

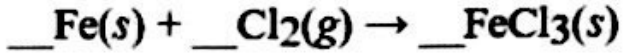
مریم جدو
دعواتکم

حلول مجلد الإعادة /
كیمياء الكویز الثاني

Total questions in exam: 25 | Answered: 0

Question No. 24

What is the coefficient of chlorine gas after balancing the following equation?



- 3
- 2
- 4
- 1

A

Question No. 19

The ionization energy of silicon is lower than the ionization energy of _____.

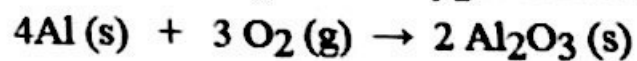
- sodium
- phosphorus
- aluminum
- magnesium

B

Total questions in exam: 25 | Answered: 0

Question No. 22

Solid aluminum and gaseous oxygen react in a combination reaction to produce Al_2O_3



The maximum amount of Al_2O_3 that can be produced from 2.5 g of Al and 2.5 g of O_2 is _____ g.

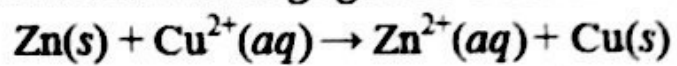
- 4.7
- 7.4
- 5.3
- 9.4

A

Total questions in exam: 25 | Answered: 0

Question No. 21

What substance is the reducing agent in the following redox reaction?



- Zn²⁺
- Cu²⁺
- Zn
- Cu

C

Question No. 20

The chemical formula of the compound is

- Mg_3Cl
- MgCl_2
- Mg_2Cl
- MgCl

Total questions in exam: 25 | Answered: 8

Question No. 3

The ionization energy of beryllium is higher than the ionization energy of _____

- boron
- lithium
- carbon
- nitrogen

B

In which of these substances are the atoms held together by polar covalent bonding?

- H_2
- Cl_2
- O_2
- H_2O

B

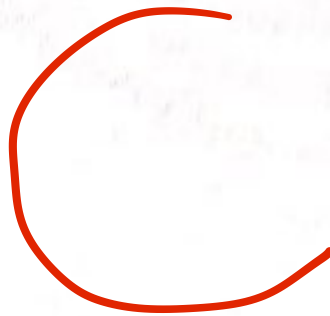
Save & Next حفظ و التالي

Total questions in exam: 25 | Answered: 8

Question No. 8

All of the following elements occurs naturally as diatomic molecule

- chlorine
- fluorine
- neon
- iodine



Total questions in exam: 25 | Answered: 8

Question No. 7

What volume (mL) of a 3.45 M lead nitrate solution must be diluted to make 450 mL of 0.99 M solution of lead nitrate?

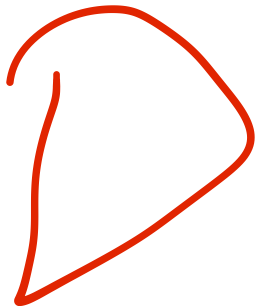
- 129 mL
- 101 mL
- 109 mL
- 56 mL

A

Question No. 6

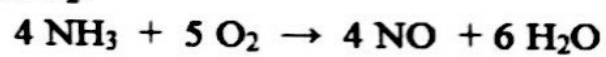
When dissolved in water, _____ can make an aqueous solution that conducts electricity.

- C_8H_{18}
- CH_3OCH_3
- sugar
- ionic salt



Question No. 5

In the reaction below, what is the theoretical yield in moles for NO when 3 moles of NH₃ react with 3 moles of O₂?



- 2.4 mol
- 2.8 mol
- 2.6 mol
- 3.0 mol

A

The type of bond in I_2 is a (an)bond

- Metallic
- Polar covalent
- Nonpolar covalent
- Ionic



Question No. 19

The molarity (M) of an aqueous solution containing 22.5 g of sucrose ($C_{12}H_{22}O_{11}$) in 35.5 mL of solution is _____.

- 0.0657
- 3.52
- 1.85
- 0.104

C

Total questions in exam: 25 | Answered: 8

Question No. 3

The ionization energy of beryllium is higher than the ionization energy of _____

- boron
- lithium
- carbon
- nitrogen

B

Question No. 17

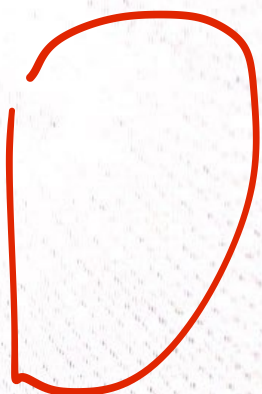
Use the periodic table to answer the following question:
The formula CCl_4 has a molar mass of ___ g/mol.

- 140
- 150
- 146
- 154

D

The most correct name for the compound NI_3 is:

- mononitrogen triiodide
- nitrogen iodide
- triiodo nitrogen
- nitrogen triiodide



Total questions in exam: 25 | Answered: 0

Question No. 13

What is the final molarity of HNO_3 solution, if 300 mL of 2M HNO_3 was diluted to a final volume of 0.2 L?

- 3.0 M
- 5.0 M
- 4.0 M
- 6.0 M

A

Total questions in exam: 25 | Answered: 9

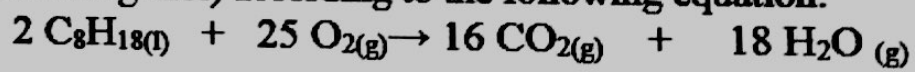
Question No. 23

Based on Lewis dot structure, the number of shared electrons in N

- 2
- 6
- 4
- 8

Question No. 24

The number of CO₂ molecules that are produced from burning of 57.11 g of C₈H₁₈ (Molar mass = 114.22 g/mol) according to the following equation:



- 8 molecules.
- 16 molecules.
- 2.41×10^{24} molecules.
- 6.02×10^{23} molecules.

Total questions in exam: 25 | Answered: 9

Question No. 25

The correct name for the acid HBr is _____

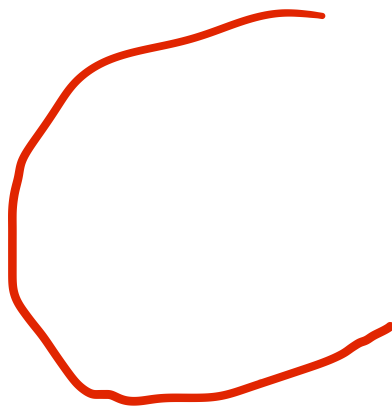
- hydrobromic
- hydrogen bromide
- hydrogen bromate
- hydrogen bromine

A

Question No. 22

How many moles and how many atoms of zinc (Zn) are in a sample weighing 34.9 g?

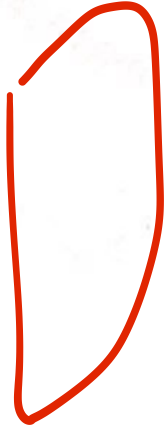
- 0.533 mol, 8.85×10^{-25} atoms
- 1.87 mol, 3.10×10^{-24} atoms
- 0.533 mol, 3.21×10^{23} atoms
- 1.87 mol, 1.13×10^{24} atoms



Question No. 20

What is the molecular formula of a compound that has a molar mass of 300 g/mol and its empirical formula is CH_2O ?

- $\text{C}_2\text{H}_4\text{O}_2$
- $\text{C}_4\text{H}_8\text{O}_4$
- $\text{C}_3\text{H}_6\text{O}_3$
- $\text{C}_{10}\text{H}_{20}\text{O}_{10}$



Question No. 22

Which of these substances gives a weak electrolyte when dissolved in water?

- weak base
- ionic salt
- strong acid
- strong base

A

Total questions in exam: 25 | Answered: 21

Question No. 14

Find out the molecular formula of a compound that has a molar mass of 138.0 g/mol and an empirical formula of NO_2 .

- N_3O_6
- NO_2
- N_2O_3
- N_2O_4

A

Question No. 20

The name of the chemical compound FeCO_3 is:

- iron carbonate
- iron(II) carbonate
- iron(III) carbonate
- iron(I) carbonate

B

The name of the chemical compound FeCO_3 is:

- iron carbonate
- iron(II) carbonate
- iron(III) carbonate
- iron(I) carbonate

B

Question No. 22

Which of these substances gives a weak electrolyte when dissolved in water?

- weak base
- ionic salt
- strong acid
- strong base

A

Find out the molecular formula of a compound that has a molar mass of 138.0 g/mol and an empirical formula of NO_2 .

- N_3O_6
- NO_2
- N_2O_3
- N_2O_4

A

Question No. 15

The molar mass of $\text{Ca}_3(\text{PO}_4)_2$ is equal to:

- 310 g/mol
- 250 g/mol
- 134 g/mol
- 215 g/mol

A

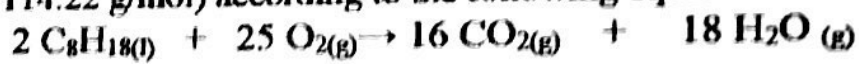
Question No. 19

Calculate the molar mass of $\text{Fe}_3(\text{PO}_4)_2$.

- 262.5 g/mol
- 525.1 g/mol
- 237.6 g/mol
- 357.5 g/mol

D

The number of CO_2 molecules that are produced from burning of 57.11 g of C_8H_{18} (Molar mass = 114.22 g/mol) according to the following equation:



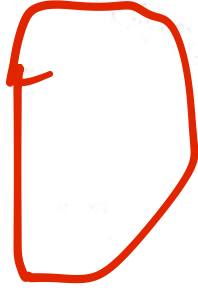
- 2.41×10^{24} molecules.
- 6.02×10^{23} molecules.
- 8 molecules.
- 16 molecules.

A

Question No. 7

The substance that causes the oxidation of another substance is called

- cathode
- reducing agent
- anode
- oxidizing agent



لأنه مفروض اختار اختزال و هو موجود
بالخيارى فالعامل المؤكسد هو نفسه اللي صار له
اختزال

Question NO. 8

For a given reaction



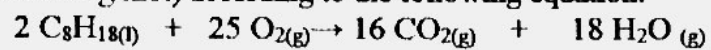
When 10 moles of V_2O_5 are mixed with 10 moles of Ca, which is the limiting reactant according to the above equation?

- V_2O_5
- CaO
- Ca
- V



Question No. 11

The number of CO₂ molecules that are produced from burning of 57.11 g of C₈H₁₈ (Molar mass = 114.22 g/mol) according to the following equation:



- 2.41 x 10²⁴ molecules.
- 6.02 x 10²³ molecules.
- 8 molecules.
- 16 molecules.

A

As you move from the top to the bottom of the periodic table:

- ionization energy increases and atomic radius decreases
- ionization energy decreases and atomic radius decreases
- ionization energy increases and atomic radius increases
- ionization energy decreases and atomic radius increases

D

Question No. 7

The substance that causes the oxidation of another substance is called

- cathode
- reducing agent
- anode
- oxidizing agent

D

Question No. 2

Write the name for FeS

- iron(I) sulfide
- iron(II) sulfate
- iron(I) sulfate
- iron(II) sulfide

D

Question No. 7

The substance that causes the oxidation of another substance is called

- cathode
- reducing agent
- anode
- oxidizing agent



Total questions in exam: 25 | Answered: 9

Question No. 25

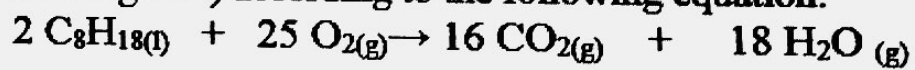
The correct name for the acid HBr is _____

- hydrobromic
- hydrogen bromide
- hydrogen bromate
- hydrogen bromine

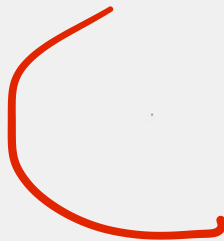
A

Question No. 24

The number of CO₂ molecules that are produced from burning of 57.11 g of C₈H₁₈ (Molar mass = 114.22 g/mol) according to the following equation:



- 8 molecules.
- 16 molecules.
- 2.41×10^{24} molecules.
- 6.02×10^{23} molecules.

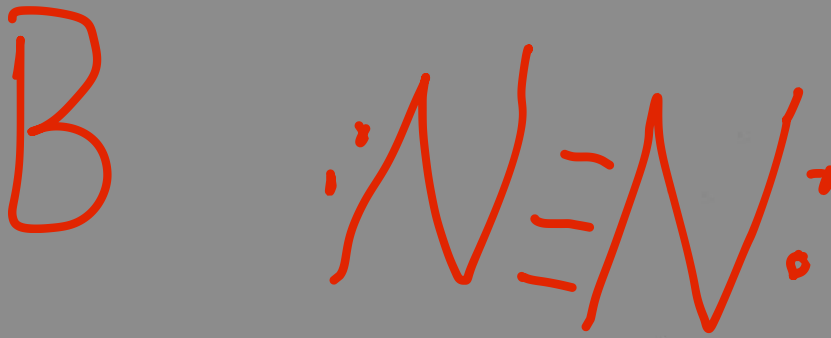


Total questions in exam: 25 | Answered: 9

Question No. 23

Based on Lewis dot structure, the number of shared electrons in N_2 molecule is _____

- 2
- 6
- 4
- 8



Question No. 19

The molarity (M) of an aqueous solution containing 22.5 g of sucrose ($C_{12}H_{22}O_{11}$) in 35.5 mL of solution is _____.

- 0.0657
- 3.52
- 1.85
- 0.104

Question No. 20

What is the molecular formula of a compound that has a molar mass of 300 g/mol and its empirical formula is CH_2O ?

- $\text{C}_2\text{H}_4\text{O}_2$
- $\text{C}_4\text{H}_8\text{O}_4$
- $\text{C}_3\text{H}_6\text{O}_3$
- $\text{C}_{10}\text{H}_{20}\text{O}_{10}$

Question No. 22

How many moles and how many atoms of zinc (Zn) are in a sample weighing 34.9 g?

- 0.533 mol, 8.85×10^{-25} atoms
- 1.87 mol, 3.10×10^{-24} atoms
- 0.533 mol, 3.21×10^{23} atoms
- 1.87 mol, 1.13×10^{24} atoms



The type of bond in I_2 is a (an) bond

- Metallic
- Polar covalent
- Nonpolar covalent
- Ionic



Question No. 17

**Use the periodic table to answer the following question:
The formula CCl_4 has a molar mass of ___ g/mol.**

- 140
- 150
- 146
- 154

The most correct name for the compound NI_3 is:

- mononitrogen triiodide
- nitrogen iodide
- triiodo nitrogen
- nitrogen triiodide

D

Total questions in exam: 25 | Answered: 8

Question No. 7

What volume (mL) of a 3.45 M lead nitrate solution must be diluted to make 450 mL of 0.99 M solution of lead nitrate?

- 129 mL
- 101 mL
- 109 mL
- 56 mL

A

Total questions in exam: 25 | Answered: 8

Question No. 8

All of the following elements occurs naturally as diatomic molecule

- chlorine
- fluorine
- neon
- iodine



Total questions in exam: 25 | Answered: 0

Question No. 13

What is the final molarity of HNO_3 solution, if 300 mL of 2M HNO_3 was diluted to a final volume of 0.2 L?

- 3.0 M
- 5.0 M
- 4.0 M
- 6.0 M

Question No. 6

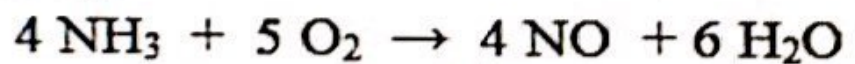
When dissolved in water, _____ can make an aqueous solution that conducts electricity.

- C_8H_{18}
- CH_3OCH_3
- sugar
- ionic salt

D

Question No. 5

In the reaction below, what is the theoretical yield in moles for NO w
NH₃ react with 3 moles of O₂?



- 2.4 mol
- 2.8 mol
- 2.6 mol
- 3.0 mol

Total questions in exam: 25 | Answered: 8

Question No. 3

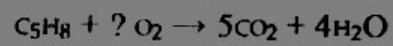
The ionization energy of beryllium is higher than the ionization energy of _____

- boron
- lithium
- carbon
- nitrogen

B

Question No. 6

What coefficient is placed in front of O_2 to complete the balancing of the following equation?



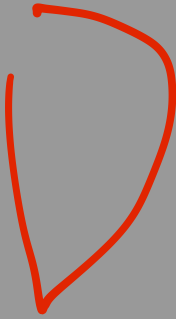
- 7
- 3
- 1
- 5

A

Question No. 5

Household sugar, sucrose, has the molecular formula $C_{12}H_{22}O_{11}$. What is the mass percent of carbon in sucrose?

- 51.4 %
- 62.8 %
- 6.5 %
- 42.1 %



Question No. 4

What is the term for the shared valence electrons in a covalent molecule?

- core electrons
- nonbonding electrons
- bonding electrons
- lone pair electrons



Question No. 1

What is the oxidation number of iron in Fe_2O_3 ?

- +3
- 6
- 3
- +6

A

Save & Next حفظ واقلی

Question No. 2

Write the name for FeS

- iron(I) sulfide
- iron(II) sulfate
- iron(I) sulfate
- iron(II) sulfide

D