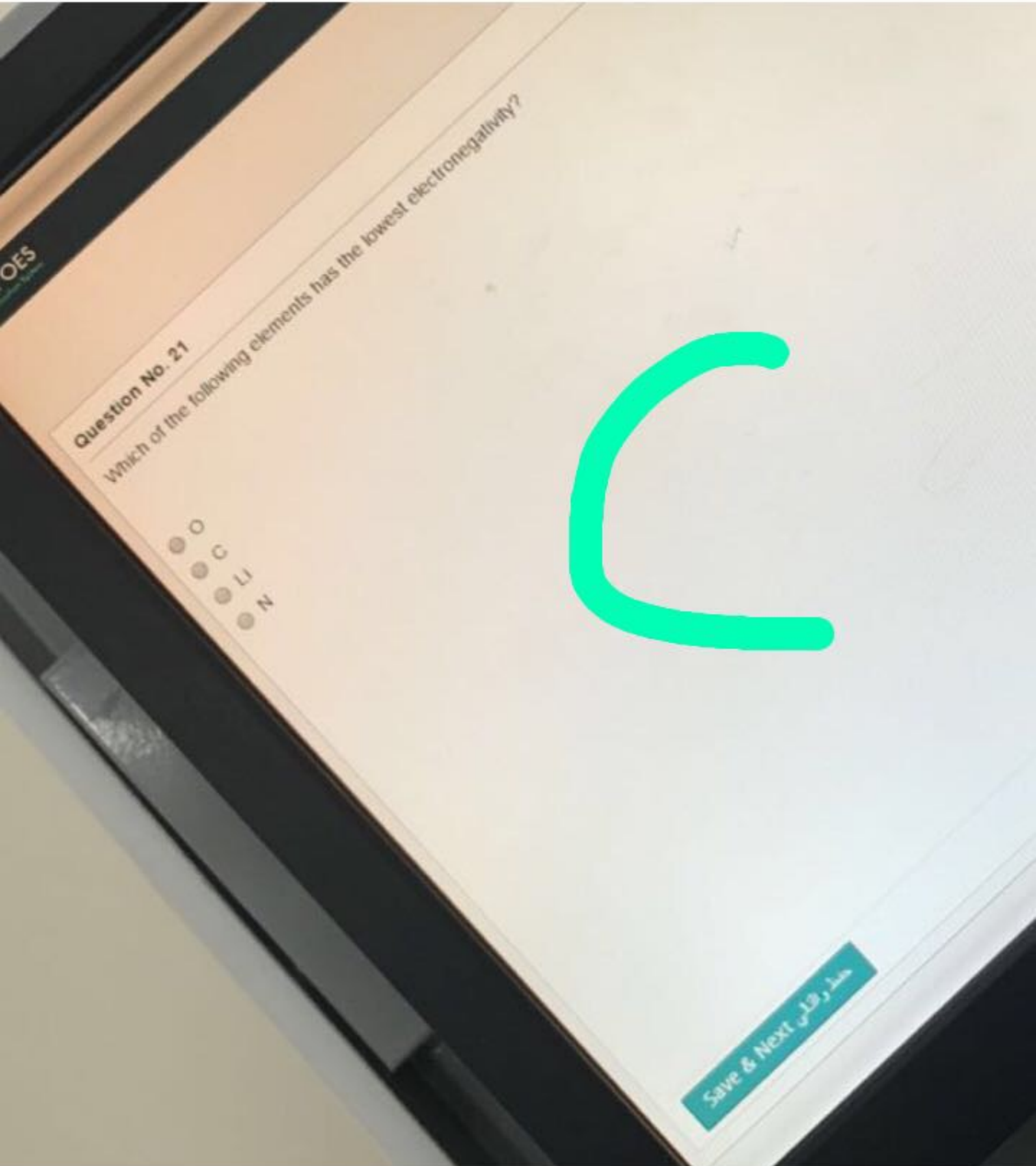
A



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Question No. 35

What is the common name for $\text{HC}\equiv\text{CH}$?

- ethylene
- acetylene
- propyne
- ethene

13

Question No. 9

The systematic name for the chemical compound NaBr is:

- sodium monobromide
- sodium(I) bromide
- sodium(II) bromide
- sodium bromide

D



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MKCL OES

Question No. 27

Lewis base is defined as

- an electron pair acceptor
- Produces OH^- ions in an aqueous solution
- a proton acceptor
- an electron pair donor



Save & Next

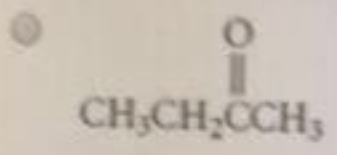
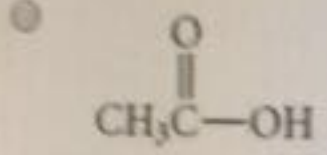
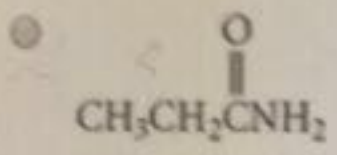
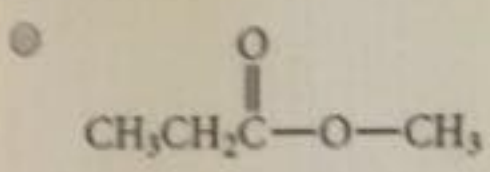
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Question No. 37

Which of the following compounds is an ester?



A

Question No. 4

Which of the following elements is a metal?

- nitrogen
- strontium
- argon
- fluorine

C

Question No. 19

Oxidation cannot occur without _____

- reduction
- water
- acid
- oxygen

A



4 📌 ❤️ @الشعبة الجميلة Days حنين ابا الخيل
انا شايفتها كلها كان

ES

Question No. 23

In the reaction below, what does the symbol \rightleftharpoons mean?
 $\text{OH}^- + \text{NH}_4^+ \rightleftharpoons \text{H}_2\text{O} + \text{NH}_3$

- It means that the rate of reverse reaction is greater than the forward reaction.
- It means that the reaction cannot decide which way to go.
- It means that the rates of forward and reverse reactions are equal.
- It means that the forward reaction does not progress



هنا ايش



Question No. 10

The molar mass of water is equal to:

- 27 g/mol
- 18 g/mol
- 36 g/mol
- 54 g/mol

B

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فيه هيدروجينة زايدة عند التفرع في بي

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Question No. 29

Which of the following pairs of species is not a conjugate acid-base pair?

- H_2O and OH^-
- HSO_4^- and SO_4^{2-}
- NH_3 and NH_2^-
- H_2SO_4 and HSO_4^-

د



Which molecule contains the weakest carbon-carbon bond?

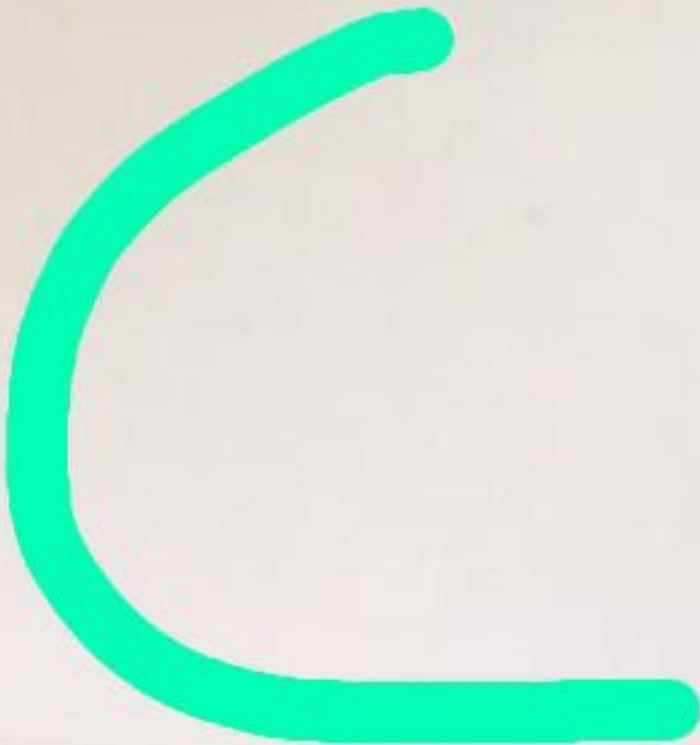
- $\text{HC}=\text{CH}$
- $\text{F}_2\text{C}-\text{CF}_2$
- $\text{H}_3\text{C}-\text{CH}_3$
- $\text{H}_2\text{C}=\text{CH}_2$

Save & Next حفظ والتالي



Which of the following is a general trend for the electronegativity of elements in the periodic table?

- decreases from left to right; decreases from bottom to top
- decreases from left to right; increases from bottom to top
- increases from left to right; increases from bottom to top
- increases from left to right; decreases from bottom to top



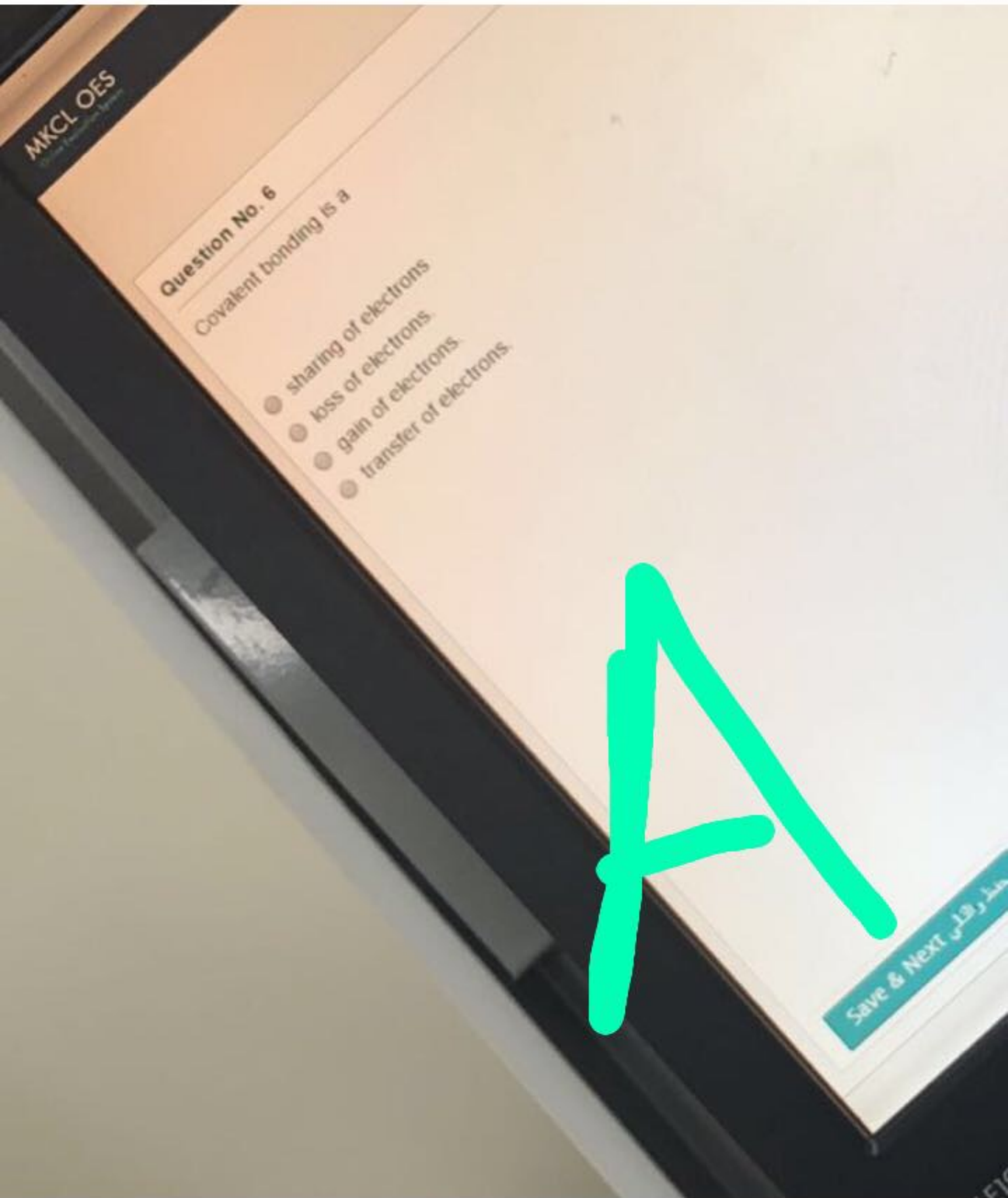
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Question No. 11

The most correct name for the compound CO is:

- carbon oxide
- monocarbon monooxide
- carbon monooxide
- carbon dioxide

C

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Question No. 12

The charge of hydroxide ion is:

- 1+
- 1-
- 2-
- 2+

B



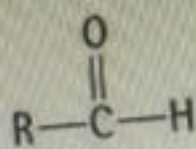
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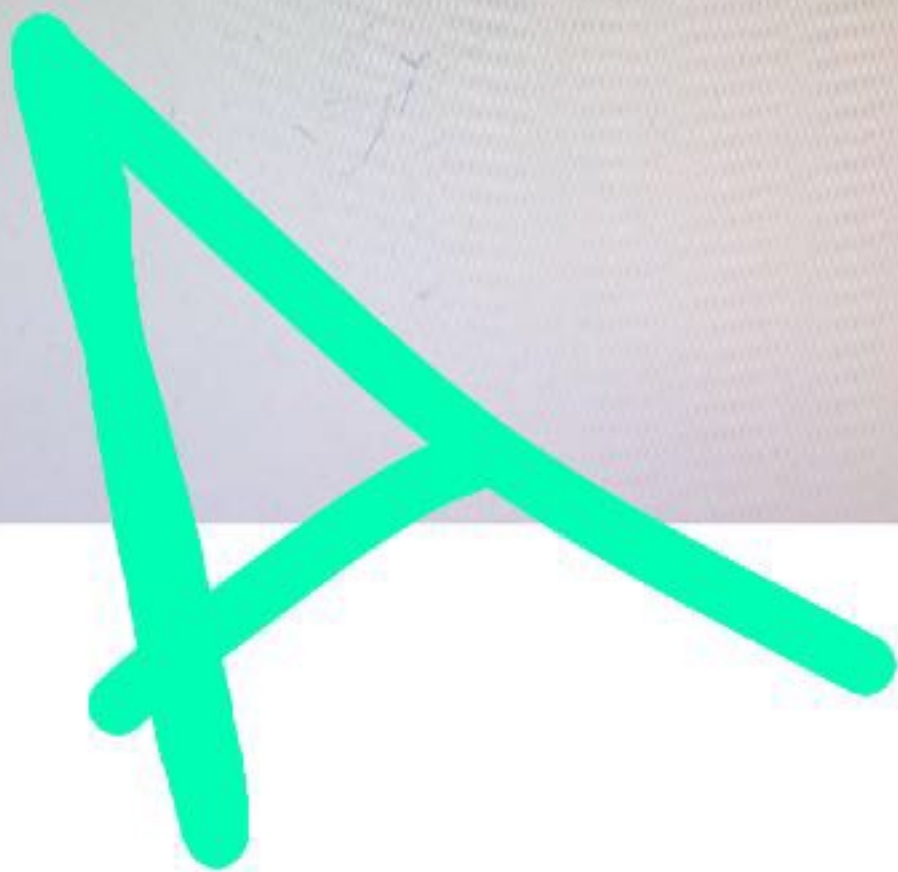
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Question No. 37

What is the name of compound has the following general formula?



- aldehyde
- carboxylic acid
- ester
- ketone

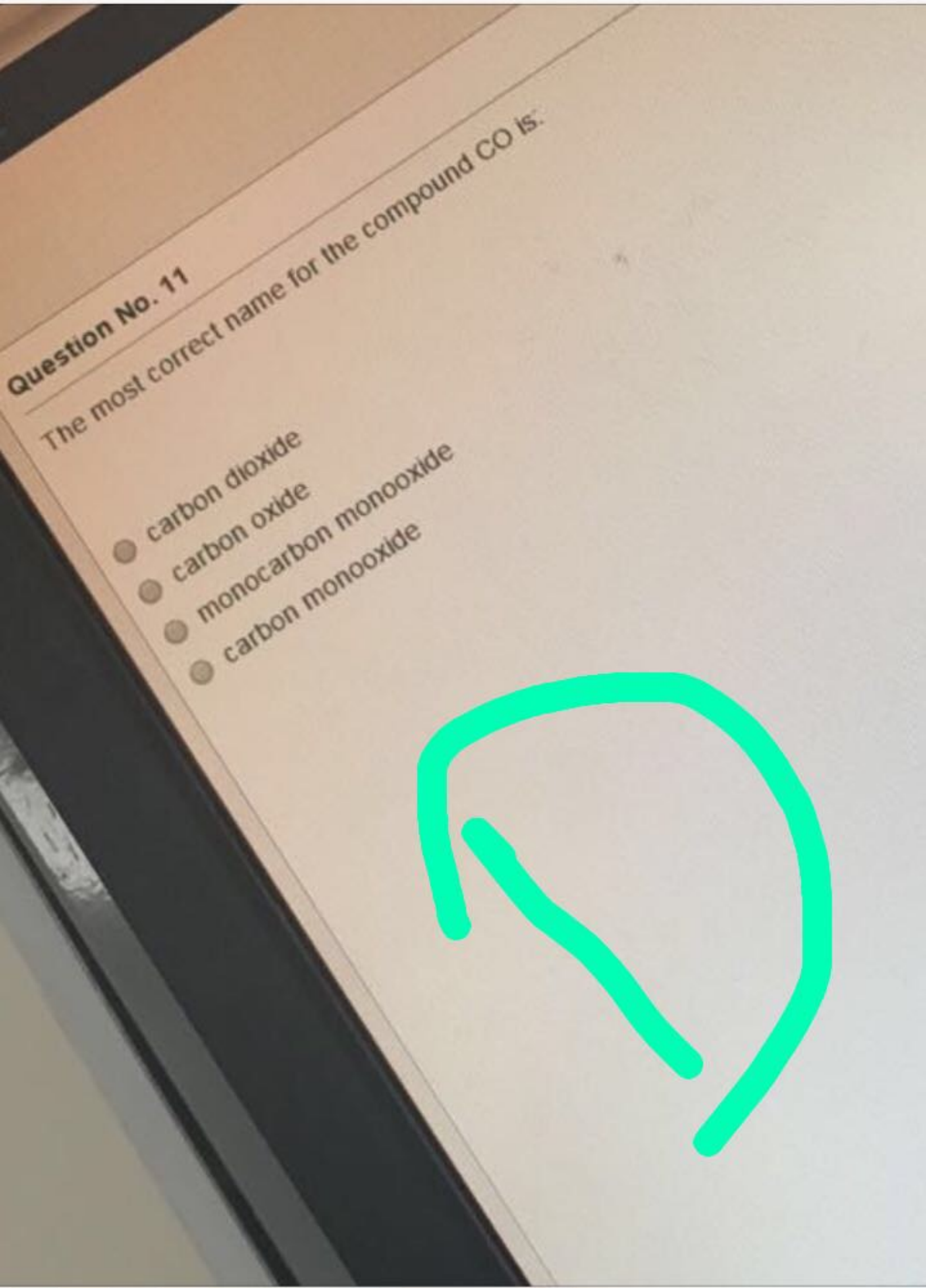




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CL OES

Question No. 12

Use the periodic table to answer the following question:
The correct name for the acid HCl is ___ acid

- hydrogen chloride
- hydrogen chlorate
- hydrochloric
- hydrogen chlorite



Question No. 9

The systematic name for the chemical compound NaBr is:

- sodium(II) bromide
- sodium monobromide
- sodium bromide
- sodium(I) bromide

C



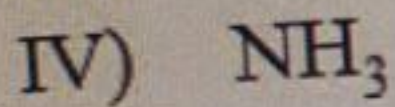
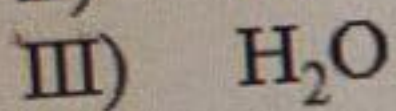
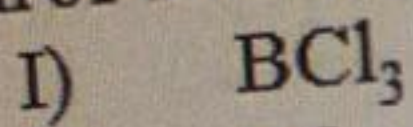
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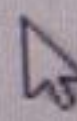
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Question No. 28

Which of the following are Lewis bases?



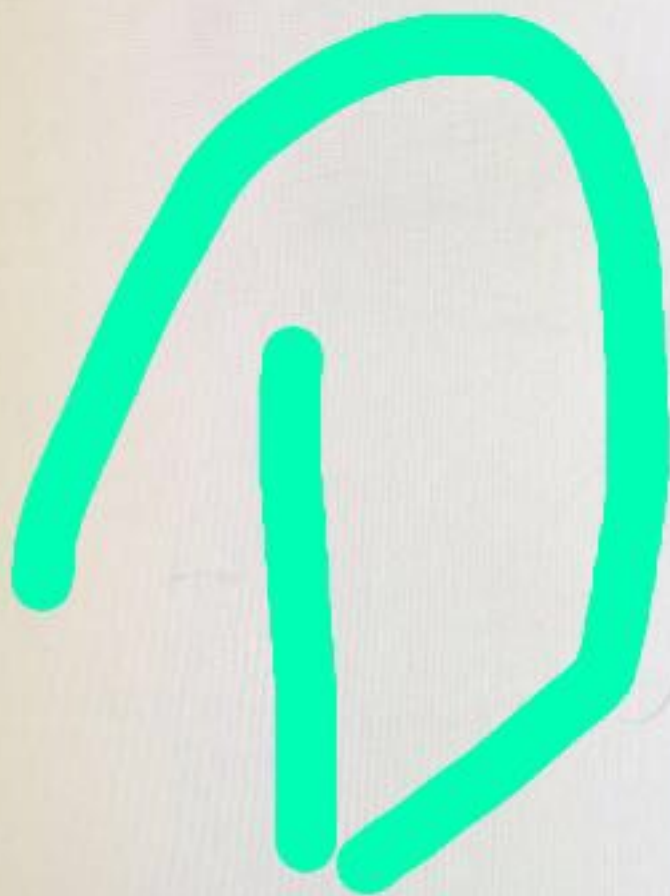
- I) and II)
- I), II), and III)
- II), III) and IV)
- III) and IV)



Question No. 31

The formula for pentyne is

- C_5H_{10}
- C_5H_{12}
- C_5H_{14}
- C_5H_8



Save & Next



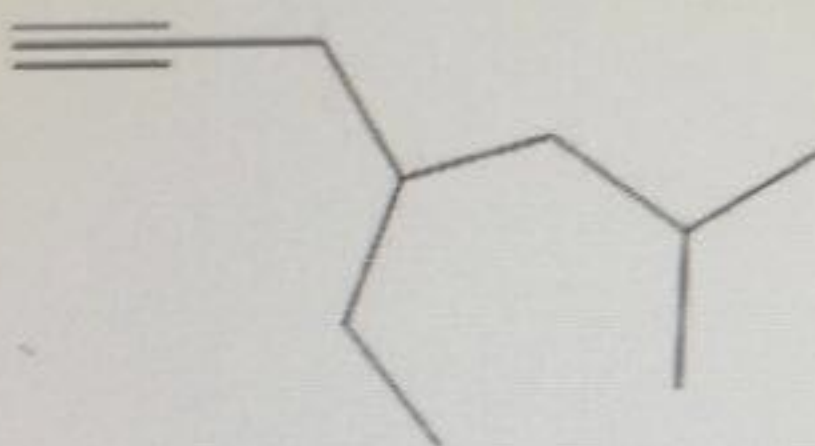
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Question No. 35

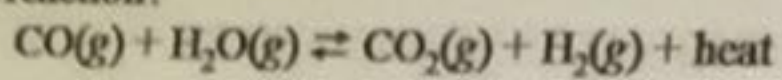
Provide the name of the compound below.




- 2-ethyl-2-methyl-6-heptyne
- 4-ethyl-6-methyl-1-heptyne
- 4-ethyl-2-methyl-6-heptyne
- 2-methyl-4-ethyl-1-heptyne

B

Which of the changes listed below will shift the equilibrium position to the *left* for the following reversible reaction?



- An increase of temperature
- A decrease of volume
- An increase of $[\text{H}_2\text{O}]$
- A decrease of $[\text{CO}_2]$

Save & Next  

4



Question No. 21

Which of the following elements has the highest electronegativity?

- Cl
- S
- Si
- Mg

A

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Question:

What is the name of the following alkyl group: $\text{CH}_3\text{-CH}_2\text{-CH}_2\text{-}$?

Options:

- ethyl
- propyl
- isopropyl
- methyl

Click on the question number to solve it.

PT2015_Chemistry

[Q01](#) [Q02](#) [Q03](#) [Q04](#) [Q05](#) [Q06](#) [Q07](#) [Q08](#) [Q09](#) [Q10](#) [Q11](#)
[Q24](#) [Q25](#) [Q26](#) [Q27](#) [Q28](#) [Q29](#) [Q30](#) [Q31](#) [Q32](#) [Q33](#) [Q34](#)**INSTRUCTION:** تعليمات Please choose the BEST answer from the given options for each**Question:**Which of the following is a chemical change?**Options:**

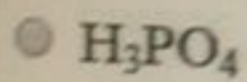
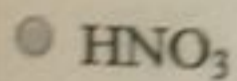
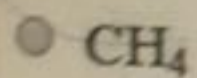
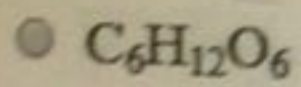
- cutting a copper wire into two pieces
- evaporation of gasoline
- rusting of iron
- dissolving salt in water

تسليم الإجابة

Submit Answer

Question No. 22

Which of the following compounds gives a strong electrolyte aqueous solution?



INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

The correct name for CaCO_3 is _____.

Options:

- calcium monocarbonate.
- calcium carbonate.
- calcium(II) carbonate.
- calcium carbon trioxide.

تسليم الإجابة
Submit Answer

تعليمات Please choose the BEST answer from the given o

the name of compound shown below?



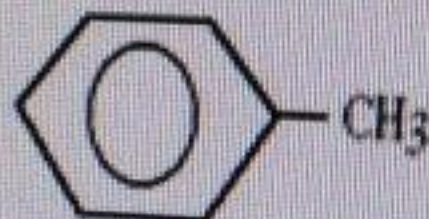
e
ne
l



INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for

Question:

What is the name of compound shown below?



Options:

- aniline
- toluene
- benzene
- phenol

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options.

Question:

Which of the following elements belongs to the halogens group?

Options:

- Fe
- Co
- Na
- Cl

Question No. 7

Which statement below accurately describes the contributions of Thomson?

- Created the modern periodic table
- Proposed the modern Atomic Theory
- Discovered the existence of electrons
- Ancient Greek philosopher who proposed that matter was continuous

Question No. 15

A triple covalent bond contains _____ of electrons.

- 3 pairs
- 4 pairs
- 2 pairs
- 1 pair

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

What is the final molarity of HCl solution, if 15 mL of 4 M HCl was diluted to a final volume of 200 mL?

Options:

- 0.3 M
- 0.08 M
- 0.16 M
- 0.2 M

Click on the question number to solve it.

Answered Questions

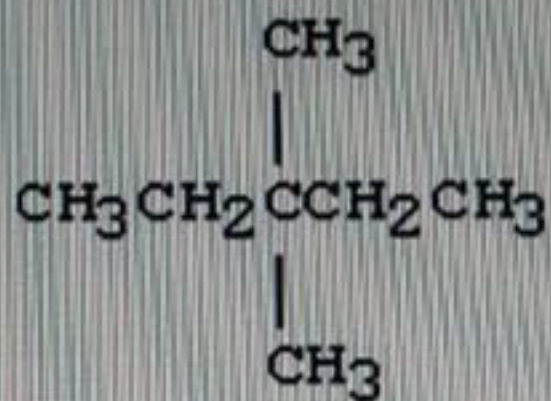
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Q01 Q02 Q03 Q04 Q05 Q06 Q07 Q08 Q09 Q10 Q11 Q12 Q13 Q14 Q15 Q16 Q17 Q18 Q19 Q20 Q21 Q22 Q23 Q24 Q25 Q26 Q27 Q28 Q29 Q30 Q31 Q32 Q33 Q34 Q35 Q36

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

What is the IUPAC name for the following compound?



Options:

- 2, 2-diethylpropane
- 2, 2-dimethyl-2-ethylpropane
- heptane
- 3, 3-dimethylpentane

تسليم الإجابة
Submit Answer

Question No. 24

Which of the following has the greatest number of nonbonding pairs of electrons?

- He
- F
- H
- C

Question:

Define specific heat capacity.

Options:

- The quantity of heat required to change a system's temperature by 1°C
- The quantity of heat required to change the temperature of 1 gram of a substance by 1°C
- The quantity of heat required to lower the temperature of 1 gram of a substance by 1°C
- The quantity of heat required to lower the temperature of 1 mole of a substance by 1°C

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

Ionic compounds often contain _____.

Options:

- nonmetallic elements only
- both metallic and nonmetallic elements
- neither metallic nor nonmetallic elements
- metallic elements only

تسليم الإجابة
Submit Answer

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Question No. 12

Which of these elements has the lowest electron affinity?

- carbon
- nitrogen
- beryllium
- boron

Click on the question number to solve it.

PT2015_Chemistry

Q01

Q02

Q03

Q04

Q05

Q06

Q07

Q08

Q09

Q24

Q25

Q26

Q27

Q28

Q29

Q30

Q31

Q32

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for

Question:

Which of the following elements belongs to the halogens group?

Options:

- Fe
- Co
- Na
- Cl

تسليم الإجابة

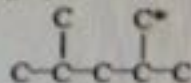
Submit Answer

All Right

INSTRUCTION: **تعليمات:** Please choose the BEST answer from the given options for each question.

Question:

The carbon skeleton of an alkane is shown below. How many hydrogen atoms are bonded to the carbon marked with a *?



Options:

- 3
- 2
- 0
- 1

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

Which of the following is not part of a nucleotide?

Options:

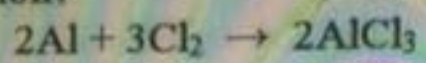
- fatty acid
- cyclic amine base
- a sugar
- phosphoric acid

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Submit Answer

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Question No. 19

How many grams of AlCl_3 could be produced when 94.5 grams of Al completely react with Cl_2 according to the reaction?

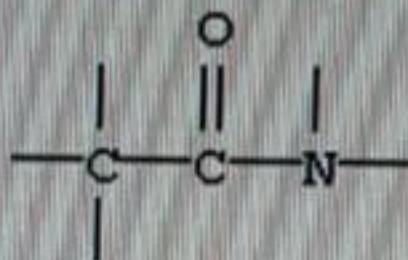


- 467 g
- 533 g
- 399 g
- 133 g

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each qu

Question:

Identify the functional group:



Options:

- alcohol
- amine
- amide
- ester

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Submit Answer

All Rights Reserve

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options

Question:

In organic chemistry, compounds are generally classified by

Options:

- taste.
- color.
- odor.
- functional group.

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

What class of hydrocarbons has the general formula C_nH_{2n} ?

Options:

- alkenes
- aromatics
- alkynes
- alkanes

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Submit Answer

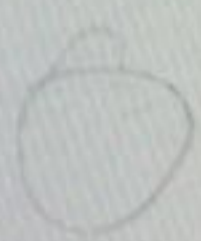
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According to the Bronsted-Lowry definition,

$\frac{2-3}{2-3}$

- a base is a proton donor.
- an acid is a proton donor
- a base produces H^+ ions in an aqueous solution.
- an acid is a proton acceptor.

2 3 4 5 6 7



~

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each

Question:

Which of the following is a heterogeneous mixture?

Options:

- pure water
- tap water
- gasoline
- ice and water

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Submit Answer

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Question No. 13

Which family is most likely to form diatomic molecules?

- Group VII-A
- Group VIII-A
- Group VI-A
- Group V-A

INSTRUCTION: Please choose the BEST answer from the given options for each question.

Question:

Calculate the volume (in mL) of a 2.75 M solution that must be used to make 1.25 L of a 0.150 M solution.

Options:

- 68.2 mL
- 0.0682 mL
- 33.0 mL
- 0.0330 mL

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for e

Question:

Sucrose, common table sugar, when hydrolyzed will form

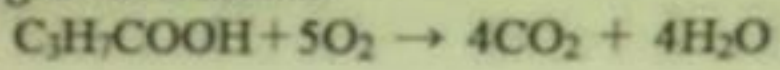
Options:

- lactose and maltose.
- glucose and fructose.
- amylose and glycogen.
- cellulose and starch.

تسليم الإجابة
Submit Answer

Question No. 19

How many grams of CO_2 could be produced when 44 grams of $\text{C}_3\text{H}_7\text{COOH}$ completely react with oxygen gas according to the reaction?



- 133 g
- 44 g
- 88 g
- 22 g

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INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for ea

Question:

Energy that is associated with the motion of an object is called

Options:

- kinetic energy
- chemical energy
- potential energy
- thermal energy

I

Click on the question number to solve it.

Answered Questions

PT2015_Chemistry

Q01 Q02 Q03 Q04 Q05 Q06 Q07 Q08 Q09 Q10 Q11 Q12 Q13 Q14
Q24 Q25 Q26 Q27 Q28 Q29 Q30 Q31 Q32 Q33 Q34 Q35 Q36 Q37

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

Compounds that have the same formula but different molecular structures are called

Options:

- isomers.
- isoelectronic.
- isotopes.
- isotones.

تسليم الإجابة
Submit Answer

Question No. 30

An acid and base react to form a salt and water in a neutralization reaction.

- neutralization
- reduction
- oxidation
- ionization

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

Which of the following compounds gives a weak electrolyte aqueous solution?

Options:

- NH_3
- HCl
- $\text{C}_6\text{H}_{12}\text{O}_6$
- HBr

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

What is the term for a family of unsaturated hydrocarbon compounds having a double bond

Options:

- alkynes
- aromatics
- alkanes
- alkenes

Click on the question number to solve it.

PT2015_Chemistry

Q01 Q02 Q03 Q04 Q05 Q06 Q07 Q08 Q09 Q10 Q11
Q24 Q25 Q26 Q27 Q28 Q29 Q30 Q31 Q32 Q33 Q34

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

Which one of the following carries no electrical charge?

Options:

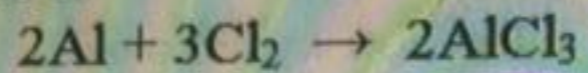
- a proton
- a neutron
- a cation
- an electron

تسليم الإجابة
Submit Answer

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Question No. 19

How many grams of AlCl_3 could be produced when 94.5 grams of Al completely react with Cl_2 according to the reaction?



- 467 g
- 533 g
- 399 g
- 133 g

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for ea

Question:

In the following reaction H_2SO_4 is _____.



Options:

- the oxidizing agent and is oxidized
- the reducing agent and is reduced
- the reducing agent and is oxidized
- the oxidizing agent and is reduced

تسليم الإجابة

Submit Answer

All Right

Question No. 16

What is the molecular formula of a compound that has a molar mass of 30 g/mol and its empirical formula is CH_2O ?

- $\text{C}_4\text{H}_8\text{O}_4$
- $\text{C}_5\text{H}_{10}\text{O}_5$
- $\text{C}_3\text{H}_6\text{O}_3$
- CH_2O

Click on the question number to solve it.

Ans

PT2015_Chemistry

Q01

Q02

Q03

Q04

Q05

Q06

Q07

Q08

Q09

Q10

Q11

Q12

Q24

Q25

Q26

Q27

Q28

Q29

Q30

Q31

Q32

Q33

Q34

Q35

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

In order to form an octet, an atom of selenium will:

Options:

- lose 2 electrons.
- gain 6 electrons.
- gain 2 electrons.
- lose 6 electrons.

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تسليم الإجابة
Submit Answer

Question No. 23

What is the oxidation number of sulfur in SO_3^{2-} ?

- +4
- +6
- +2
- 2

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

Another term for alkanes is

Options:

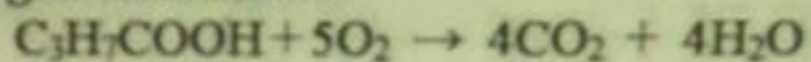
- saturated hydrocarbons.
- unsaturated hydrocarbons.
- alkenes.
- alkynes.

تسليم الإجابة
Submit Answer

All Rights Reserved

Question No. 19

How many grams of CO_2 could be produced when 44 grams of $\text{C}_3\text{H}_7\text{COOH}$ completely react with oxygen gas according to the reaction?



- 133 g
- 44 g
- 88 g
- 22 g

f n s

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

The number of electron lone pairs in the water molecule is _____.

Options:

- 2
- 4
- 8
- 3

Question:

Define specific heat capacity.

Options:

- The quantity of heat required to change a system's temperature by 1°C
- The quantity of heat required to change the temperature of 1 gram of a substance by 1°C
- The quantity of heat required to lower the temperature of 1 gram of a substance by 1°F
- The quantity of heat required to lower the temperature of 1 mole of a substance by 1°C

Question No. 20

What is the percent yield for a reaction if its theoretical yield is 135 g and its actual yield is 97 g?

- 53%
- 72%
- 63%
- 86%

Question:

What is the common name for $\text{HC}\equiv\text{CH}$?

Options:

- propyne
- ethene
- acetylene
- ethylene

Click on the question number to solve it.

Unanswered Questions Current Question

PT2015_Chemistry

Q01	Q02	Q03	Q04	Q05	Q06	Q07	Q08	Q09	Q10	Q11	Q12	Q13	Q14	Q15	Q16	Q17	Q18	Q19	Q20	Q21	Q22	Q23	Q24	Q25	Q26	Q27	Q28	Q29	Q30	Q31	Q32	Q33	Q34	Q35	Q36	Q37	Q38	Q39	Q40
-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----	-----

INSTRUCTION: Please choose the BEST answer from the given options for each question.

Question:

What is the term for a triglyceride (triacylglycerol) from an animal source that has mostly saturated fatty acids?

Options:

- a) wax
- b) phospholipid
- c) oil
- d) fat

تسليم الإجابة

Submit Answer

Question No. 8

Any p orbital can hold up to 2 electrons.

18

2

8

6

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

Which statement about electronegativity is incorrect?

Options:

- Fluorine is the most electronegative atom of all the elements.
- Metals generally have higher electronegativity values than nonmetals.
- Within a group, electronegativity increases from bottom to top.
- Within a row, electronegativity increases from left to right.

تسليم الإجابة

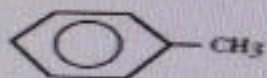
Submit Answer

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INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for

Question:

What is the name of compound shown below?



Options:

- aniline
- toluene
- benzene
- phenol

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each

Question:

A solution with a pH of 2.0 is _____.

Options:

- strongly acidic
- strongly basic
- weakly acidic
- weakly basic

تسليم الإجابة

Submit Answer

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Question No. 2

Which of the following items is a mixture?

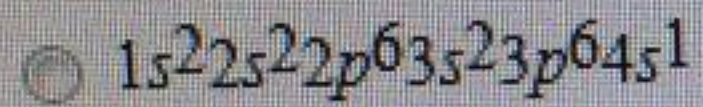
- sulfur
- cereal and milk
- baking soda
- table salt

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each

Question:

What is the electron configuration for potassium?

Options:



Question No. 14

Which of the following indicates a liquid in a chemical equation?

(g)

(l)

(aq)

(s)

Click on the question number to solve it.

■ Answered Question ■ Current Question

PT2015_Chemistry

Q01	Q02	Q03	Q04	Q05	Q06	Q07	Q08	Q09	Q10	Q11	Q12	Q13	Q14	Q15	Q16	Q17	Q18
Q24	Q25	Q26	Q27	Q28	Q29	Q30	Q31	Q32	Q33	Q34	Q35	Q36	Q37	Q38	Q39	Q40	

INSTRUCTION: **تعليمات:** Please choose the BEST answer from the given options for each question.**Question:**

A(n) _____ is a chemical bond that results from the sharing of a pair of electrons between two atoms.

Options:

- covalent bond
- metallic bond
- ionic bond
- molecular bond

تسليم الإجابة

Submit Answer

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Question No. 21

How many liters of a 0.20 M NaOH solution contain 2.0 moles of NaOH?

- 5.0 L
- 0.5 L
- 10 L
- 20 L

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

The first law of thermodynamics is sometimes called the law of

Options:

- conservation of energy.
- creation of energy.
- conservation of mass.
- destruction of energy.

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

In a solution, the solvent is _____.

Options:

- always water
- the substance present in the greatest amount
- always a liquid
- the substance being dissolved

تسليم الإجابة
Submit Answer

Question No. 9

Nickel is an example of an element that is a _____

- metalloid
- noble gas
- nonmetal
- transition element



INSTRUCTION: **تعليمات:** Please choose the BEST answer from the given options for each question.

Question:

What is the percent yield for a reaction if its theoretical yield is 65 g and its actual yield is 56 g?

Options:

- 9%
- 60%
- 86%
- 56%

Click on the question number to solve it.

PT2015_Chemistry

Q01

Q02

Q03

Q04

Q05

Q06

Q07

Q24

Q25

Q26

Q27

Q28

Q29

Q30

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given

Question:

Which of the following elements is a metalloid?

Options:

- Si
- Al
- Mg
- N

Question No. 3

Which of the following statements about energy is FALSE?

- Energy cannot be transformed from one form to another.
- Energy is the capacity to do work.
- Energy can neither be created nor destroyed.
- An object possessing energy can do work on another object.

INSTRUCTION: تعليمات Please choose the BEST answer from the given

Question:

Glucose can be classified as a(an)

Options:

- disaccharides.
- animal polysaccharide.
- monosaccharide.
- plant polysaccharide.

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each

Question:

Which of the following is a non-electrolyte?

Options:

- KOH
- HCl
- NaBr
- C₆H₁₂O₆

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Submit Answer

Question No. 28

Which of the following is a diprotic acid?

- H_3PO_4
- HNO_3
- H_2SO_4
- $\text{HC}_2\text{H}_3\text{O}_2$

Handwritten scribbles and numbers, possibly "2-3" and "2-3".

Handwritten scribbles and lines.

Handwritten scribbles and numbers, possibly "2 3 4 5 6 7".

Handwritten scribble, possibly a circle.

Handwritten scribble, possibly a wavy line.

INSTRUCTION: **تعليمات** Please choose the BEST answer from the

Question:

Which of the following is a compound?

Options:

- copper
- magnesium
- iron
- carbon dioxide

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options.

Question:

In organic chemistry, compounds are generally classified by

Options:

- taste.
- color.
- odor.
- functional group.

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

In an Arrhenius acid-base context, the compounds KOH, H₂SO₄, and HNO₃ when dissolved in water, function respectively as a (n) _____.

Options:

- acid, base and base
- base, base and acid
- base, acid and base
- base, acid and acid



INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

In an Arrhenius acid-base context, the compounds KOH, H₂SO₄, and HNO₃ when dissolved in water, function respectively as a (n) _____.

Options:

- acid, base and base
- base, base and acid
- base, acid and base
- base, acid and acid

Question No. 16

What is the empirical formula of the compound that has a composition by mass of 34.6% C, 3.9% H and 61.5% O?

- CH₂O
- CH₂O₃
- C₃H₄O₄
- CHO

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T
P N S ✓

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INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

What is the term for a triglyceride (triacylglycerol) from an animal source that has a high melting point?

Options:

- wax
- phospholipid
- oil
- fat

تسليم الإجابة
Submit Answer

Question No. 4

Which of the following is matter?

- the air you are breathing
- the sunlight
- the sound of music
- the light of a flame

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

The bomb calorimeter is used to:

Options:

- measure changes in internal energy for formation reaction
- measure changes in external energy for combustion reaction
- measure changes in internal energy for combustion reaction
- measure changes in external energy for formation reaction

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

Which experiment led to the notion that the atom contains an extremely small, positively

Options:

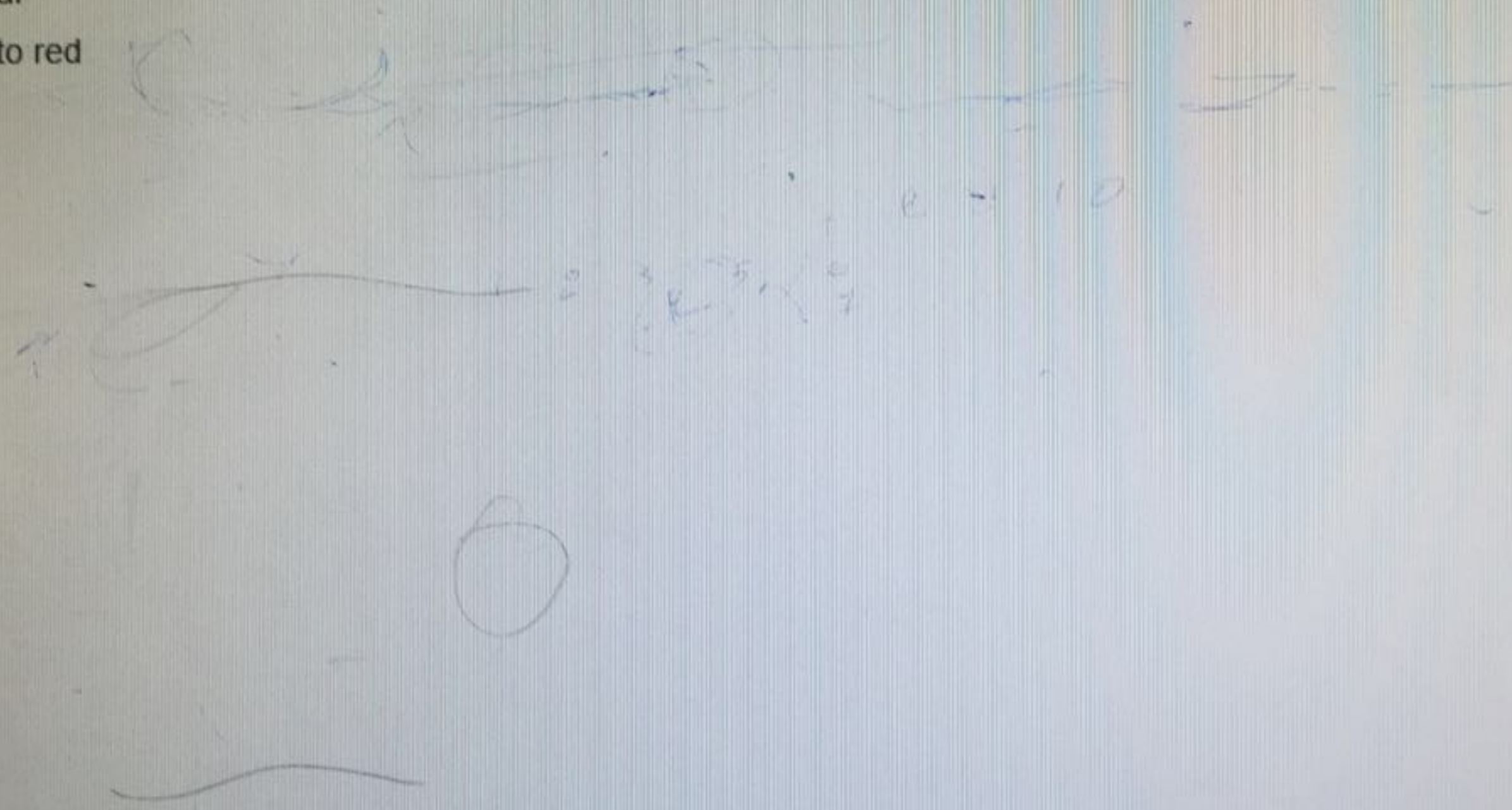
- Dalton's atomic experiment
- Rutherford's gold foil experiment
- Millikan's oil drop experiment
- Thomson's cathode ray experiment

تسليم الإجابة
Submit Answer

Question No. 29

CO_2 acts as a Lewis acid in the reaction $\text{CaO}(s) + \text{CO}_2 \rightarrow \text{CaCO}_3(s)$ because it _____.

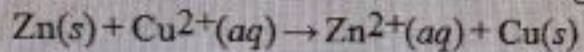
- is a proton donor
- is an electron-pair acceptor
- reacts with a metal
- turns blue litmus to red



INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

What substance is oxidized in the following redox reaction?



Options:

- Zn
- Zn^{2+}
- Cu^{2+}
- Cu

INSTRUCTION: **تعليمات:** Please choose the BEST answer from the given options for each question.

Question:

Which of the following would not be considered matter?

Options:

- light
- cheese
- blood
- plants

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Submit Answer

Question No. 17

The systematic name for the chemical compound RbCl is:

- rubidium(I) chloride
- rubidium(II) chloride
- rubidium chloride
- rubidium monochloride

INSTRUCTION: **تعليمات:** Please choose the BEST answer from the given options for each question.

Question:

What is the IUPAC name for $\text{CH}_3\text{-CH=CH-CH}_2\text{-CH}_2\text{-CH}_3$?

Options:

- 2-hexene
- 1-hexene
- 3-hexene
- hexene

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question.

Question:

Liquids are characterized as having _____ volume which means they are _____.

Options:

- indefinite, incompressible
- indefinite, compressible
- definite, incompressible
- definite, compressible

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Submit Answer

Question No. 6

A patient has a temperature of 38.5°C . What is the temperature in degrees Fahrenheit?

- 126.95 $^{\circ}\text{F}$
- 101.3 $^{\circ}\text{F}$
- 311.0 $^{\circ}\text{F}$
- 70.5 $^{\circ}\text{F}$

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for e

Question:

Which of the following is not a pure substance?

Options:

- sodium bicarbonate
- distilled water
- blood
- calcium

تسليم الإجابة
Submit Answer

All Right

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

What is the IUPAC name for $\text{CH}_3\text{-CH=CH-CH}_2\text{-CH}_2\text{-CH}_3$?

Options:

- 2-hexene
- 1-hexene
- 3-hexene
- hexene

Click on the question number to solve it.

■ Answered Question

PT2015_Chemistry

Q01 Q02 Q03 Q04 Q05 Q06 Q07 Q08 Q09 Q10 Q11 Q12 Q13 Q14 Q15
Q24 Q25 Q26 Q27 Q28 Q29 Q30 Q31 Q32 Q33 Q34 Q35 Q36 Q37 Q38

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

An ionic bond forms between two atoms through _____.

Options:

- sharing of electron pairs
- transferring protons from the nucleus of the nonmetal to the nucleus of the metal
- transferring of electrons from metallic atoms to nonmetallic atoms
- each atom acquiring a negative charge

تسليم الإجابة
Submit Answer

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for e

Question:

Which is a characteristic property of nonmetals?

Options:

- thermal conductivity
- brittle
- luster
- malleability

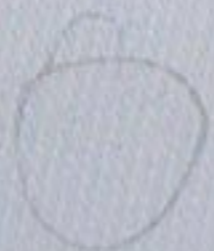
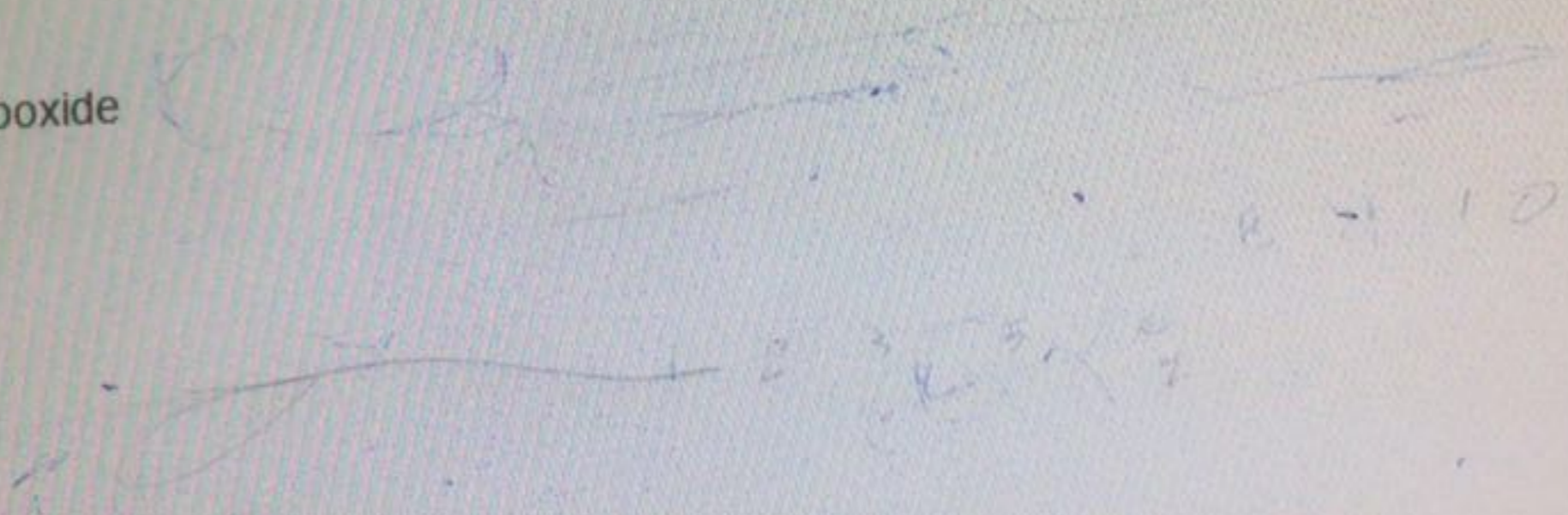
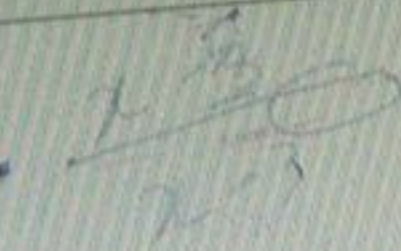
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Question No. 18

The most correct name for the compound CO is:

- carbon monoxide
- carbon dioxide
- carbon oxide
- monocarbon monoxide



Question:

One element that has 5 valence electrons is _____.

Options:

- nitrogen
- sulfur
- carbon
- lithium

INSTRUCTION: **تعليمات:** Please choose the BEST answer from the given options for each question.

Question:

What is the formula mass of C_4H_6O is -----?

Options:

- 57.10 amu
- 70.10 amu
- 54.10 amu
- 64.00 amu

تسليم الإجابة

Submit Answer

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Question No. 5

Which of the following is a chemical change?

- burning sugar
- melting gold
- shine of silver
- odor of paint thinner

INSTRUCTION: *تعليمات* Please choose the BEST answer from the given options for each question.

Question:

The bomb calorimeter is used to:

Options:

- measure changes in internal energy for formation reaction
- measure changes in external energy for combustion reaction
- measure changes in internal energy for combustion reaction
- measure changes in external energy for formation reaction

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options.

Question:

Glucose can be classified as a(an)

Options:

- disaccharides.
- animal polysaccharide.
- monosaccharide.
- plant polysaccharide.

Question No. 11

0.100 mole of lithium weighs

- 6.94 g.
- 0.694 g.
- 0.300 g.
- 3.00 g.

INSTRUCTION: Please choose the BEST answer from the given options for each question.

Question:

What is the term for a compound that contains only hydrogen and carbon atoms?

Options:

- saturated hydrocarbon
- aromatic hydrocarbon
- hydrocarbon
- unsaturated hydrocarbon

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each

Question:

The number of neutrons in an atom is equal to:

Options:

- atomic number - mass number
- mass number - atomic number
- the number of protons
- the number of electrons

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Submit Answer

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Question No. 25

When a system is at equilibrium?

- the reaction rate of the reverse reaction is small compared to forward.
- the amount of product and reactant is exactly equal.
- the reaction rate of the forward reaction is small compared to the reverse.
- the reaction rate of the forward reaction is equal to the rate of the reverse.

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each

Question:

Which of the following is a physical process?

Options:

- the explosion of nitroglycerin
- the baking of a potato
- the rusting of iron
- the condensation of water vapor

تسليم الإجابة
Submit Answer

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Question:

One element that has 5 valence electrons is _____.

Options:

- nitrogen
- sulfur
- carbon
- lithium

INSTRUCTION: **تعليمات** Please choose the BEST answer from

Question:

The sugar in the nucleotides of RNA is

Options:

- ribose.
- sucrose
- glucose.
- lactose.

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question

Question:

Which one of the following amine bases does not appear in DNA?

Options:

- guanine
- adenine
- uracil
- cytosine

تسليم الإجابة
Submit Answer

Question No. 5

Which of the following is a chemical change?

- burning sugar
- melting gold
- shine of silver
- odor of paint thinner

INSTRUCTION: تعليمات Please choose the BEST answer from the given options for each question

Question:

Which of the following is a heterogeneous mixture?

Options:

- pure water
- tap water
- gasoline
- ice and water

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Question No. 1

A solution has a density of 1.22 g/mL. What volume of the solution has a mass of 55.0 g?

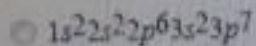
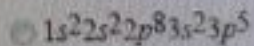
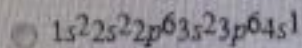
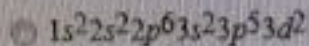
- 45.1 mL
- 49.4 mL
- 58.8 mL
- 0.00253 mL

INSTRUCTION: Please choose the BEST answer from the given options for each

Question:

What is the electron configuration for potassium?

Options:



INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question

Question:

What is the name of the following alkyl group: $\text{CH}_3\text{-CH}_2\text{-CH}_2\text{-}$?

Options:

- ethyl
- propyl
- isopropyl
- methyl

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

How much heat must be absorbed by 250 grams of water to raise its temperature by 3.0°C, the Specific Heat for water is 4.186 (J/g °C) [$q = m \cdot c_s \cdot \Delta T$]

Options:

- 21.31 KJ
- 56.62 KJ
- 31.40 KJ
- 18.31 KJ

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for

Question:

Which of the following is not a strong base?

Options:

- NaOH
- NH₃
- KOH
- Ca(OH)₂

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Submit Answer

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Question No. 5

Which of the following is an example of chemical property

- density.
- color.
- flammability.
- taste.

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Question No. 11

Which statement below accurately describes the contributions of Thomson?

- Ancient Greek philosopher who proposed that matter was continuous
- Proposed the modern Atomic Theory
- Discovered the existence of electrons
- Created the modern periodic table



Question No. 25

As you move up and to the right on the periodic table

- atomic radius increases and ionization energy increases
- atomic radius decreases and ionization energy decreases
- atomic radius decreases and ionization energy increases
- atomic radius increases and ionization energy decreases

C

Question No. 18

The Group 8A (18) elements

- are unreactive.
- are liquids at room temperature.
- are good conductors of electricity.
- melt at high temperatures.

A

Question No. 15

According to the Pauli exclusion principle, any orbital can hold at most _____ electrons.

- 6
- 18
- 2
- 8



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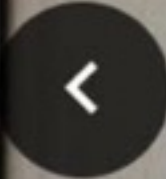


Question No. 24

Which of these elements has the lowest first ionization energy?

- aluminum
- sulfur
- sodium
- phosphorus

C





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MKCL OES

Question No. 7

One of the following elements exists as a diatomic molecule:

- chlorine
- iron
- argon
- neon

A



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Question No. 1

Which statement about an atom is false?

- The volume occupied by electrons is referred to as an electron cloud.
- The nucleus accounts for a small portion of the mass of an atom.
- The nucleus is the center region of an atom.
- The space occupied by the electrons is primarily empty.

B





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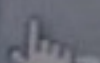


Question No. 19

Which of these elements has the highest electron affinity?

- Rubidium
- Potassium
- Lithium
- Caesium

C



Question No. 12

The charge of electron was established by

- Millikan's oil drop experiment.
- Rutherford's gold foil experiment.
- Dalton's atomic experiment
- Thomson's cathode ray tube experiment.

A

Question No. 14

One element that has ONE valence electrons is _____.

- lithium
- carbon
- sulfur
- nitrogen

A

Question No. 21

How many sulfur dioxide molecules (SO_2) are there in 1.80 mol of sulfur dioxide?

- 1.08×10^{23}
- 6.02×10^{24}
- 1.08×10^{24}
- 1.80×10^{24}

C

Question No. 18

The elements rubidium, sodium, and potassium _____

- are isotopes of each other
- are in the same period of elements
- are in the same group
- have the same number of neutrons



Question No. 15

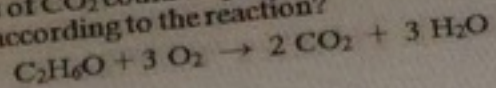
One element that has 4 valence electrons is _____

- lithium
- carbon
- sulfur
- nitrogen

B

Question No. 16

How many moles of CO_2 could be produced when 2 moles of $\text{C}_2\text{H}_6\text{O}$ completely react with oxygen gas according to the reaction?



- 2 mol
- 6 mol
- 4 mol
- 8 mol

C

Question No. 23

Which of the following elements is a metal?

- nitrogen
- fluorine
- strontium
- argon

C



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Question No. 18

What is the correct formula for a potassium ion with 18 electrons?

- P⁺
- P⁻
- K⁺
- K⁻

D



Question No. 19

When an atom gains an electron, the resulting particle is called:

- a cation.
- an isotope.
- a proton.
- an anion.

D



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Question No. 15

All of the following statements about different elements are true EXCEPT:

- Barium is an alkaline earth metal.
- Krypton is one of the noble gases.
- Lithium is considered a metalloid.
- Manganese is a transition metal.

C





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MKCL QES

Question No. 14

Metalloids are located where on the periodic table?

- left side
- between metals and nonmetals
- middle
- right side

B



MKCL OES

Question No. 22

What element has the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^2$?

- carbon
- silicon
- sulfur
- oxygen

B





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Question No. 25

The atomic size (radius) of atoms

- increases going across a period.
- does not change going across a period.
- decreases going down within a group.
- decreases going across a period.

D





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Question No. 25

The atomic radius of potassium is smaller than the atomic radius of _____

- hydrogen
- sodium
- lithium
- cesium

D





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Question No. 24

Which of these elements has the highest electron affinity?

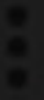
- silicon
- aluminum
- sulfur
- magnesium

C

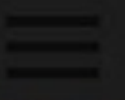




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تجميعات ... الثاني



Question No. 13

The maximum number of electrons that can go in the main energy level 3 is?

- 32
- 18
- 10
- 8

B





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Question No. 23

Which of the following elements is a metal?

- strontium
- argon
- fluorine
- nitrogen

A



Question No. 12

Which of the following is NOT a correct name, symbol combination?

- potassium, P
- manganese, Mn
- magnesium, Mg
- iron Fe

A

Question No. 24

Ionization energy is

- the energy an ion acquires from an electron.
- higher for potassium than for lithium.
- the energy needed to remove the least tightly bound electron.
- highest for metals in Group 1A(1).

C

Question No. 2

Helium has _____ valence electrons.

- 4
- 2
- 6
- 8

B

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Question No. 3

The elements in a column of the periodic table are known as

- a period.
- Metalloids.
- Noble gases.
- a group.

D

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Question No. 5

Which of these elements has the lowest first ionization energy?

- phosphorus
- sodium
- aluminum
- sulfur

B

Question No. 2

What is the electron configuration for calcium atom?

- $1s^2 2s^2 2p^8 3s^2 3p^6$
- $1s^2 2s^2 2p^6 3s^2 3p^{10}$
- $1s^2 2s^2 2p^6 3s^2 3p^5 3d^3$
- $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$

D



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MKCL OES

Question No. 21

How many moles of potassium are contained in 449 g of potassium?

- 23.9 moles
- 69.2 moles
- 17.6 moles
- 11.5 moles

D





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Question No. 22

How many Li atoms are contained in 97.9 g of Li?

- 4.27×10^{22} Li atoms
- 8.49×10^{24} Li atoms
- 7.09×10^{21} Li atoms
- 5.90×10^{25} Li atoms

B





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Question No. 16

Which of the following is a characteristic of the modern periodic table?

- The elements in each group have different chemical properties.
- The elements in each group have similar chemical properties.
- A period is a column on the periodic table.
- A group is a horizontal row on the periodic table.

B



Question No. 22

The molecular formula of aspirin is $C_9H_8O_4$. How many aspirin molecules are present in one 500-milligram tablet?

- 1.67×10^{24} molecules
- 1.67×10^{21} molecules
- 2.77 molecules
- 2.77×10^{-3} molecules

B

Question No. 19

What is the term for a substance that accepts a proton in an acid-base reaction?

- Arrhenius base
- Brønsted-Lowry base
- Brønsted-Lowry acid
- Arrhenius acid

B

Question No. 7

An ionic bond is best described as

- the attraction between 2 nonmetal atoms.
- the attraction between 2 metal atoms.
- the transfer of electrons from one atom to another.
- the sharing of electrons.

C



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Question No. 19

A cation of +3 indicates that an element has

- gained three protons.
- gained three electrons.
- lost three protons.
- lost three electrons.

D



Question No. 17

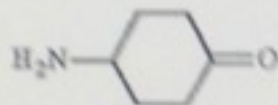
How many moles and how many atoms of zinc are in a sample weighing 34.9 g?

- 0.533 mol, 8.85×10^{-25} atoms
- 1.87 mol, 1.13×10^{-24} atoms
- 1.87 mol, 3.10×10^{-24} atoms
- 0.533 mol, 3.21×10^{23} atoms

D

Question No. 36

The functional groups in the molecule below are



- aldehyde and ketone
- ketone, and amine
- carboxylic acid, and amine
- aldehyde and amine.

D

Question No. 8

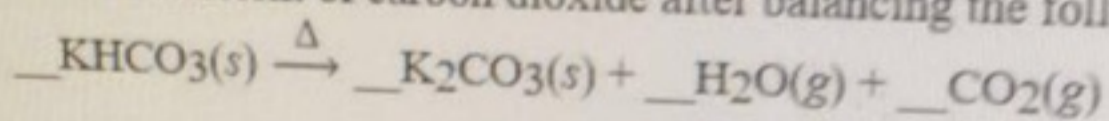
The name of the Cu^{2+} ion is _____.

- copper(III)
- copper
- copper(I)
- copper(II)

D

Question No. 6

What is the coefficient of carbon dioxide after balancing the following equation?



- 3
- 1
- 4
- 2

D

Question No. 1

A solution has a density of 1.22 g/mL. What volume of the solution has a mass of 55.0 g?

- 0.00253 mL
- 49.4 mL
- 58.8 mL
- 45.1 mL

D

Question No. 11

One element that has 5 valence electrons is _____.

- lithium
- sulfur
- nitrogen
- carbon

C

Question No. 18

Non Metals are located where on the periodic table?

- left side
- between metals and nonmetals
- middle
- right side

D

Question No. 3

The elements in a column of the periodic table are known as

- a period.
- Metalloids.
- Noble gases.
- a group.

D

Question No. 2

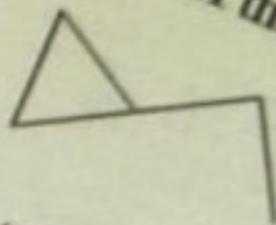
Helium has _____ valence electrons.

- 4
- 2
- 6
- 8

D

Question No. 35

Provide the name of the compound below.



- 2-cyclopropylethane
- ethylcyclopropane
- isopropylcyclopropane
- methylcyclopropane

B

Question No. 1

The maximum number of electrons that can go in the main energy level 3 is?

- 10
- 18
- 32
- 8

B

Question No. 1

Which of the following is NOT a correct name, symbol combination?

- magnesium, Mg
- Beryllium, Be
- calcium, Ca
- manganese, Mn



Question No. 1

Which statement below accurately describes the contributions of Josef Proust?

- ancient Greek philosopher who proposed that matter was continuous
- proposed the modern Atomic Theory
- discovered the law of definite proportions
- created the modern periodic table

C

Determine the molecular formula of a compound that has a molar mass of 146 g/mol and an empirical formula of $C_3H_5O_2$.

- $C_6H_{15}O_4$
- $C_3H_5O_2$
- $C_9H_{15}O_6$
- $C_6H_{10}O_4$

D

Which family is most likely to form diatomic molecules?

- Group V-A
- Group VII-A
- Group VIII-A
- Group VI-A

B

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What is the correct formula for the iron(III) ion?

- Fe^+
- Fe^{3+}
- Fe^{2-}
- Fe^{2+}

B

Save & Next حفظ والتالي

Which one of the following elements is a poor conductor of heat and electricity?

- copper
- iron
- lead
- fluorine

D

Question No. 6

Which of these elements is most likely to be a good conductor of electricity?

- S
- N
- He
- Fe

D

Save & Next

Question No. 18

How many liters of a 0.025 M NaOH solution contain 0.5 mole of NaOH?

- 10 L
- 0.5 L
- 20 L
- 5 L

C

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Question No. 3

Which of the following properties is NOT a characteristic of the Group 1A(1) elements (alkali metals)?

- They are good conductors of heat.
- They are shiny.
- They are good conductors of electricity.
- Most of them are liquids at room temperature.

D

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Question No. 20

Which of these compounds is a *strong electrolyte*?

- O₂
- C₆H₁₂O₆ (glucose)
- H₂O
- H₂SO₄

D

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Question No. 15

The name of the chemical compound CuOH is

- copper hydroxide
- copper(II) hydroxide
- copper(I) hydroxide
- copper(III) hydroxide



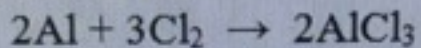
Question No. 14

The most correct name for the compound CS_2 is:

- carbon trisulfide
- carbon disulfide
- monocarbon disulfide
- carbon sulfide

D

How many moles of AlCl_3 could be produced when 284 grams of Cl_2 completely react with aluminum according to the reaction?



- 2.67 mol
- 3.92 mol
- 4.60 mol
- 3.33 mol

A

Question No. 25

The number of electron lone pairs in the water molecule is _____.

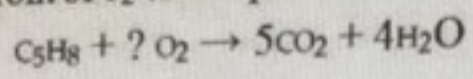
- 4
- 2
- 8
- 3

B

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Question No. 8

What coefficient is placed in front of O_2 to complete the balancing of the following equation?



- 7
- 1
- 3
- 5

A

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Question No. 7

Which group of elements is mostly found as diatomic molecules?

- alkaline earth metals
- alkali metals
- noble gases
- halogens

D

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What is the chemical formula for the binary compound composed of Mg^{2+} and O^{2-} ions?

- MgO
- MgO_2
- Mg_2O_2
- Mg_2O

A

Question No. 24

Which of the following elements has the lowest electronegativity?

- chlorine
- bromine
- fluorine
- iodine

D

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Question No. 10

What is the empirical formula of the compound that has a composition by mass of 34.6% C, 3.9% H and 61.5% O?

- CH_2O_3
- CHO
- CH_2O
- $\text{C}_3\text{H}_4\text{O}_4$

D

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HP LE1901w

Question No. 13

In a solution, the solvent is _____

- always a solid
- the substance present in the greatest amount
- always water
- substance present in a small amount

B

Question No. 7

An ionic bond is best described as

- the transfer of electrons from one atom to another.
- the attraction between 2 nonmetal atoms.
- the attraction between 2 metal atoms.
- the sharing of electrons.

A

Question No. 22

The conjugate base of H_2CO_3 is _____.

- HCO_3^-
- CO_3^{2-}
- H_2O
- OH^-

A

Question No. 14

The distinguishing characteristic of all electrolyte solutions is that th

- conduct electricity.
- react with other solutions.
- conduct heat
- contain molecules.

A

Question No. 12

What is the percent yield for a reaction if its theoretical yield is 115 g and its actual yield is 58 g?

- 56%
- 80%
- 64%
- 50%

D

Question No. 22

The conjugate base of H_2SO_4 is

- H_2SO_4
- HSO_4^-
- OH^-
- HSO_4^+

B

Question No. 24

Define potential energy.



- energy associated with the motion of an object
- energy associated with the force of an object
- energy associated with the position or composition of an object
- energy associated with the temperature of an object

Question No. 39

Which of the following types of linkage is found in a lipid?

- peptide linkage
- glycoside linkage
- ester linkage
- phosphate linkage

A

Question No. 9

What is the chemical formula for the binary compound composed of Mg^{2+} and O^{2-} ions?

- Mg_2O_2
- MgO_2
- MgO
- Mg_2O

C

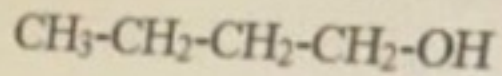
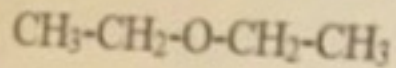
Question No. 21

The compound CO_2 can be described as _____.

- Lewis acid
- Lewis base
- Bronsted-Lowry acid
- Arrhenius acid

A

The two molecules represented below are examples of



- isomers.
- geometric isomers.
- identical.
- stereoisomers.

B

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Question No. 30

..... are the most reactive hydrocarbons.

- cycloalkanes
- alkanes
- alkenes
- alkynes

D



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Question No. 20

Which of the following compounds gives a non electrolyte aqueous solution?

- H_3PO_4
- H_2CO_3
- Na_2CO_3
- $C_{12}H_{22}O_{11}$

D

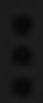
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Chem Q2 sem-2



What is the mass% of calcium in calcium chloride CaCl_2 ?

- 52.4%
- 63.9%
- 50.0%
- 36.1%

D

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Question No. 23

Short bonds are usually , while the long bonds are

- Weak, strong
- Strong, weak
- Strong, strong
- Weak, weak

B





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Question No. 24

Which of these atoms is the most electronegative?

- Al
- Mg
- Na
- P

D





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Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Matter is conserved.
- Matter is rearranged.
- Atoms react with other atoms.
- Atoms are changed into other atoms.

D





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Question No. 21

What is the term for a compound that releases hydrogen ions when dissolved in water?

- base
- acid
- aqueous solution
- aqueous salt

B





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Question No. 22

What term describes a reactant that gains electrons from another reactant?

- reducing agent
- base
- precipitate
- oxidizing agent

D





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Chem Q2 sem-2



The most correct name for the compound P_4S_{10} is:

- tetraphosphorus decasulfide
- tetraphosphorus nonasulfide
- tetraphosphorus octasulfide
- triphosphorus decasulfide





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An ionic bond is best described as

- the sharing of electrons.
- the attraction between 2 nonmetal atoms.
- the attraction between 2 metal atoms.
- the transfer of electrons from one atom to another.

D

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What is the oxidation number of sulfur in H_2SO_4 ?

- 2
- +4
- +6
- +2

What is the oxidation number of sulfur in H_2SO_4 ?

- 2
- +4
- +6
- +2

C

Which of the following elements has the lowest electronegativity?

- Ba
- Mg
- Sr
- Be

D

Question No. 23

Which molecule contains the shortest carbon-carbon bond?

- $\text{H}_3\text{C}-\text{CH}_3$
- $\text{F}_2\text{C}=\text{CF}_2$
- $\text{HC}\equiv\text{CH}$
- $\text{H}_2\text{C}=\text{CH}_2$

C

Save & Next حفظ والتالي

Which of the following is the strongest acid?

- H_3PO_4
- NaOH
- H_2CO_3
- HNO_3

D

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Question No. 6

A neutron has an electrical charge of _____.

- 2
- +1
- 1
- 0

D

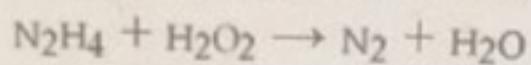
What is the final molarity of HCl solution, if 25 mL of 2 M HCl was diluted to a final volume of 400 mL?

- 0.125 M
- 0.16 M
- 0.08 M
- 0.25 M

A

Question No. 8

What are the correct coefficients needed to balance the reaction below?



- 1, 4, 1, 4
- 2, 4, 2, 8
- 1, 1, 1, 1
- 1, 2, 1, 4

D

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Question No. 10

What is the empirical formula of the compound that has a composition by mass of 84.2% C and 15.8% H?

- C_3H_9
- C_4H_{10}
- C_3H_8
- C_4H_9

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Which of the following compounds gives a non electrolyte aqueous solution?

- H_2CO_3
- $\text{C}_{12}\text{H}_{22}\text{O}_{11}$
- Na_2CO_3
- H_3PO_4

B



Question No. 12

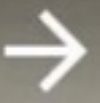
The chemical formula of the compound formed between sodium and fluorine is:

- NaF₃
- Na₂F
- NaF
- NaF₂





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Question No. 7

One of the following elements exists as a diatomic molecule:

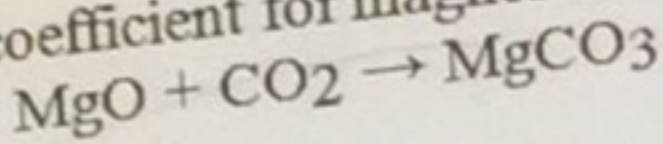
- neon
- argon
- iron
- chlorine

D



Question No. 8

In this reaction, what is the coefficient for magnesium oxide?



- 3
- 4
- 1
- 2

C

Question No. 15

Which of the following polyatomic ions has a positive charge?

- ammonium ion
- hydroxide ion
- sulfate ion
- carbonate ion

A

Question No. 19

What is the molarity of a hydrochloric acid solution prepared by diluting 500.0 mL of 1.00 M HCl to a total volume of 2.50 L?

- 0.500 M
- 2.00 M
- 0.100 M
- 0.200 M

D

Question No. 11

The name of the Cu^+ ion is _____.

- copper(III)
- copper
- copper(II)
- copper(I)

D

Question No. 14

The most correct name for the compound SBr_6 is:

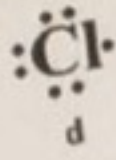
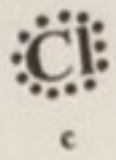
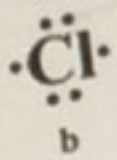
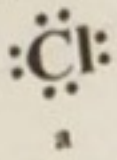
- monosulfur heptabromide
- sulfur bromide
- monosulfur hexabromide
- sulfur hexabromide

D

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Question No. 25

Which of the following is the correct electron dot structure for chlorine (atomic no. = 17)



الإختيار B الرسمة D

- c
- d
- a
- b

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Question No. 19

What is the concentration of HCl in the final solution when 65 mL of a 12 M HCl solution is diluted with distilled water to a total volume of 0.15 L?

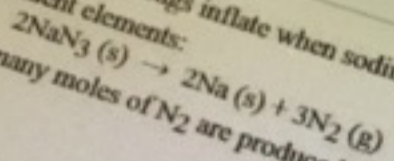
- 5.2 M
- 2.8×10^{-2} M
- 5.2×10^3 M
- 28 M

A



Question No. 16

Automotive air bags inflate when sodium azide (NaN_3) decomposes explosively to its constituent elements:



How many moles of N_2 are produced by the decomposition of 3.55 mol of sodium azide?


- 10.7
- 2.37
- 1.18
- 5.33

D

Question No. 23

What is the term for a bond composed of three electron pairs shared between two atoms?

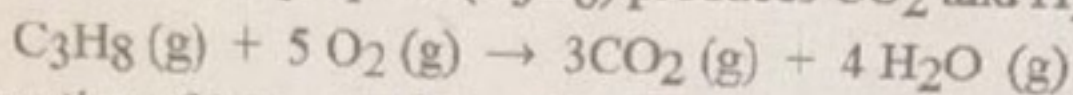
- electrovalent bond
- single bond
- double bond
- triple bond

DSave & Next 



Question No. 16

The combustion of propane (C₃H₈) produces CO₂ and H₂O:



The reaction of 5.5 mol of O₂ will produce _____ mol of H₂O.

- 4.4
- 2
- 5
- 5.5

A

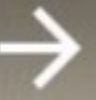




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Question No. 14

The most correct name for the compound SCl_8 is:

- sulfur nonachloride
- sulfur chloride
- sulfur octachloride
- monosulfur octachloride

C



Question No. 18

If 4.0 moles of KBr is dissolved in enough water to make 12.0 L of solution, what is the molarity of this solution?

- 0.25 M
- 0.75 M
- 0.33 M
- 4.0 M

C

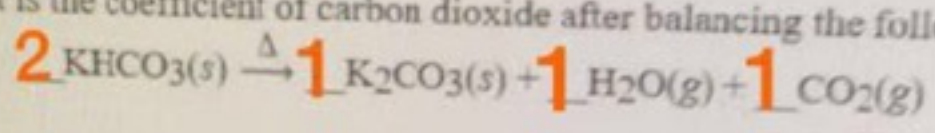


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Question No. 6

What is the coefficient of carbon dioxide after balancing the following equation?



- 3
- 1
- 4
- 2

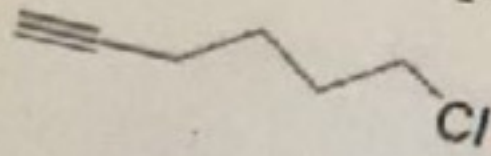
B



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Question No. 34

What is the IUPAC name for the following compound?



- 6-chloro-1-pentyne
- 6-chloro-2-pentyne
- 1-chloro-5-hexyne
- 6-chloro-1-hexyne

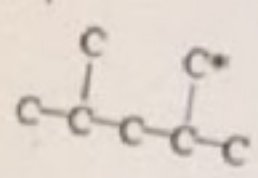
D

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Question No. 32

The carbon skeleton of an alkane is shown below. How many hydrogen atoms are bonded to the carbon marked with a *?



- 1
- 0
- 2
- 3



D

Save & Next

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MKCL OES

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Question No. 39

Which of the following types of linkage is found in a lipid?

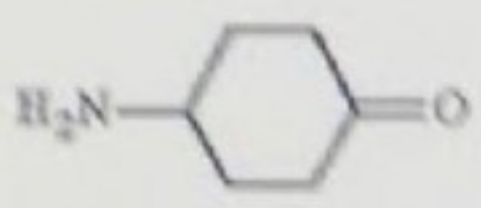
- peptide linkage
- glycoside linkage
- ester linkage
- phosphate linkage

C



Question No. 36

The functional groups in the molecule below are



- aldehyde and ketone
- ketone, and amine
- carboxylic acid, and amine
- aldehyde and amine.

B

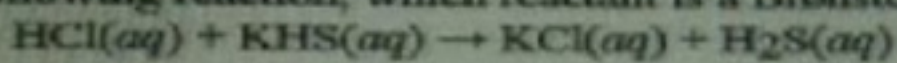


MKCL OES

Chem

Question No. 21

In the following reaction, which reactant is a Brønsted-Lowry base?



- KCl
- KHS
- H₂S
- HCl

B



Question No. 6

Which of the following is a general guideline for balancing an equation?

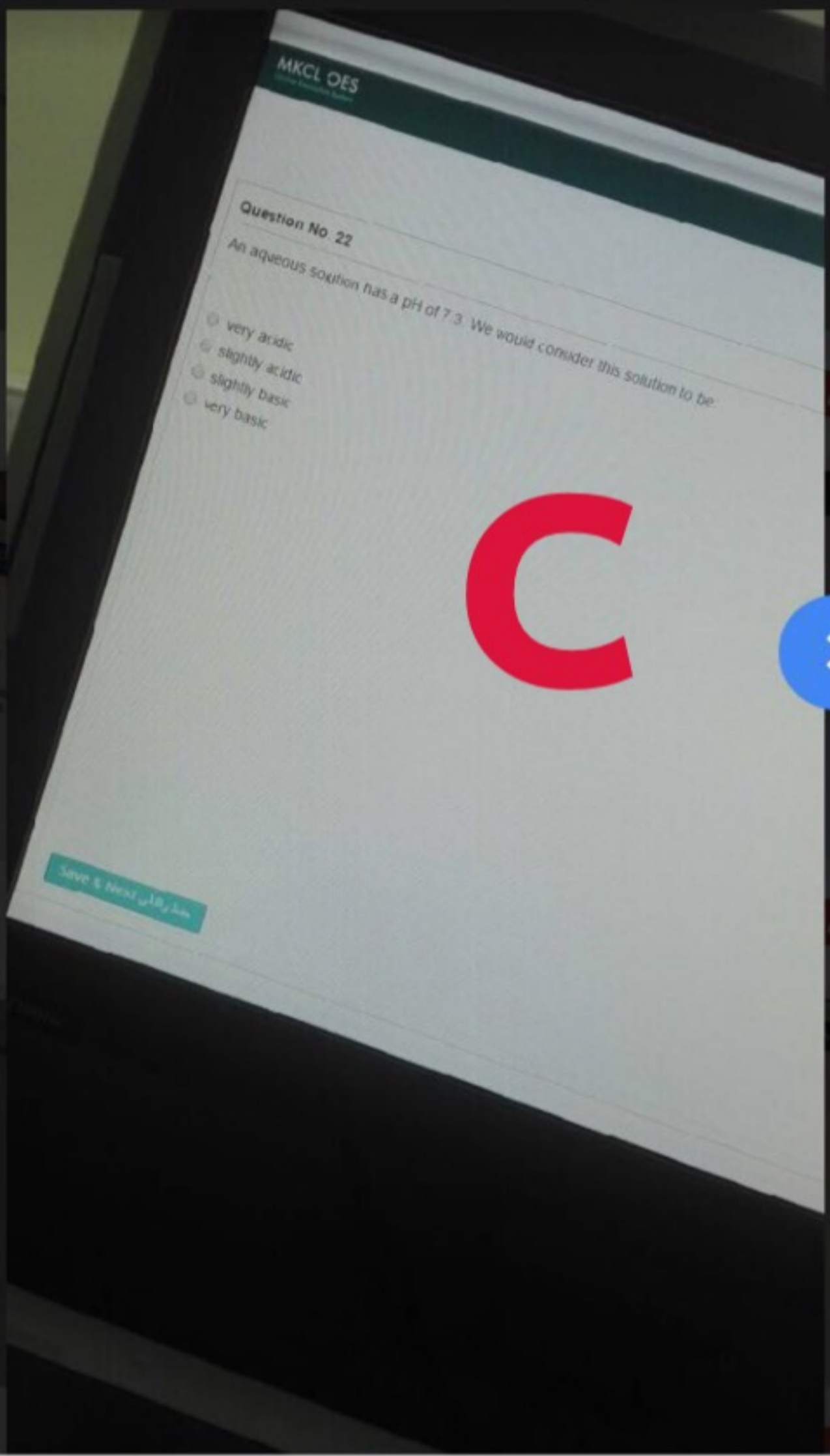
- All are correct
- Check each reactant and product to verify the coefficients.
- Balance polyatomic ions as a single unit.
- Write correct formulas for reactants and products.

A

* organic compounds that contain a benzene ring
possess certain properties similar to those of benzene
are called ----- compounds.

- ① saturated .
- ② acidic .
- ③ aromatic .
- ④ carboxylic .

C



Question No. 18

Which of the following is true before a reaction reaches chemical equilibrium?

- The rate of the forward reaction is increasing, and the rate of the reverse reaction is decreasing.
- The rates of the forward and reverse reactions are decreasing.
- The rate of the forward reaction is decreasing, and the rate of the reverse reaction is increasing.
- The rates of the forward and reverse reactions are increasing.

A



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IMG_9996.JPG



Question No. 23

Which of the following is true if the hydronium ion concentration of an aqueous solution increases?

- K_w increases
- pH increases
- K_w decreases
- pH decreases

D



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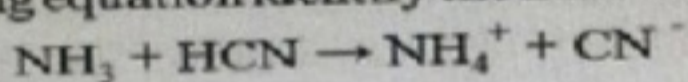
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IMG_9996.JPG

Question No. 21

In the following equation identify the Brønsted-Lowry acid:



- CN^-
- NH_3
- HCN
- NH_4^+

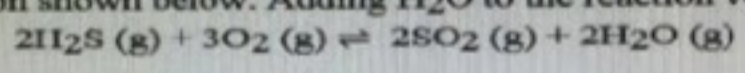


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Question No. 20

Refer to the reaction shown below. Adding H₂O to the reaction vessel will _____.



- cannot be determined, since the temperature is unknown
- shift the reaction to the right
- shift the reaction to the left
- have no effect



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Question No. 13

In a solution, the solute is _____

- the substance present in the greatest amount
- always water
- always a solid
- the substance present in the smallest amount





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Question No. 25

Identify the conjugate base of HPO_4^{2-} in the reaction
 $\text{HCO}_3^- + \text{HPO}_4^{2-} \leftrightarrow \text{H}_2\text{CO}_3 + \text{PO}_4^{3-}$

- H_2O
- HCO_3^-
- PO_4^{3-}
- H_2CO_3





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Question No. 19

When the reverse reaction is favored

- The rate of the forward reaction is greater than the reverse reaction.
- The rate of the reverse reaction is less than the forward reaction.
- The equilibrium constant is much less than one; that is, $K_{eq} \ll 1$
- The equilibrium constant is much greater than one; that is, $K_{eq} \gg 1$

C





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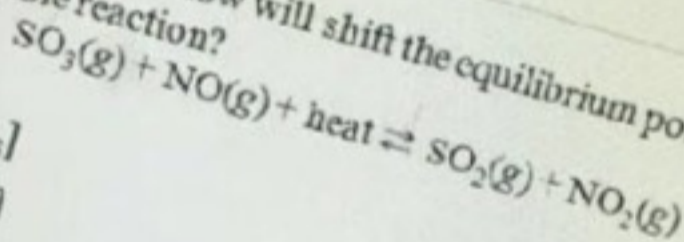


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Chemistry_

Question No. 20

Which of the changes listed below will shift the equilibrium position to the right for the following reversible reaction?



- An increase of $[\text{SO}_3]$
- An increase of $[\text{SO}_2]$
- A decrease of temperature
- An increase of volume

B





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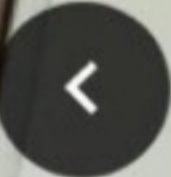
MKCL OES

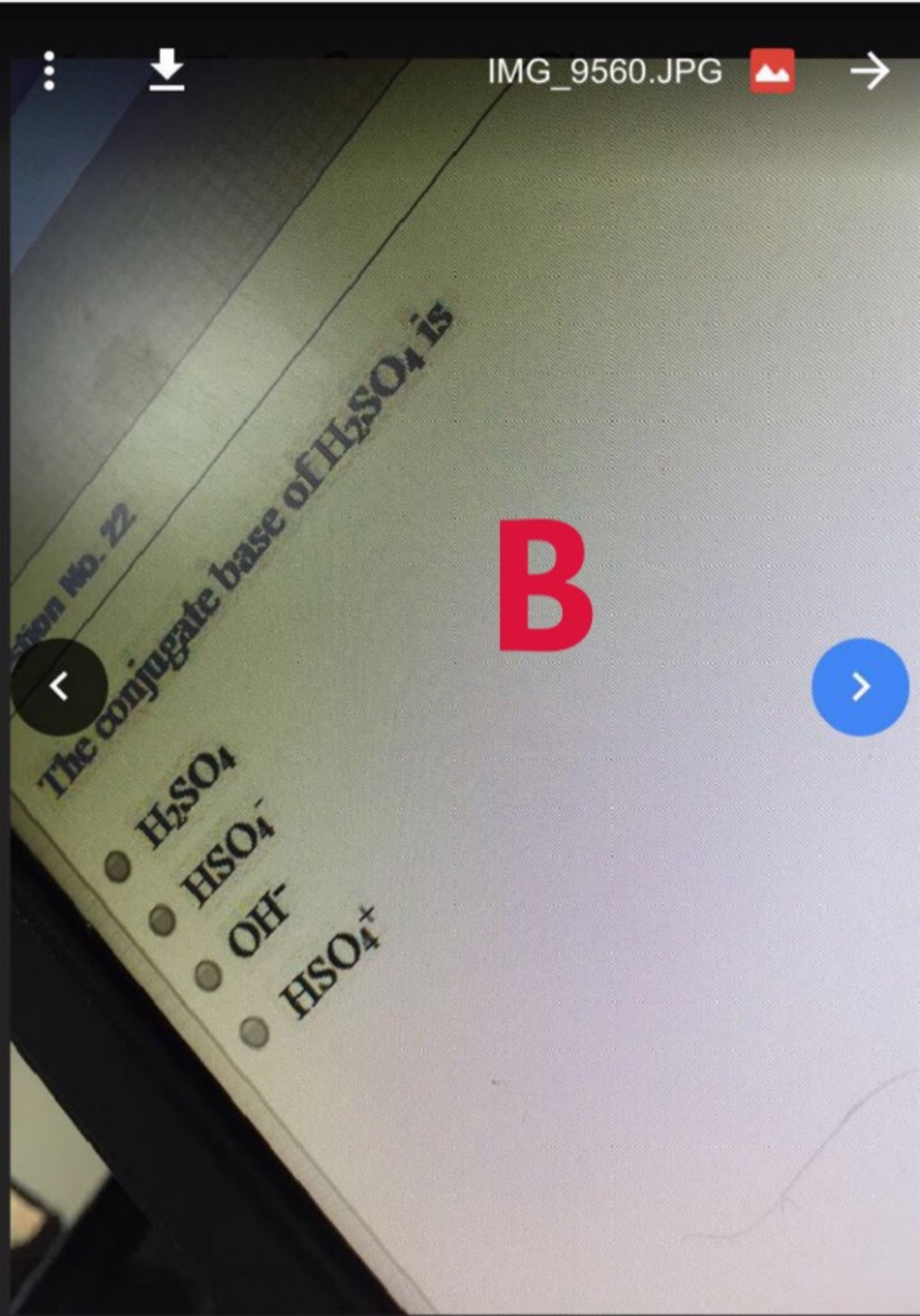
Question No. 39

The term "carbohydrates" refers to a large class of polyhydroxylated _____

- alcohols and carboxylic acids
- amines and amides
- aldehydes and ketones
- ethers and esters

C





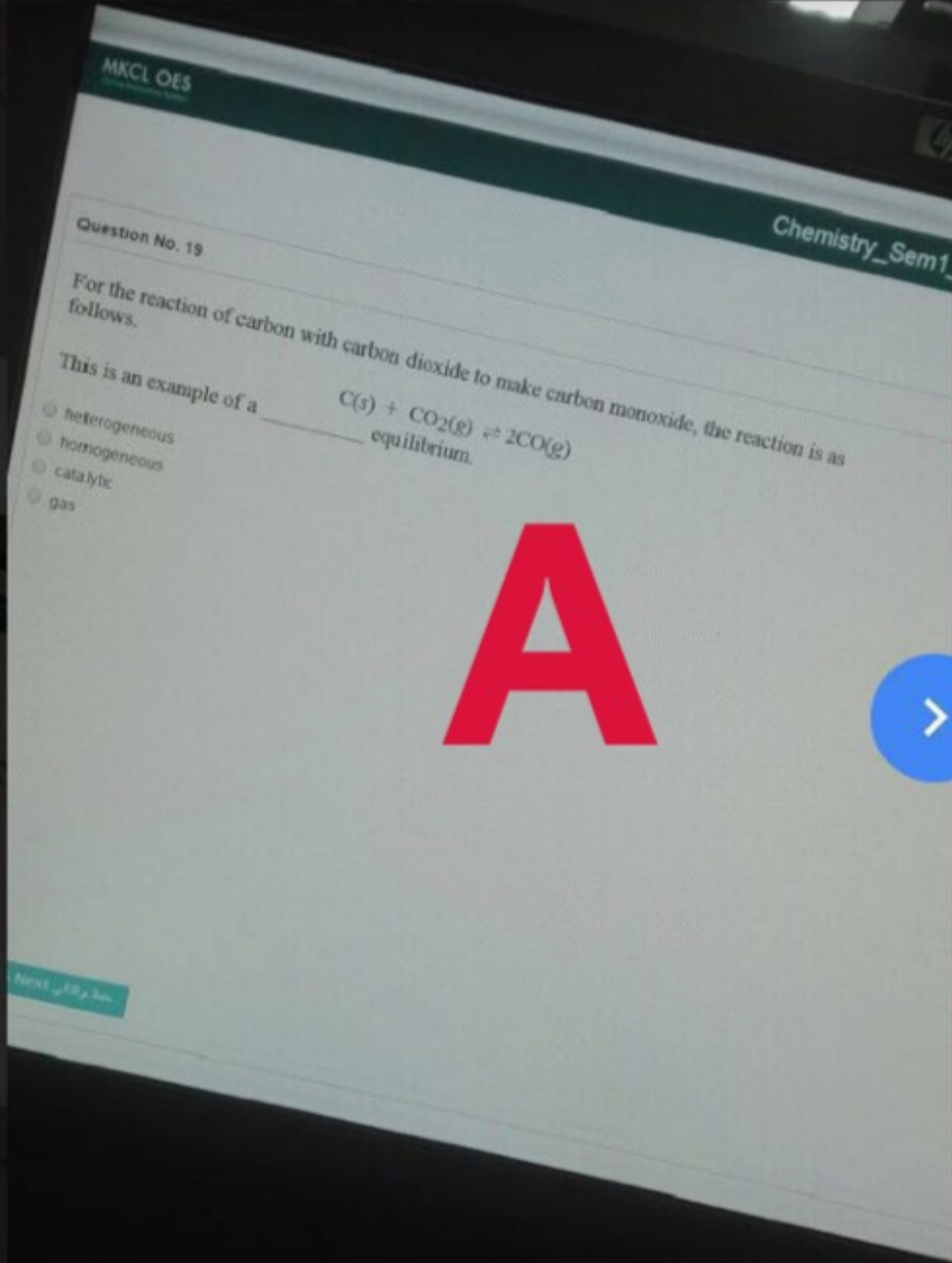
Question No. 30

... are the most reactive hydrocarbons.

- cycloalkanes
- alkanes
- alkenes
- alkynes

D







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Question No. 22

What is the conjugate acid of OH^- ?

- O_2^{2-}
- H_3O^+
- H_2O
- O^-





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Question No. 16

The ability of an atom to attract the shared electrons in a covalent bond is its

- polarity.
- ionic character.
- bonding ability.
- electronegativity.

D



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Question No. 14

Which of these compounds is a *non-electrolyte*?

- CH_3COOH (acetic acid)
- $\text{C}_6\text{H}_{12}\text{O}_6$ (glucose)
- HNO_3
- NaF

B

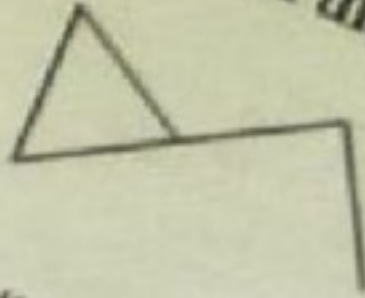




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Question No. 35

Provide the name of the compound below.



- 2-cyclopropylethane
- ethylcyclopropane
- isopropylcyclopropane
- methylcyclopropane

B



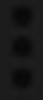


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Chem Final



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Question No. 19

What is the term for a substance that accepts a proton in an acid-base reaction?

- Arrhenius base
- Brønsted-Lowry base
- Brønsted-Lowry acid
- Arrhenius acid

B



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Question No. 39

Which of the following is a monosaccharide?

- amylose
- sucrose
- starch
- galactose

D





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MKCL OES
Online Evaluation System

Question No. 11

The most correct name for the compound SBr_6 is:

- monosulfur heptabromide
- sulfur bromide
- sulfur hexabromide
- monosulfur hexabromide

C





Question No. 21

The compound CO_2 can be described as _____

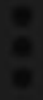
- Lewis acid
- Lewis base
- Bronsted-Lowry acid
- Arrhenius acid

A





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Chem Final



Question No. 13

In a solution, the solvent is _____

- always a solid
- the substance present in the greatest amount
- always water
- substance present in a small amount





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Chem Final



Question No. 1

The maximum number of electrons that can go in the main energy level 3 is?

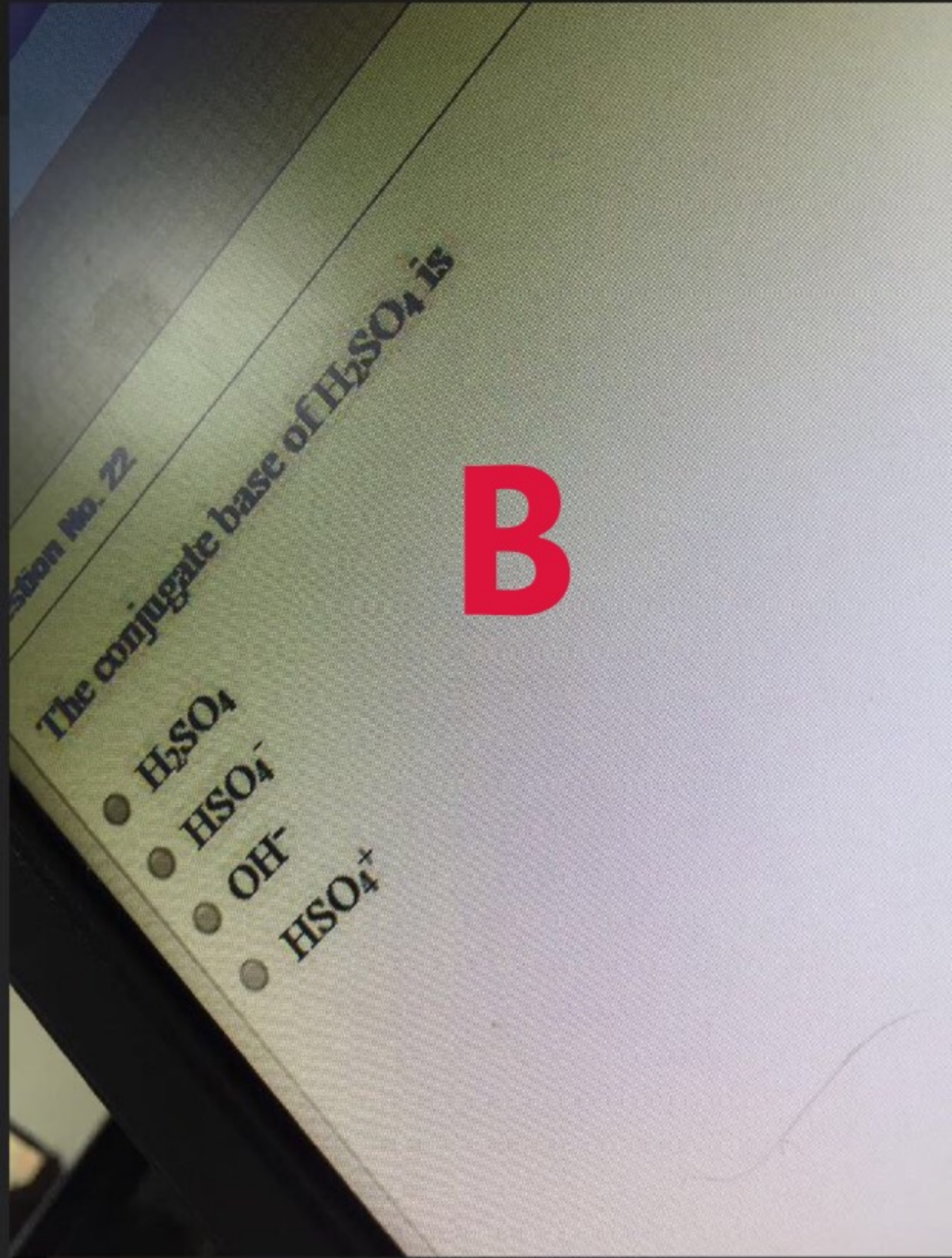
- 10
- 18
- 32
- 8

B





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B



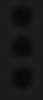


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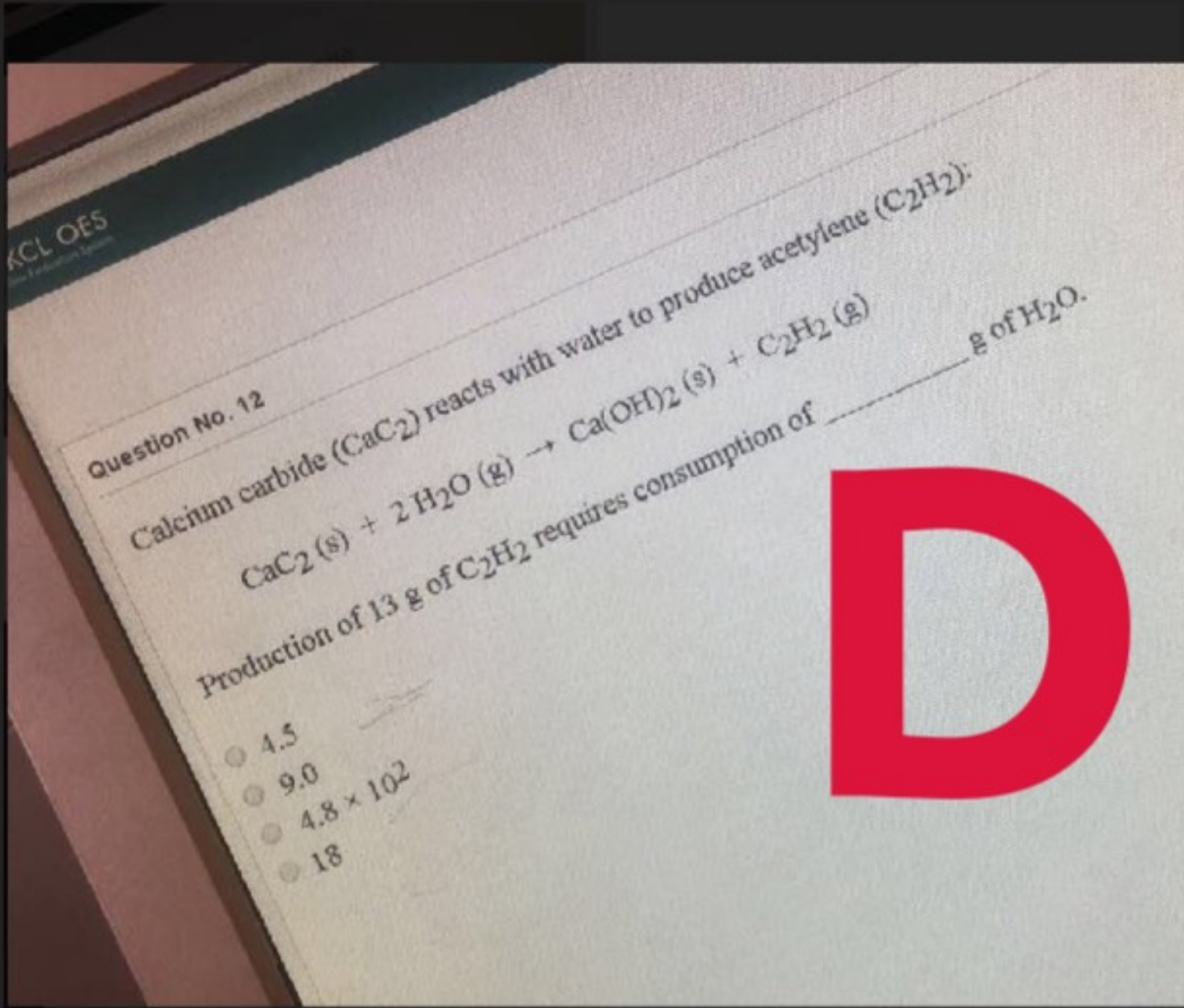
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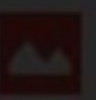
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Question No. 21

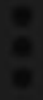
The compound BF_3 can be described as _____

- Arrhenius acid
- Bronsted-Lowry acid
- Lewis acid
- Lewis base





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Chem Final



Question No. 20

Which of the following is true if the pH of a solution changes from 5 to 7?

- [H⁺] decreases
- K_w decreases
- K_w increases
- [H⁺] increases

A



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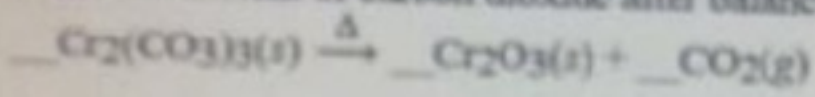




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Question No. 8

What is the coefficient of carbon dioxide after balancing the following equation?



- 2
- 4
- 3
- 1

C





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Question No. 4

The nucleus of an atom consists mainly of

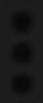
- protons and electrons.
- protons and neutrons.
- neutrons and electrons.
- protons, neutrons, and electrons.

B





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Chem Final



Question No. 21

Which of the following are Lewis bases?

- I) BCl_3
- II) H^-
- III) H_2O
- IV) NH_3

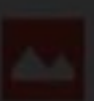
- I) and II)
- I), II), and III)
- III) and IV)
- II), III) and IV)



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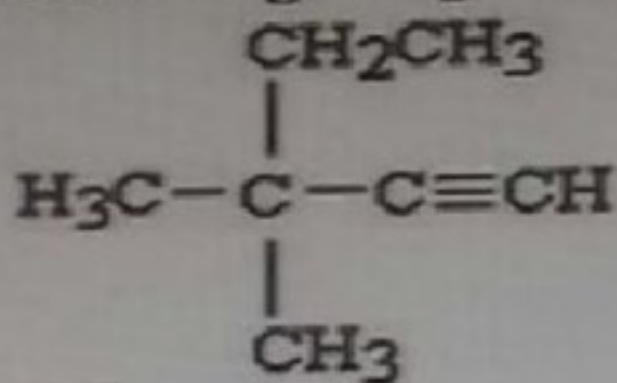


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MRCL QES

Question No. 34

Name the following compound.



- 2-ethyl-2-methyl-3-butyne
- 3, 3-dimethyl-1-pentyne
- 3-ethyl-3-methyl-1-butyne
- 1-butyne

B





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Chem Final



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Question No. 19

What is the term for a substance that accepts a proton in an acid-base reaction?

- Brønsted-Lowry base
- Arrhenius base
- Brønsted-Lowry acid
- Arrhenius acid



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Chem Final



MKCL OES

Question No. 40

An amino acid is a compound that contains at least

- one carboxylic acid group and one amino group.
- two amino groups and one carboxylic acid group.
- one hydroxyl group and one methyl group.
- one amino group and one amide group

A



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Question No. 25

A process or reaction which absorb heat from the surroundings is said to be

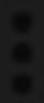
- exothermic.
- conservative.
- isothermal.
- endothermic.

D





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Chem Final



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Question No. 20

An aqueous solution has a pH of 7.3. We would consider this solution to be:

- slightly basic
- slightly acidic
- very basic
- very acidic

A

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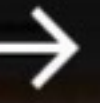
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MKCL OES

Question No. 31

The formula for butane is

- C_4H_8
- C_5H_{12}
- C_6H_{14}
- C_4H_{10}





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Question No. 6

Which of the following is a general guideline for balancing an equation?

- Check each reactant and product to verify the coefficients.
- Write correct formulas for reactants and products.
- Balance polyatomic ions as a single unit.
- All are correct

D





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AKCL OES

Question No. 33

Which of the following molecular formulas corresponds to an alkene?

- C_7H_{18}
- C_7H_{14}
- C_7H_{16}
- C_7H_{12}

B



Question No. 37

Organic compounds that contain a benzene ring or possess certain properties similar to those of benzene are called _____ compounds.

- saturated
- acidic
- aromatic
- carboxylic

C



Question No. 38

Sucrose is hydrolyzed to what two monosaccharides?

- glucose and galactose
- glucose and fructose
- maltose and fructose
- galactose and lactose

B





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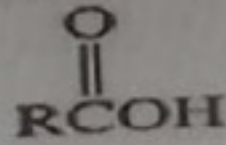
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Chem Final

MKCL QES

Question No. 36

What is the name of compound has the following general formula?



- phenol
- carboxylic acid
- ester
- aldehyde

B





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MKCL OES

Question No. 30

The names of compounds with carbon-carbon triple bonds contain the suffix _____

- one
- ene
- yne
- ane



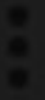


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Question No. 38

What is the term for a lipid composed of glycerol and three fatty acids?

- wax
- oils
- triglyceride
- cholesterol





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MKCL QES

Question No. 39

Which of the following is a monosaccharide?

- amylose
- sucrose
- starch
- galactose

D



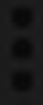


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Question No. 15

In a neutralization reaction

- an acid and a base react to form a salt and water.
- water and a salt react to form an acid and a base.
- an acid and a salt react to form water and a base.
- two acids react to form water.



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Question No. 24





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Chem Final



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Question No. 15

In a neutralization reaction

- an acid and a base react to form a salt and water.
- water and a salt react to form an acid and a base.
- an acid and a salt react to form water and a base.
- two acids react to form water.

A



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Question No. 24





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Chem Final



Question No. 7

What is the empirical formula of P_4O_{10} ?

- $P_1O_{2.5}$
- P_2O_5
- P_4O_{10}
- PO

B

Question No. 24



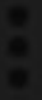


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Chem Final



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Question No. 10

The molar mass of $C_3H_8O_2$ is

- 69 g/mol
- 76 g/mol
- 32 g/mol.
- 42 g/mol

B



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Question No. 24





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What is the term for a substance that releases hydroxide ions (OH^-) in water?

- Arrhenius base
- Arrhenius acid
- Brønsted-Lowry acid
- Brønsted-Lowry base

A

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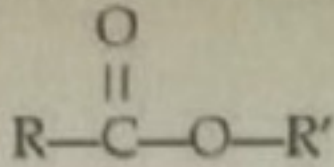




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Question No. 36

What is the name of compound has the following general formula?



- carboxylic acid
- aldehyde
- ketone
- ester

D





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Question No. 3

All of the following statements about different elements are true EXCEPT

- Krypton is one of the noble gases.
- Sulfur is considered a metalloid.
- Barium is an alkaline earth metal.
- Manganese is a transition metal.

B





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Question No. 38

Which of the following is a polysaccharide?

- fructose
- glucose
- galactose
- starch

D





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Question No. 39

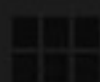
Which suffix on a ketone or aldehyde indicates a sugar?

- ene
- ane
- ose
- one





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Chem Final



Question No. 18

Which of the following is true before a reaction reaches chemical equilibrium?

- The rate of the forward reaction is increasing, and the rate of the reverse reaction is decreasing.
- The rates of the forward and reverse reactions are decreasing.
- The rate of the forward reaction is decreasing, and the rate of the reverse reaction is increasing.
- The rates of the forward and reverse reactions are increasing.

A



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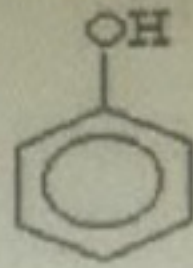




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Question No. 35

What is the name of compound shown below?



- aniline
- benzene
- phenol
- toluene





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Question No. 19

Which substance can be called an Arrhenius base?

- HBr
- CH₃OH
- NaCl
- KOH

D





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Question No. 16

All of the following compounds has a covalent bond except.

- AlCl_3
- BrF
- ICl
- HBr

A

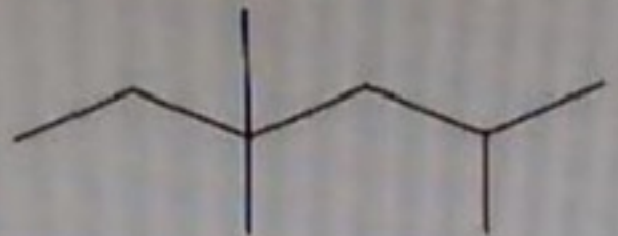


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Question No. 32

What is the IUPAC name of the following structure?



- 1,1,3,3-tetramethylpentane
- 2,2,5-trimethylhexane
- 2,4,4-trimethylhexane
- 3,3,5-trimethylhexane

C





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Question No. 15

Which one of the following is a weak acid?

- HI
- HCl
- HNO₃
- HF

D





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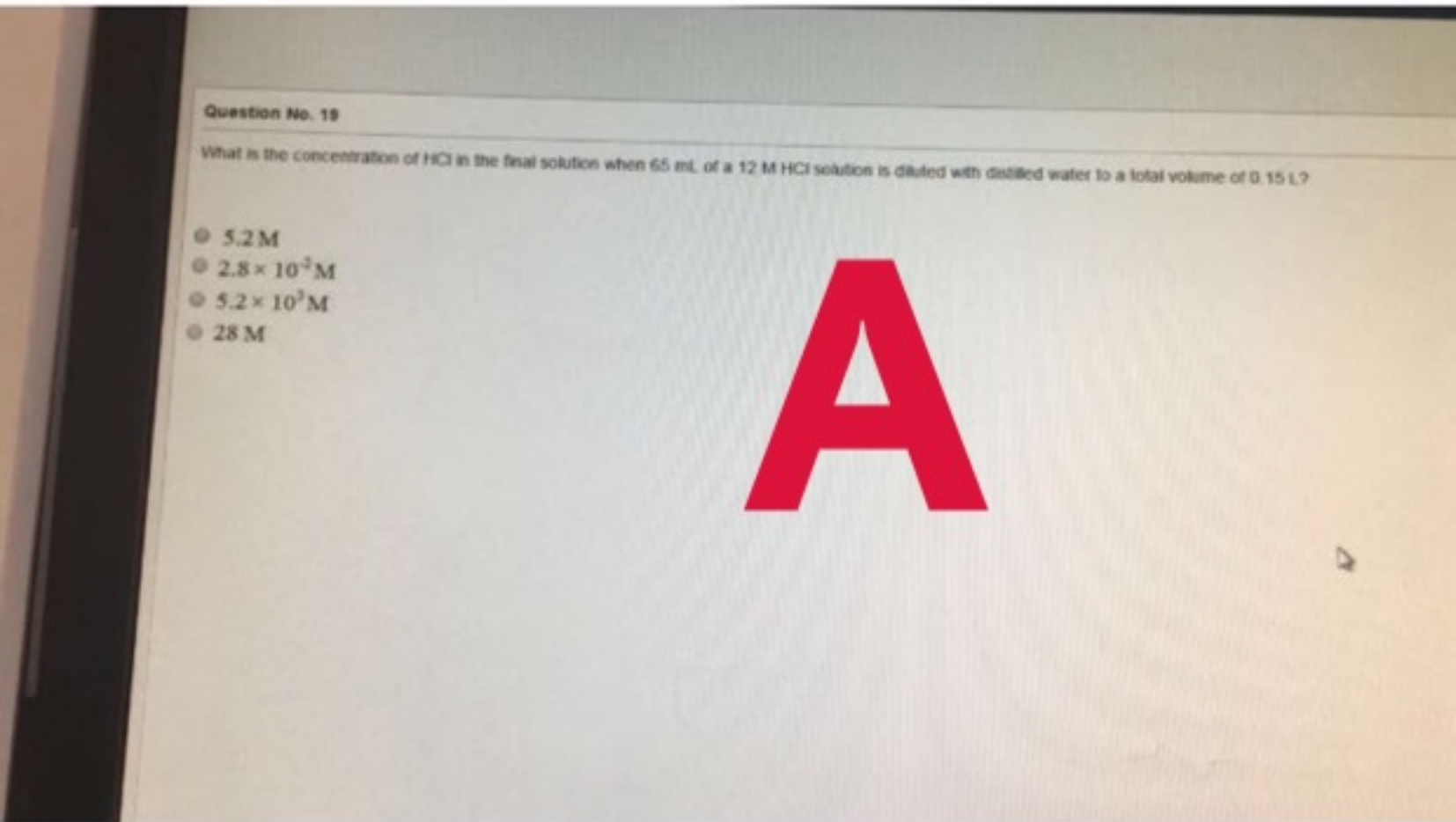
Question No. 33

What is the IUPAC name for the following compound?



- 2-pentene
- 1-pentene
- 1-methylbutene
- 3-pentene

A





Question No. 19

What is the molarity of a hydrochloric acid solution prepared by diluting 500.0 mL of 1.00 M HCl to a total volume of 2.50 L?

- 0.500 M
- 2.00 M
- 0.100 M
- 0.200 M

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