

Question No. 17

All of the following statements about different elements are true EXCEPT:

- Barium is an alkaline earth metal.
- Krypton is one of the noble gases.
- Manganese is a transition metal.
- Lithium is considered a metalloid.

D

Question No. 25

As you move up and to the right on the periodic table

- atomic radius increases and ionization energy increases
- atomic radius decreases and ionization energy increases
- atomic radius decreases and ionization energy decreases
- atomic radius increases and ionization energy decreases

C

Question No. 18

The Group 8A (18) elements

- are unreactive.
- are liquids at room temperature.
- are good conductors of electricity.
- melt at high temperatures.

A

Question No. 3

Which of the following is an example of a heterogeneous mixture?

- milk
- oil
- soda
- chicken soup

D

Question No. 5

How many inches are in 25.8 cm? [1 in = 2.54 cm]

- 0.1
- 10.2
- 28.3
- 0.0984

B

Question No. 21

0.100 mole of lithium weighs

- 3.00 g.
- 0.694 g.
- 0.300 g.
- 6.94 g.

D

Question No. 17

How many moles and how many atoms of zinc are in a sample weighing 34.9 g?

- 0.533 mol, 8.85×10^{-25} atoms
- 1.87 mol, 1.13×10^{24} atoms
- 1.87 mol, 3.10×10^{-24} atoms
- 0.533 mol, 3.21×10^{23} atoms

D

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Question No. 12

What is the atomic symbol for silicon?

- S
- Si
- Ag
- Au

B

Question No. 1

A solution has a density of 1.22 g/mL. What volume of the solution has a mass of 55.0 g?

- 0.00253 mL
- 49.4 mL
- 58.8 mL
- 45.1 mL

D

Question No. 11

Which statement below accurately describes the contributions of Thomson?

- Ancient Greek philosopher who proposed that matter was continuous
- Proposed the modern Atomic Theory
- Discovered the existence of electrons
- Created the modern periodic table

C

Question No. 24

Ionization energy is

- the energy an ion acquires from an electron.
- higher for potassium than for lithium.
- the energy needed to remove the least tightly bound electron.
- highest for metals in Group 1A(1).

C

Question No. 7

The state of matter that has fixed volume and unfixed shape is called?

- liquid
- plasma
- solid
- gas

A

Question No. 4

Which of the following items is a mixture?

- table salt
- cereal and milk
- baking soda
- sulfur

B

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Question No. 10

What is the value of 335 K on the Celsius temperature scale?

- 62
- 66.4
- 167
- 608

A

Question No. 19

What is the charge on the ion formed by aluminum?

- 5-
- 3+
- 3-
- 13+

B

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Question No. 16

0.100 mole of lithium weighs

- 3.00 g
- 0.694 g
- 6.94 g
- 0.300 g

C

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Question No. 19

Which of these elements has the highest electron affinity?

- magnesium
- silicon
- aluminum
- sulfur

D

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Question No. 9

What is the atomic symbol for gold?

- Ar
- As
- Au
- Ag

C

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Question No. 4

9.31 g is the same mass as _____

- 931 μg
- 9310 mg
- 93.1 cg
- 931 kg

B

Question No. 13

The elements rubidium, sodium, and potassium _____

- are isotopes of each other
- are in the same group
- have the same number of neutrons
- are in the same period of elements

B

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Question No. 11

One element that has 5 valence electrons is _____

- lithium
- sulfur
- nitrogen
- carbon

C

Question No. 13

The elements rubidium, sodium, and potassium _____

- are isotopes of each other
- are in the same group
- have the same number of neutrons
- are in the same period of elements

B

Question No. 6

Which state of matter has atomic spacing that is close together and fixed shape?

- liquid
- plasma
- gas
- solid

D

Question No. 9

Condensation refers to which conversion?

- solid -> liquid
- liquid -> gas
- gas -> liquid
- solid -> gas

C

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Question No. 14

Which one of these represents a physical change?

- bleach turns hair yellow
- sugar, when heated, becomes brown
- milk turns sour
- water, when heated, forms water vapor

D

Question No. 25

Chlorine has more metallic character than _____

- iodine
- bromine
- fluorine
- astatine

Question No. 4

Choose the heterogeneous mixture from the list below.

- carbon (graphite)
- chlorine gas
- chicken noodle soup
- black coffee

C

Question No. 9

Which of the following is NOT an example of matter?

- a dust particle
- a pencil eraser
- a balloon full of helium
- heat from a burning candle

D

Question No. 17

One element that has 8 valence electrons is _____.

- Argon
- Helium
- Carbon
- Oxygen

D

If matter is uniform throughout and cannot be separated into other substances by physical

- either an element or a compound
- a homogeneous mixture
- a heterogeneous mixture
- a compound

A

Question No. 1
Which type of energy is associated with motion of objects?

- electrical energy
- kinetic energy
- potential energy
- chemical energy

B

Question No. 7

What is the prefix multiplier used to represent the factor 10^{-6} .

- micro
- mega
- kilo
- nano

A

Question No. 4

Which of the following is an element?

- salt
- gold
- water
- sugar

B

Question No. 12

Which one of these represents a chemical change?

- melting butter
- boiling water to form water vapor
- grilling a steak
- mixing carbon and oxygen (at room temperature)

C

Which type of energy is associated with motion of objects?

- electrical energy
- kinetic energy
- potential energy
- chemical energy

B

Which family is most likely to form diatomic molecules?

- Group VII-A
- Group VIII-A
- Group V-A
- Group VI-A

A

Question No. 16

The elements xenon, helium and argon

- are isotopes of each other.
- are in the same period of elements.
- have the same number of neutrons.
- are in the same group.

D



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Question No. 25

The atomic radius of krypton is smaller than the atomic radius of _____

- xenon
- neon
- helium
- argon

A



Question No. 5

Salt water is classified as

- pure substance-element
- mixture-homogeneous
- pure substance-compound
- mixture-heterogeneous

B



Question No. 25

The ionization energy of strontium is higher than the ionization energy of _____

- magnesium
- beryllium
- barium
- calcium



C

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Question No. 12

The charge of electron was established by

- Millikan's oil drop experiment.
- Rutherford's gold foil experiment.
- Dalton's atomic experiment
- Thomson's cathode ray tube experiment.



A

Question No. 3

Which one of the following is homogeneous mixture?

- Alloys.
- Noodles soup.
- Vegetable salad.
- Sand in water.

A



Question No. 23

Which of the following elements is a metal?

- nitrogen
- fluorine
- strontium
- argon

C



**Question No. 7**

Which of the following is a property of both solids and liquids?

- definite shape
- indefinite shape
- indefinite volume
- definite volume

D





Question No. 6

A decimeter equals:

- 1 m
- 1000 mm
- 10 m
- 10 cm



C

Question No. 9

Which of the following is a chemical change?

- burning natural gas
- melting ice
- cutting paper
- cutting a tomato

A



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Question No. 5

The correct prefix for the multiplier 1,000,000,000 is

- milli.
- giga.
- mega.
- tera.

B



Question No. 8

Which of the following is matter?

- time
- ice cream
- sunlight
- sound



B

**Question No. 4**

Choose the heterogeneous mixture from the list below.

- carbon (graphite)
- chlorine gas
- chicken noodle soup
- black coffee



C

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Question No. 11

Gases and liquids share the property of _____

- unfixed shape
- compressibility
- fixed volume
- rigid shape

A

Question No. 24

Which of the following has the greatest number of nonbonding pairs of electrons?

- He
- F
- H
- C

B

Question No. 13

According to Hund's Rule

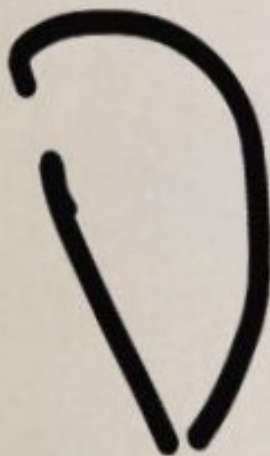
- no two electrons in the same atom can have the same four quantum numbers.
- electrons found in the same orbital must have the same spin.
- electrons fill orbitals of the same energy singly and pair up only after the orbitals are
- electrons pair first before filling other orbitals with same energy



Question No. 13

When filling orbitals that have same energy with electrons, fill them singly first, with parallel spins.

- Aufbau principle.
- Heisenberg uncertainty principle.
- Pauli exclusion principle.
- Hund's rule.



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Question No. 23

Where is most of the mass of an atom concentrated?

- nucleus
- neutrons
- electrons
- protons

A





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Question No. 24

Which of these elements has the highest electron affinity?

- Rubidium
- Cesium
- Lithium
- Potassium



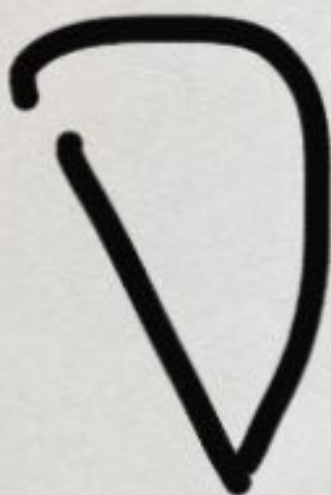
C



Question No. 13

When filling orbitals that have same energy with electrons, fill them singly first, with parallel spins, is known as

- Aufbau principle.
- Heisenberg uncertainty principle.
- Pauli exclusion principle.
- Hund's rule.





Question No. 21

0.100 mole of lithium weighs

- 6.94 g
- 0.094 g
- 0.300 g
- 3.00 g

B



Question No. 13

The elements rubidium, sodium, and potassium _____

- are isotopes of each other
- are in the same group
- have the same number of neutrons
- are in the same period of elements

B

Question No. 4

9.31 g is the same mass as _____

- 931 μg
- 9310 mg
- 93.1 cg
- 931 kg

B

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Question No. 10

An s orbital can hold _____ electrons.

- 6
- 2
- 18
- 10

B

Question No. 6

Which state of matter has atomic spacing that is close together and fixed shape?

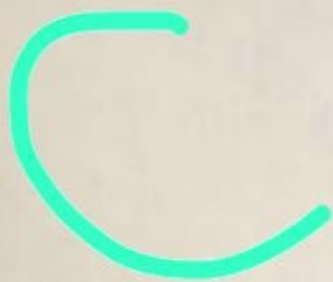
- liquid
- plasma
- gas
- solid

D

Question No. 5

A solid form of matter in which there is long range repeating order is called _____.

- amorphous
- fixed
- crystalline
- rigid



Question No. 7

Which of the following is a chemical change?

- burning natural gas
- cutting paper
- cutting a tomato
- melting ice

A

Question No. 12

Francium is an example of an element that is a _____.

- metalloïd
- noble gas
- nonmetal
- metal

D



INSTRUCTION: Please choose the BEST answer from the given options for each question.

Question:

Which of the following is a heterogeneous mixture?

Options:

- pure water
- tap water
- gasoline
- ice and water

D





Question No. 19

Which of these elements has the highest electron affinity?

- Rubidium
- Potassium
- Lithium
- Caesium

C



INSTRUCTION: **تعليمات:** Please choose the BEST answer from the given options for each

Question:

Which of the following is a physical process?

Options:

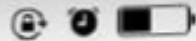
- the explosion of nitroglycerin
- the baking of a potato
- the rusting of iron
- the condensation of water vapor

D



تسليم الإجابة

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Today

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11:09 PM ✓✓

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11:10 PM ✓✓

Gold in a ring is a
-compound
-element
-heterogeneous
-homogeneous

11:10 PM ✓✓

Homogenous

11:11 PM

لأنه سبيكة alloy

11:12 PM

لو قالي فضه كمان نفس الشئ ؟

11:13 PM ✓✓

لا اعتقد

11:15 PM

يعني لو فضه احط عنصر صح 🙄

11:16 PM ✓✓

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11:17 PM



Question No. 24

Which of these elements has the lowest first ionization energy?

- aluminum
- sulfur
- sodium
- phosphorus

C



Question No. 18

Non Metals are located where on the periodic table?

- left side
- between metals and nonmetals
- middle
- right side



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Question No. 23

Which of the following elements is a metal?

- fluorine
- strontium
- nitrogen
- argon

B



Question No. 7

Which of the following is a chemical change?

- burning natural gas
- cutting paper
- cutting a tomato
- melting ice

A



Question No. 12

The names of the elements whose symbols are Si, P, Mn, and S are respectively.

- silver, phosphorus, magnesium, and sulfur.
- silicon, phosphorus, magnesium, and sulfur.
- silicon, potassium, magnesium, and sulfur.
- silicon, phosphorus, manganese, and sulfur.



P

Question No. 24

Which of these elements has the lowest electron affinity?

- carbon
- boron
- nitrogen
- beryllium

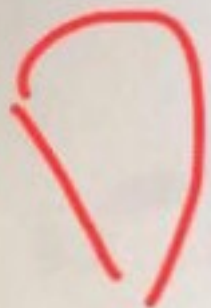
D



Question No. 20

The ionization energy of silicon is lower than the ionization energy for _____

- magnesium
- aluminum
- sodium
- phosphorus



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Question No. 11

One element that has 5 valence electrons is _____

- lithium
- sulfur
- nitrogen
- carbon

C




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Question No. 8

If the temperature is 76°F , what is the temperature in degrees Celsius?

- 25.6 $^{\circ}\text{C}$
- 145 $^{\circ}\text{C}$
- 262.4 $^{\circ}\text{C}$
- 401 $^{\circ}\text{C}$

A

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The final step in forming the modern atomic theory was done by:

A. Rutherford

B. Dalton

C. Proust

D. Lavoisier



Question No. 7

The state of matter that has fixed volume and unfixed shape is called?

- liquid
- plasma
- solid
- gas

A





Question No. 1

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B



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Question No. 10

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Question No. 24

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C