

تجميعات " ريفان و ود "

دعواتكم 🙏❤️❤️.

# CHEMISTRY

2020

الميد الأول

Question No. 4

The electron configuration of Ne is .....

- $1s^2 2s^2 2p^6$
- $1s^2 2s^2 2p^5$
- $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
- $1s^2 2s^2 2p^3$

Save & Next

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Total questions in exam: 25 | Answered: 2

**Question No. 2**

How many inches are there in 0.5 km? (1 inch = 2.54 cm)

- 127000 inch
- 19685 inch
- 1968.5 inch
- 12700 inch



Question No. 1

If the temperature is 128 °F, what is the temperature in degrees Celsius?

- 53.33 °C
- 262.40 °C
- 401.15 °C
- 145.15 °C



Liquids have a ..... mass, a ..... volume, and a ..... shape.

- constant, fixed, variable
- constant, fixed, definite
- constant, variable, definite
- variable, variable, definite

Save & Next

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TABLE 101 SUPPLEMENTAL INFORMATION

$C = K - 273$ $1 \text{ mol} = 6.022 \times 10^{23}$ $n = \frac{m}{M}$	$K = (C) + 273$ $pH = -\log [H^+]$ $\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ Avogadro's No. = $6.022 \times 10^{23}$
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$n = \frac{m}{M}$

$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

Avogadro's No. =  $6.022 \times 10^{23}$

The electronic distribution of the electrons on the principal energy levels of  $^{17}\text{Cl}$  is .....

- 2, 8, 7
- 2, 2, 8, 5
- 2, 8, 2, 5
- 2, 2, 5, 8

Save & Next

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Total questions in exam: 25 | Attempted: 16

Question No. 20

Which of the following properties is **NOT** a characteristic of the group 8A (18) in the periodic table?

- They are poor conductors of electricity.
- They are poor conductors of heat.
- They are gases at room temperature.
- They are shiny.

Save &amp; Next

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Question No. 1

If the temperature is 128 °F, what is the temperature in degrees Celsius?

- 53.33 °C
- 262.40 °C
- 401.15 °C
- 145.15 °C



Total questions in exam: 25 | Answered: 7

## Question No. 3

Which one of the following species has same number of electrons as  $\text{Mg}^{2+}$  cation?

- $\text{K}^+$
- $\text{Cl}^-$
- $\text{S}^{2-}$
- $\text{F}^-$

Save &amp; Next

The building-block of a compound, whose properties are the same as those of the compound is the .....

- element.
- atom.
- molecule.
- mixture.

Next >



Question No. 2

If the density of chloroform is  $1.492 \text{ g/mL}$ , what is the mass of  $225 \text{ mL}$ ?

- $226.49 \text{ g}$
- $150.80 \text{ g}$
- $335.70 \text{ g}$
- $6.63 \times 10^{-3} \text{ g}$

Save & Next

REFER TO SUPPLEMENTAL INFORMATION

PERIODIC TABLE OF THE ELEMENTS													
1	2	3	4	5	6	7	8	9	10	11	12	13	14
$1s^2$		$2s^2$	$2p^6$		$3s^2$	$3p^6$		$3d^{10}$	$4s^2$	$4p^6$		$4d^{10}$	$4f^{14}$
$1s^2$	$2s^2$	$2p^6$	$3s^2$	$3p^6$	$3d^{10}$	$4s^2$	$4p^6$	$4d^{10}$	$4f^{14}$	$5s^2$	$5p^6$	$5d^{10}$	$5f^{14}$
$1s^2$	$2s^2$	$2p^6$	$3s^2$	$3p^6$	$3d^{10}$	$4s^2$	$4p^6$	$4d^{10}$	$4f^{14}$	$5s^2$	$5p^6$	$5d^{10}$	$5f^{14}$
$1s^2$	$2s^2$	$2p^6$	$3s^2$	$3p^6$	$3d^{10}$	$4s^2$	$4p^6$	$4d^{10}$	$4f^{14}$	$5s^2$	$5p^6$	$5d^{10}$	$5f^{14}$
$1s^2$	$2s^2$	$2p^6$	$3s^2$	$3p^6$	$3d^{10}$	$4s^2$	$4p^6$	$4d^{10}$	$4f^{14}$	$5s^2$	$5p^6$	$5d^{10}$	$5f^{14}$
$1s^2$	$2s^2$	$2p^6$	$3s^2$	$3p^6$	$3d^{10}$	$4s^2$	$4p^6$	$4d^{10}$	$4f^{14}$	$5s^2$	$5p^6$	$5d^{10}$	$5f^{14}$
$1s^2$	$2s^2$	$2p^6$	$3s^2$	$3p^6$	$3d^{10}$	$4s^2$	$4p^6$	$4d^{10}$	$4f^{14}$	$5s^2$	$5p^6$	$5d^{10}$	$5f^{14}$
$1s^2$	$2s^2$	$2p^6$	$3s^2$	$3p^6$	$3d^{10}$	$4s^2$	$4p^6$	$4d^{10}$	$4f^{14}$	$5s^2$	$5p^6$	$5d^{10}$	$5f^{14}$
$1s^2$	$2s^2$	$2p^6$	$3s^2$	$3p^6$	$3d^{10}$	$4s^2$	$4p^6$	$4d^{10}$	$4f^{14}$	$5s^2$	$5p^6$	$5d^{10}$	$5f^{14}$

$^\circ\text{C} = \text{K} - 273$

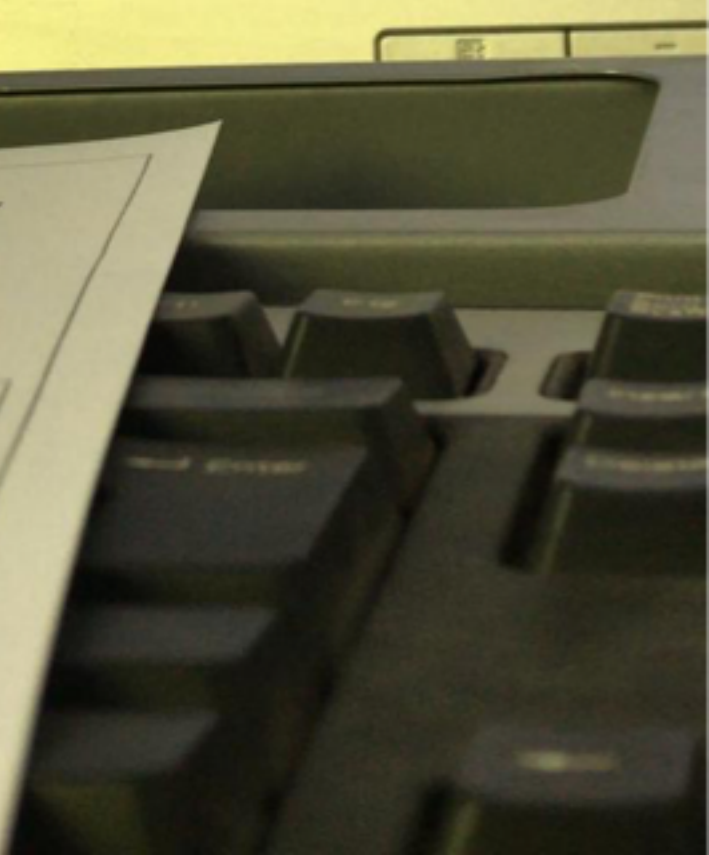
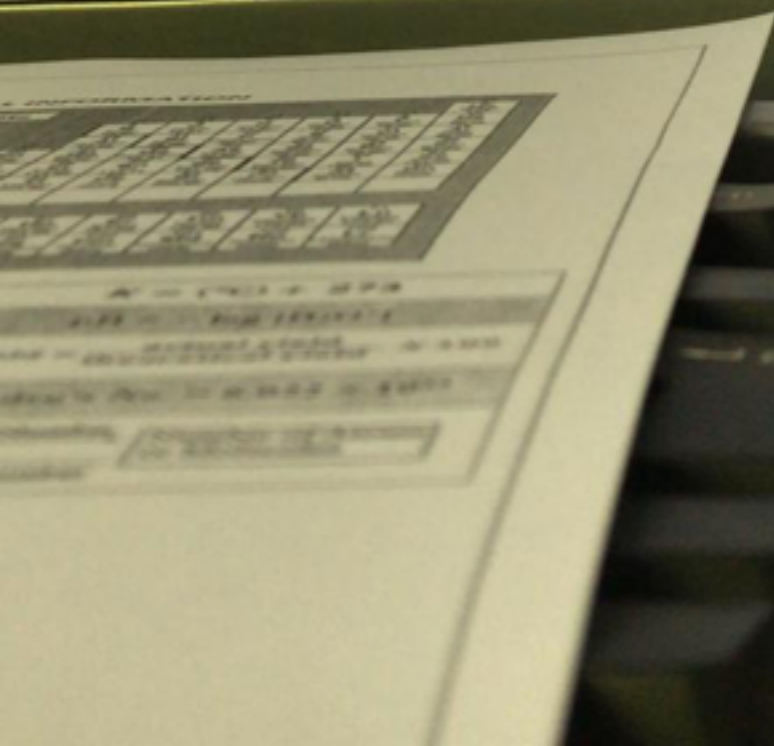
$\text{K} = (^\circ\text{C}) + 273$

The Rutherford gold foil experiment demonstrated that atoms .....

- are visible to the naked eye
- are homogeneous
- consist of an almost empty nucleus surrounded by a dense cloud of electrons.
- consist of a dense nucleus surrounded by mostly empty space.

Save & Next

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Total questions in exam: 25 | Answered: 0

Question No. 1

Which of the following pairs are isotopes?

- H-2 and He-2
- Pa-141 and Po-209
- U-238 and U-240
- C-12 and Co-59

Save & Next

Question No. 5

An iodine (I) atom has a mass number,  $A = 131$ .  
How many protons, neutrons, and electrons does it have?

- 53 protons, 131 neutrons, 54 electrons.
- 53 protons, 78 neutrons, 53 electrons.
- 39 protons, 78 neutrons, 39 electrons.
- 53 protons, 78 neutrons, 54 electrons.

Save &amp; Next





Total questions in exam: 25 | Answered: 2

Question No. 1

The "f" sublevel can hold up to ..... electrons.

- 18
- 8
- 14
- 6

Save &amp; Next

Element	Symbol	Atomic Number	Group	Period	Block	Classification
Hydrogen	H	1	1	1	s	Non-metal
Helium	He	2	18	1	s	Noble gas
Lithium	Li	3	1	2	s	Alkali metal
Beryllium	Be	4	2	2	s	Alkaline earth metal
Boron	B	5	13	2	p	Metalloid
Carbon	C	6	14	2	p	Non-metal
Nitrogen	N	7	15	2	p	Non-metal
Oxygen	O	8	16	2	p	Non-metal
Fluorine	F	9	17	2	p	Non-metal
Neon	Ne	10	18	2	p	Noble gas
Sodium	Na	11	1	3	s	Alkali metal
Magnesium	Mg	12	2	3	s	Alkaline earth metal
Aluminum	Al	13	13	3	p	Metal
Silicon	Si	14	14	3	p	Metalloid
Phosphorus	P	15	15	3	p	Non-metal
Sulfur	S	16	16	3	p	Non-metal
Chlorine	Cl	17	17	3	p	Non-metal
Argon	Ar	18	18	3	p	Noble gas
Potassium	K	19	1	4	s	Alkali metal
Calcium	Ca	20	2	4	s	Alkaline earth metal
Scandium	Sc	21	3	4	d	Transition metal
Titanium	Ti	22	4	4	d	Transition metal
Vanadium	V	23	5	4	d	Transition metal
Chromium	Cr	24	6	4	d	Transition metal
Manganese	Mn	25	7	4	d	Transition metal
Iron	Fe	26	8	4	d	Transition metal
Cobalt	Co	27	9	4	d	Transition metal
Nickel	Ni	28	10	4	d	Transition metal
Copper	Cu	29	11	4	d	Transition metal
Zinc	Zn	30	12	4	d	Transition metal
Gallium	Ga	31	13	4	p	Metal
Germanium	Ge	32	14	4	p	Metalloid
Arsenic	As	33	15	4	p	Metalloid
Selenium	Se	34	16	4	p	Non-metal
Bromine	Br	35	17	4	p	Non-metal
Krypton	Kr	36	18	4	p	Noble gas
Rubidium	Rb	37	1	5	s	Alkali metal
Strontium	Sr	38	2	5	s	Alkaline earth metal
Yttrium	Y	39	3	5	d	Transition metal
Zirconium	Zr	40	4	5	d	Transition metal
Niobium	Nb	41	5	5	d	Transition metal
Molybdenum	Mo	42	6	5	d	Transition metal
Technetium	Tc	43	7	5	d	Transition metal
Ruthenium	Ru	44	8	5	d	Transition metal
Rhodium	Rh	45	9	5	d	Transition metal
Palladium	Pd	46	10	5	d	Transition metal
Silver	Ag	47	11	5	d	Transition metal
Cadmium	Cd	48	12	5	d	Transition metal
Indium	In	49	13	5	p	Metal
Tin	Sn	50	14	5	p	Metal
Antimony	Sb	51	15	5	p	Metalloid
Tellurium	Te	52	16	5	p	Metalloid
Iodine	I	53	17	5	p	Non-metal
Xenon	Xe	54	18	5	p	Noble gas
Cesium	Cs	55	1	6	s	Alkali metal
Barium	Ba	56	2	6	s	Alkaline earth metal
Lanthanum	La	57	3	6	f	Lanthanide
Cerium	Ce	58	4	6	f	Lanthanide
Praseodymium	Pr	59	5	6	f	Lanthanide
Neodymium	Nd	60	6	6	f	Lanthanide
Europium	Eu	63	7	6	f	Lanthanide
Gadolinium	Gd	64	8	6	f	Lanthanide
Terbium	Tb	65	9	6	f	Lanthanide
Dysprosium	Dy	66	10	6	f	Lanthanide
Ytterbium	Yb	70	14	6	f	Lanthanide
Lutetium	Lu	71	15	6	f	Lanthanide
Hafnium	Hf	72	4	6	d	Transition metal
Tantalum	Ta	73	5	6	d	Transition metal
Tungsten	W	74	6	6	d	Transition metal
Rhenium	Re	75	7	6	d	Transition metal
Osmium	Os	76	8	6	d	Transition metal
Iridium	Ir	77	9	6	d	Transition metal
Rhodium	Rh	78	10	6	d	Transition metal
Platinum	Pt	78	10	6	d	Transition metal
Gold	Au	79	11	6	d	Transition metal
Mercury	Hg	80	12	6	d	Transition metal
Thallium	Tl	81	13	6	p	Metal
Lead	Pb	82	14	6	p	Metal
Bismuth	Bi	83	15	6	p	Metalloid
Polonium	Po	84	16	6	p	Radioactive
Astatine	At	85	17	6	p	Radioactive
Francium	Fr	87	1	7	s	Alkali metal
Radium	Ra	88	2	7	s	Alkaline earth metal
Actinium	Ac	89	3	7	f	Actinide
Thorium	Th	90	4	7	f	Actinide
Protactinium	Pa	91	5	7	f	Actinide
Uranium	U	92	6	7	f	Actinide
Neptunium	Np	93	7	7	f	Actinide
Plutonium	Pu	94	8	7	f	Actinide
Americium	Am	95	9	7	f	Actinide
Curium	Cm	96	10	7	f	Actinide
Berkelium	Bk	97	11	7	f	Actinide
Californium	Cf	98	12	7	f	Actinide
Einsteinium	Es	99	13	7	f	Actinide
Mendelevium	Md	101	15	7	f	Actinide
Nobelium	No	102	16	7	f	Actinide
Lr	Lr	103	17	7	f	Actinide
Rutherfordium	Rf	104	4	8	d	Transition metal
Dubnium	Du	105	5	8	d	Transition metal
Seaborgium	Sg	106	6	8	d	Transition metal
Bh	Bh	107	7	8	d	Transition metal
Hassium	Hs	108	8	8	d	Transition metal
Mt	Mt	109	9	8	d	Transition metal
Darmstadtium	Ds	110	10	8	d	Transition metal
Roentgenium	Rg	111	11	8	d	Transition metal
Copernicium	Cn	112	12	8	d	Transition metal
Nihonium	Nh	113	13	8	p	Metal
Flerovium	Fl	114	14	8	p	Metal
Moscovium	Mc	115	15	8	p	Metalloid
Livermorium	Lv	116	16	8	p	Non-metal
Tennessine	Ts	117	17	8	p	Non-metal
Oganesson	Og	118	18	8	p	Noble gas

Question No. 7

In the periodic table, group 7A (17) elements are also called

.....

- alkaline earth metals.
- alkali metals.
- noble gases.
- halogens.

Save & Next

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ADDITIONAL INFORMATION

$R = (^\circ C) + 273$
$pH = -\log [H^+] ]$
$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$
Avogadro's No. = $6.022 \times 10^{23}$



Total questions in exam: 25 | Answered: 12

Question No. 20

How many unpaired electrons are there in an Iron atom (Fe) in its ground state?

- 2
- 4
- 6
- 0

Save & Next

Total questions in exam: 25 | Answered: 2

Question No. 14

Compounds are composed of .....

- bonded atoms of two or more elements.
- only one type of atoms.
- bonded atoms of the same element.
- two or more mixed substances.

Save & Next



Total questions in exam: 25 | Answered: 3

Question No. 1

A<sup>-</sup>

A

A<sup>+</sup>

How many micrometers are there in 6.0 km?

- $6.0 \times 10^6$  micrometers
- $1.7 \times 10^{-9}$  micrometers
- $6.0 \times 10^9$  micrometers
- $1.7 \times 10^{-4}$  micrometers

Save &amp; Next



Total questions in exam: 25 | Answered: 11

Question No. 2

A<sup>-</sup>

A

A pure substance is .....

- composed of two or more different types of atoms or molecules combined in variable proportions.
- composed of only one type of atoms or molecules.
- composed of two or more regions with different compositions.
- composed of two or more different types of atoms or molecules homogenously mixed together.

Save &amp; Next



Total questions in exam: 25 | Answered: 11

**Question No. 6**

How would you classify sugar?

- A heterogeneous mixture.
- A homogeneous mixture.
- An element.
- A compound.

Save & Next

10.05.7.215

HP LE1851w

MKCL

Total questions in exam: 25 | Answered: 11

Question No. 20

Which of the following statements is **NOT** correct?

- Atoms are made of subatomic particles.
- An atom has mass.
- Atoms are invisible.
- Each element is formed of different atoms.

Save & Next

10.05.7.215

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MKCL OES Exam C



Total questions in exam: 25 | Answered: 11

Question No. 7

When an atom gains electron, the resulting particle is called .....

- an isotope.
- a proton.
- a cation.
- an anion.

Save & Next

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HP LE1851w

MKCL OES Exam Client V

Total questions in exam: 25 | Answered: 11

Question No. 5

Which of these is a physical property?

- Lead (Pb) becomes a liquid when heated to 601 degree Celsius
- Acidity of sulfuric acid
- Flammability of gasoline
- Toxicity of cyanide

Save & Next

10 86 7 215

P LE1851w

MKCL OES Exam Client

Total questions in exam: 25 | Answered: 11

Question No. 16

..... has one kind of atoms while ..... has more than one kind.

- A compound, an element
- An element, a compound
- A mixture, an element
- A mixture, a compound

Save & Next

2/215



Total questions in exam: 25 | Answered: 11

Question No. 10

Isotopes are .....

- atoms of the same element that have the same number of neutrons.
- atoms of the same element that have different number of neutrons.
- atoms of the same element that have different number of electrons.
- atoms of the same element that have different number of protons.

Save & Next

## Question No. 11

Which orbital-filling diagram does NOT obey Hund's rule?

- $\begin{array}{c} \uparrow \\ 4s \end{array} \quad \begin{array}{c} \uparrow \\ \uparrow \\ \uparrow \\ 3d \end{array} \quad \begin{array}{c} \uparrow \\ \uparrow \\ \uparrow \\ \uparrow \\ \uparrow \end{array}$
- $\begin{array}{c} \uparrow\downarrow \\ 4s \end{array} \quad \begin{array}{c} \uparrow \\ \uparrow \\ \uparrow \\ 3d \end{array} \quad \begin{array}{c} \uparrow \\ \uparrow \\ \uparrow \\ \uparrow \\ \uparrow \end{array}$
- $\begin{array}{c} \uparrow\downarrow \\ 4s \end{array} \quad \begin{array}{c} \uparrow\downarrow \\ \uparrow \\ \uparrow \\ 3d \end{array} \quad \begin{array}{c} \uparrow \\ \uparrow \\ \uparrow \\ \uparrow \\ \uparrow \end{array}$
- $\begin{array}{c} \uparrow\uparrow \\ 4s \end{array} \quad \begin{array}{c} \uparrow \\ \uparrow \\ \uparrow \\ 3d \end{array} \quad \begin{array}{c} \uparrow \\ \uparrow \\ \uparrow \\ \uparrow \\ \uparrow \end{array}$

Save &amp; Next

10/06/2015

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Total questions in exam: 25 | Answered: 11

## Question No. 21

Potassium element (K) is .....

- a noble gas.
- an alkaline earth metal.
- an alkali metal.
- a halogen.

Save &amp; Next

10.09.2015

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Total questions in exam: 25 | Answered: 11

## Question No. 19

A-

A

Which of the following properties is **NOT** a characteristic of the group 8A (18) in the periodic table?

- They are gases at room temperature.
- They are shiny.
- They are poor conductors of electricity.
- They are poor conductors of heat.

Previous &amp; Next

Total questions in exam: 25 | Answered: 11

## Question No. 22

Which of the following can NOT be considered as matter?

- wood
- sand
- paper
- light

Save &amp; Next

10.66.7.215

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Total questions in exam: 25 | Answered: 11

## Question No. 23

A<sup>-</sup>

How many unpaired electrons are there in an Iron atom (Fe) in its ground state?

- 0
- 6
- 2
- 4

Save &amp; Next

10/05/2016

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Total questions in exam: 25 | Answered: 12

Question No. 23

Nonmetals exist as .....

- liquids.
- all of them.
- solids.
- gases.

Save & Next



Total questions in exam: 25 | Answered: 11

## Question No. 24

A

Which one of the following elements is a poor conductor of heat and electricity?

- Copper (Cu)
- Fluorine (F)
- Iron (Fe)
- Lead (Pb)

Save &amp; Next

10.65.7.215

Nitrogen gas has a(n) \_\_\_\_\_ volume and a(n) \_\_\_\_\_ shape.

- definite; indefinite
- definite; definite
- indefinite; definite
- indefinite; indefinite



Total questions in exam: 25 | Answered: 12

Question No. 25

H<sub>2</sub> is an example on .....

- atomic elements
- molecular elements
- compounds
- homogeneous mixtures

Save & Next

The building-block of a compound, whose properties are the same as those of the compound is the .....

- element.
- atom.
- molecule.
- mixture.

Next >

Question No. 20

Which of the following properties is **NOT** a characteristic of the group 8A (18) in the periodic table?

- They are poor conductors of electricity.
- They are poor conductors of heat.
- They are gases at room temperature.
- They are shiny.

Save &amp; Next



Question No. 1

If the temperature is 128 °F, what is the temperature in degrees Celsius?

- 53.33 °C
- 262.40 °C
- 401.15 °C
- 145.15 °C

Question No. 14



Consider the Bromine isotope Br-81, select the combination which lists respectively the atomic number, neutrons number, and mass number.

- 46, 81, 35
- 35, 46, 81
- 35, 81, 46
- 81, 46, 35

Save &amp; Next



Question No. 12

An element cannot .....

- combine with other elements to form compound.
- be part of a heterogeneous mixture.
- be part of a homogeneous mixture.
- be broken down into simpler substances in a chemical process.

Save & Next

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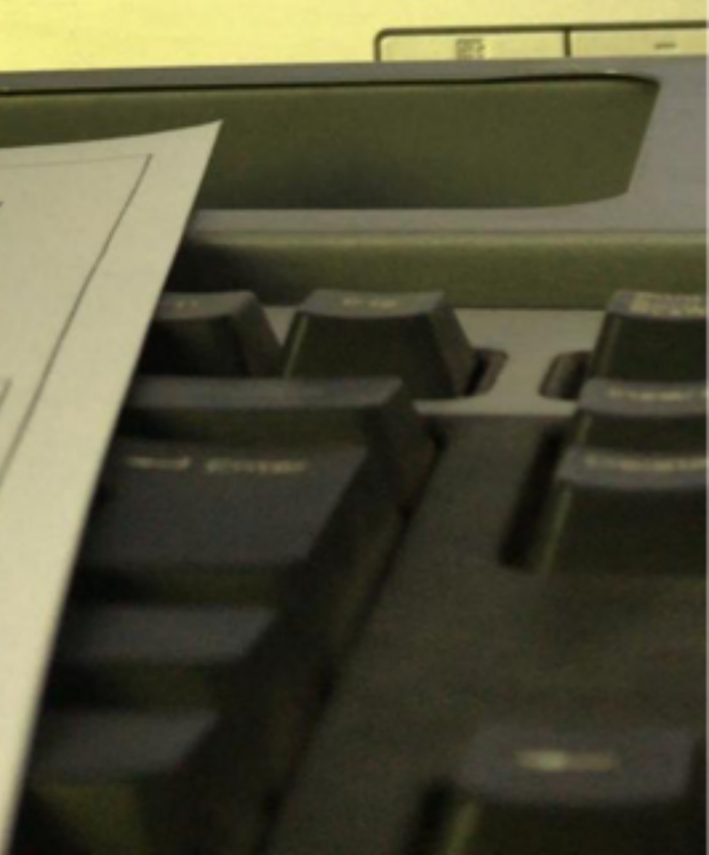
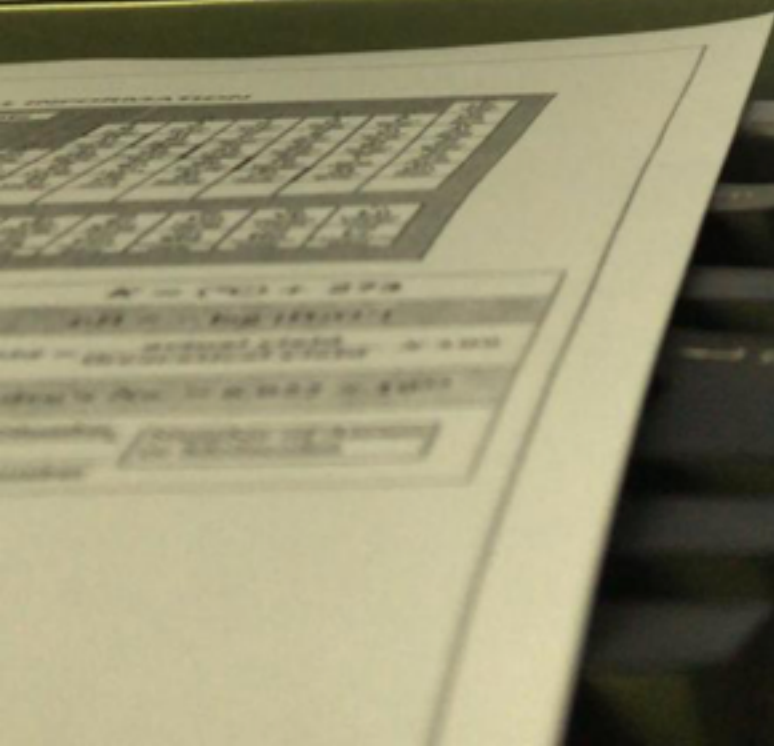


The Rutherford gold foil experiment demonstrated that atoms .....

- are visible to the naked eye
- are homogeneous
- consist of an almost empty nucleus surrounded by a dense cloud of electrons.
- consist of a dense nucleus surrounded by mostly empty space.

Save & Next

Compaq (E171)



## Question No. 15

An anion is defined as \_\_\_\_\_

- a group of stable atoms.
- a charged atom or group of atoms with a net negative charge.
- an atom or group of atoms with a net positive charge.
- a stable atom.

Save & Next

Total questions in exam: 25 | Answered: 12

## Question No. 24

What is the class of  $P_4$ ?

- Compound
- Homogeneous mixture
- Molecular element
- Heterogeneous mixture

[Save & Next](#)



Total questions in exam: 25 | Answered: 0

Question No. 1

Which of the following pairs are isotopes?

- H-2 and He-2
- Pa-141 and Po-209
- U-238 and U-240
- C-12 and Co-59

Save & Next

Question No. 23

Which of the following measurements are NOT equivalent?

- 24 dL = 2.4 L
- 150 msec = 0.150 sec
- 84 cm = 8.4 mm
- 25 mg = 0.025 g

Save & Next

## Question No. 15

An anion is defined as .....

- a group of stable atoms.
- a charged atom or group of atoms with a net negative charge.
- an atom or group of atoms with a net positive charge.
- a stable atom.

Save & Next



Total questions in exam: 25 | Answered: 12

## Question No. 17

Which of the following statements about energy is FALSE?

- Energy is the capacity to do work.
- Energy can neither be created nor destroyed.
- Energy cannot be transformed from one form to another.
- Energy is a fundamental part of chemical and physical changes.

Save &amp; Next

Total questions in exam: 25 | Answered: 3

Question No. 1

A<sup>-</sup>

A

A<sup>+</sup>

How many micrometers are there in 6.0 km?

- $6.0 \times 10^6$  micrometers
- $1.7 \times 10^{-9}$  micrometers
- $6.0 \times 10^9$  micrometers
- $1.7 \times 10^{-4}$  micrometers

Save &amp; Next

Question No. 20

Which of these elements is a good conductor of electricity?

- He
- Ag
- S
- N

Save & Next



What question is to answer: 20 / Assessment: 3

Question No. 10

What decimal power does the abbreviation "pico" represent?

- $1 \times 10^{-1}$
- $1 \times 10^{-12}$
- $1 \times 10^{-9}$
- $1 \times 10^6$

[Save & Next](#)

Question No. 24

The gaseous state is characterized by .....

- hardness
- fixed shape
- compressibility
- fixed volume

Save & Next

Question No. 17

The value of the electrical charge of electron was measured by

- J. Chadwick.
- R. Millikan.
- E. Rutherford.
- J. Dalton.

Save & Next



## Question No. 14

Compounds are composed of .....

- bonded atoms of two or more elements.
- only one type of atoms.
- bonded atoms of the same element.
- two or more mixed substances.

Question No. 11

Which of the following determines the identity of an atom?

- The number of neutrons.
- The number of electrons.
- The number of protons.
- The total number of protons and neutrons.

Save & Next

**Question No. 8**

Which of the following is NOT a way of changing the state of matter?

- Melting
- Freezing
- Mixing
- sublimation



Question No. 19

The fact that Carbon and Oxygen can combine together in different ratio of 1:1 and 1:2 is a representation to the Law of

.....

- definite proportions.
- conservation of matter.
- conservation of energy.
- multiple proportions.

**Question No. 18**

The maximum number of electrons that may occupy the second energy level ( $n = 2$ ) in the atom

- 2
- 8
- 10
- 18

Melting is a process in which ..... changes into ....., and it requires ..... the temperature.

- a solid, a gas, increasing
- a liquid, a gas, decreasing
- a liquid, a solid, increasing
- a solid, a liquid, increasing

### Question No. 3

---

Which of these elements is chemically similar to Chlorine (Cl)?

- Sulfur
- Bromine
- Krypton
- Germanium



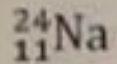
**Question No. 1**

The drops of water that appear on the outside of a glass of cold juice are an example of .....

- sublimation
- freezing
- evaporation
- condensation



How many protons, neutrons, and electrons respectively are there in an atom of Sodium?



- 24, 11, 12
- 11, 13, 11
- 13, 12, 24
- 11, 11, 24

Question No. 4

Which of these is a physical property?

- Acidity of sulfuric acid
- Lead (Pb) becomes a liquid when heated to 601 degree Celsius
- Flammability of gasoline
- Toxicity of cyanide

Melting is a process in which ..... changes into  
....., and it requires ..... the temperature.

- a solid, a gas, increasing
- a liquid, a gas, decreasing
- a liquid, a solid, increasing
- a solid, a liquid, increasing



**Question No. 14**

Which of the following quantities represents the largest mass?

- $2.0 \times 10^2$  mg
- 0.0010 kg
- $1.0 \times 10^5$  mg
- $2.0 \times 10^2$  g.

Question No. 13

An isotope represented as Fe-58

- is an isotope of fluorine.
- has 32 protons.
- contains 32 neutrons.
- must have 58 electrons.

The element vanadium (V) is one of the .....

noble gases.

halogens.

transition metals.

alkaline earth metals.

## Question No. 11

The charge of the Indium (In) ion is

1+

2+

4+

3+



Express 7.5 nm as picometers.

$7.5 \times 10^3 \text{ pm}$

$750 \text{ pm}$

$7.5 \times 10^{-6} \text{ pm}$

$75.0 \text{ pm}$

**Question No. 15**

All of the following statements about different elements are true, EXCEPT ..

- Barium is an alkaline earth metal.
- Bromine is a metalloid.
- Krypton is one of the noble gases.
- Manganese is a transition metal.

Question No. 18

The element vanadium (V) is one of the .....

- noble gases.
- transition metals.
- halogens.
- alkaline earth metals.

An object has a mass of 30.0 g and a volume of 5.0 cm<sup>3</sup>. What is its density?

- 6.00 g/cm<sup>3</sup>
- 17.00 g/cm<sup>3</sup>
- 6.00 g
- 7.00 g



A chloride ion ( $\text{Cl}^-$ ) has the same electron configuration as

.....

- an argon atom.
- a sulfur atom.
- a neon atom.
- a fluorine atom.

What is the prefix multiplier used to represent the factor  $10^{-6}$ ?

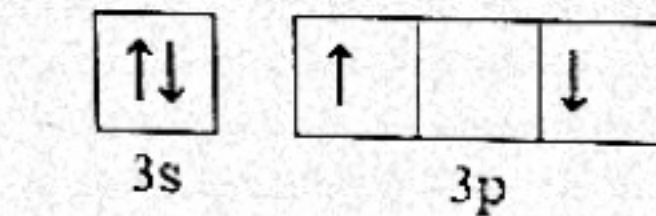
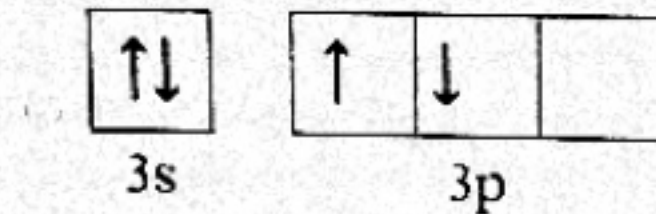
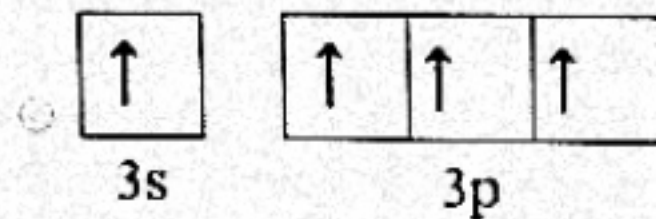
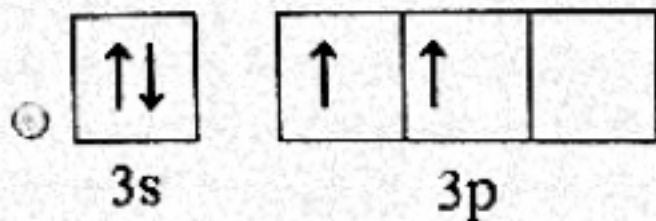
- Mega
- nano
- kilo
- micro

What is the density of a solid sample in g/mL if its volume was found to be 3.05 mL and its mass was 7.03 g.

- 0.043 g/mL
- 0.43 g/mL
- 21.44 g/mL
- 2.30 g/mL

Question No. 13

The orbital diagram of Silicon that **violates** Aufbau principal ...





What is the density of a solid sample in g/mL if its volume was found to be 3.05 mL and its mass was 7.03 g.

- 0.043 g/mL
- 0.43 g/mL
- 21.44 g/mL
- 2.30 g/mL

## Question No. 2

Which statement about an atom is FALSE?

- The nucleus is at the center of an atom.
- The nucleus accounts for a small portion of the mass of an atom.
- The volume occupied by electrons is referred to as an electron cloud.
- The space occupied by the electrons is almost empty.

Question No. 1

The charge of the nucleus of a Calcium atom is ...

+ 2+

+ 40+

+ 20-

+ 20+

### Question No. 5

What Kelvin temperature is equal to 75 °F?

- 43.00 K
- 167.00 K
- 297.04 K
- 23.88 K



Question No. 17

Combustion of fuel in car engine is considered as a .....

- physical change.
- physical property.
- chemical property.
- chemical change.

Question No. 2

Which of the following statements is true?

- Energy cannot be converted between potential and kinetic energies.
- As we increase the temperature of a substance, its kinetic energy decreases.
- All substances have both potential and kinetic energies, regardless to their physical states.
- Solids have the greatest kinetic energy and gases have the lowest kinetic energy.

Which of the following are isologous for each other?

$\frac{21}{19}A$ ,  $\frac{22}{20}B$ ,  $\frac{23}{21}C$  and  $\frac{24}{22}D$

- A and C
- A and B
- B and D
- C and D

**Question No. 16**

Which of these represents a chemical change?

- Mixing sugar and water at room temperature.
- Boiling water to form steam.
- Digestion of food in the stomach.
- Melting butter at room temperature.

**Question No. 4**

What volume of a liquid with a density of 13.53 g/mL is needed to provide 155 grams of this liquid?

- 2097.155 mL
- 141.47 mL
- 0.087 mL
- 11.45 mL



**Question No. 9**

Neon (Ne) has ..... valence electrons.

2

8

6

4

اي من الاتي مستحيل يكون درجة حرارة

200F

35F

100C

-273K

✓✓ ٩:٣٩ ص

جاب تحويل 2.21lbs الى kg بالاله ع طول

طلعت 5.6

٤:٥٤ م

The Number of proton in the nucleus of an atom?

- 1) is unique for each element
- 2) is call the atomic number
- 3) is the same for all isotopes of an element
- 4 ) all answers is correct?

The majority of the elements in the periodic table are:

م ٤:٥٤

فلزات-لافلزات-اشباه فلزات-غازات نبيله م ٤:٥٤

م ٤:٥٤ جاني وحتيت فلزات

\* جاني تحولي من فهرنهايت الي كلفن وکمان  
من فهرنهايت الي سيليز  
\*تعريف التکثف

ايبي ممايلي صحيح عن الماده؟

\*ايبي ممايلي خاصيه فزيائيه وکمان سوال  
ثاني جايب خصائص فيزيايه ويوقولك ايت  
فيهم كمياييه؟

ايبي ممايلي صحيح عن الطاقه؟ انها توجد  
بكل المواد

\*جاني عناصر وکان يبغا ايت فيهم نظير؟  
ايبي ممايلي مو من خصائص الفلزات؟

\*جائبي وش تعريف الايون؟ يفقد او يكتسب  
الكثرون

\* كمان ابي ممايلي يخالف الاوربتل لباولي؟

\* عناصر و يكتب اسم المجموعه حقتها

ويقولك ايت فيهم الي غلط؟

\*كمان يجيب لك عنصر الكربون والعدد

الكتلي قدامه تحديي البرتون والنيترون

والكثرون؟

ابي ممايلي غلط عن الطاقه؟ الي هي ماتنقل

من شكل لآخر

\*يجيبك لك رموزر عناصر وحدوريها

بالجدول الدوري

وجاني سوال كان يبغا الكثافه ومره يبغا

الحجم



Which doesn't chemical change ?

- 1) Liquid water became ice when cooled
- 2) zink react with an acid
- 3) milk turns sour in standing in room temperature
- 4) backing off bread.

جاني سؤال كاتب ايون الصوديوم موجب  
واحد وكاتب ان النيوترون ١٢ والاكترون ١٠  
وطالب العدد الكتلي والعدد الذري حطيت  
الذري ١١ والكتلي ٢٣