
اختبار اعادة
الكيمياء الكويز
الثاني ⭐
مريم جدو 💗 دعواتكم

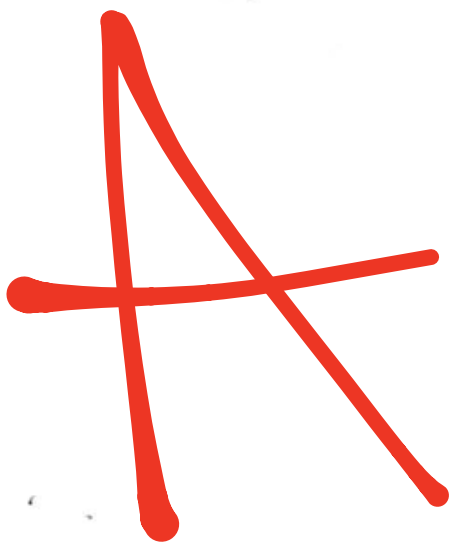
Total questions in exam: **25** | Answered: **0**

Question No. 24

What is the coefficient of chlorine gas after balancing the following equation?

$$\_\_Fe(s) + \_\_Cl_2(g) \rightarrow \_\_FeCl_3(s)$$

- 3
- 2
- 4
- 1



Question No. 19

The ionization energy of silicon is lower than the ionization energy of _____.

- sodium
- phosphorus
- aluminum
- magnesium

B

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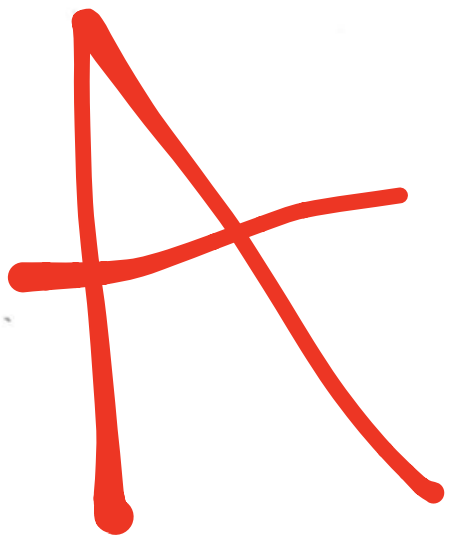
Question No. 22

Solid aluminum and gaseous oxygen react in a combination reaction to produce Al_2O_3

$$4Al\ (s) + 3\ O_2\ (g) \rightarrow 2\ Al_2O_3\ (s)$$

The maximum amount of Al_2O_3 that can be produced from 2.5 g of Al and 2.5 g of O_2 is _____ g.

- 4.7
- 7.4
- 5.3
- 9.4



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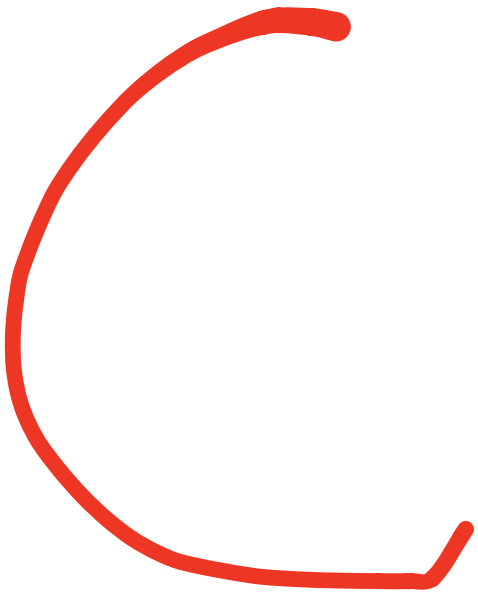
Question No. 21

What substance is the reducing agent in the following redox reaction?

$$Zn(s) + Cu^{2+}(aq) \rightarrow Zn^{2+}(aq) + Cu(s)$$

- Zn^{2+}
- Cu^{2+}
- Zn
- Cu



Question No. 20

The chemical formula of the compound f

- Mg_3Cl
- $MgCl_2$
- Mg_2Cl
- $MgCl$

Total questions in exam: 25 | Answered: 8

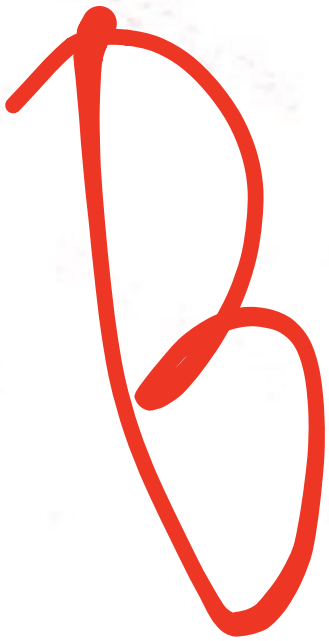
Question No. 3

The ionization energy of beryllium is higher than the ionization energy of _____

- boron
- lithium
- carbon
- nitrogen

In which of these substances are the atoms held together by polar covalent bonding?

- $SrCl_2$
- ClF
- CsCl
- Si



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Question No. 8

All of the following elements occurs naturally as diatomic molecu

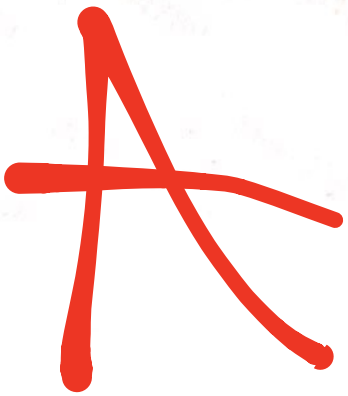
- chlorine
- fluorine
- neon
- iodine

Total questions in exam: **25** | Answered: **8**

Question No. 7

What volume (mL) of a 3.45 M lead nitrate solution must be diluted to make 450 mL of 0.99 M solution of lead nitrate?

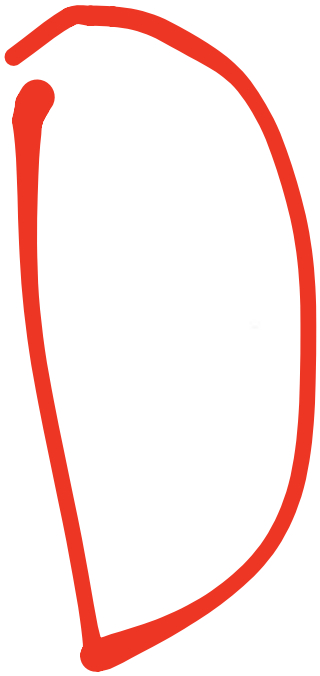
- 129 mL
- 101 mL
- 109 mL
- 56 mL



Question No. 6

When dissolved in water, _____ can make an aqueous solution that conducts electricity.

- C_8H_{18}
- CH_3OCH_3
- sugar
- ionic salt

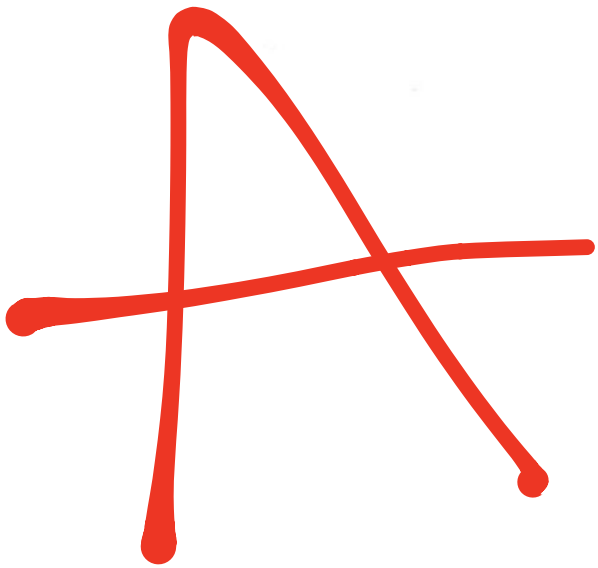


Question No. 5

In the reaction below, what is the theoretical yield in moles for NO when 3 moles of NH_3 react with 3 moles of O_2?

$$4\ NH_3 + 5\ O_2 \rightarrow 4\ NO + 6\ H_2O$$

- 2.4 mol
- 2.8 mol
- 2.6 mol
- 3.0 mol



The type of bond in I_2 is a (an) .....bond

- Metallic
- Polar covalent
- Nonpolar covalent
- Ionic

Question No. 19

The molarity (M) of an aqueous solution containing 22.5 g of sucrose ($C_{12}H_{22}O_{11}$) in 35.5 mL of solution is _____.

- 0.0657
- 3.52
- 1.85
- 0.104

Total questions in exam: 25 | Answered: 8

Question No. 3

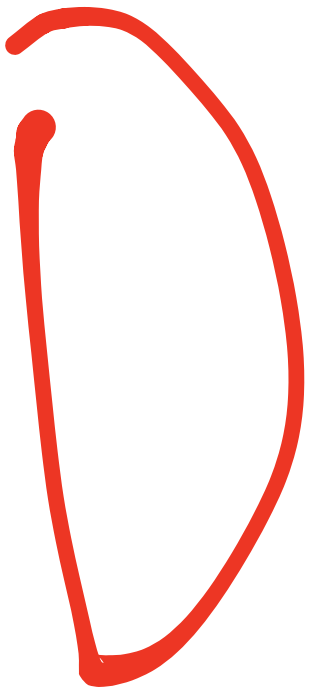
The ionization energy of beryllium is higher than the ionization energy of _____

- boron
- lithium
- carbon
- nitrogen

Question No. 17

Use the periodic table to answer the following question:
The formula CCl_4 has a molar mass of ___ g/mol.

- 140
- 150
- 146
- 154



The most correct name for the compound NI_3 is:

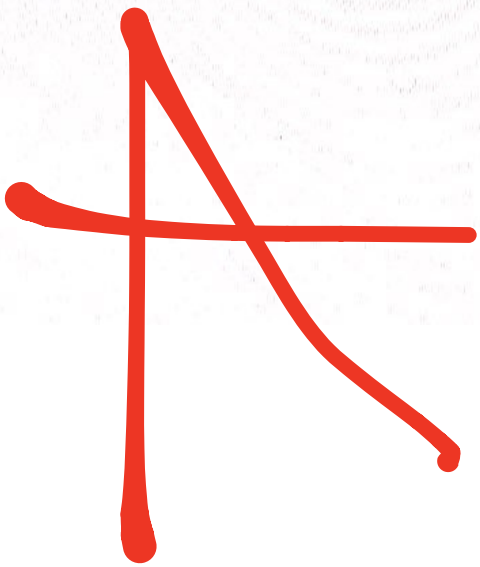
- mononitrogen triiodide
- nitrogen iodide
- triiodo nitrogen
- nitrogen triiodide

Total questions in exam: 25 | Answered: 8

Question No. 13

What is the final molarity of HNO_3 solution, if 300 mL of 2 M HNO_3 was diluted to a final volume of 0.2 L?

- 3.0 M
- 5.0 M
- 4.0 M
- 6.0 M



Total questions in exam: **25** | Answered: **9**

Question No. 23

Based on Lewis dot structure, the number of shared electrons in N

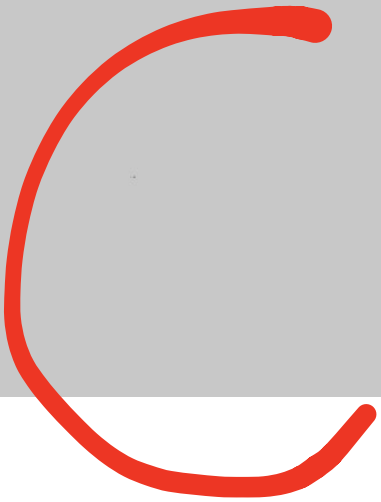
- 2
- 6
- 4
- 8

Question No. 24

The number of CO_2 molecules that are produced from burning of 57.11 g of C_8H_{18} (Molar mass = 114.22 g/mol) according to the following equation:

$$2\ C_8H_{18(l)} + 25\ O_{2(g)} \rightarrow 16\ CO_{2(g)} + 18\ H_2O_{(g)}$$

- 8 molecules.
- 16 molecules.
- 2.41×10^{24} molecules.
- 6.02×10^{23} molecules.



Total questions in exam: 25 | Answered: 9

Question No. 25

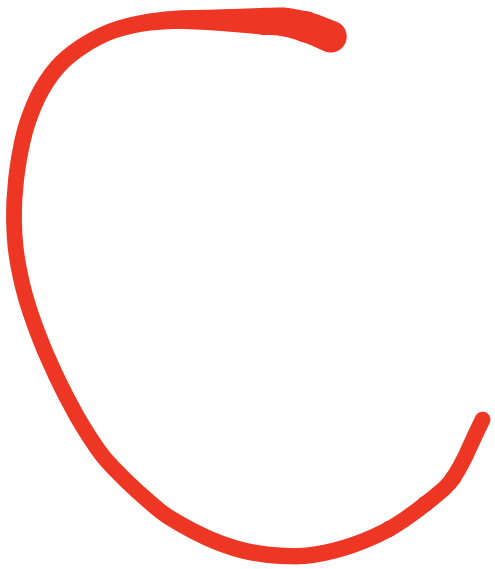
The correct name for the acid HBr is _____

- hydrobromic
- hydrogen bromide
- hydrogen bromate
- hydrogen bromine

Question No. 22

How many moles and how many atoms of zinc (Zn) are in a sample weighing 34.9 g?

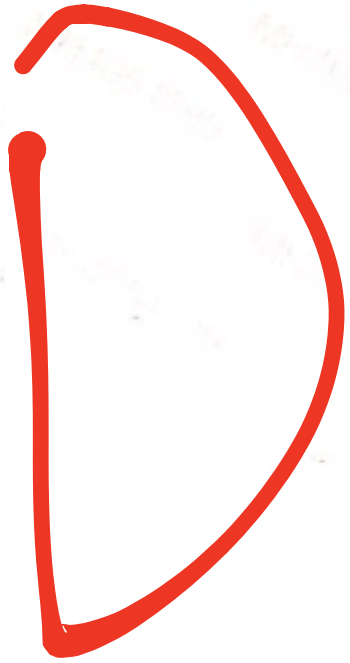
- 0.533 mol, 8.85×10^{-25} atoms
- 1.87 mol, 3.10×10^{-24} atoms
- 0.533 mol, 3.21×10^{23} atoms
- 1.87 mol, 1.13×10^{24} atoms



Question No. 20

What is the molecular formula of a compound that has a molar mass of 300 g/mol and its empirical formula is CH_2O?

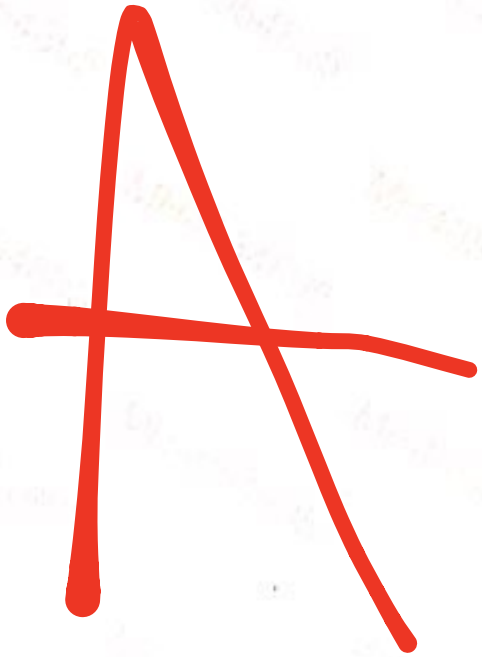
- $C_2H_4O_2$
- $C_4H_8O_4$
- $C_3H_6O_3$
- $C_{10}H_{20}O_{10}$



Question No. 22

Which of these substances gives a weak electrolyte when dissolved in water?

- weak base
- ionic salt
- strong acid
- strong base



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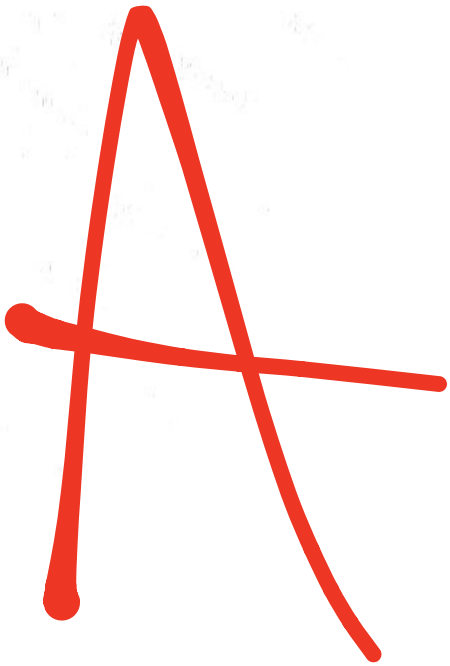
Total questions in exam: 25 | Answered: 21

Question No. 14

Find out the molecular formula of a compound that has a molar mass of 138.0 g/mol and an empirical formula of NO_2.

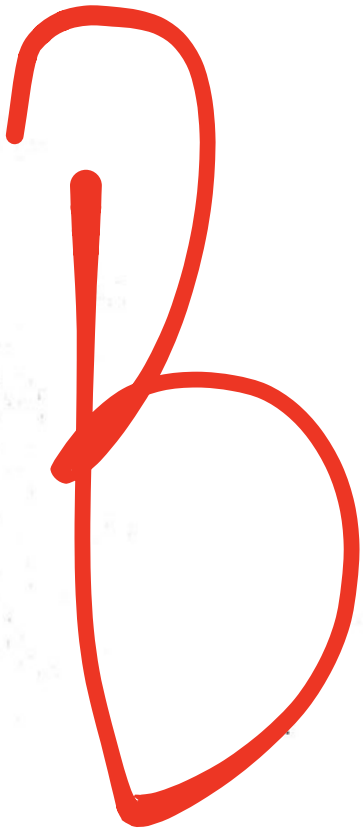
- N_3O_6
- NO_2
- N_2O_3
- N_2O_4



Question No. 20

The name of the chemical compound $FeCO_3$ is:

- iron carbonate
- iron(II) carbonate
- iron(III) carbonate
- iron(I) carbonate



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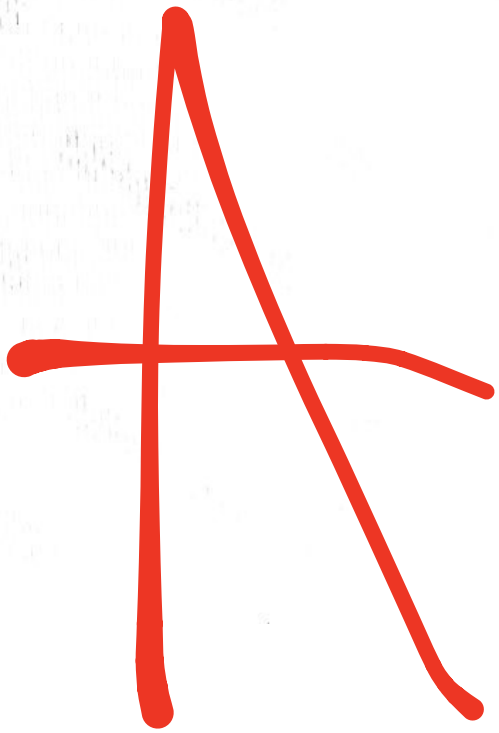

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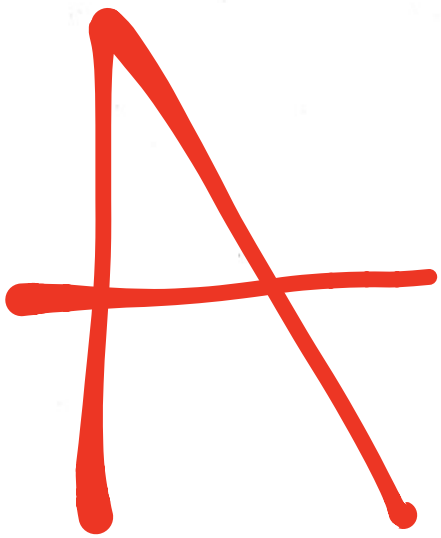
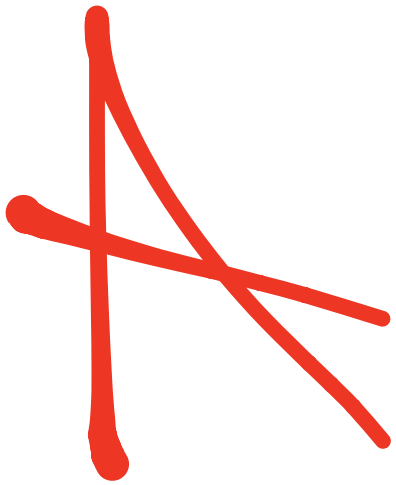
- N_3O_6
- NO_2
- N_2O_3
- N_2O_4



Question No. 15

The molar mass of $Ca_3(PO_4)_2$ is equal to:

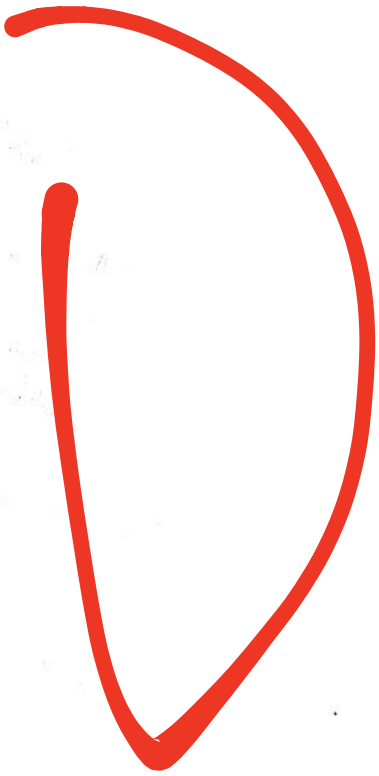
- 310 g/mol
- 250 g/mol
- 134 g/mol
- 215 g/mol



Question No. 19

Calculate the molar mass of $Fe_3(PO_4)_2$.

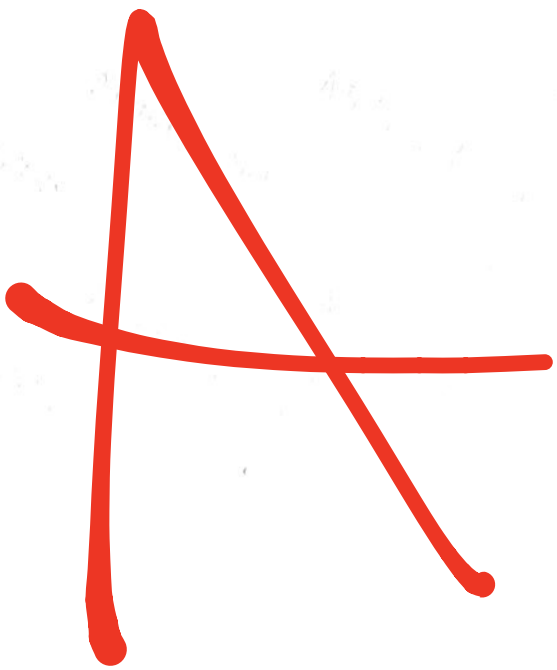
- 262.5 g/mol
- 525.1 g/mol
- 237.6 g/mol
- 357.5 g/mol



The number of CO_2 molecules that are produced from burning of 57.11 g of C_8H_{18} (Molar mass = 114.22 g/mol) according to the following equation:

$$2\ C_8H_{18(l)} + 25\ O_{2(g)} \rightarrow 16\ CO_{2(g)} + 18\ H_2O_{(g)}$$

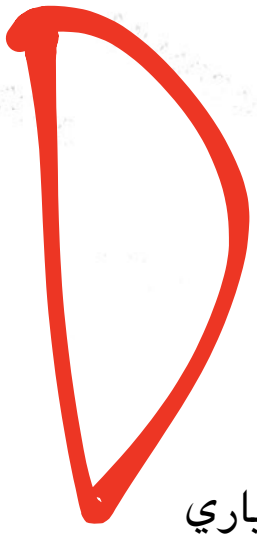
- 2.41×10^{24} molecules.
- 6.02×10^{23} molecules.
- 8 molecules.
- 16 molecules.



Question No. 7

The substance that causes the oxidation of another substance is called

- cathode
- reducing agent
- anode
- oxidizing agent



لأنه مفروض أختار اختزال بس مو موجود بالخياري
فالعامل المؤكسد هو نفسه اللي صارله اختزال

Question No. 8

For a given reaction

$$V_2O_5(s) + 5\,Ca(l) \rightarrow 2\,V(l) + 5\,CaO(s)$$

When 10 moles of V_2O_5 are mixed with 10 moles of Ca, which is the limiting reactant according to the above equation?

- V_2O_5
- CaO
- Ca
- V

Question No. 11

The number of CO_2 molecules that are produced from burning of 57.11 g of C_8H_{18} (Molar mass = 114.22 g/mol) according to the following equation:

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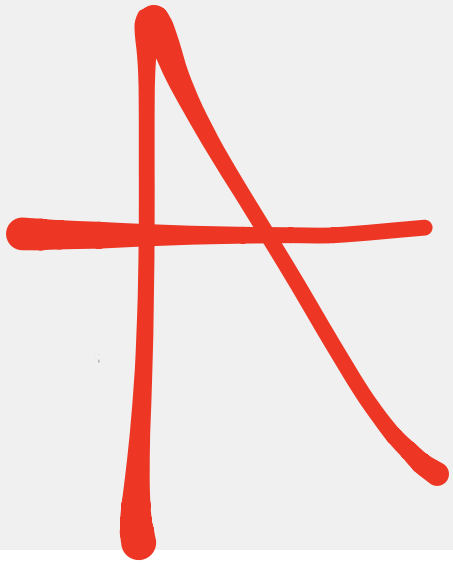
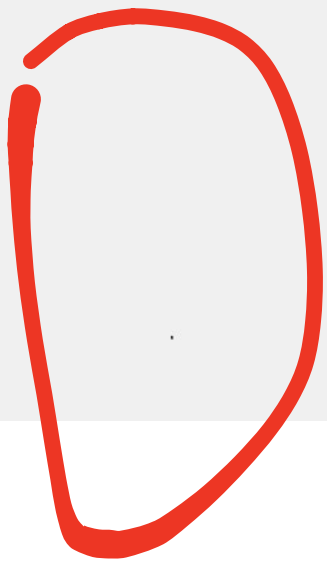


As you move from the top to the bottom of the periodic table:

- ionization energy increases and atomic radius decreases
- ionization energy decreases and atomic radius decreases
- ionization energy increases and atomic radius increases
- ionization energy decreases and atomic radius increases



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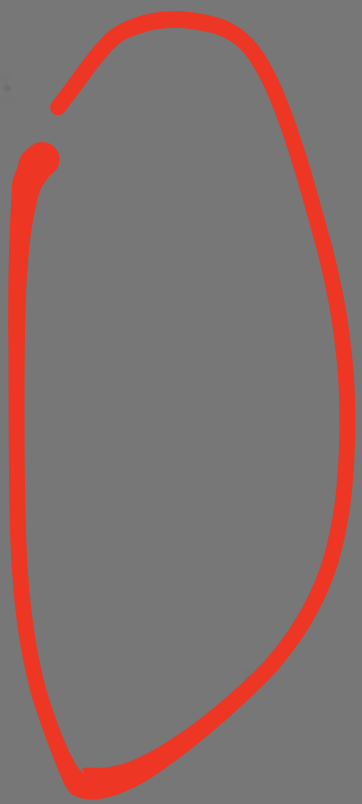
Total questions in exam: **25** | Answered: *0*

Question No. 2

Write the name for FeS

- iron(I) sulfide
- iron(II) sulfate
- iron(I) sulfate
- iron(II) sulfide



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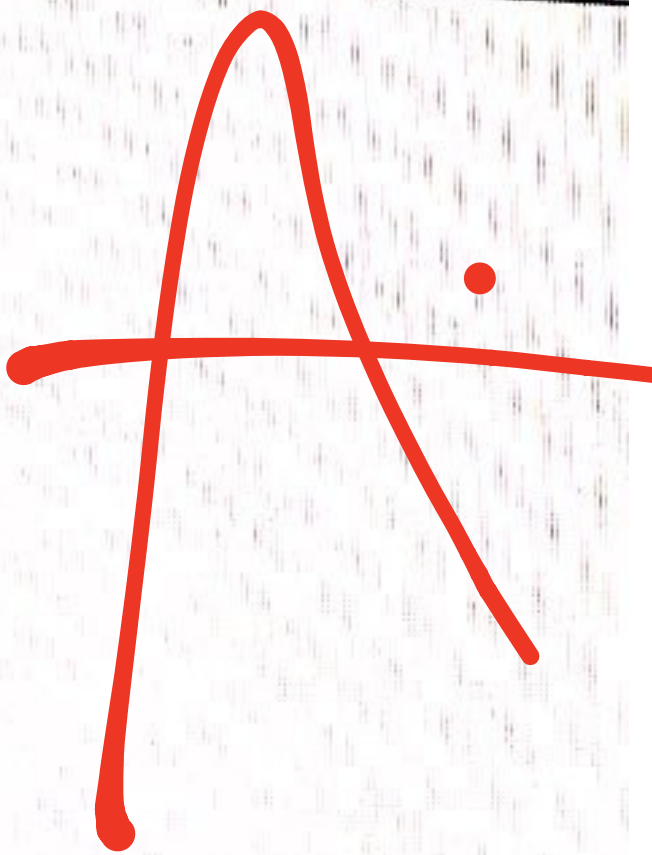

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Based on Lewis dot structure, the number of shared electrons in N_2 molecule is _____

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- 6
- 4
- 8

The type of bond in I_2 is a (an) .....bond

- Metallic
- Polar covalent
- Nonpolar covalent
- Ionic

The most correct name for the compound NI_3 is:

- mononitrogen triiodide
- nitrogen iodide
- triiodo nitrogen
- nitrogen triiodide

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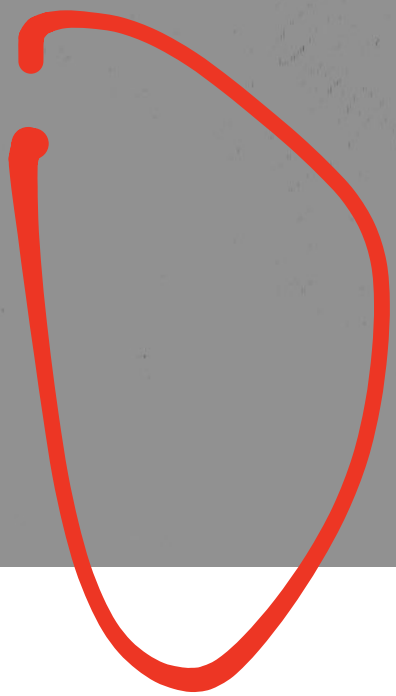

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Question No. 3

The ionization energy of beryllium is higher than the ionization energy of _____

- boron
- lithium
- carbon
- nitrogen

B

Total questions in exam: 25 | Answered: 7

Question No. 6

What coefficient is placed in front of O_2 to complete the balancing of the following equation?

$$C_5H_8 + ?\ O_2 \rightarrow 5CO_2 + 4H_2O$$

- 7
- 3
- 1
- 5

Question No. 5

Household sugar, sucrose, has the molecular formula $C_{12}H_{22}O_{11}$. What is the mass percent of carbon in sucrose?

- 51.4 %
- 62.8 %
- 6.5 %
- 42.1 %

Answered: 5

Question No. 4

What is the term for the shared valence electrons in a covalent molecule?

- core electrons
- nonbonding electrons
- bonding electrons
- lone pair electrons

Question No. 1

What is the oxidation number of iron in Fe_2O_3?

- +3
- -6
- -3
- +6

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Write the name for FeS

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- iron(II) sulfate
- iron(I) sulfate
- iron(II) sulfide