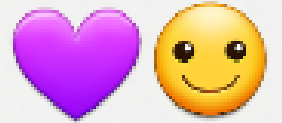


## تجميعات كيمياء 2016

معظم الاسئلة اللي جات باختبارات الكيمياء كويز 1  
وكويز 2 صح التجميع مو مرتب بس ان شاء الله تستفيدوا



Please choose the BEST answer from the given options

**Question:**

What is the common name for  $\text{HC}\equiv\text{CH}$ ?

**Options:**

- propyne
- ethene
- acetylene
- ethylene

- 24) Which statement below is true?  
 A) Metals gain electrons to have a positive charge. B) Metals gain electrons to have a negative charge.  
 C) Metals lose electrons to have a positive charge. D) Nonmetals lose electrons.
- 25) The ground state electron configuration of phosphorus is \_\_\_\_\_.  
 A)  $1s^2 2s^2 2p^4 3p^6 3d^1$  C)  $1s^2 2s^2 2p^6 3s^2 3p^3$   
 B)  $1s^2 2s^2 2p^6 3s^4 3p^1$  D)  $1s^2 2s^2 2p^6 3s^2 2d^3$
- 26) 57.7 g Ni contains how many atoms?  
 A)  $6.13 \times 10^{23}$  atoms C)  $3.47 \times 10^{23}$  atoms  
 B)  $5.92 \times 10^{23}$  atoms D)  $1.24 \times 10^{24}$  atoms
- 27) Which of the following two atoms are isotopes?  
 A)  $^{40}_{18}\text{Ar}$  and  $^{40}_{20}\text{Ca}$  C)  $^{12}_6\text{C}$  and  $^{13}_6\text{C}$   
 B)  $^{35}_{17}\text{Cl}$  and  $^{80}_{35}\text{Br}$  D)  $^{24}_{12}\text{Mg}$  and  $^{12}_6\text{C}$
- 28) Which is not an ionic compound?  
 A) KCl B)  $\text{Ca}(\text{NO}_3)_2$  C)  $\text{SO}_3$  D)  $\text{AlBr}_3$
- 29) Covalent bond formed .....  
 A) when an electrically positive species interacts with an electrically negative species  
 B) when two nonmetallic elements interact to form a compound  
 C) when two electrically negative species interact  
 D) when two electrically positive species interact
- 30) A bond where the electrons are shared unequally is called \_\_\_\_\_.  
 A) polar covalent B) metallic C) nonpolar covalent D) ionic

*Good Luck*



Form B

Chem 101 Unit 2 1<sup>st</sup> Semester 2015 - 1497  
Select the correct single choice of the following MCQs, then mark the proper square in the answer sheet.

- 1) The electron-dot symbol for aluminum is \_\_\_\_\_ D) Al<sup>+</sup>  
 A)  $\cdot\text{Al}\cdot$  B)  $\cdot\text{Al}$  C)  $\cdot\text{Al}$
- 2) Which of the following is the conjugate acid of water? D) H<sub>3</sub>O<sup>+</sup>  
 A) HCl B) NH<sub>4</sub><sup>+</sup> C) OH<sup>-</sup>
- 3) Which of the following is the correct formula of the compound formed between the calcium ion and the sulfate ion? D) CaSO<sub>4</sub>  
 A) Ca(SO<sub>4</sub>)<sub>2</sub> B) CuSO<sub>4</sub> C) Ca(SO<sub>4</sub>)<sub>2</sub>
- 4) What is the correct name for Cl<sub>2</sub>O<sub>7</sub>? D) dichlorine heptoxide  
 A) chloric acid B) chlorine oxide C) dichloride oxide
- 5) Calculate the percentage composition of oxygen in Ca(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub>. D) 46.39%  
 A) 18.51% B) 40.50% C) 29.74%
- 6) What coefficient is placed in front of O<sub>2</sub> to complete the balancing of the following equation?  
 $\text{C}_5\text{H}_8 + ? \text{O}_2 \rightarrow 5 \text{CO}_2 + 4 \text{H}_2\text{O}$   
 A) 1 B) 3 C) 5 D) 7
- 7) Given the following reaction:  $2 \text{Fe}(s) + 3 \text{Cl}_2(g) \rightarrow 2 \text{FeCl}_3(s)$  How many moles of FeCl<sub>3</sub> are obtained when 4.6 mol of Cl<sub>2</sub> reacts with excess Fe?  
 A) 3.07 mol B) 4.6 mol C) 1.5 mol D) 2.3 mol
- 8) How many moles of solute are contained in 500 mL of 0.50 M KCl?  
 A) 0.25 B) 0.50 C) 1.0 D) 2.0
- 9) Consider the following chemical system at equilibrium.  
 $\text{Heat} + 6 \text{H}_2\text{O}(g) + 2 \text{N}_2(g) \rightleftharpoons 4 \text{NH}_3(g) + 3 \text{O}_2(g)$   
 Which of the following stresses would shift the equilibrium to the left?  
 A) increasing the concentration of O<sub>2</sub> C) increasing the reaction temperature  
 B) increasing the concentration of H<sub>2</sub>O D) decreasing the concentration of NH<sub>3</sub>
- 10) Given the following reaction, the equilibrium expression will be:  
 $4 \text{CuO}(s) + \text{CH}_4(g) \rightleftharpoons \text{CO}_2(g) + 4 \text{Cu}(s) + 2 \text{H}_2\text{O}(g)$   
 A)  $[\text{CH}_4]/[\text{CO}_2][\text{H}_2\text{O}]^2$  C)  $[\text{CuO}]^4/[\text{Cu}]^4$   
 B)  $[\text{Cu}]^4/[\text{CuO}]^4$  D)  $[\text{CO}_2][\text{H}_2\text{O}]^2/[\text{CH}_4]$
- 11) How many lone pairs of electrons are in this Lewis structure shown here?  
 $\text{O}::\text{C}::\text{O}$   
 A) 2 B) 4 C) 6 D) 8

- 12) A Bronsted-Lowry acid is defined as a substance that \_\_\_\_\_.
- A) increases the  $[H^+]$  concentration when placed in water
  - B) decreases the  $[H^+]$  concentration when placed in water
  - C) acts as a proton donor
  - D) acts as a proton acceptor
- 13) Which of the following is a weak acid?
- A) HF
  - B)  $H_2SO_4$
  - C) HBr
  - D)  $HNO_3$
- 14) What is the  $[H_3O^+]$  in a solution with  $[OH^-] = 1.0 \times 10^{-12} M$ ?
- A)  $1.0 \times 10^{-12} M$
  - B)  $1.0 \times 10^{-8} M$
  - C)  $1.0 \times 10^2 M$
  - D)  $1.0 \times 10^{-2} M$
- 15) The empirical formula of a compound is  $CH_2$ , and its molar mass is about 42.g. What is its molecular formula?
- A)  $CH_2$
  - B)  $C_2H_4$
  - C)  $C_3H_6$
  - D)  $C_4H_8$
- 16) What is the mass in gram of NaCl solution in a 0.500 L bottle of 2.00 M NaCl?
- A) 29.3 g
  - B) 58.5 g
  - C) 117 g
  - D) 234 g
- 17) The substance NaOH in water solution is a:
- A) nonelectrolyte
  - B) weak electrolyte
  - C) weak base
  - D) strong electrolyte
- 18) In the balanced redox reaction:  $2 C_2H_6(g) + 7 O_2(g) \rightarrow 4 CO_2(g) + 6 H_2O(g)$ , which species is reduced?
- A)  $C_2H_6(g)$
  - B)  $O_2(g)$
  - C)  $CO_2(g)$
  - D)  $H_2O(g)$
- 19) When an acid reacts with a base, the reaction is called which of the following?
- A) neutralization
  - B) precipitation
  - C) redox
  - D) single displacement
- 20) Under which of the following conditions is the A-B bond considered to be ionic?
- A) when A and B have the same electronegativity
  - B) when the electronegativity difference between the two atoms is 1.0
  - C) when the electronegativity difference between the two atoms is 1.5
  - D) when the electronegativity difference between the two atoms is  $\gg 2.0$
- 21) Which element is not an alkali metal?
- A) Li
  - B) K
  - C) Rb
  - D) Kr
- 22) Which element has the smallest ionization energy?
- A) K
  - B) Ga
  - C) Br
  - D) Ca
- 23) How many neutrons are in the isotope  ${}_{92}^{238}U$ ?
- A) 238
  - B) 146
  - C) 92
  - D) 330

12/17/2010

تعليق

أعجبني

اكتب تعليقا...



Chem 101 Quiz-2 1<sup>st</sup> Semester 2018 - 1437

- 24) Which statement below is true?  
 A) Metals gain electrons to have a positive charge. B) Metals gain  
 C) Metals lose electrons to have a positive charge. D) Nonmetals
- 25) The ground state electron configuration of phosphorus is  
 A)  $1s^2 2s^2 2p^4 3p^6 3d^1$  C)  $1s^2 2s^2 2p^6 3s^2 3p^3$   
 B)  $1s^2 2s^2 2p^6 3s^4 3p^1$  D)  $1s^2 2s^2 2p^6 3s^2 2d^3$
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 A)  $^{40}_{18}\text{Ar}$  and  $^{40}_{20}\text{Ca}$  C)  $^{12}_6\text{C}$  and  $^{13}_6\text{C}$   
 B)  $^{35}_{17}\text{Cl}$  and  $^{80}_{35}\text{Br}$  D)  $^{24}_{12}\text{Mg}$  and  $^{12}_6\text{C}$
- 28) Which is not an ionic compound?  
 A) KCl B)  $\text{Ca}(\text{NO}_3)_2$  C)  $\text{SO}_3$
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 A) when an electrically positive species interacts with an  
 B) when two nonmetallic elements interact to form a com  
 C) when two electrically negative species interact  
 D) when two electrically positive species interact
- 30) A bond where the electrons are shared unequally is calle  
 A) polar covalent B) metallic C)

21) Which element is not an alkali metal?

تعليق

أعجبني

اكتب تعليقا...



20) Given the following



- A)  $[\text{CH}_4]/[\text{CO}_2][\text{H}_2\text{O}]^2$
- B)  $[\text{Cu}]^4/[\text{CuO}]^4$

- C)  $[\text{CuO}]^4/[\text{CO}_2]$
- D)  $[\text{CO}_2][\text{H}_2\text{O}]^2$

21) Which element is not an alkali metal?

A) Li

B) K

C) Rb

22) Which element has the smallest ionization energy?

A) K

B) Ga

C) Br

23) How many neutrons are in the isotope  ${}^{238}_{92}\text{U}$ ?

A) 238

B) 146

C) 9

3

تعليق

أعجبني

اكتب تعليقا...



CHEM 101 Quiz-2 1<sup>st</sup> Semester 2018 - 1437

24) Which statement below is true?

- A) Metals gain electrons to have a positive charge.
- B) Metals gain electrons to have a negative charge.
- C) Metals lose electrons to have a positive charge.
- D) Nonmetals gain electrons to have a negative charge.

25) The ground state electron configuration of phosphorus is

A)  $1s^2 2s^2 2p^4 3p^6 3d^1$

C)  $1s^2 2s^2 2p^6 3s^2 3p^3$

اكتب تعليقاً...



A) chloric acid

**B) chlorine oxide**

C) dichloride of

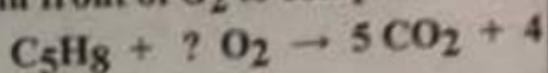
15) Calculate the percentage composition of oxygen in  $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$

A) 18.51%

B) 40.50%

C) 29.74%

16) What coefficient is placed in front of  $\text{O}_2$  to complete the balanced equation?



A) 1

B) 3

C) 5

17) Given the following reaction:  $2 \text{Fe}(\text{s}) + 3 \text{Cl}_2(\text{g}) \rightarrow 2 \text{FeCl}_3$   
 How many moles of  $\text{FeCl}_3$  are obtained when 4.6 mol of  $\text{Cl}_2$  reacts with excess  $\text{Fe}$ ?

A) 3.07 mol

B) 4.6 mol

C) 1.5 mol

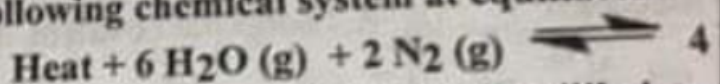
18) How many moles of solute are contained in 500 mL of 0.5 M  $\text{NaCl}$  solution?

A) 0.25

B) 0.50

C) 1.0

19) Consider the following chemical system at equilibrium.

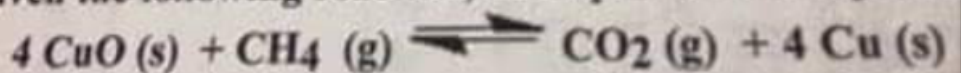


Which of the following stresses would shift the equilibrium to the right?

A) increasing the concentration of  $\text{O}_2$

B) increasing the concentration of  $\text{H}_2\text{O}$

20) Given the following reaction, the equilibrium expression is:



A)  $[\text{CH}_4]/[\text{CO}_2][\text{H}_2\text{O}]^2$

B)  $[\text{Cu}]^4/[\text{CuO}]^4$

C)  $[\text{Cu}]^4/[\text{CO}_2]$

D)  $[\text{CO}_2]/[\text{CH}_4]$

تم النسخ إلى الحافظة.

12/16/2015

تعليق

أعجبني



اكتب تعليقاً...



- 3) Which of the following is a weak acid?  
A) HF  
B) H<sub>2</sub>SO<sub>4</sub>  
C) HBr
- 4) What is the [H<sub>3</sub>O<sup>+</sup>] in a solution with [OH<sup>-</sup>] = 1.0 × 10<sup>-12</sup> M?  
A) 1.0 × 10<sup>-12</sup> M  
B) 1.0 × 10<sup>-8</sup> M  
C) 1.0 × 10<sup>2</sup> M  
D) 1.0 × 10<sup>-2</sup> M
- 5) The empirical formula of a compound is CH<sub>2</sub>, and its molar mass is 42 g/mol. What is its molecular formula?  
A) CH<sub>2</sub>  
B) C<sub>2</sub>H<sub>4</sub>  
C) C<sub>3</sub>H<sub>6</sub>
- 6) What is the mass in gram of NaCl solution in a 0.500 L bottle?  
A) 29.3 g  
B) 58.5 g  
C) 117 g
- 7) The substance NaOH in water solution is a:  
A) nonelectrolyte  
B) weak electrolyte  
C) weak electrolyte
- 8) In the balanced redox reaction: 2 C<sub>2</sub>H<sub>6</sub>(g) + 7 O<sub>2</sub>(g) → 4 CO<sub>2</sub>(g) + 6 H<sub>2</sub>O(l)  
Which species is reduced?  
A) C<sub>2</sub>H<sub>6</sub>(g)  
B) O<sub>2</sub>(g)  
C) CO<sub>2</sub>(g)
- 9) When an acid reacts with a base, the reaction is called which type of reaction?  
A) neutralization  
B) precipitation  
C) redox
- 10) Under which of the following conditions is the A-B bond most ionic?  
A) when A and B have the same electronegativity  
B) when the electronegativity difference between the two atoms is small  
C) when the electronegativity difference between the two atoms is large

ملف طباعة البريد الإلكتروني نسخ على الفرض المتصفحة

Windows - عرض الصور في Windows

Question:

In the following reaction, what is the effect of adding more  $\text{NO}_2$  to the starting reaction mixture?

$$2\text{NO}_2(\text{g}) \rightleftharpoons \text{N}_2\text{O}_4(\text{g})$$

Options:

- It would decrease the final quantity of products.
- It would increase the final quantity of products.
- It would make the reaction more exothermic.
- It would make the reaction more endothermic.

١٢/١٦/٢٠١٥

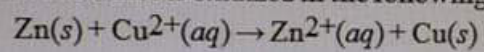
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 C) when two electrically negative species interact  
 D) when two electrically positive species interact
- 30) A bond where the electrons are shared unequally is called  
 A) polar covalent B) metallic C) ionic

12/17/2010

**INSTRUCTION:** تعليمات Please choose the BEST answer from the given options for each question

**Question:**

What substance is oxidized in the following redox reaction?



**Options:**

- Zn
- $\text{Zn}^{2+}$
- $\text{Cu}^{2+}$
- Cu



**INSTRUCTION:** تعليمات Please choose the BEST answer from the given options for each question.

**Question:**

What is the term for a family of unsaturated hydrocarbon compounds having a double bond?

**Options:**

- alkynes
- aromatics
- alkanes
- alkenes

١٢/١٦/٢٠١٥

تم النسخ إلى الحافظة.

**Question:**

The systematic name for the chemical compound  $\text{RbCl}$  is:

**Options:**

- rubidium monochloride
- rubidium(I) chloride
- rubidium(II) chloride
- rubidium chloride

12/17/2010



**Question:**

The molar mass of  $\text{H}_2\text{SO}_4$  is equal to:

**Options:**

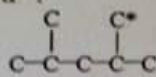
- 108 g/mol
- 98 g/mol
- 88 g/mol
- 77 g/mol

١٢/١٦/٢٠١٥

**INSTRUCTION:** **تعليمات** Please choose the BEST answer from the given options for each question.

**Question:**

The carbon skeleton of an alkane is shown below. How many hydrogen atoms are bonded to the carbon marked with a \*?



**Options:**

- 3
- 2
- 0
- 1

١٢/١٦/٢٠١٥





**Question:**

What is the common name for  $\text{HC}\equiv\text{CH}$ ?

**Options:**

- propyne
- ethene
- acetylene
- ethylene

١٢/١٦/٢٠١٥



**Question:**

What is the empirical formula of the compound that has a composition by mass of 84.2% C and 15.8% H?

**Options:**

- C<sub>4</sub>H<sub>9</sub>
- C<sub>3</sub>H<sub>8</sub>
- C<sub>4</sub>H<sub>10</sub>
- C<sub>3</sub>H<sub>9</sub>

١٢/١٦/٢٠١٥



**Question:**

Which of the following is not a strong acid?

**Options:**

- $\text{HNO}_3$
- $\text{HCl (aq)}$
- $\text{H}_2\text{SO}_4$
- $\text{HF}$



**INSTRUCTION:** **تعليمات** Please choose the BEST answer from the given options for each question.

**Question:**

What is the term for a family of unsaturated hydrocarbon compounds having a double bond?

**Options:**

- alkynes
- aromatics
- alkanes
- alkenes

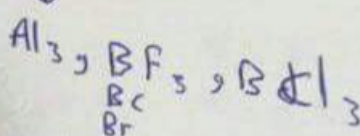
١٢/١٦/٢٠١٥

تم النسخ إلى الحافظة.

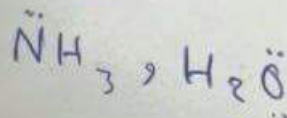
Which of the following is a Lewis acid?

- A)  $\text{AlF}_3$
- B)  $\text{H}_2\text{O}$
- C)  $\text{SiF}_4$
- D)  $\text{C}_5\text{H}_{12}$

صيف لوييس



Lewis base





Select the correct single choice of the following MCQS and Mark with Circle, then mark the proper square in the answer sheet.

1. one the following elements has the gretest electronegativity :  
A) Beryllium  
B) lithium.  
C) Flurine.  
D) Crbon.
2. To prepare 1 L of 1M HCl solution how many Lof 2M HCl solution should be used ?  
A) 0.5L  
B) 0L  
C) 3.5L.  
D) 2.5L.
3. Arrhenius acid is the substance that ?  
A) gives  $H^+$  in queous solution  
B) accepts  $OH^-$   
C) gives  $OH^-$  in queous solution  
D) donates an electron pair
4. In a particular experiment ,if the theoretical yield of the reaction of is 4.0 g and the rea produced 2.0 g .The yield % would be ?  
A) 74%  
B) 37%  
C) 50%  
D) 66%

١٢/١٦/٢٠١٥

5. The oxidation number of chlorine in HCl is ?  
A) -1  
B) -3  
C) +2  
D) +1
6. In the reaction  $4\text{Fe} + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3(\text{s})$ , which species is oxidized ?  
A) Fe.  
B)  $\text{O}_2$ .  
C)  $\text{Fe}^{+3}$ .  
D)  $\text{O}^{-2}$ .
7. Consider the following chemical equation :  $\text{NH}_4^+ + \text{OH}^- \leftrightarrow \text{H}_2\text{O} + \text{NH}_3$ , which the statement is correct ?  
A)  $\text{NH}_4^+$  is an acid and  $\text{NH}_3$  is conjugate base  
B)  $\text{NH}_4^+$  is an acid and  $\text{OH}^-$  is conjugate base  
C)  $\text{OH}^-$  is an acid and  $\text{H}_2\text{O}$  is conjugate base  
D)  $\text{NH}_3$  is base and  $\text{H}_2\text{O}$  is conjugate acid
8. The molar mass of  $\text{Ca}(\text{OH})_2$ ?  
A) 70g/mol  
B) 78g/mol  
C) 68g/mol  
D) 74g/mol
9. Solution A is 2 mol in 1 L, solution B is 3mol in 1L. Which solution has a higher molarity?  
A) Solution A.  
B) Solution B.  
C) the molarity of the solution cannot be calculated.  
D) both solutions have equal concentration.
10. For the following reaction, calculate the mass in gram of the product  $\text{Al}_2\text{O}_3$  if 6 g of oxygen is consumed  $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ ? (Al = 27, O = 16)  
A) 0.5g.  
B) 1.1g.  
C) 3.4g  
D) 6.8g.

12/17/2010

**Question:**

How many moles of  $\text{CO}_2$  could be produced when 5 moles of  $\text{C}_2\text{H}_6\text{O}$  completely react with oxygen gas according to the reaction?



**Options:**

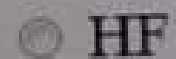
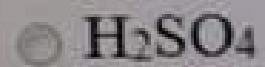
- 10 mol
- 4 mol
- 6 mol
- 8 mol



**Question:**

Which of the following is not a strong acid?

**Options:**



4 اي من التالي له electronegativity اقل وجابوا الهالوجينات  
والإجابة كانت Iodine

5 اي من التالي له electron affinity اقل

1 C 2 N 3 O 4 B

6 اي من التالي غير صحيح عند تقليل التركيز  
الإجابة ان عدد المولات لا يتغير

7 اي من التالي Strong electrolyte  
والإجابات كلها مركبات تساهمية ما عدا LiCl وهي الإجابة

8 تسمية CuO

9 اين تقع Non-metal

The right side

وأغلب الأسئلة معادلات بسيطة جدا ما عدا معادلة وحدة كان شوي تلخبط



وكان معطيك عدد ذرات H ويبقى مولات HCl طبعا تحول عدد الذرات  
إلى مولات بعدين تحولها إلى مولات

1 من صاحب atomic theory

John Dalton

تم النسخ إلى الحافظة.

جاني تحويل من جرام لمول

و العائد النظري

و عدد مولات اذا اضفنا 3 مول من النشادر و 4 مول من ثاني اكسيد الكربون كم هيصير المركب .....(ناسيه ايش  
المركب )

عدد الكترونات في الاكسجين lone pairs

معادلة التخفيف

و المولاريه تعويض مباشر

و ما هو العنصر الذي لديه اعداد اكسدة مختلفه و هوا الحديد

HCl كم lone pair فيها

الجواب 3

التوزيع الالكتروني لعنصر الفلور

لما يزيد تركيز ايون الهيدروجين ايش يصير لل pH؟ ينخفض

ايش قاعدة لويس؟ جاء بين الخيارات BF3 واعتقد هي الصح

جاء اي واحد حمض؟ بين الخيارات ايون الهيدرونيوم واعتقد كمان هو الصح

---

اختبار الدوري الاول كيمياء :

● تعريف الرابطة الأيونية

● اي من الآتي أيون موجب :

Sulfate ion , carbonate ion , hydrogen ion , ammonium ion

● استنادا على قاعدة بلولي أقصى عدد من الإلكترونات في كل orbital اثنين

● تطلعي ال empirical formula

C , 10.4% H , 27.5 % O 62%

● حولي من كيلو جرام الى باوند

● سؤال عن الكثافة ، تطلع الحجم



● حولي من كيلو جرام الى باوند

● سؤال عن الكثافة ، تطلع الحجم

● اذا نزلنا من فوق الى تحت في الجدول الدوري ايش يصير في radius و ionization energy

● تعريف المعادلة الموزونة

● عدد إلكترونات - : 36 br

● اي من العناصر الاتيه metal

● تحويل من C الى F

● يعطيك الكتلة ويبغى عدد الجزئيات : اول شيء التحويل الى مول و بعدين الضرب في عدد افوجادرو

● which one of the following is pure substance : diamond

● تسمية المركب P3O5

وايش شحنه الالكترون اللي هو نيقتيف

اقولكم جاني سؤال هاتي امبركل فورملا للمركب  $C_4H_{10}$  و  
حطيت  $C_2H_{10}$  ها صح ولا طمنوني

---

و الاعلى الكهروسالبية من بين العناصر و تختاري العنصر الاكثر سالبية

و كمان حمض لويس

و كمان نصف قطر الذره

و العالم صاحب قطرة الزيت

و كمان عدد الالكترونات و البروتونات و النيوترونات

ووزن معادله لعنصر الاكسجين فقط

العناصر الانتقاليه

و اين تقع اشباه الفلزات بالجدول الدوري

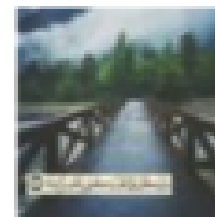
و اي العناصر يعتبر حمض برونستدلوري

2 اي من التالي Lewis Base عشان يلخبطونك  
جابوا  $NH_3$  و  $NH_4$  والإجابة طبعا 3

3 اي من التالي molecular element وهي السبعة عناصر كان  
الوحيد منها الهيدروجين فالأجوبة

atomic element

**Jody Al Har**



هیدوجین

نیتروجین

بروبین

ایودین ...

**Question:**

What is the empirical formula of the compound that has a composition by mass of 84.2% C and 15.8% H?

**Options:**

- C<sub>4</sub>H<sub>9</sub>
- C<sub>3</sub>H<sub>8</sub>
- C<sub>4</sub>H<sub>10</sub>
- C<sub>3</sub>H<sub>9</sub>

**Question:**

The molar mass of  $\text{H}_2\text{SO}_4$  is equal to:

**Options:**

- 108 g/mol
- 98 g/mol
- 88 g/mol
- 77 g/mol



موقع النن ميتل بالجدول  
جاب C-14 كم عدد النيترونات  
احول من متر الى انش  
احول من فهرنيات الى سيليز  
الصوديوم كم شحنته  
كم عدد الذرات في واحد مول  
ايت واحد مادة نقيه الثلج ولا ماء البحر و  
berss والهواء  
ايت واحد رمز العنصر مو صح حيكون  
البوتاسيوم  
سم العنصر يود 2 فلور 6  
ايت واحد اسمه CO<sub>2</sub> صح كوبر  
(II) ...نسيت  
سوال ايت واحد الكلور اكبر منه الخيارات

## < Notes

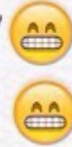
مسائل الفصل الثالث جاني على حل حاجه  
الصيفه الجزئيه والبريتكال ونسبة عنصر من  
مركب كل حاجه والتسميات تقريبا جاء على كل  
حاجه

CuCl<sub>2</sub>

حيسير 11) cabor( Clorooid

حولي من -١٩٦٥ الى كالفن

٧٧ حاولت احفظلكم والله بس اكلت ونسيت



متجانس وغير متجانس جا عليها بس ناسيه  
احسبي عدد المولات احسبي الكتله الموليه جا  
يقولك شحنه الاكسجين (-٢)  
عنصر التكافؤ حقو Na

انا ماجاني علماء بس صحبتي جاها جا  
لصحتي مسأله من شابترون فيها طرح وكذا  
جاني يقولي اوجدي الكتله ومعطيني الحجم  
والكتافه

وبس هذا اللي اتذكر ربنا يوفقكم  
❤️❤️❤️❤️❤️❤️❤️❤️❤️❤️❤️❤️



## < Notes

١٩ مارس، ٢٠١٥ ١٠:٠٤ ص

تحت. ال٢ صغيره s6Br2

sulphar hexa browid

وال٢ صغيره ال٢(يقول يقول شحنة) Fe2O3

Fe=?? سيكون موجب ثلاثه??

ماده نقيه جانبي اثنين مرا غلط وواحد الالكاس

وشي سويتر اخترت الالماس

وجانبي يقولك عنصر يرتبط بشجنه او امثر اللي

هوا انا اخترت Fe وزن المعادله جانبي شي مرا

تافهه وسهل AlO2-----Al+O3 يقولك ايش

نحط قبل الAlO2

الكلور اكبر من فلور برون ايود

مسائل الفصل الثالث جانبي على كل حاجه

الصيغه الجزئيه والبريتكال ونسبة عنصر من

مركب كل حاجه والتسميات تقريبا جاء على كل

حاجه

CuCl2

حيسير 11( cabor) Clorooid

حولي من -١٩٦٢ الى كالفن

٧٧ حاولت احفظلكم والله بس اكلت ونسيت 🤔



☆ جاني سؤال يطلب weak acid

الي هو  $\text{CH}_3\text{COOH}$  vinegar (خل) .

☆ وجا سؤال اقوى حمض ويكون  $\text{H}_3\text{O}^+$ .

☆ سؤال يبغى عدد المولات ومديني عدد جزيئات  
احولها لمولات واجيب عدد المولات.

15- The atomic number of an element is

- 1) The number of protons in an atom
- 2) Equal to the positive charge of an atom's nucleus
- 3) Equal to the number of electrons in a neutral atom
- 4) All of the above are valid definition

20- A d sublevel can hold a maximum of

- 1) 2 electrons
  - 2) 6 electrons
  - 3) 10 electrons
  - 4) 14 electrons
-

☆  $M_1V_1 = M_2V_2$  المطلوب  $V_2 >$  .

☆  $K_{eq}$  اعطاني معادله ويبغى

☆ اربع عناصر من مجموعه  $2A$  ويبغى اعلى واحد في  
.electronaffinities

☆ صيفه ويبغى empirical formulas

☆ اعطاني معادله ويبغى الوزن

☆  $A+B+heat \rightarrow C+D$ .

ويبغى يروح لليساار

-ينقص تركيز مادة من النواتج نسبت المادة .

-زيادة حرارة

-زيادة تركيز وحده من المتفاعلات .

-نسبت الاخير):

☆ مسئله يبغى percent yield.

☆ كم long pere في 0

☆  $K \gg 1$  is favored

الاجابه Forward reaction ص 127

☆ حمض

☆ يبغى سترونق اسد

☆ جاب خيارات قيم ph ايت واحد حمض

☆ HCl كم long per

Heterogeneous mixtures-مخاليط غير متجانسة:

1- vegetable soup & chicken soup

( شوربة دجاج - شوربة خضار )

2- water and sand ( تراب فّ مويّا )

3- pizza ( بيتزا )

4- fruit salad ( سلطة فواكة )

5- blood ( الدم )

Homogeneous mixtures - مخاليط متجانسة :

1- water and sugar / sugar in tea

( السكر في الماء- السكر في الشاي )

2- air ( الهواء )

3- Mouthwash ( غسول الفم )

4- blood plasma ( عينة دم )

5- abbott of vinegar ( خل )





الملاحظات &gt;

5 مارس، 2017، 3:43 م

دكتورة فتحية راجعت لنا اليوم وذكرت بعض الأشياء:

▲ Boiling point in water is 100c →

property.

▲ When there's any "ability" in the question ( toxicity, Acidity..etc) →

property.

▲ there will be many questions about 1s, 2s, 2p, ...etc.

▲ salt + water is mixture ( NOT pure substance ).

▲ الكترولونات التكافؤ →  $N=7$  ,  $N_2=14$ .

▲ Metal in acid → chemical change.

Question No. 3

Which of the following is an example of a heterogeneous mixture?

- milk
- oil
- soda
- chicken soup

2 والسبب. انو الزيت  
غير متجانس

استناداً لسرعة التفاعل  
الزهراني التفاعلية لدى سرعة التفاعل  
تسبب رطوبة سرعة التفاعل  
تسبب سرعة التفاعل  
تسبب سرعة التفاعل  
تسبب سرعة التفاعل

EN

**Question No. 5**

Salt water is classified as

- pure substance-compound
- mixture-heterogeneous
- mixture-homogeneous
- pure substance-element

Save & Next حفظ والتالي

## Question No. 9

---

Which of the following is a chemical change?

- burning natural gas
- melting ice
- cutting paper
- cutting a tomato

Save & Next حفظ والتالي

### Question No. 8

---

Which of the following is matter?

- time
- ice cream
- sunlight
- sound

Save & Next >

**Question No. 7**

Which of the following is a property of both solids and liquids?

- definite shape
- indefinite shape
- indefinite volume
- definite volume

Save & Next حفظ والذلي

### Question No. 6

---

A decimeter equals:

- 1 m
- 1000 mm
- 10 m
- 10 cm

Question No. 12

The charge of electron was established by

- Millikan's oil drop experiment.
- Rutherford's gold foil experiment.
- Dalton's atomic experiment
- Thomson's cathode ray tube experiment.



### Question No. 23

Which of the following elements is a metal?

- nitrogen
- fluorine
- strontium
- argon

Save & Next حفظ و التالي

**Question No. 5**

A solid form of matter in which there is long range repeating order is called \_\_\_\_\_

- amorphous
- fixed
- crystalline
- rigid

Question No. 4

Which of the following sets is a mixture?

- a. table salt
- b. cereal and milk
- c. baking soda
- d. sugar

### Question No. 16

The elements xenon, helium and argon

- are isotopes of each other.
- are in the same period of elements.
- have the same number of neutrons.
- are in the same group.

Save & Next حفظ والتالي

**Question No. 3**

Which one of the following is homogeneous mixture?

- Alloys.
- Noodles soup.
- Vegetable salad.
- Sand in water.

Question No. 5

Salt water is classified as

- pure substance-element
- mixture-homogeneous
- pure substance-compound
- mixture-heterogeneous

## Question No. 25

The atomic radius of krypton is smaller than the atomic radius of \_\_\_\_\_

- xenon
- neon
- helium
- argon

Save & Next حفظ والتالي

## Question No. 21

0.100 mole of lithium weighs

- 6.94 g.
- 0.694 g.
- 0.300 g.
- 3.00 g.

Save & Next حفظ و التالي



Question No. 23

---

Where is most of the mass of an atom concentrated?

- nucleus
- neutrons
- electrons
- protons



Question No. 13

When filling orbitals that have same energy with electrons, fill them singly first, with parallel spins, is known as

- Aufbau principle.
- Heisenberg uncertainty principle.
- Pauli exclusion principle.
- Hund's rule.

Save & Next حفظ والتالي



Question No. 22

The molecular formula of aspirin is  $C_9H_8O_4$ . How many aspirin molecules are present in one 500-milligram tablet?

- $1.67 \times 10^{24}$  molecules
- $1.67 \times 10^{21}$  molecules
- 2.77 molecules
- $2.77 \times 10^{-3}$  molecules

Save & Next حفظ والتالي

## Question No. 22

The molecular formula of aspirin is  $C_9H_8O_4$ . How many aspirin molecules are present in one 500-milligram tablet?

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- $1.67 \times 10^{21}$  molecules
- 2.77 molecules
- $2.77 \times 10^{-3}$  molecules

Save & Next حفظ و التالي

**Question No. 24**

Which of these elements has the highest electron affinity?



- Rubidium
- Cesium
- Lithium
- Potassium



Question No. 13

When filling orbitals that have same energy with electrons, fill them singly first, with parallel spins, is known as

- Aufbau principle.
- Heisenberg uncertainty principle.
- Pauli exclusion principle.
- Hund's rule.

## Question No. 13

According to Hund's Rule

- no two electrons in the same atom can have the same four quantum numbers.
- electrons found in the same orbital must have the same spin.
- electrons fill orbitals of the same energy singly and pair up only after the orbitals are all half filled.
- electrons pair first before filling other orbitals with same energy

Save & Next حفظ و التالي

**Question No. 4**

Choose the heterogogeneous mixture from the list below.

- carbon (graphite)
- chlorine gas
- chicken noodle soup
- black coffee



ناسيتها

جاتني ايش الشي اللي مو صح عن التفاعل الكيميائي  
وحاطط لي انو المادة محفوظة والمادة لا تخلق من العيب  
والذرات تتفاعل مع الذرات  
والذرات تتغير مع الذرات  
اخترت تتغير

Question No. 24

Which of the following has the greatest number of nonbonding pairs of electrons?

- He
- F
- H
- C



## Question No. 11

Which statement below accurately describes the contributions of Thomson?

- Ancient Greek philosopher who proposed that matter was continuous
- Proposed the modern Atomic Theory
- Discovered the existence of electrons
- Created the modern periodic table

Save & Next حفظ و التالي

**Question No. 17**

All of the following statements about different elements are true EXCEPT:

- Barium is an alkaline earth metal.
- Krypton is one of the noble gases.
- Manganese is a transition metal.
- Lithium is considered a metalloid.

حفظ والتالى Save & Next

Question No. 25

As you move up and to the right on the periodic table

- atomic radius increases and ionization energy increases
- atomic radius decreases and ionization energy increases
- atomic radius decreases and ionization energy decreases
- atomic radius increases and ionization energy decreases

Question No. 17

All of the following statements about different elements are true EXCEPT:

- Barium is an alkaline earth metal.
- Krypton is one of the noble gases.
- Manganese is a transition metal.
- Lithium is considered a metalloid.

Save & Next حفظ و التالي

**Question No. 18**

The Group 8A (18) elements

- are unreactive.
- are liquids at room temperature.
- are good conductors of electricity.
- melt at high temperatures.

Question No. 3

Which of the following is an example of a heterogeneous mixture?

- milk
- oil
- soda
- chicken soup

Save & Next حفظ والتالي



Question No. 5

How many inches are in 25.8 cm? [ 1 in = 2.54 cm]

- 0.1
- 10.2
- 28.3
- 0.0984

Save & Next حفظ والتالي

### Question No. 21

---

0.100 mole of lithium weighs

- 3.00 g.
- 0.694 g.
- 0.300 g.
- 6.94 g.

Question No. 12

What is the atomic symbol for silicon?

- S
- Si
- Ag
- Au

Save & Next حفظ والتالي

Question No. 1

A solution has a density of 1.22 g/mL. What volume of the solution has a mass of 55.0 g?

- 0.00253 mL
- 49.4 mL
- 58.8 mL
- 45.1 mL

## Question No. 2

A solution has a density of 1.22 g/mL. What volume of the solution has a mass of 48.2 g?

- 0.00253 mL
- 49.4 mL
- 58.8 mL
- 39.5 mL

Save & Next حفظ والتالي

## Question No. 24

Ionization energy is

- the energy an ion acquires from an electron.
- higher for potassium than for lithium.
- the energy needed to remove the least tightly bound electron.
- highest for metals in Group 1A(1).

Question No. 7

The state of matter that has fixed volume and unfixed shape is called?

- liquid
- plasma
- solid
- gas

Save & Next حفظ و التالي

**Question No. 4**

Which of the following items is a mixture?

- table salt
- cereal and milk
- baking soda
- sulfur

Save & Next حفظ التالي



## Question No. 10

What is the value of 335 K on the Celsius temperature scale?

- 62
- 66.4
- 167
- 608

Save & Next حفظ والتالي

Question No. 19

What is the charge on the ion formed by aluminum?

- 5-
- 3+
- 3-
- 13+

Save & Next حفظ والتالي

## Question No. 16

0.100 mole of lithium weighs

- 3.00 g
- 0.694 g
- 6.94 g
- 0.300 g

Save & Next حفظ التالي

## Question No. 9

What is the atomic symbol for gold?

- Ar
- As
- Au
- Ag

Save & Next حفظ والتالي

## Question No. 4

9.31 g is the same mass as \_\_\_\_\_.

- 931  $\mu\text{g}$
- 9310 mg
- 93.1 cg
- 931 kg

Save & Next حفظ التالي

## Question No. 11

One element that has 5 valence electrons is \_\_\_\_\_

- lithium
- sulfur
- nitrogen
- carbon

Save & Next حفظ و اگلی

## Question No. 7

Which of the following is a chemical change?

- burning natural gas
- cutting paper
- cutting a tomato
- melting ice

Save & Next حفظ و التالي

The final step in forming the modern atomic theory was done by:

A. Rutherford

B. Dalton

C. Proust

D. Lavoisier



## Question No. 8

If the temperature is  $78^{\circ}\text{F}$ , what is the temperature in degrees celsius?

- $25.6^{\circ}\text{C}$
- $-145^{\circ}\text{C}$
- $262.4^{\circ}\text{C}$
- $401^{\circ}\text{C}$

Save & Next حفظ و التالي

## Question No. 20

The ionization energy of silicon is lower than the ionization energy for \_\_\_\_\_.

- magnesium
- aluminum
- sodium
- phosphorus

Save & Next حفظ والتالي

**Question No. 24**

Which of these elements has the lowest first ionization energy?

- aluminum
- sulfur
- sodium
- phosphorus

## Question No. 24

Which of these elements has the lowest electron affinity?

- carbon
- boron
- nitrogen
- beryllium

Save & Next حفظ واقتلي

# heterogeneous

غیر یکسان  
ترکیب

مادہ

- Salt water
- Chicken soup
- Hydrogen peroxide
- Ice



Question No. 7

Which of the following is NOT an example of matter?

- a balloon full of helium
- heat from a burning candle
- a pencil eraser
- a dust particle

Save & Next حفظ والتالي

HP LE1901w



Question No. 12

The names of the elements whose symbols are Si, P, Mn, and S are respectively.

- silver, phosphorus, magnesium, and sulfur.
- silicon, phosphorus, magnesium, and sulfur.
- silicon, potassium, magnesium, and sulfur.
- silicon, phosphorus, manganese, and sulfur.



Question No. 23

Which of the following elements is a metal?

- fluorine
- strontium
- nitrogen
- argon



## Question No. 22

How many moles of  $(\text{NH}_4)_2\text{S}$  are there in 75 g of  $(\text{NH}_4)_2\text{S}$ ?

- 1.5
- 1.04
- 1.56
- 1.1

Save & Next حفظ والتالي



Question No. 3

Which of the following is not an element?

- gold
- carbon
- water
- silver

Save & Next حفظ و التالي

## Question No. 5

The correct prefix for the multiplier 1,000,000,000 is

- milli.
- giga.
- mega.
- tera.

Save & Next حفظ و التالي

- 11] Who formulated the modern atomic theory
- 12] How many electrons in a-2 of S
- 13] The group FA we can name it
- 14] The physical property is
- 15] How many moles in  $(NH_4)_2S$  9 150g  $g_{mole}$
- 16] Solids  $lattice$  is distinct volume and shape
- 17] Which of the following is a heterogeneous
- 18] Which of the following is a homogeneous
- 19] The symbol of Strontium
- 110] Density =  $^{\circ}$
- 111] Convert between  $f$  to  $^{\circ}$ , and  $^{\circ}$  to  $f$
- 112] How many electrons in the f orbital
- 113] How many electrons on the 4<sup>th</sup> level
- 114] atomic radius
- 115] ionization



Today

السلام عليكم ...

11:09 PM ✓✓

دكتوراه لو سمحتي ايش اجابة هذا السؤال ؟

11:10 PM ✓✓

Gold in a ring is a  
-compound  
-element  
-heterogeneous  
-homogeneous

11:10 PM ✓✓

Homogenous

11:11 PM

لأنه سبيكة alloy

11:12 PM

لو قالي فضه كمان نفس الشي ؟

11:13 PM ✓✓

لا اعتقد

11:15 PM

يعني لو فضه احط عنصر صح 🙈

11:16 PM ✓✓

ايوة

11:17 PM



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## Question No. 19

Which of these elements has the highest electron affinity?

- Rubidium
- Potassium
- Lithium
- Caesium

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HP LE1901w

## Question No. 20

The ionization energy of silicon is lower than the ionization energy for \_\_\_\_\_.

- magnesium
- aluminum
- sodium
- phosphorus

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Question No. 18

Non Metals are located where on the periodic table?

- left side
- between metals and nonmetals
- middle
- right side

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Question No. 19

Which of these elements has the highest electron affinity?

- magnesium
- silicon
- aluminum
- sulfur

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## Question No. 16

0.100 mole of lithium weighs

- 3.00 g
- 0.694 g
- 6.94 g
- 0.300 g

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## Question No. 17

How many moles and how many atoms of zinc are in a sample weighing 34.9 g?

- 0.533 mol,  $8.85 \times 10^{-25}$  atoms
- 1.87 mol,  $1.13 \times 10^{24}$  atoms
- 1.87 mol,  $3.10 \times 10^{-24}$  atoms
- 0.533 mol,  $3.21 \times 10^{23}$  atoms

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Question No. 15

What is the charge on the ion formed by sodium?

- 1-
- 2+
- 1+
- 2-

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## Question No. 14

Isotopes are:

- atoms of the same element that have different number of neutrons.
- atoms of the same element that have the same number of neutrons.
- atoms of the same element that have different number of electrons.
- atoms of the same element that have different number of protons.

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## Question No. 13

The elements rubidium, sodium, and potassium \_\_\_\_\_

- are isotopes of each other
- are in the same group
- have the same number of neutrons
- are in the same period of elements

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## Question No. 12

Francium is an example of an element that is a \_\_\_\_\_.

- metalloid
- noble gas
- nonmetal
- metal

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## Question No. 11

One element that has 5 valence electrons is \_\_\_\_\_

- lithium
- sulfur
- nitrogen
- carbon

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## Question No. 10

An s orbital can hold \_\_\_\_\_ electrons.

- 6
- 2
- 18
- 10

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## Question No. 6

Which state of matter has atomic spacing that is close together and fixed shape?

- liquid
- plasma
- gas
- solid

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## Question No. 5

A solid form of matter in which there is long range repeating order is called \_\_\_\_\_.

- amorphous
- fixed
- crystalline
- rigid

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## Question No. 4

9.31 g is the same mass as \_\_\_\_\_

- 931  $\mu\text{g}$
- 9310 mg
- 93.1 cg
- 931 kg

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## Question No. 9

What is the atomic symbol for gold?

- Ar
- AS
- Au
- Ag

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## Question No. 7

Which of the following is a chemical change?

- burning natural gas
- cutting paper
- cutting a tomato
- melting ice

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## Question No. 16

0.100 mole of lithium weighs

- 3.00 g
- 6.94 g
- 6.94 g
- 0.300 g

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Question No. 2

In a solution, the solvent is \_\_\_\_\_.

- always water
- the substance being dissolved
- the substance present in the greatest amount
- always a liquid



**INSTRUCTION:** تعليمات Please choose the BEST answer from the given options for each question

**Question:**

Which of the following is a heterogeneous mixture?

**Options:**

- pure water
- tap water
- gasoline
- ice and water

تسليم الإجابة  
Submit Answer

**INSTRUCTION:** **تعليمات** Please choose the BEST answer from the given options for each

**Question:**

Which of the following is a physical process?

**Options:**

- the explosion of nitroglycerin
- the baking of a potato
- the rusting of iron
- the condensation of water vapor

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Submit Answer

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