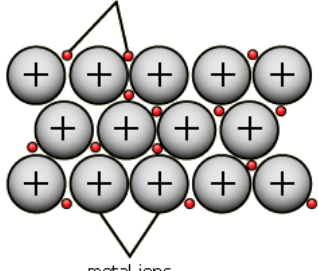
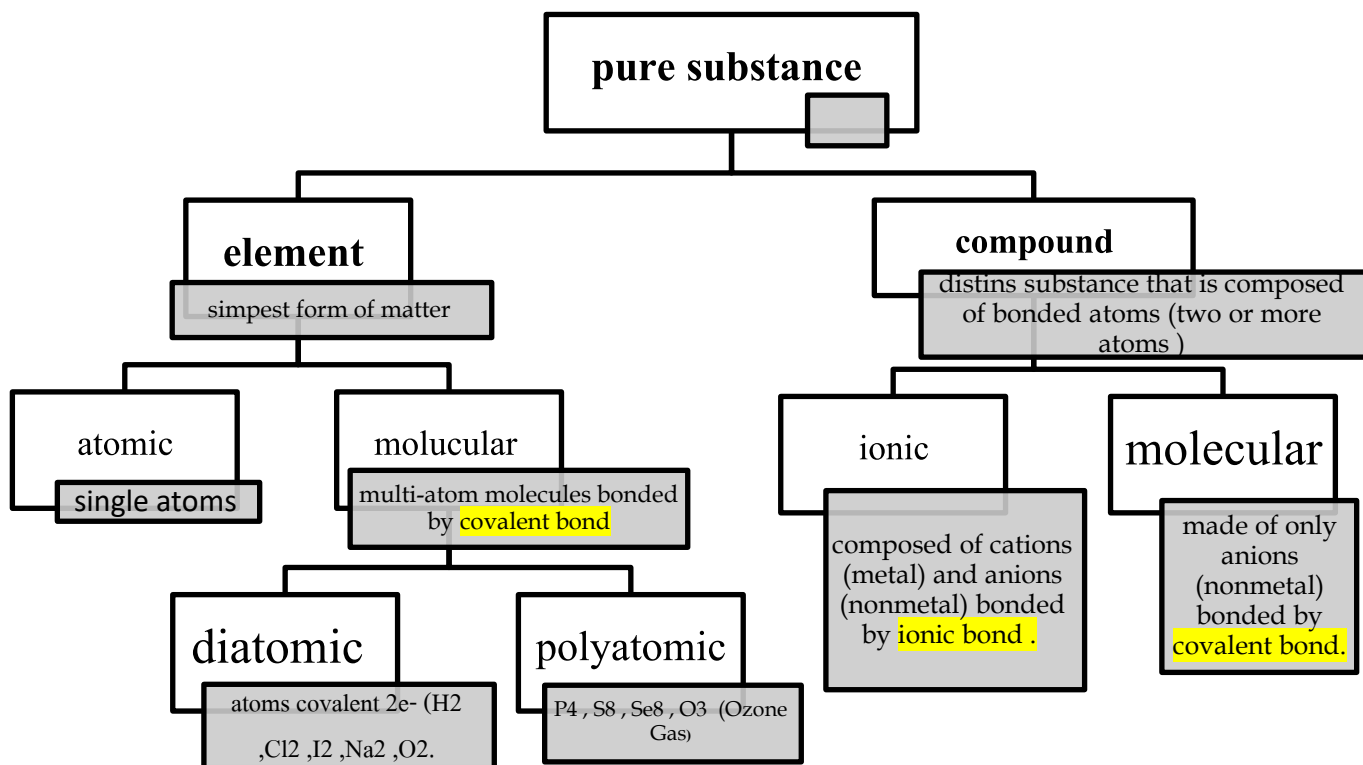
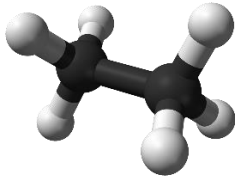


	Ionic bonds	Covalent bonds	Metallic Bond:
Occur between :	Metal + nonmetal Cations + anions (+) charge + (-) charge	Two nonmetal Anions (-) charge	occurs in metals (cations)
Definition Method:	Transferred of (e-)	Sharing of (e-)	electron pooling (electron sea of delocalized electrons)
Ex:-	<p style="text-align: center;">NaCl</p> <pre> graph TD NaCl[NaCl] --- Na[Na: metal] NaCl --- Cl[Cl: nonmetal] Na --- Cation[=cation (+)] Cl --- Anion[=anions (-)] </pre> <p>*e- transferred from Na to Cl</p>	<p>single bonds (C-H)</p> <ul style="list-style-type: none"> • 2e- = 1 pair <p>double bonds (C=O)</p> <ul style="list-style-type: none"> • 4e- = 2 pairs <p>triple bonds (C≡N)</p> <ul style="list-style-type: none"> • 6e- = 3 pairs <p>triple bond (≡) is short and strong, while single bond (—) is long and weak.</p>	<p>free electrons from outer shells of metal atoms</p>  <p>metals tend to lose electrons easily because they have low ionization</p>



Compound : represented

			Molecular Compounds:	
			used to represent the molecules of compounds	
Molecular Formula:	Empirical Formula:	Structural Formula:	Ball-and-Stick Model	Space-Filling Model
gives the actual number of atoms of each element in a molecule of a compound. Ethane C ₂ H ₆ MOLECULAR FORMULA : IS C₂H₆	gives the relative number (simplest whole-number) of atoms . *The ionic compounds are usually represented using their Empirical Formulas Ethane C ₂ H ₆ EMPIRICAL FORMULA : IS CH₃	a sketch or diagram of how the atoms are bonded to each other. It uses lines to represent covalent bonds ("Molecular Compounds") <div style="text-align: center;"> $\begin{array}{c} \text{H} & & \text{H} \\ & & \\ \text{H}-\text{C} & - & \text{C}-\text{H} \\ & & \\ \text{H} & & \text{H} \\ \text{ETHANE} \end{array}$ </div>	Ethane 	Ethane 