

The nucleus of an atom consists mainly of

- protons and neutrons.
- protons, neutrons, and electrons.
- protons and electrons.
- neutrons and electrons.

A

One decimeter is equal to _____.

- 10 cm
- 10 m
- 1000 mm
- 1 m

A

Which of the following is NOT a correct (name, symbol) combination?

- iron, I
- manganese, Mn
- magnesium, Mg
- beryllium, Be

A

What is the charge on the ion formed by sodium?

- 1+
- 2+
- 2-
- 1-

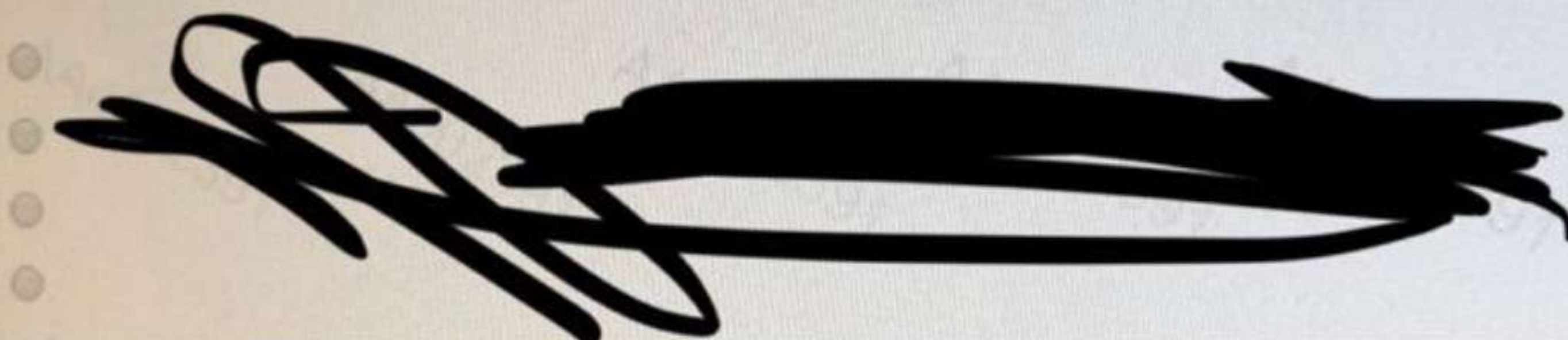
A

Determine the volume of an object that has a mass of 455.6 g and a density of 19.3 g/cm^3 .

- 18.5 mL
- 23.6 mL
- 42.4 mL
- 87.9 mL

B

What is the ion with the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^6$?



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A +

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Atoms react with other atoms.
- Matter is conserved.
- Matter is rearranged.

A

The number of electrons in the outer energy level of a neutral atom of boron is _____

- 5
- 8
- 3
- 2



Which of the following can be broken down into simpler substances?

MgO

Fe

Co

Al

A

Which one of these represents a physical change?

- sugar, when heated, becomes brown
- milk turns sour
- bleach turns hair yellow
- water, when heated, forms water vapor

D

Which of the following is an element?

salt

sugar

water

gold

D

Question No. 22

In the periodic table, group 1A (1) elements are also called

- alkaline earth metals
- alkali metals
- halogens
- noble gases

B

The outside air temperature is 30 °F, what is the temperature in Kelvin?

- 274 K
- 272 K
- 307 K
- 303 K

B

Semiconductors are located in the periodic table on (or in) the

- line dividing metals and nonmetals.
- first period.
- right side.
- left side.



If the density chloroform is 1.492 g/mL, what is the mass of 225 mL?

- 225 g
- 1.60 g
- 151 g
- 336 g

D

The mass number of an atom can be calculated from the _____

- number of protons
- number of electrons + protons
- number of protons + neutrons
- number of electrons



If the temperature of an object is $-68\text{ }^{\circ}\text{C}$, What is the corresponding reading on the Kelvin scale?

- 91 K
- 48 K
- 127 K
- 205 K



D

Which state of matter has atomic spacing that is close together and unixed shape?

- gas
- plasma
- solid
- liquid

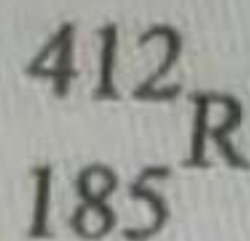
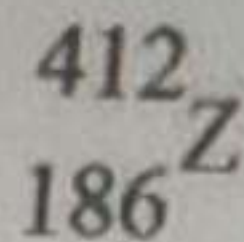
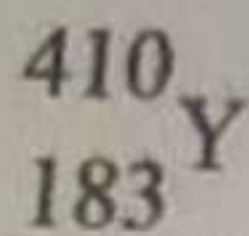
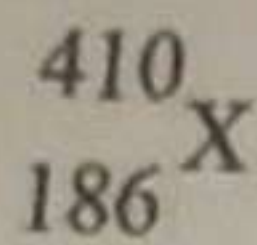


A specific isotope of an element is known to have 15 protons and 16 neutrons. Which symbol would properly represent this isotope?

- $^{16}_{15}\text{X}$
- $^{31}_{16}\text{S}$
- $^{31}_{15}\text{P}$
- $^{15}_{31}\text{Ga}$



Atoms X, Y, Z, and R have the following nuclear compositions:



Which two are isotopes?

- Z & R
- X & Z
- Y & R
- X & Y

B

Which of the following can NOT be considered as matter?

- heat
- paper
- wood
- sand

A

All of the following can be considered physical properties EXCEPT _____

- flammability
- boiling point
- color
- taste

A

Which of the following statements is incorrect?

- Evaporation of water from a fish tank is evidence of a physical change.
- The fact that sulfur powder is yellow is a physical property.
- The burning of a piece of charcoal to a white powder is an example of a physical change.
- The condensation of steam on a mirror is an example of a physical change.



When an atom loses an electron, the resulting particle is called

- a cation.
- an isotope.
- an anion.
- a proton.

A

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Question No. 3

Compounds are pure substances that by definition consist of _____

- a single element
- two or more combined elements
- two or more mixed elements
- solids

Save & Next

B



A تجميعات 2018

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Question No. 3

Which of the following statements about energy is FALSE?

- Energy cannot be transformed from one form to another.
- Energy is the capacity to do work.
- Energy can neither be created nor destroyed.
- An object possessing energy can do work on another object.

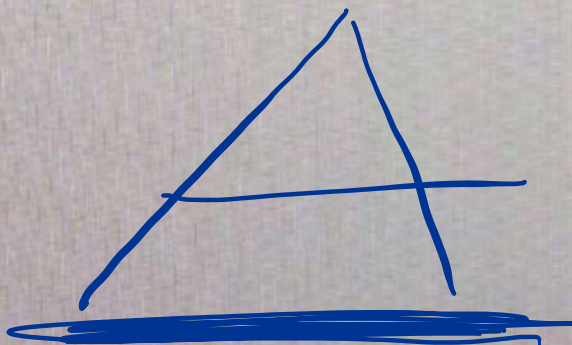
A

Question No. 17

One element that has 8 valence electrons is _____

- Argon
- Helium
- Carbon
- Oxygen

~~D~~



Question No. 15

Helium atom has _____ valence electrons.

- 2
- 4
- 8
- 6

A

Question No. 21

Which of the following statements is incorrect?

- All atoms of an element have the same atomic number.
- All atoms of an element must have the same mass.
- Atoms of an element may have different numbers of neutrons.
- All atoms of an element have the same number of electrons.

B

Question No. 18

How many electrons are in Rf^{3+} ?

- 4
- 34
- 07
- 56

36

D



1	2	3
8	9	1
15	16	2
22	23	3

Question No. 4

Which of the following is an example of a homogeneous mixture?

- sand mixed with oil
- salad dressing
- salt dissolved in water
- gasoline mixed with water



جاني نفس السؤال بس كآن غير متجنس و حطية

١:٥٢ م

a

Question No. 2

What is the mass of 3.5 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 0.023 kg
- 4.03 kg
- 0.015 kg

C

Save & Next حفظ والتالي

Question No. 20

The smallest subatomic particles is _____.

- neutron
- nucleus
- electron
- proton

C

Question No. 15

Which of these elements is chemically similar to Sulfur?

- S
- O
- Cl
- P

B

Number of qu

15 Assessed

6 Not Valued

1	2
8	9
15	16
22	

Question No. 17

Semiconductors are located in the periodic table on (or in) the _____.

- left side.
- line dividing metals and nonmetals.
- right side.
- first period.

B

Question No. 21

How many neutrons are there in an atom of lead (Pb) whose mass number is 207?

- 82
- 207
- 125
- 290

125

Question No. 18

An atom of ^{14}C contains _____ protons.

- 6
- 8
- 14
- 12

A

Question No. 12

Which of the following is the highest temperature?

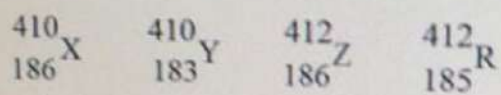
- 40 °C
- 200 °F
- 323 K
- freezing point of water



B

Question No. 25

Atoms X, Y, Z, and R have the following nuclear compositions:



Which two are isotopes?

- X & Y
- X & Z
- Y & R
- Z & R

B

Question No. 21

The mass number of an atom can be calculated from the _____.

- number of electrons
- number of protons + neutrons
- number of electrons + protons
- number of protons

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B

Question No. 14

What element has the electron configuration $1s^2 2s^2 2p^6 3s^2$?

- Be
- Mg
- Ca
- Na

B

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Question No. 3

Compounds are pure substances that by definition consist of _____

- a single element
- two or more combined elements
- two or more mixed elements
- solids

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B



Question No. 15

A chloride ion (Cl^-) has the same electron configuration as a(n) _____.

- Argon atom
- Neon atom
- Fluorine atom
- Sulfur atom

A



Question No. 14

When placing the electrons in orbitals of the same energy, the electrons are first placed singly, the rule is known as:

- Lavoisier's principal
- Mendeleev's rule
- Dalton's atomic theory
- Hund's rule



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Question No. 6

In which of the following is the metric unit paired with its correct abbreviation?

- microgram / mg
- milliliter / mL
- kilogram / cg
- centimeter / km

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B

Question No. 10

A physical change _____.

- occurs when glucose is converted into energy inside your cells
- occurs when iron rusts
- occurs when sugar is burnt
- occurs when water is evaporated

D

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Question No. 10

When an egg is fried, what type of process is happening?

- melting
- a physical change
- evaporation
- a chemical change

D

User: AA400233

Number of ques

9 Answered
15 Not Visited

1	2	3
8	9	10
15	16	17
22	23	24

Question No. 3

One of the following is an example of a pure substance.

- lemon juice and tea
- sea water
- diamond
- oil and water

11

Question No. 4

One of the following is an example of a mixture

- lemon juice and tea
- sodium
- table salt
- water

A

Question No. 5

How many kg are in 12.5 pounds? (1 kg = 2.21 lb)

- 5.65 kg
- 27.62 kg
- 2.76 kg
- 0.565 kg

A

Question No. 10

Which of the following is a physical change?

- rusting of iron
- Sublimation of dry ice
- flammability of propane gas
- baking the flour to make a cake

Save & Next حفظ والتالي

B

Question No. 5

_____ is the metric unit of the temperature of an object.

- Fahrenheit
- Celsius
- Quart
- Kelvin

Save & Next حفظ والتالي

B

Question No. 1

What is the density of a solid sample in g/mL if its volume was found to be 3.05 mL, and its mass was 7.03 g

- 2.30 g/mL
- 21.4 g/mL
- 0.043 g/mL
- 0.43 g/mL

Save & Next حفظ و التالي

A

Question No. 8

Gases and liquids share the property of _____.

- compressibility
- fixed volume
- rigid shape
- unfixed shape

Save & Next حفظ والتالي

D

Question No. 16

Which one of the following elements is a poor conductor of heat and electricity?

- copper
- fluorine
- lead
- iron

Save & Next حفظ و التالي

B

Question No. 9

Which of the following is a chemical property?

- Density
- Viscosity
- Melting point
- Acidity

Save & Next حفظ و التالي

2

Question No. 6

931 g is the same mass as _____

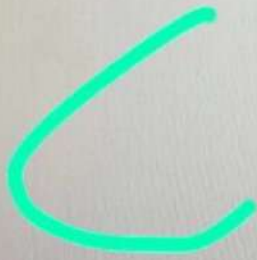
- 93.1 cg
- 931 kg
- 9310 mg
- 931 μ g

✓

Question No. 4

Which of the following is an example of a homogeneous mixture?

- sand mixed with oil
- salad dressing
- salt dissolved in water
- gasoline mixed with water



Question No. 24

Which of the following correctly matches the ion with the total number of electrons in the ion?

- Br⁻, 34 electrons
- Mg²⁺, 14 electrons
- P³⁻, 15 electrons
- Zn²⁺, 28 electrons

Save & Next حفظ و التالي

D

Question No. 7

Which of the following items is an example of matter?

- light
- heat
- time
- air



Question No. 18

When an atom gains an electron, the

- a cation.
- an anion.
- an isotope.
- a proton.

B

Question No. 5

For which of the following can the composition vary?

- both homogeneous and heterogeneous mixtures
- heterogeneous mixture
- homogeneous mixture
- pure substance

B

Question No. 14

Each electron occupies the lowest energy orbital available is known as the _____.

- Pauli exclusion principle.
- Aufbau principle.
- Hund's rule.
- Heisenberg uncertainty principle.

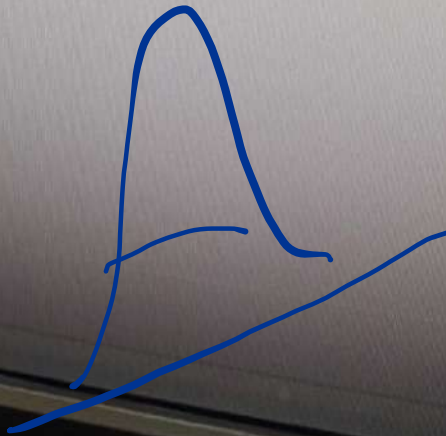
B

Question No. 13

Who formulated the modern atomic theory?

- John Dalton
- Nivaldo Tro
- Antoine Lavoisier
- Joseph Thomson

Submit & Next



Question No. 6

Which of the following measurements are NOT equivalent?

- 25 mg = 0.025 g
- 150. msec = 0.150 sec
- 24 dL = 2.4 L
- 84 cm = 8.4 mm

2

Question No. 4

How would you classify sugar?

- pure substance-compound
- mixture-heterogeneous
- mixture-homogeneous
- pure substance-element

A

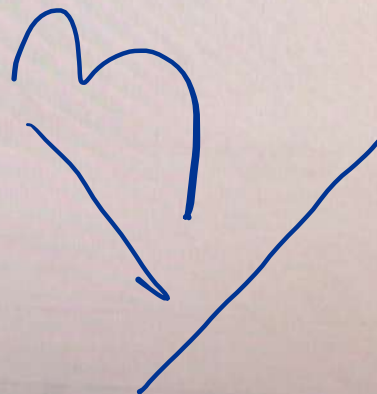
Save & Next حفظ و التالي

Question No. 10

Which of the following is a physical change?

- rusting of iron
- Sublimation of dry ice
- flammability of propane gas
- baking the flour to make a cake

3



Save & Next حفظ والتالي



Question No. 19

In which of the following sets do all species have the same number of electrons?

- F⁻, Ne, Mg²⁺
- Ge, Se²⁻, Br⁻
- Br, Br⁻, Br⁺
- K⁺, Rb⁺, Cs⁺

A

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Question No. 4

Which of the following is a homogeneous mixture?

- gasoline mixed with water
- sugar in tea
- salad dressing
- orange juice with pulp

B

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Question No. 3

Choose the heterogeneous mixture from the list below.

- air
- carbon (graphite)
- chlorine gas
- water in gasoline





Question No. 25

Which of the following elements would you expect to have very similar properties to those of the element Radon, Rn?

- Ra
- Al
- Ar
- Rb

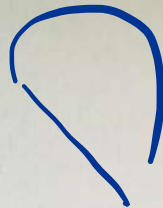
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Question No. 14

When placing the electrons in orbitals of the same energy, the electrons are first placed singly, the rule is known as:

- Lavoisier's principal
- Mendeleev's rule
- Dalton's atomic theory
- Hund's rule



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Question No. 3

Compounds are pure substances that by definition consist of _____

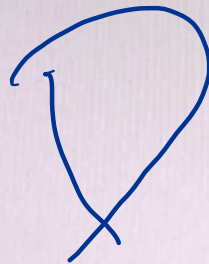
- a single element
- two or more combined elements
- two or more mixed elements
- solids



Question No. 8

Nitrogen gas has a(n) _____ volume and a(n) _____ shape.

- definite; indefinite
- definite; definite
- indefinite; definite
- indefinite; indefinite



Save & Next حفظ والتالي

Total questions in exam: 25 | Answered: 0

Question No. 2

Which of the following correctly matches the ion with the total number of electrons in the ion?

- Mg^{2+} , 14 electrons
- Br^- , 34 electrons
- Zn^{2+} , 28 electrons
- P^{3-} , 15 electrons

Question No. 17

The nucleus of an atom consists mainly of

- protons and electrons.
- protons, neutrons, and electrons.
- neutrons and electrons.
- protons and neutrons.

d

Total questions in exam: 25 | Answered: 15

Question No. 21

Iron has a density of 7.86 g/cm^3 . The volume occupied by 55.85 g of iron is

- 2.8 cm^3
- 0.141 cm^3
- 7.11 cm^3
- 439 cm^3

C

Question No. 7

Which of the following statements is incorrect?

- Evaporation of water from a fish tank is evidence of a physical change.
- The condensation of steam on a mirror is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.
- The burning of a piece of charcoal to a white powder is an example of a physical change.

d

Total questions in exam: 26 | Answered: 1

Question No. 17

The density of glycerin is 1.26 g/mL . Calculate the mass in grams of 800.0 mL glycerin.

- 1.008 Kg
- 1.01 g
- 100 g
- 1008 g

d

Question No. 9

Which of the following is NOT an example of a physical property?

- Solid ice can be very brittle
- Water can break into hydrogen and oxygen gases under electrolysis conditions
- Water can freeze solid at 32 degrees Fahrenheit
- Liquid water can turn into steam in a heated tea kettle

b

Total questions in exam: 25 | Answered: 1

Question No. 1

Which one of the following does NOT occur in a chemical reaction?

- Atoms are changed into other atoms.
- Matter is rearranged.
- Atoms react with other atoms.
- Matter is conserved.

A

Question No. 5

The maximum number of electrons that may occupy the second energy level in the atom is _____

- 18
- 10
- 2
- 8

8

Total questions in exam: 25 | Answered: 1

Question No. 10

How many milliliters are in 0.005 L? [1 L = 1000 mL]

- 0.5 mL
- 0.50 mL
- 5 mL
- 0.000005 mL

C

Question No. 7

A pure substance is:

- composed of two or more different types of atoms or molecules that has constant composition.
- composed of two or more different types of atoms or molecules combined in variable proportions.
- composed of two or more regions with different compositions.
- composed of only one type of atoms or molecules.

Question No. 1

The elements cesium, potassium, and lithium _____.

- are isotopes of each other.
- have the same number of neutrons.
- are in the same period of elements.
- are of the same family.

d

Total questions in exam: 25 | Answered: 1

Question No. 7

Deposition is the process in which _____ changes into _____.

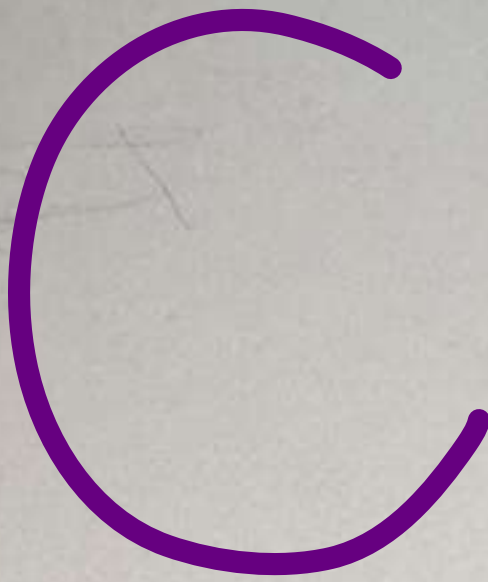
- a gas, a solid
- a solid, a gas
- a gas, a liquid
- a liquid, a solid

A

Question No. 3

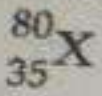
One element that has 4 valence electrons is _____

- lithium
- nitrogen
- carbon
- sulfur



Question No. 6

What does "X" represent in the following symbol?



- iron
- scandium
- chlorine
- bromine

d

Question No. 5

Which of the following items is a mixture?

- table salt
- sulfur
- baking soda
- cereal and milk

d

Question No. 11

Which of the following is a compound?

- mercury (Hg) in a thermometer
- aluminum (Al) in a can
- salt (NaCl) in a shaker
- pure gold (Au) in ring

C

Question No. 12

Which of the following elements is not a metal?

- Ba
- Xe
- Mg
- Ag

b

Total questions in exam: 25 | Answered: 1

Question No. 4

In which of the following is the metric unit paired with its correct abbreviation?

- microgram / mg
- centimeter / km
- milliliter / mL
- kilogram / cg

C

میلی لیٹر / mL

Questions in exam: 25 | Answered: 1

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Question No. 13

What is the density of a solid sample in g/mL if its volume was found to be 3.05 mL and its mass was 7.03 g

- 2.30 g/mL
- 0.43 g/mL
- 21.4 g/mL
- 0.043 g/mL

A

Question No. 8

Which of the following statements is incorrect?

- The condensation of steam on a mirror is an example of a physical change.
- Evaporation of water from a fish tank is evidence of a physical change.
- The burning of a piece of charcoal to a white powder is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.

C

Question No. 2

Atoms of the same element with different mass numbers are called _____

- ions
- chemical families
- neutrons
- isotopes

d

Question No. 9

Which of the following is NOT an example of a physical property?

- Solid ice can be very brittle
- Water can break into hydrogen and oxygen gases under electrolysis conditions
- Water can freeze solid at 32 degrees Fahrenheit
- Liquid water can turn into steam in a heated tea kettle



Total questions in exam: 25 | Answered: 1

Question No. 14

Which one of the following does NOT occur in a chemical reaction?

- Matter is rearranged.
- Atoms are changed into other atoms.
- Matter is conserved.
- Atoms react with other atoms.

b

Question No. 16

A chemical change _____

- occurs when methane gas is flamed
- occurs when water is vaporized
- occurs when paper is shredded (cut)
- occurs when sugar is dissolved in water

A

Question No. 15

Semiconductors are located in the periodic table on (or in) the _____.

- left side.
- line dividing metals and nonmetals.
- right side.
- first period.

b

Question No. 18

What is the charge on the ion formed by aluminum?

- 5-
- 13+
- 3-
- 3+

d

Question No. 20

The number of protons in the nucleus of an atom

- is called the atomic number.
- is the same for all isotopes of an element.
- identifies the atom as a particular element.
- All the answers are correct

d

QUESTION No. 20

The energy stored in the chemical bonds of a molecule is

- kinetic energy
- work
- specific heat
- potential energy

d



Total questions in exam: 25 | Answered: 8

Question No. 12

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The space occupied by the electrons is almost empty.
- The volume occupied by electrons is referred to as an electron cloud.

b

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Question No. 16

Which of the following elements is a liquid at room temperature?

- helium
- sodium
- chlorine
- mercury

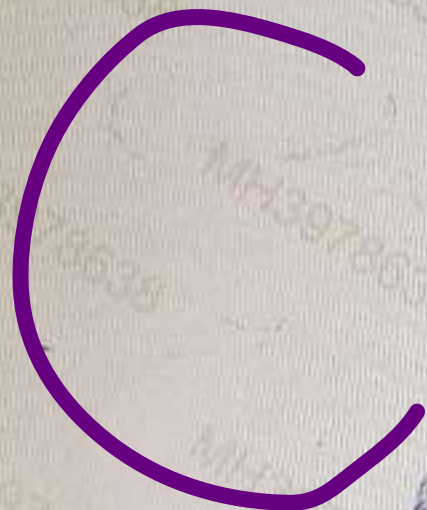
d

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Question No. 4

Bromine is the only nonmetal that is a liquid at room temperature. Consider the isotope Br-81. Select the combination which lists the correct atomic number, neutron number, and mass number, respectively.

- 35, 81, 46
- 81, 46, 35
- 35, 46, 81
- 46, 81, 35



User: MH397

Number of m
Number of qu

2 Answered

20 Not Visited

1 2

8 9

15 16

22 23

Atoms are neutral because _____

- neutrons neutralize the (+) and (-) charges in the atom
- equal numbers of protons and neutrons are present
- all atoms contain neutrons
- the atomic number equals the number of electrons

d

Question No. 9

Which one of the following is a homogeneous mixture?

- Alloys
- Vegetable salad
- water
- Sand in water

A

Question No. 17

The nucleus of an atom consists mainly of

- protons and electrons.
- protons, neutrons, and electrons.
- neutrons and electrons.
- protons and neutrons.

d

Total questions in exam: 25 | Answered: 8

Question No. 11

One decimeter is equal to _____

- 1 m
- 10 cm
- 1000 mm
- 10 m

A

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Total questions in exam: 25 | Answered: 8

Question No. 8

Which of the following is NOT a correct (name; symbol) combination?

- beryllium, Be
- calcium, Ca
- magnesium, Mg
- manganese, Mn

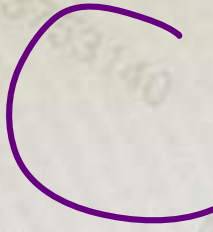
D

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Question No. 10

According to Hund's Rule _____.

- electrons pair first before filling other orbitals with same energy
- no two electrons in the same atom can have the same four quantum numbers.
- electrons fill orbitals of the same energy singly and pair up only after the orbitals are all half filled.
- electrons found in the same orbital must have the same spin.



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Question No. 3

Scandium belongs to _____

- transition metals
- alkaline earth metals
- noble gases
- halogens

A

Question No. 4

Which one of these represents a physical change?

- water, when heated, forms water vapor
- sugar, when heated, becomes brown
- bleach turns hair yellow
- milk turns sour

A

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Total questions in exam: 25 | Answered: 8

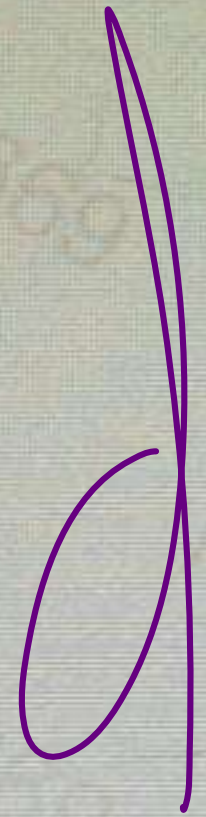
Question No. 12

Consider the atoms of Co-59 and Co-60, both of these atoms have the same _____

- number of protons
- atomic mass number
- number of neutrons
- none of these is true

What is the atomic symbol for manganese?

- Mo
- Mg
- Md
- Mn



Question No. 7

Which of the following statements is incorrect?

- Evaporation of water from a fish tank is evidence of a physical change.
- The condensation of steam on a mirror is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.
- The burning of a piece of charcoal to a white powder is an example of a physical change.

D

Question No. 3

Which one of the following does NOT occur in a chemical reaction?

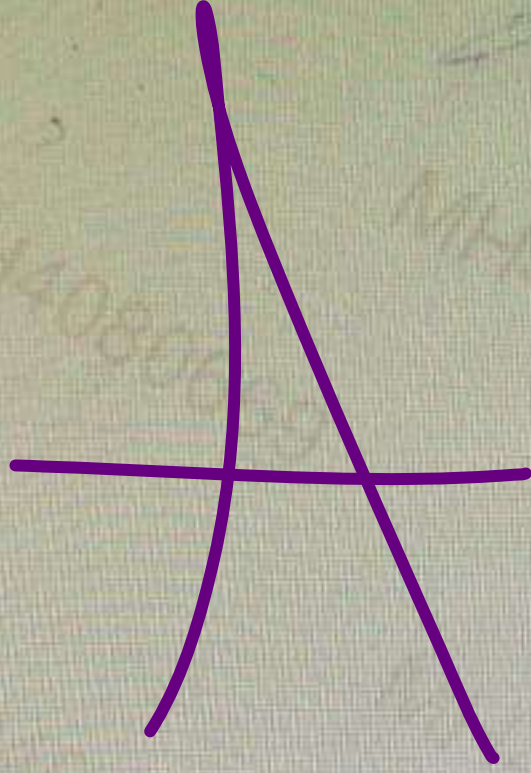
- Matter is conserved
- Matter is rearranged
- Atoms are changed into other atoms.
- Atoms react with other atoms.

C

Question No. 6

Condensation refers to which conversion?

- gas \rightarrow liquid
- solid \rightarrow gas
- liquid \rightarrow gas
- solid \rightarrow liquid



Question No. 8

How many protons and electrons are present in O^{2-} ?

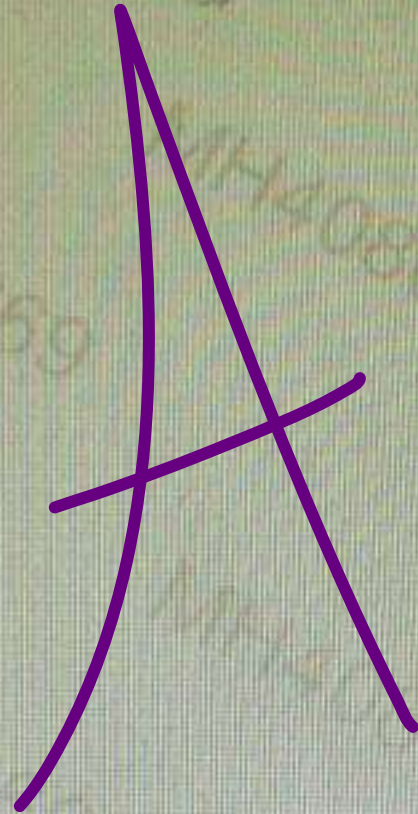
- 8 protons and 10 electrons
- 8 protons and 8 electrons
- 16 protons and 8 electrons
- 10 protons and 8 electrons



Question No. 3

Which of the following is an example of chemical property?

- flammability
- viscosity
- density
- color



Question No. 2

Which of the following gives the correct numbers of protons, neutrons, and electrons in a neutral atom of $^{118}_{50}\text{Sn}$?

- 50 protons, 68 neutrons, 50 electrons
- 118 protons, 50 neutrons, 118 electrons
- 68 protons, 68 neutrons, 50 electrons
- 118 protons, 118 neutrons, 50 electrons

A

Question No. 1

Which of the following is an example of matter?

- ice cream
- time
- sound
- sunlight

A

Question No. 3

Which of the following is an example of chemical property?

- flammability
- viscosity
- density
- color

A

Question No. 2

Which of the following gives the correct numbers of protons, neutrons, and electrons in a neutral atom of $^{118}_{50}\text{Sn}$?

- 50 protons, 68 neutrons, 50 electrons
- 118 protons, 50 neutrons, 118 electrons
- 68 protons, 68 neutrons, 50 electrons
- 118 protons, 118 neutrons, 50 electrons

A

Question No. 9

Which one of the following is a homogeneous mixture?

- Alloys
- Vegetable salad
- water
- Sand in water

A

Question No. 7

Which of the following statements is incorrect?

- Evaporation of water from a fish tank is evidence of a physical change.
- The condensation of steam on a mirror is an example of a physical change.
- The fact that sulfur powder is yellow is a physical property.
- The burning of a piece of charcoal to a white powder is an example of a physical change.

d

Question No. 4

What is the atomic symbol for manganese?

- Mo
- Mg
- Md
- Mn



Question No. 4

Bromine is the only nonmetal that is a liquid at room temperature. Consider the isotope Br-81. Select the combination which lists the correct atomic number, neutron number, and mass number, respectively.

- 35, 81, 46
- 81, 46, 35
- 35, 46, 81
- 46, 81, 35



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2 Answered

20 Not Visited

1

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8

9

15

16

22

23

Total questions in exam: 25 | Answered: 8

Question No. 12

Consider the atoms of Co-59 and Co-60, both of these atoms have the same _____

- number of protons
- atomic mass number
- number of neutrons
- non of these is true

Question No. 4

Which one of these represents a physical change?

- water, when heated, forms water vapor
- sugar, when heated, becomes brown
- bleach turns hair yellow
- milk turns sour

A

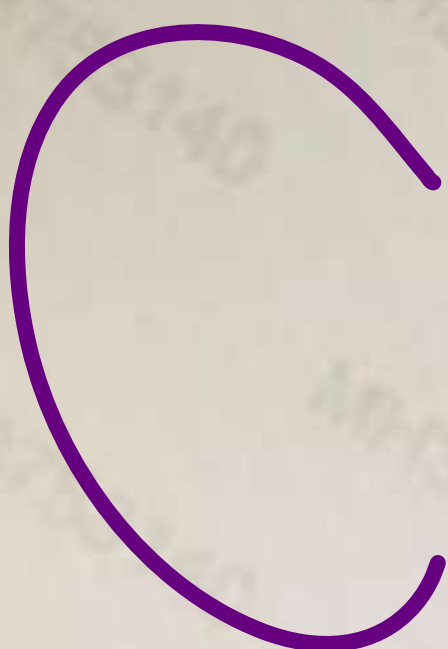
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Question No. 10

According to Hund's Rule _____

- electrons pair first before filling other orbitals with same energy
- no two electrons in the same atom can have the same four quantum numbers.
- electrons fill orbitals of the same energy singly and pair up only after the orbitals are all half filled.
- electrons found in the same orbital must have the same spin.



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Total questions in exam: 25 | Answered: 8

Question No. 12

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The space occupied by the electrons is almost empty.
- The volume occupied by electrons is referred to as an electron cloud.

b

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Question No. 7

A pure substance is:

- composed of two or more different types of atoms or molecules that has constant
- composed of two or more different types of atoms or molecules combined in vari
- composed of two or more regions with different compositions.
- composed of only one type of atoms or molecules.

Question No. 14

Express 7.5 nm as picometers.

- 750 pm
- 7.5×10^6 pm
- 75.0 pm
- 7.5×10^3 pm



Question No. 21

Melting is a process in which _____ changes into _____, and it requires _____ temperature.

- a solid, a liquid, increasing
- a liquid, a gas, decreasing
- a liquid, a solid, increasing
- a solid, a gas, increasing

A

Total questions in exam: 25 | Answered: 8

Question No. 12

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The space occupied by the electrons is almost empty.
- The volume occupied by electrons is referred to as an electron cloud.

b



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Total questions in exam: 25 | Answered: 1

Question No. 16

A chemical change _____

- occurs when methane gas is flamed
- occurs when water is vaporized
- occurs when paper is shredded (cut)
- occurs when sugar is dissolved in water

A

Question No. 17

All of the following statements about different elements are true EXCEPT:

- A Barium is an alkaline earth metal.
- B Krypton is one of the noble gases.
- C Manganese is a transition metal.
- D Lithium is considered a metalloid.

D

Question No. 25

As you move up and to the right on the periodic table

- atomic radius increases and ionization energy increases
- atomic radius decreases and ionization energy increases
- atomic radius decreases and ionization energy decreases
- atomic radius increases and ionization energy decreases

C

Question No. 18

The Group 8A (18) elements

- are unreactive.
- are liquids at room temperature.
- are good conductors of electricity
- melt at high temperatures.

A

Question No. 3

Which of the following is an example of a heterogeneous mixture?

- milk
- oil
- soda
- chicken soup

D

Question No. 5

How many inches are in 25.8 cm? [1 in = 2.54 cm]

- 0.1
- 10.2
- 28.3
- 0.0984

B

Question No. 12

What is the atomic symbol for silicon?

- S
- Si
- Ag
- Au

B

Question No. 1

A solution has a density of 1.22 g/mL. What volume of the solution has a mass of 55.0 g?

- 0.00253 mL
- 49.4 mL
- 58.8 mL
- 45.1 mL

D

Question No. 11

Which statement below accurately describes the contributions of Thomson?

- Ancient Greek philosopher who proposed that matter was continuous
- Proposed the modern Atomic Theory
- Discovered the existence of electrons
- Created the modern periodic table

C

Question No. 7

The state of matter that has fixed volume and unfixed shape is called?

- liquid.
- plasma
- solid
- gas

A

Question No. 4

Which of the following items is a mixture?

- table salt
- cereal and milk
- baking soda
- sulfur

B

Save & Next 

Question No. 10

What is the value of 335 K on the Celsius temperature scale?

- 62
- 66.4
- 167
- 608

A

Question No. 19

What is the charge on the ion formed by aluminum?

- 5-
- 3+
- 3-
- 13+

B

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Question No. 9

What is the atomic symbol for gold?

- Ar
- As
- Au
- Ag

C

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Question No. 4

9.31 g is the same mass as _____

- 931 μg
- 9310 mg
- 93.1 cg
- 931 kg

B

Question No. 13

The elements rubidium, sodium, and potassium _____

- are isotopes of each other
- are in the same group
- have the same number of neutrons
- are in the same period of elements

B

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Question No. 11

One element that has 5 valence electrons is _____

- lithium
- sulfur
- nitrogen
- carbon

C

Question No. 6

Which state of matter has atomic spacing that is close together and fixed shape?

- liquid
- plasma
- gas
- solid

D

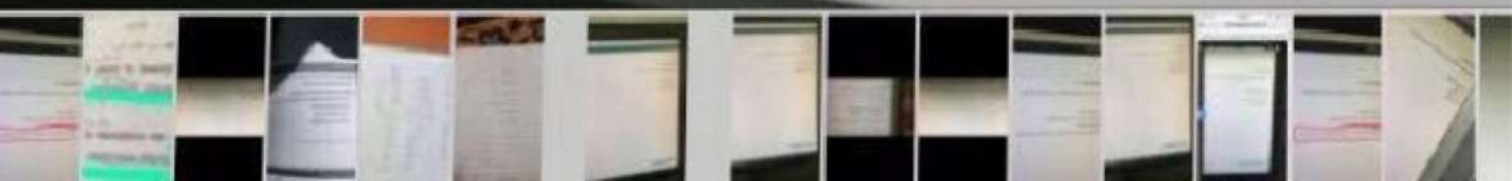
Question No. 9

Condensation refers to which conversion?

- solid -> liquid
- liquid -> gas
- gas -> liquid
- solid -> gas

C

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Question No. 14

Which one of these represents a physical change?

- bleach turns hair yellow
- sugar, when heated, becomes brown
- milk turns sour
- water, when heated, forms water vapor



D

Question No. 25

Chlorine has more metallic character than _____

- iodine
- bromine
- fluorine
- astatine

Question No. 4

Choose the heterogeneous mixture from the list below.

- carbon (graphite)
- chlorine gas
- chicken noodle soup
- black coffee

C

Question No. 9

Which of the following is NOT an example of matter?

- a dust particle
- a pencil eraser
- a balloon full of helium
- heat from a burning candle

D

If matter is uniform throughout and cannot be separated into other substances by physical

- either an element or a compound
- a homogeneous mixture
- a heterogeneous mixture
- a compound

A

Question No. 7
Which type of energy is associated with motion of objects?

- electrical energy
- kinetic energy
- potential energy
- chemical energy

B

Question No. 7

What is the prefix multiplier used to represent the factor 10^{-6} .

- micro
- mega
- kilo
- nano

A

Question No. 4

Which of the following is an element?

- salt
- gold
- water
- sugar

B

Question No. 12

Which one of these represents a chemical change?

- melting butter
- boiling water to form water vapor
- grilling a steak
- mixing carbon and oxygen (at room temperature)

C

Question No. 16

The elements xenon, helium and argon

- are isotopes of each other.
- are in the same period of elements.
- have the same number of neutrons.
- are in the same group.

D



Question No. 5

Salt water is classified as

- pure substance-element
- mixture-homogeneous
- pure substance-compound
- mixture-heterogeneous

B



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Question No. 12

The charge of electron was established by

- Millikan's oil drop experiment.
- Rutherford's gold foil experiment.
- Dalton's atomic experiment.
- Thomson's cathode ray tube experiment.



A



Question No. 23

Which of the following elements is a metal?

- nitrogen
- fluorine
- strontium
- argon

C





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Question No. 7

Which of the following is a property of both solids and liquids?

- definite shape
- indefinite shape
- indefinite volume
- definite volume

D



Question No. 6

A decimeter equals:

- 1 m
- 1000 mm
- 10 m
- 10 cm



C

Question No. 9

Which of the following is a chemical change?

- burning natural gas
- melting ice
- cutting paper
- cutting a tomato

A



ANSWER KEY



Question No. 5

The correct prefix for the multiplier 1,000,000,000 is

- milli.
- giga.
- mega.
- tera.

B



Question No. 8

Which of the following is matter?

- time
- ice cream
- sunlight
- sound



B



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Question No. 4

Choose the heterogeneous mixture from the list below

- carbon (graphite)
- chlorine gas
- chicken noodle soup
- black coffee



C



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Question No. 24

Which of the following has the greatest number of nonbonding pairs of electrons?

- He
- F
- H
- C

3



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Question No. 23

Where is most of the mass of an atom concentrated?

- nucleus
- neutrons
- electrons
- protons

A



Question No. 13

The elements rubidium, sodium, and potassium _____

- are isotopes of each other
- are in the same group
- have the same number of neutrons
- are in the same period of elements

B

Question No. 4

9.31 g is the same mass as _____

- 931 μg
- 9310 mg
- 93.1 cg
- 931 kg

B

Question No. 10

An s orbital can hold _____ electrons.

- 6
- 2
- 18
- 10

B

Question No. 6

Which state of matter has atomic spacing that is close together and fixed shape?

- liquid
- plasma
- gas
- solid

D

Question No. 5

A solid form of matter in which there is long range repeating order is called _____.

- amorphous
- fixed
- crystalline
- rigid

C

Question No. 7

Which of the following is a chemical change?

- burning natural gas
- cutting paper
- cutting a tomato
- melting ice

A

Question No. 12

Francium is an example of an element that is a _____

- metalloïd
- noble gas
- nonmetal
- metal

D



INSTRUCTION: Please choose the BEST answer from the given options for each question.

Question:

Which of the following is a heterogeneous mixture?

Options:

- pure water
- tap water
- gasoline
- ice and water

D





INSTRUCTION → Please choose the BEST answer from the given options for each

Question:

Which of the following is a physical process?

Options:

- the explosion of nitroglycerin
- the baking of a potato
- the rusting of iron
- the condensation of water vapor

D





Today

السلام عليكم ...

11:09 PM ✓

دكتوراه لو سمحتي ايش اجابة هذا السؤال ؟

11:10 PM ✓

Gold in a ring is a
-compound
-element
-heterogeneous
-homogeneous

11:10 PM ✓

Homogenous

11:11 PM

لأنه سبيكة alloy

11:12 PM

لو قالي فضه كمان نفس الشئ ؟

11:13 PM ✓

لا اعتقد

11:15 PM

يعني لو فضه احط عنصر صح 🙄

11:16 PM ✓

ايوة

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Question No. 18

Non Metals are located where on the periodic table?

- left side
- between metals and nonmetals
- middle
- right side



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Question No: 7

Which of the following is a chemical change?


- burning natural gas
- cutting paper
- cutting a tomato
- melting ice

A



The final step in forming the modern atomic theory was done by:

A. Rutherford

 B. Dalton

C. Proust

D. Lavoisier

B



Question No. 7

The state of matter that has fixed volume and unfixed shape is called?

- liquid
- plasma
- solid
- gas

A





Question No. 1

A solution has a density of 1.22 g/mL. What volume of the solution has a mass of 55.0 g?

- 0.00253 mL
- 49.4 mL
- 58.8 mL
- 45.1 mL



Question No. 10

What is the value of 335 K on the Celsius temperature scale?

- 62
- 66.4
- 167
- 608

A



Question No. 10

Condensation refers to which conversion?

- solid -> gas
- solid -> liquid
- liquid -> gas
- gas -> liquid



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Question No. 4

Which of the following is a compound?

- mercury (Hg) in a thermometer
- aluminum (Al) in a can
- pure gold (Au) in ring
- salt (NaCl) in a shaker

Question No. 8

The correct prefix for the multiplier 0.000001 is _____

- mega
- milli
- nano
- micro

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Question No. 3

Which of the following items is a pure substance?

- air
- brass (zinc and copper)
- ice
- sea water

Question No. 2

What is the mass of 2.00 L of a solution with a density of 1.15 g/mL?

- 1.15 kg
- 0.015 kg
- 2.30 kg



D

Question No. 1

What is the density of a solid sample in g/mL if its volume was found to be 3.05 mL and its mass was 7.03 g

- 2.30 g/mL
- 0.43 g/mL
- 0.043 g/mL
- 21.4 g/mL

Question No. 6

_____ is the metric unit of the temperature of an object.

- Celsius
- Fahrenheit
- Quart
- Kelvin

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Question No. 7

Convert 4 micrometers to meters. [1 m = 1000,000 μm]

- 4×10^{-6} m
- 4×10^{-9} m
- 4×10^6 m
- 4×10^{-3} m

Question No. 8

Which of the following quantities represents the largest mass?

- 0.0010 kg
- 2.0×10^2 mg
- 2.0×10^2 g
- 1.0×10^5 μ g

Question No. 4

One of the following is an example of a pure substance:

- sea water
- oil and water
- diamond
- lemon juice and tea

Question No. 5

Which of the following is an example of a homogeneous mixture?

- salt dissolved in water
- sand mixed with oil
- gasoline mixed with water
- salad dressing

Question No. 3

Choose the pure substance from the list below.

- sugar
- milk
- lemonade
- air

Question No. 9

Identify the crystalline solid from the following.

- plastic
- water
- table salt
- glass

Question No. 1

What is the mass of 36.0 mL of a liquid if its density is 2.70 g/mL?

0.0750 g

13.3 g

0.748 g

97.2 g

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Question No. 10

Sublimation refers to which conversion?

liquid -> gas

gas -> liquid

solid -> gas

solid -> liquid

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Question No. 11

Which one of these represents a chemical change?

- mixing carbon and oxygen (at room temperature)
- melting butter
- boiling water to form water vapor
- grilling a steak

Question No. 5

How would you classify sugar?

- mixture-heterogeneous
- mixture-homogeneous
- pure substance-element
- pure substance-compound

Question No. 25

The smallest subatomic particles is _____.

- electron
- proton
- nucleus
- neutron

Question No. 22

What is the charge on the ion formed by sodium?

2-

2+

1+

1-

Question No. 23

A cation of +1 indicates that an element has _____

- lost one proton.
- gained one proton.
- gained one electron.
- lost one electron.

Question No. 24

A neutron has an electrical charge of _____.

-1

-2

+1

0

Question No. 21

The elements lithium, beryllium, and boron _____

- are in the same group
- are in the same period
- have the same number of neutrons
- are isotopes of each other

Question No. 20

The elements beryllium, carbon, lithium _____

- have the same mass number.
- are in the same group.
- have the same number of neutrons.
- are in the same period of elements.

Question No. 19

Valence electrons are electrons located _____

- in the nucleus of an atom.
- in the first three shells of an atom.
- in the innermost energy level of an atom.
- in the outermost energy level of an atom.

Question No. 16

Who formulated the modern atomic theory?

- Antoine Lavoisier*
- Joseph Thomson*
- Nivaldo Tro*
- John Dalton*

Question No. 17

What element has the electron configuration, $1s^2 2s^2 2p^5$?

- Na
- S
- F
- Mg

Chem

Question No. 18

The "p" orbital can hold up to _____ electrons.

- 8
- 2
- 18
- 6

Question No. 1

Determine the volume of an object that has a mass of 455.6 g and a density of 19.3 g/cm³.

- 42.4 mL
- 18.5 mL
- 87.9 mL
- 23.6 mL

Question No. 8

In which of the following is the metric unit paired with its correct abbreviation?

- microgram / mg
- centimeter / km
- kilogram / cg
- milliliter / mL

Question No. 16

In the winter, if the outdoor temperature is -12°F , what is the temperature on the Kelvin scale?

- 262 K
- 273 K
- 284 K
- 249 K

Question No. 17

"Every orbital in a subshell is singly occupied with one electron before any one orbital is doubly occupied" is known as _____

- Hund's rule
- Pauli exclusion principle
- Heisenberg uncertainty principle
- Aufbau principle

Question No. 18

What element has the electron configuration $1s^2 2s^2 2p^5$?

Na

Mg

S

F

Question No. 3

Choose the pure substance from the list below.

- sugar
- air
- milk
- lemonade

Question No. 15

Which of the following is the lowest temperature?

35 °F

0 °C

-10 °C

100 K

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Question No. 16

Which of the following is NOT a correct (name, symbol) combination?

- magnesium, Mg
- beryllium, Be
- manganese, Mn
- calcium, Ca

Question No. 11

A chemical change _____.

- occurs when paper is shredded (cut)
- occurs when methane gas is flamed
- occurs when sugar is dissolved in water
- occurs when water is vaporized

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Question No. 12

All of the following can be considered physical properties EXCEPT _____.

- color
- taste
- boiling point
- flammability

Question No. 13

Which one of the following choices includes only chemical changes?

- rusting, evaporation, and deposition
- decomposition, digestion, and burning
- melting, evaporation, and sublimation
- melting, freezing, and digestion

Question No. 9

The _____ is the *building block of matter*.

- atom
- element
- compound
- molecule

Save & Next

Question No. 10

Identify a solid.

- definite volume and no definite shape
- definite volume and definite shape
- no definite shape and no definite volume
- definite shape and compressible

Question No. 6

Which of the following is a molecular element?

- barium
- argon
- xenon
- fluorine

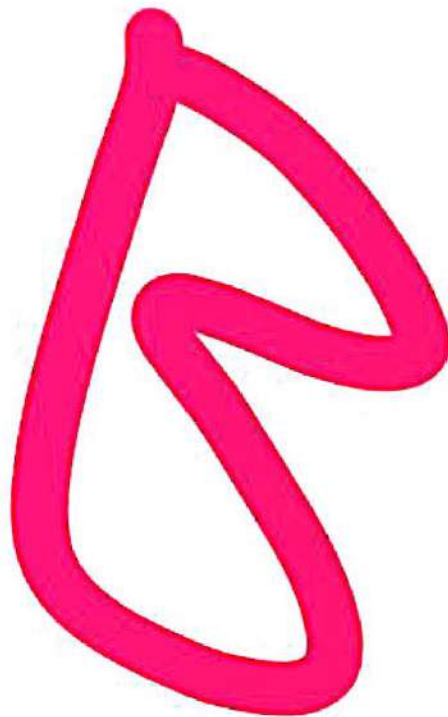
D

Question No. 4

Non Metals are located where on the periodic table?

- left side
- right side
- middle
- between metals and nonmetals

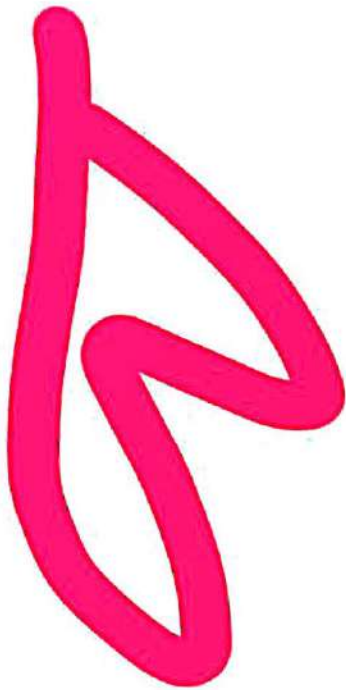
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Question No. 2

Group 7A elements are also called:

- noble gases.
- halogens.
- alkaline earth metals.
- alkali metals.



Question No. 6

Fluorine is considered which of the following?

- atomic element
- molecular element
- molecular compound
- ionic compound

B

Question No. 15

According to the Pauli exclusion principle, any orbital can hold at most _____ electrons.

- 6
- 18
- 2
- 8

C



Question No. 18

What is the correct formula for a potassium ion with 18 electrons?

- P⁻
- P⁺
- K⁻
- K⁺

D

Question No. 22

What element has the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^2$?

- carbon
- silicon
- sulfur
- oxygen

B

Question No. 1

Which of the following is NOT a correct name, symbol combination?

- magnesium, Mg
- Beryllium, Be
- calcium, Ca
- manganese, Mn

C

Question No. 19

A cation of +3 indicates that an element has

- gained three protons.
- gained three electrons.
- lost three protons.
- lost three electrons.

D

Question No. 12

Which of the following is NOT a correct name, symbol combination?

- potassium, P
- manganese, Mn
- magnesium, Mg
- iron Fe

A

Question No. 4

Semiconductors are usually

- noble gases.
- nonmetals.
- metals.
- metalloids.

D

Question No. 4

Metalloids are located where on the periodic table?

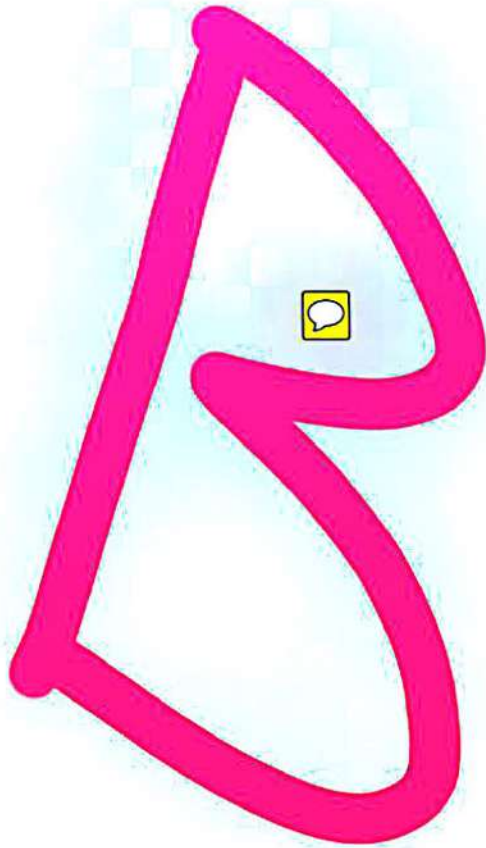
- middle
- between metals and nonmetals
- right side
- left side

B

Question No. 4

What is the charge on the oxygen ion?

- 1+
- 2-
- 2+
- 1-



Question No. 4

When an atom gains an electron, the resulting particle is called.

- a cation.
- an isotope.
- a proton.
- an anion.



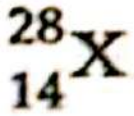
Question No. 2

The elements beryllium, carbon, lithium

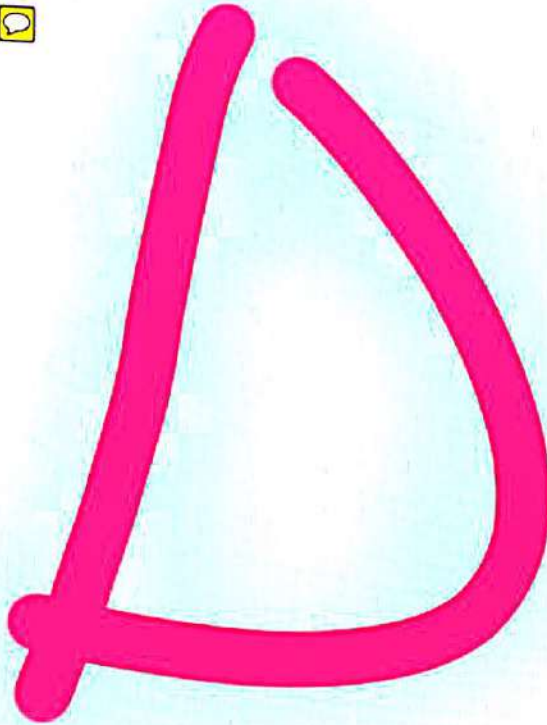
- are in the same group.
- are in the same period of elements.
- have the same mass number.
- have the same number of neutrons.

B

What does "X" represent in the following symbol?



- zinc
- sulfur
- nickel
- silicon



Question No. 3

The elements lithium, beryllium, and boron _____

- have the same number of neutrons
- are in the same period
- are isotopes of each other
- are in the same group



B

Question No. 2

Which describes the valance electrons of calcium?



- $4s^2$
- $4s^2d^1$
- $4s^0$
- $4s^1$

A

Question No. 6

Which of the following is a molecular element?

- calcium
- argon
- oxygen
- sodium



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Question No. 1

Each electron occupies the lowest energy orbital available is known as the

- Heisenberg uncertainty principle.
- Aufbau principle.
- Pauli exclusion principle.
- Hund's rule.



Nickel is an example of an element that is a _____.

- metalloid
- noble gas
- transition element
- nonmetal

C

Question No. 2

The maximum number of electrons that may occupy the fourth electron energy level is _____

- 32
- 2
- 8
- 18

A 32

Question No. 1

Who formulated the modern atomic theory?

- Nivaldo Tro
- John Dalton
- Antoine Lavoisier
- John Dalton and Antoine Lavoisier



The number of electrons in the outer energy level of a neutral atom of beryllium is

- 8
- 2
- 5
- 3

B

Question No. 6

Which of the following is a molecular element?

- calcium
- argon
- oxygen
- sodium

C

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Question No. 1

What is the element with the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^5$?

- F
- Be
- S
- Cl

D

The elements cesium, potassium, and lithium

- have the same number of neutrons.
- are in the same group.
- are isotopes of each other.
- are in the same period of elements.

3

Question No. 4

Which of these elements is most likely to be a good conductor of electricity?

- S
- Fe
- N
- He



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Question No. 4

Where is most of the mass of an atom concentrated?

- nucleus
- neutrons
- electrons
- protons

A



Question No. 1

The maximum number of electrons that can go in the main energy level 3 is?

- 18
- 32
- 8
- 10

A

3s² 3p²

INSTRUCTION: Please choose the BEST answer from the given options for each question.

Question:

All of the following statements about different elements are true EXCEPT:

Options:

- Manganese is a transition metal.
- Barium is an alkaline earth metal.
- Sulfur is considered a metalloid.
- Krypton is one of the noble gases.

C

Which of the following statements about ions is INCORRECT?

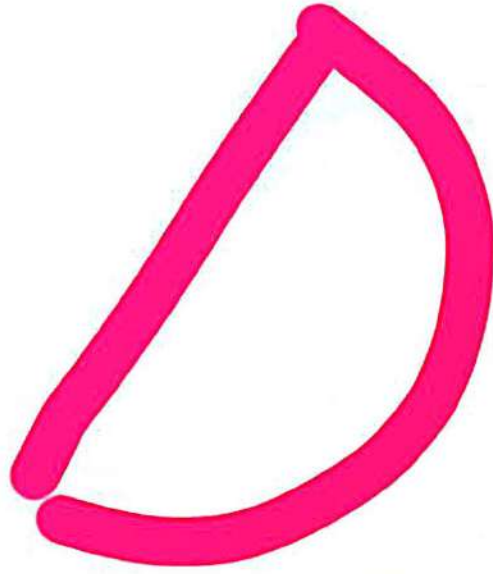
- Cations always have the same number of protons as electrons.
- Cations are formed when an atom loses electrons.
- Cations are positive ions and anions are negative ions.
- Anions are formed when an atom gains electrons.

A

Question No. 1

The number of electrons in the outer energy level of a neutral atom of phosphorous is

- 2
- 8
- 3
- 5



Question No. 1

What is the element with the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^4$?

- F
- S
- Be
- Mg

B

Question No. 1

Neon has _____ valence electrons.

- 8
- 6
- 2
- 4

A

Question No. 5

Which of the following is an example of chemical property

- density.
- color.
- flammability.
- taste.

C

Question:

Liquids are characterized as having _____ volume which means they are _____.

Options:

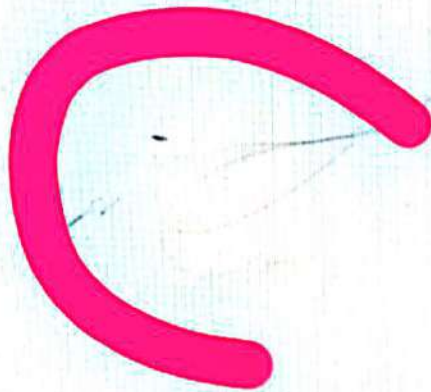
- indefinite, incompressible
- indefinite, compressible
- definite, incompressible
- definite, compressible



Question No. 7

Which statement below accurately describes the contributions of Thomson?

- Created the modern periodic table
- Proposed the modern Atomic Theory
- Discovered the existence of electrons
- Ancient Greek philosopher who proposed that matter was continuous



Question No. 2

Which valence orbital-filling diagram represents the ground state of oxygen?

- $\frac{\uparrow\downarrow}{2s} \quad \frac{\uparrow\downarrow}{2p} \quad \frac{\uparrow}{2p} \quad \frac{\uparrow}{2p}$
- $\frac{\uparrow\uparrow}{2s} \quad \frac{\uparrow\downarrow}{2p} \quad \frac{\uparrow}{2p} \quad \frac{\uparrow}{2p}$
- $\frac{\uparrow\downarrow}{2s} \quad \frac{\uparrow\downarrow}{2p} \quad \frac{\uparrow\downarrow}{2p} \quad \text{---}$
- $\frac{\uparrow}{2s} \quad \frac{\uparrow\downarrow}{2p} \quad \frac{\uparrow\downarrow}{2p} \quad \frac{\uparrow}{2p}$

A

Question No. 6

Which state of matter has atomic spacing that is close together and fixed shape?

- liquid
- plasma
- gas
- solid

D

Question No. 13

The elements rubidium, sodium, and potassium _____

- are isotopes of each other
- are in the same group
- have the same number of neutrons
- are in the same period of elements

B

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Question No. 15

Which of the following elements is a metalloid?

- Mg
- Al
- N
- Si

A

Question No. 12

Francium is an example of an element that is a _____

- metalloid
- noble gas
- nonmetal
- metal

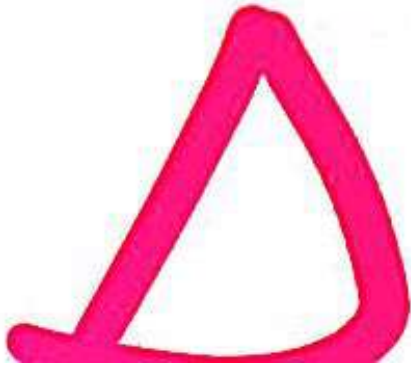


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Question No. 4

What is the charge on the ion formed by *magnesium*?

- 1-
- 2-
- 1+
- 2+



Question No. 3

Zinc is a member of which family?

- noble gases
- halogens
- alkaline earth metals
- Transitional Metals



Question No. 1

What is the atomic symbol for silver?

- S
- Ag
- Au
- Si

B

Question No. 3

All of the following statements about different elements are true **EXCEPT**:

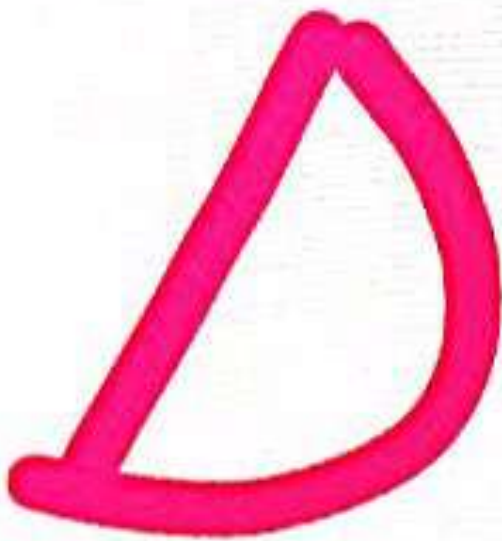
- Barium is an alkaline earth metal.
- Manganese is a transition metal.
- Lithium is considered a metalloid.
- Krypton is one of the noble gases.

C

Question No. 1

Which statement below accurately describes the contributions of Josef Proust?

- ancient Greek philosopher who proposed that matter was continuous
- created the modern periodic table
- proposed the modern Atomic Theory
- discovered the law of definite proportions



INSTRUCTION: Please choose the BEST answer from the given options for each question.

Question:

Which of the following would not be considered matter?

Options:

- light
- cheese
- blood
- plants

A

Question:

The number of neutrons in an atom is equal to:

Options:

- atomic number - mass number
- mass number - atomic number
- the number of protons
- the number of electrons



INSTRUCTION: تعليمات Please choose the BEST answer from

Question:

Which of the following is a compound?

Options:

- copper
- magnesium
- iron
- carbon dioxide

D

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each

Question:

Isotopes of a given element _____

Options:

- have the same mass number but different chemical properties
- have the same atomic number but different chemical properties
- have the same number of protons, but different mass numbers
- have the same mass number but different numbers of protons

C

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each question.

Question:

Which experiment led to the notion that the atom contains an extremely small, positively

Options:

- Dalton's atomic experiment
- Rutherford's gold foil experiment
- Millikan's oil drop experiment
- Thomson's cathode ray experiment

2

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each qu

Question:

Which one of the following carries no electrical charge?

Options:

- a proton
- a neutron
- a cation
- an electron

ب

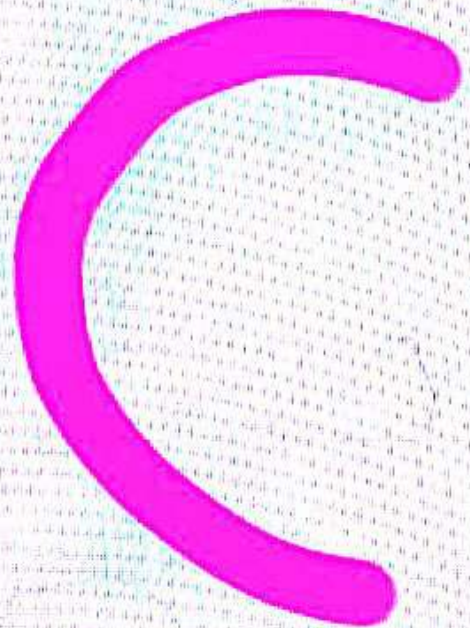
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Question No. 3

How many protons and neutrons are in Cl-37?

- 17 protons, 37 neutrons
- 20 protons, 17 neutrons
- 17 protons, 20 neutrons
- 37 protons, 17 neutrons



Question No. 17

How many valence electrons are in an oxygen atom and an oxide ion?

- 2 and 8, respectively
- 8 and 10, respectively
- 6 and 8, respectively
- 8 and 6, respectively



INSTRUCTION: تعليمات Please choose the BEST answer from the given options for ea

Question:

Energy that is associated with the motion of an object is called

Options:

- kinetic energy
- chemical energy
- potential energy
- thermal energy

A

I

Question No. 3

Which of the following elements belongs to the halogens group?

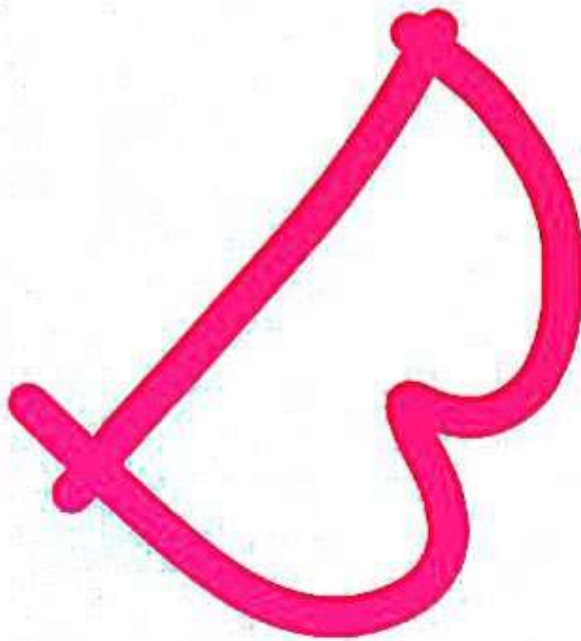
- Cl
- Na
- Fe
- Co

A

Question No. 23

Which of the following elements is a metal?

- fluorine
- strontium
- nitrogen
- argon



Question No. 12

The names of the elements whose symbols are Si, P, Mn, and S are respectively.

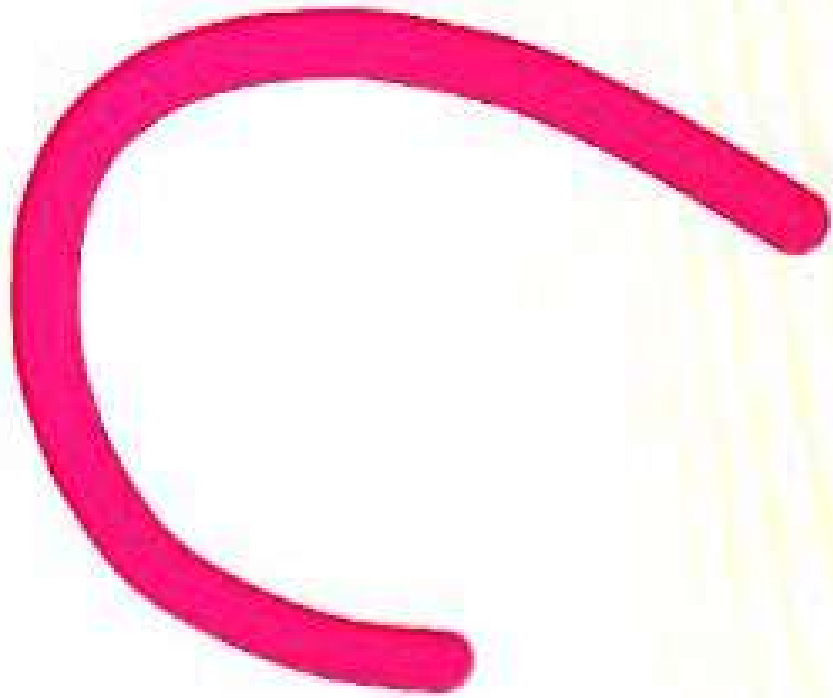
- silver, phosphorus, magnesium, and sulfur
- silicon, phosphorus, magnesium, and sulfur
- silicon, potassium, magnesium, and sulfur
- silicon, phosphorus, manganese, and sulfur



Question No. 9

What is the atomic symbol for gold?

- Ar
- AS
- Au
- Ag

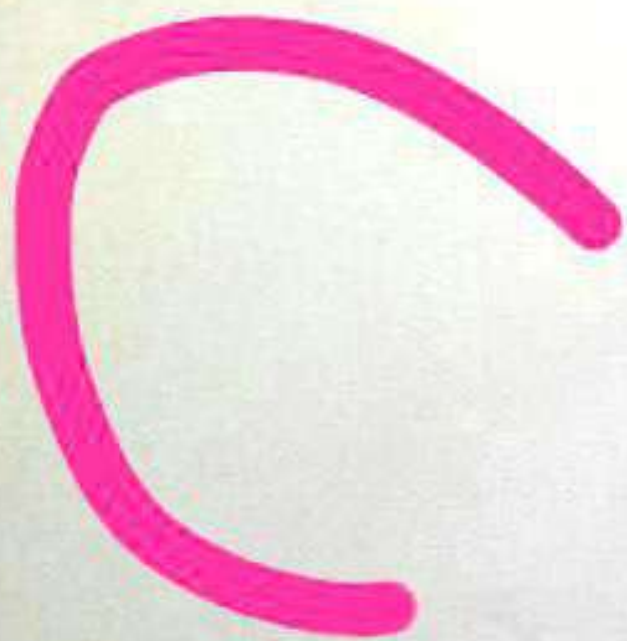




Question No. 3

Which of the following is not an element?

- gold
- carbon
- water
- silver



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Question No. 18

The Group 8A (18) elements

- are unreactive.
- are liquids at room temperature.
- are good conductors of electricity.
- melt at high temperatures.

A

Question:

What is the charge on the cesium ion?

Options:

- 1+
- 2+
- 1-
- 2-

A

INSTRUCTION: **تعليمات** Please choose the BEST answer from the given options for each

Question:

What is the correct electron configuration for the beryllium atom?

Options:

$1s^1 2s^2$

$1s^2$

$1s^2 2s^2$

$1s^3$



INSTRUCTION: **تعليمات** Please choose the BEST answer from

Question:

What is the electron configuration for potassium?

Options:

- $1s^2 2s^2 2p^6 3s^2 3p^5 3d^2$
- $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
- $1s^2 2s^2 2p^8 3s^2 3p^5$
- $1s^2 2s^2 2p^6 3s^2 3p^7$

B

Question:

One element that has 5 valence electrons is _____.

Options:

- nitrogen
- sulfur
- carbon
- lithium

A

Question No. 5

How many inches are in 25.8 cm? [1 in = 2.54 cm]

- 0.1
- 10.2
- 28.3
- 0.0984



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Question No. 7

The state of matter that has fixed volume and unfixed shape is called?

- liquid
- plasma
- solid
- gas

A

Save & Next

Question No. 4

Which of the following items is a mixture?

- table salt
- cereal and milk
- baking soda
- sulfur

B

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Question No. 5

A solid form of matter in which there is long range repeating order is called _____.

- amorphous
- fixed
- crystalline
- rigid



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Question No. 7

Which of the following is a chemical change?

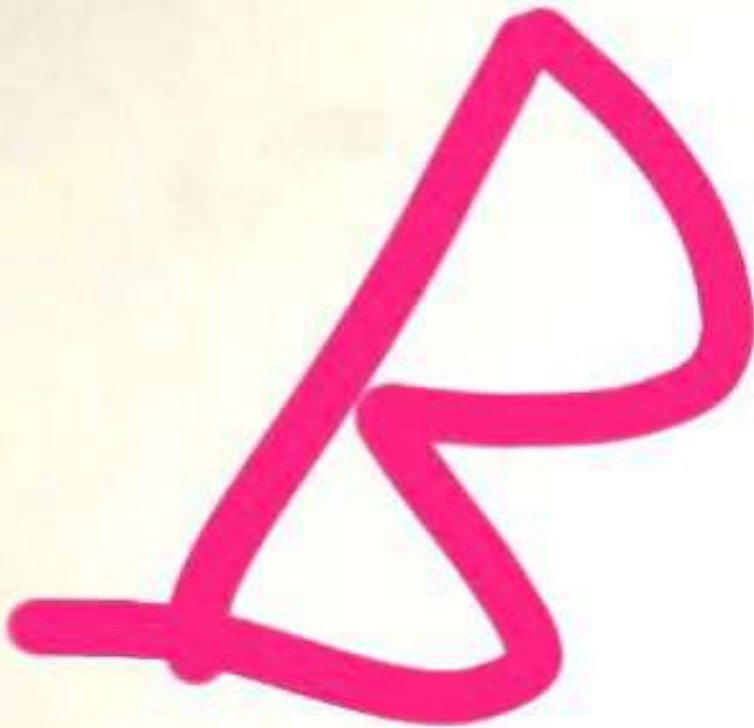
- burning natural gas
- cutting paper
- cutting a tomato
- melting ice

A

Question No. 4

9.31 g is the same mass as _____.

- 931 μg
- 9310 mg
- 93.1 cg
- 931 kg



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Question No. 5

The correct prefix for the multiplier 1,000,000,000 is

- milli
- giga
- mega
- tera

B

Question No. 1

Only two electrons, with opposing spins, are allowed in each orbital is known as the

- Hund's rule.
- Heisenberg uncertainty principle.
- Aufbau principle.
- Pauli exclusion principle.



Question No. 3

How many protons and neutrons are in Cl-37?

- 17 protons, 37 neutrons
- 20 protons, 17 neutrons
- 17 protons, 20 neutrons
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Question No. 1

Each electron occupies the lowest energy orbital available is known as the

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B

Question No. 1

Which statement below accurately describes the contributions of Josef Proust?

- ancient Greek philosopher who proposed that matter was continuous
- created the modern periodic table
- proposed the modern Atomic Theory
- discovered the law of definite proportions



Question No. 2

Semiconductors are located in the periodic table on (or in) the



- left side of the table.
- first period of the table.
- right side of the table.
- line dividing metals from nonmetals in the table.

INSTRUCTION: → Please choose the BEST answer from the given options for

Question:

Which is a characteristic property of nonmetals?

Options:

- thermal conductivity
- brittle
- luster
- malleability

B

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