MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 4

How many grams of NaF are there in 120 mL of a 1.8 M aqueous solution of NaF?

- 3.12 g
- 1.34 g
- 9.07 g
- 13.43 g

Save & Next جतन व पुढचे

KCL OES

Question No. 10

How many valence electrons are in an oxygen atom and an oxide ion?

- 8 and 10, respectively
- 8 and 6, respectively
- 2 and 8, respectively
- 6 and 8, respectively

Total questions in exam: 25 | Answered: 0

Question No. 2

The coefficients (a,b,c) needed to balance the equation below are:

$$a\,H_2 + b\,O_2 \rightarrow c\,H_2O$$

- (1,2,1)
- (2,2,1)
- (1,2,2)
- (2,1,2)

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 4

How many grams of NaF are there in 120 mL of a 1.8 M aqueous solution of NaF?

- 3.12 g
- 1.34 g
- 9.07 g
- 13.43 g

Save & Next جतन व पुढचे

Question No. 3

What is the percent yield for a reaction if its theoretical yield is 130 g and its actual yield is 74 g?

- 57%
- 53%
- 80%
- 64%

Total questions in exam: 25 | Answered: 0

Question No. 2

The coefficients (a,b,c) needed to balance the equation below are:

$$a\,H_2 + b\,O_2 \rightarrow c\,H_2O$$

- (1,2,1)
- (2,2,1)
- (1,2,2)
- (2,1,2)

Question No. 20

In an oxidation-reduction reaction, the oxidized substance always _____

- shows loss of electrons.
- shows gain of electrons.
- shows gain of neutrons.
- gives up hydrogen atoms.

ler appropriate conditions, nitrogen and hydrogen undergo a combination reaction
ield ammonia:

$N_2\,(g) + 3\,H_2\,(g) \rightarrow 2NH_3\,(g)$

_____ g sample of N_2 requires 3.0 g of H_2 for complete reaction.

Save & Next حفظ والتالي

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$N_2\,(g) + 3\,H_2\,(g) \rightarrow 2NH_3\,(g)$

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Save & Next حفظ والتالي

Question No. 3

What is the percent yield for a reaction if its theoretical yield is 130 g and its actual yield is 74 g?

- 57%
- 53%
- 80%
- 64%

Total questions in exam: 25 | Answered: 0

Question No. 7

Calculate the molar mass for $Mg(ClO_4)_2$.

- 123 g/mol
- 223 g/mol
- 119 g/mol
- 247 g/mol

Save & Next حفظ والتالي

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 8

How many valence electrons are in an oxygen atom and an oxide ion?

- 6 and 8, respectively
- 2 and 8, respectively
- 8 and 6, respectively
- 8 and 10, respectively

Save & Next

Total questions in exam: 25 | Answered: 0

Question No. 7

Calculate the molar mass for $Mg(ClO_4)_2$.

- 123 g/mol
- 223 g/mol
- 119 g/mol
- 247 g/mol

Save & Next حفظ والتالي

Question No. 23

Which of the following has the greatest number of unshared pairs of electrons?

- H
- F
- C
- He

Save & Next حفظ والتالي

ler appropriate conditions, nitrogen and hydrogen undergo a combination reaction
ield ammonia:

$N_2\,(g) + 3\,H_2\,(g) \rightarrow 2NH_3\,(g)$

_____ g sample of N_2 requires 3.0 g of H_2 for complete reaction.

Save & Next حفظ والتالي

Question No. 1

When the following equation is balanced, the coefficient of wate

$$N_2O_5\,(g) + H_2O\,(l) \rightarrow HNO_3\,(aq)$$

- 1
- 5
- 3
- 2

MKCL OES

Question No. 6

Which molecule contains the weakest carbon-carbon bond?

- $H_3C{-}CH_3$
- $HC{\equiv}CH$
- $F_3C{-}CF_3$
- $H_3C{-}CH_3$

Question No. 1

The most correct name for the compound CS_2 is:

- carbon trisulfide
- monocarbon disulfide
- carbon sulfide
- carbon disulfide

Total questions in exam: 25 | Answered: 0

Question No. 2

The coefficients (a,b,c) needed to balance the equation below are:

$$a\,H_2 + b\,O_2 \rightarrow c\,H_2O$$

- (1,2,1)
- (2,2,1)
- (1,2,2)
- (2,1,2)

Total questions in exam: 25

Question No. 1

What is the oxidation number of manganese in MnO_2?

- +4
- +1
- 0
- +2

MKCL OES

Total questions in exam: 25 | Answered: 1

Question No. 2

What is the reducing agent in the following reaction?

$$Mg + NiO_2 + 2H_2O \rightarrow Mg(OH)_2 + Ni(OH)_2$$

- O
- Ni
- Mg
- H

Question No. 3

What is the percent yield for a reaction if its theoretical yield is 130 g and its actual yield is 74 g?

- 57%
- 53%
- 80%
- 64%

Total questions in exam: 25 | Answered: 0

Question No. 7

Calculate the molar mass for $Mg(ClO_4)_2$.

- 123 g/mol
- 223 g/mol
- 119 g/mol
- 247 g/mol

Save & Next حفظ والتالي

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 8

How many valence electrons are in an oxygen atom and an oxide ion?

- 6 and 8, respectively
- 2 and 8, respectively
- 8 and 6, respectively
- 8 and 10, respectively

Save & Next

Total questions in exam: 25 | Answered: 0

Question No. 7

Calculate the molar mass for $Mg(ClO_4)_2$.

- 123 g/mol
- 223 g/mol
- 119 g/mol
- 247 g/mol

Save & Next حفظ والتالي

Question No. 23

Which of the following has the greatest number of unshared pairs of electrons?

- H
- F
- C
- He

Save & Next حفظ والتالي

ler appropriate conditions, nitrogen and hydrogen undergo a combination reaction
ield ammonia:

$N_2\,(g) + 3\,H_2\,(g) \rightarrow 2NH_3\,(g)$

_____ g sample of N_2 requires 3.0 g of H_2 for complete reaction.

Save & Next حفظ والتالي

Question No. 1

When the following equation is balanced, the coefficient of wate

$$N_2O_5\,(g) + H_2O\,(l) \rightarrow HNO_3\,(aq)$$

- 1
- 5
- 3
- 2

MKCL OES

Question No. 6

Which molecule contains the weakest carbon-carbon bond?

- $H_3C{-}CH_3$
- $HC{\equiv}CH$
- $F_3C{-}CF_3$
- $H_3C{-}CH_3$

Question No. 1

The most correct name for the compound CS_2 is:

- carbon trisulfide
- monocarbon disulfide
- carbon sulfide
- carbon disulfide

Total questions in exam: 25 | Answered: 0

Question No. 2

The coefficients (a,b,c) needed to balance the equation below are:

$$a\,H_2 + b\,O_2 \rightarrow c\,H_2O$$

- (1,2,1)
- (2,2,1)
- (1,2,2)
- (2,1,2)

Total questions in exam: 25

Question No. 1

What is the oxidation number of manganese in MnO_2?

- +4
- +1
- 0
- +2

Total questions in exam: 25 | Answered: 0

Question No. 7

Calculate the molar mass for $Mg(ClO_4)_2$.

- 123 g/mol
- 223 g/mol
- 119 g/mol
- 247 g/mol

Save & Next حفظ والتالي

MKCL OES

Total questions in exam: 25 | Answered: 1

Question No. 2

What is the reducing agent in the following reaction?

$$Mg + NiO_2 + 2H_2O \rightarrow Mg(OH)_2 + Ni(OH)_2$$

- O
- Ni
- Mg
- H

Total questions in exam: 25 | Answered: 0

Question No. 5

The name of the chemical compound Cu_2O is

- Copper(II) oxide
- Copper(I) oxide
- Copper(III) oxide
- Copper oxide

MKCL OES

Total questions in exam: 25 | Answered: 25

Question No. 23

What is the percent yield for a reaction if its theoretical yield is 135 g and its actual yield is 97 g?

- 72%
- 53%
- 86%
- 63%

Save & Next حفظ والتالي

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Question No. 25

The name of the chemical compound Cu_2S is:

- Copper(III) sulfide
- Copper(II) sulfide
- Copper(I) sulfide
- Copper sulfide

Save & Next حفظ والتالي

ler appropriate conditions, nitrogen and hydrogen undergo a combination reaction
ield ammonia:

$N_2\,(g) + 3\,H_2\,(g) \rightarrow 2NH_3\,(g)$

_____ g sample of N_2 requires 3.0 g of H_2 for complete reaction.

Save & Next حفظ والتالي

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 3

How many grams of $AlCl_3$ could be produced when 1.5 moles of Cl_2 completely react with aluminum according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 533 g
- 267 g
- 134 g
- 333 g

Question No. 23

Which of the following has the greatest number of unshared pairs of electrons?

- H
- F
- C
- He

Save & Next حفظ والتالي

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 3

How many grams of $AlCl_3$ could be produced when 1.5 moles of Cl_2 completely react with aluminum according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 533 g
- 267 g
- 134 g
- 333 g

Question No. 25

What is the molecular formula of a compound that has a molar mass of 70 g/mol and its empirical formula is CH_2?

- $C_{10}H_{10}$
- C_5H_{10}
- C_7H_{10}
- $C_{20}H_{40}$

Save & Next حفظ والتالي

Question No. 22

What is the term for a bond in which a pair of electrons is shared equally?

- electrovalent bond
- polar covalent bond
- ionic bond
- nonpolar covalent bond

Save & Next حفظ والتالي

Question No. 21

Molarity is defined as

- moles of solute per liters of solution
- moles of solute per kilograms of solute
- grams of solute per moles of solvent
- grams of solute per liters of solution

Save & Next حفظ والتالي

Chemistry_Quiz2_Sem1_2018

Total questions in exam: 25 | Answered: 0

Question No. 1

The number which is located on the LEFT of a chemical formula that helps to balance a chemical equation is called _____

- exponent
- coefficient
- superscript
- subscript

Save & Next حفظ والتالي

Question No. 22

What is the term for a bond in which a pair of electrons is shared equally?

- electrovalent bond
- polar covalent bond
- ionic bond
- nonpolar covalent bond

Save & Next حفظ والتالي

Question No. 21

Molarity is defined as

- moles of solute per liters of solution
- moles of solute per kilograms of solute
- grams of solute per moles of solvent
- grams of solute per liters of solution

Save & Next حفظ والتالي

Question No. 20

A chemical reaction produces 7.25 moles of barium sulfate, $BaSO_4$. What mass (in grams) of barium sulfate is produced?

- 253 g
- 1.69×10^3 g
- 185 g
- 31.1 g

Save & Next

Chemistry_Quiz2_Sem1_2018

Total questions in exam: 25 | Answered: 0

Question No. 1

The number which is located on the LEFT of a chemical formula that helps to balance a chemical equation is called _____

- exponent
- coefficient
- superscript
- subscript

Save & Next حفظ والتالي

Total questions in exam: 25

Question No. 1

What is the oxidation number of manganese in MnO_2?

- +4
- +1
- 0
- +2

Question No. 1

Which of the following gives the balanced equation for this reaction?

$$Na_3PO_4 + Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + NaNO_3$$

- $Na_3PO_4 + 2Ca(NO_3)_2 \rightarrow 2Ca_3(PO_4)_2 + NaNO_3$
- $2Na_3PO_4 + Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + 6NaNO_3$
- $Na_3PO_4 + Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + 3NaNO_3$
- $2Na_3PO_4 + 3Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + 6NaNO_3$

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 1

When the following equation is balanced, the coefficient of wate

$$N_2O_5\,(g) + H_2O\,(l) \rightarrow HNO_3\,(aq)$$

- 1
- 5
- 3
- 2

Question No. 21

Write the name for FeS

- iron(II) sulfate
- iron(I) sulfate
- iron(II) sulfide
- iron(I) sulfide

Total questions in exam: **25** | Answered: **1**

Question No. 6

How many phosphorus (P) atoms are contained in 158 kg of phosphorus?

- 1.18×10^{24} phosphorus atoms
- 3.07×10^{27} phosphorus atoms
- 3.25×10^{28} phosphorus atoms
- 2.95×10^{27} phosphorus atoms

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Total questions in exam: 25 | Answered: 2

Question No. 1

Which of the following gives the balanced equation for this reaction?

$$Na_3PO_4 + Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + NaNO_3$$

- $Na_3PO_4 + 2Ca(NO_3)_2 \rightarrow 2Ca_3(PO_4)_2 + NaNO_3$
- $2Na_3PO_4 + Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + 6NaNO_3$
- $Na_3PO_4 + Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + 3NaNO_3$
- $2Na_3PO_4 + 3Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + 6NaNO_3$

Save & Next حفظ والتالي

HP Compaq LE1711

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 3

How many grams of $AlCl_3$ could be produced when 1.5 moles of Cl_2 completely react with aluminum according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 533 g
- 267 g
- 134 g
- 333 g

ler appropriate conditions, nitrogen and hydrogen undergo a combination reaction
ield ammonia:

$N_2\,(g) + 3\,H_2\,(g) \rightarrow 2NH_3\,(g)$

_____ g sample of N_2 requires 3.0 g of H_2 for complete reaction.

Save & Next حفظ والتالي

For a given reaction

$$V_2O_5(s) + 5\,Ca(l) \rightarrow 2\,V(l) + 5\,CaO(s)$$

When 10 moles of V_2O_5 are mixed with 10 moles of Ca, which is the limiting reactant acco to the above equation?

- V_2O_5
- Ca
- V
- CaO

Question No. 1

When the following equation is balanced, the coefficient of wate

$$N_2O_5\,(g) + H_2O\,(l) \rightarrow HNO_3\,(aq)$$

- 1
- 5
- 3
- 2

MKCL OES

Total questions in exam: 25 | Answered: 3

Question No. 7

The number of lone electron pairs in the N_2 molecule is ___.

- 5 pairs
- 2 pairs
- 1 pair
- 3 pairs

Save & Next जतन व पुढील

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 3

How many grams of $AlCl_3$ could be produced when 1.5 moles of Cl_2 completely react with aluminum according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 533 g
- 267 g
- 134 g
- 333 g

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 3

How many grams of $AlCl_3$ could be produced when 1.5 moles of Cl_2 completely react with aluminum according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 533 g
- 267 g
- 134 g
- 333 g

Total questions in exam: **25** | Answered: **14**

Question No. 12

In which of these substances are the atoms held together by polar covalent bonding?

- $CsCl$
- $SrCl_2$
- ClF
- S_8

ler appropriate conditions, nitrogen and hydrogen undergo a combination reaction
ield ammonia:

$N_2\,(g) + 3\,H_2\,(g) \rightarrow 2NH_3\,(g)$

_____ g sample of N_2 requires 3.0 g of H_2 for complete reaction.

Save & Next حفظ والتالي

Total questions in exam: 28 | Answered: 0

Question No. 18

What is the final molarity of H_2SO_4 solution, if 85 mL of 4M H_2SO_4 was diluted to a final volume of 0.5 L?

- 0.60 M
- 0.52 M
- 0.68 M
- 0.76 M

Question No. 15

What is the molecular formula of a compound that has a molar mass of 30 g/mol and its empirical formula is CH_2O?

- $C_3H_6O_3$
- CH_2O
- $C_4H_8O_4$
- $C_5H_{10}O_5$

Total questions in exam: **25** | Answered: **4**

Question No. 8

What is the final molarity of KOH solution, if 0.064 L of 3.4 M KOH was diluted to a final volume of 0.4 L?

- 0.27 M
- 0.61 M
- 0.54 M
- 0.68 M

Save & Next جطر وقلي

MKCL OES

Chemistry_Quiz2_Sem1_2018

Total questions in exam: 25 | Answered: 4

Question No. 6

Oxidation is the _____ and reduction is the _____.

- gain of oxygen, loss of electrons
- loss of electrons, gain of electrons
- loss of oxygen, gain of electrons
- gain of electrons, loss of electrons

Save & Next جतन करा पुढे

HP Compaq LE1711

Question No. 1

Which of the following gives the balanced equation for this reaction?

$$Na_3PO_4 + Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + NaNO_3$$

- $Na_3PO_4 + 2Ca(NO_3)_2 \rightarrow 2Ca_3(PO_4)_2 + NaNO_3$
- $2Na_3PO_4 + Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + 6NaNO_3$
- $Na_3PO_4 + Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + 3NaNO_3$
- $2Na_3PO_4 + 3Ca(NO_3)_2 \rightarrow Ca_3(PO_4)_2 + 6NaNO_3$

Save & Next جمع و التالي

HP Compaq LE1711

Question No. 2

Which element is *reduced* in the following reaction?

$$Cu + 2H_2SO_4 \rightarrow CuSO_4 + SO_2 + 2H_2O$$

- H
- S
- O
- Cu

Total questions in exam: 25 | Answered: 5

Question No. 8

Which one of the following is oxidized in the following equation?

$$2Al(s) + 3H_2SO_4(aq) \rightarrow Al_2(SO_4)_3(aq) + 3H_2(g)$$

- $Al_2(SO_4)_3(aq)$
- $H_2(g)$
- $Al(s)$
- $H_2SO_4(aq)$

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Total questions in exam: **25** | Answered: **0**

Question No. 9

The mass percent composition of oxygen in the acid H_2SO_3 is:

- 58.5%
- 2.5%
- 39.1%
- 65.3%

Save & Next

Question No. 12

Calcium carbide (CaC_2) reacts with water to produce acetylene (C_2H_2):

$$CaC_2\,(s) + 2\,H_2O\,(g) \rightarrow Ca(OH)_2\,(s) + C_2H_2\,(g)$$

Production of 13 g of C_2H_2 requires consumption of _____ g of H_2O.

- 4.5
- 18
- 4.8×10^2
- 9.0

Save & Next

For a given reaction

$$V_2O_5(s) + 5\,Ca(l) \rightarrow 2\,V(l) + 5\,CaO(s)$$

When 10 moles of V_2O_5 are mixed with 10 moles of Ca, which is the limiting reactant according to the above equation?

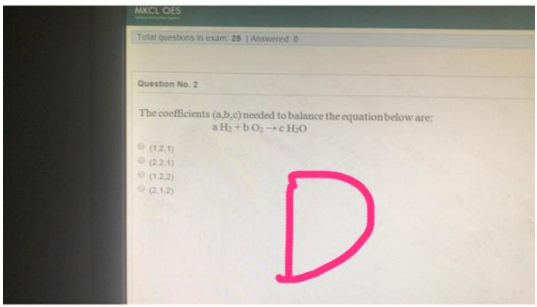
- V_2O_5
- Ca
- V
- CaO

Question No. 23

Which of the following has the greatest number of unshared pairs of electrons?

- H
- F
- C
- He

Save & Next حفظ والتالي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Total questions in exam: 25 | Answered: 0
Question No. 2
The coefficients (a,b,c) needed to balance the equation below are:
a H_2 + b O_2 → c H_2O
(1,2,1)
(2,2,1)
(1,2,2)
(2,1,2)
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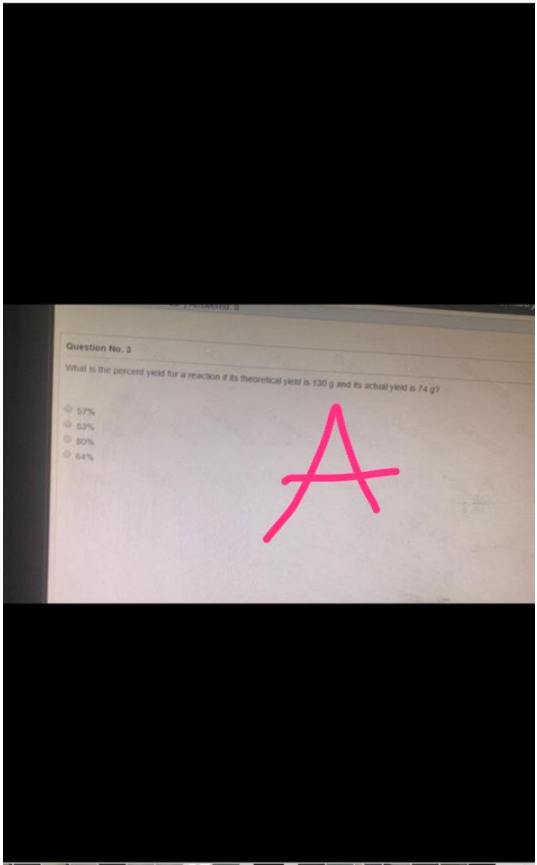
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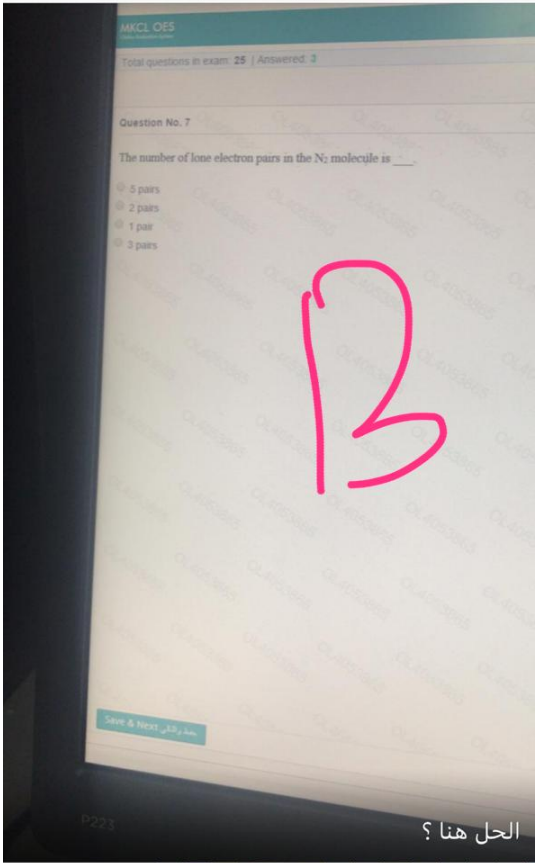
57%
53%
80%
64%
A


MKCL OES
Total questions in exam: 25 | Answered: 3
Question No. 7
The number of lone electron pairs in the N_2 molecule is ___.
5 pairs
2 pairs
1 pair
3 pairs
B
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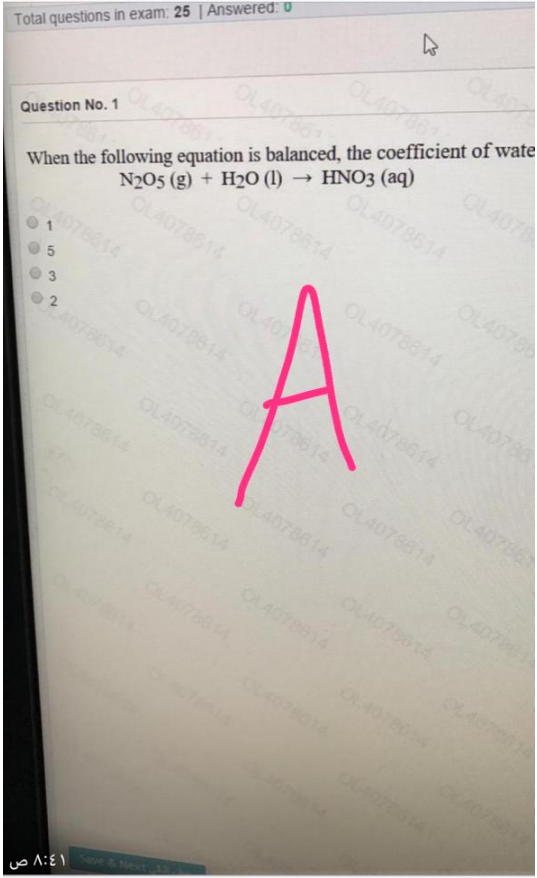

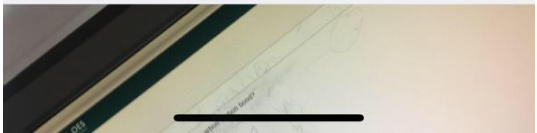

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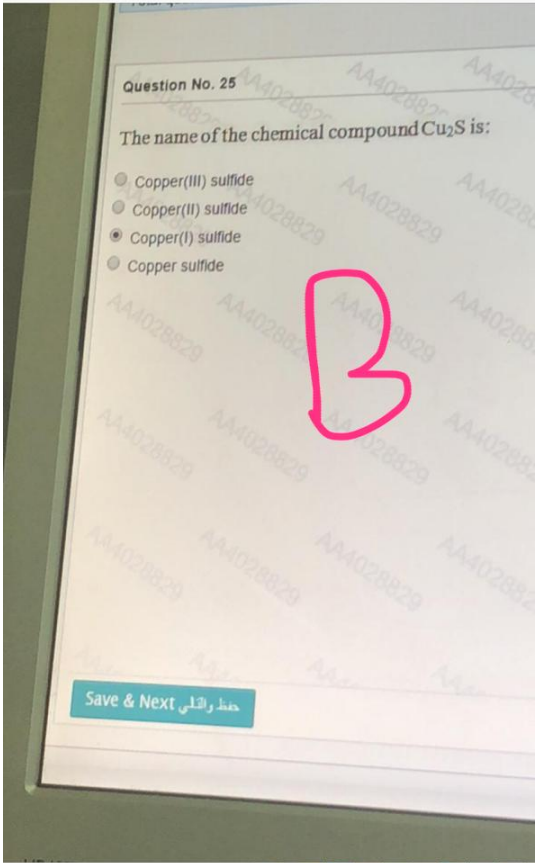

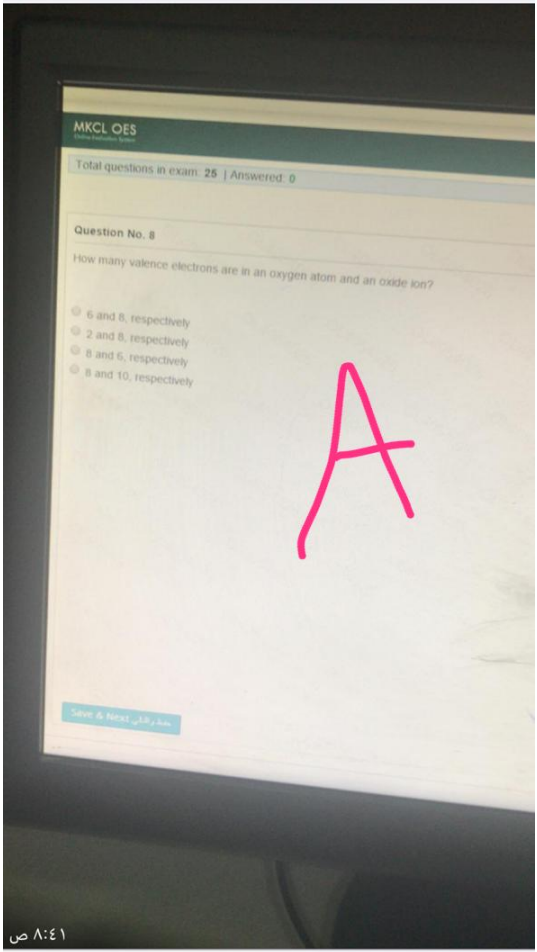
MKCL OES
Total questions in exam: 25 | Answered: 0
Question No. 8
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6 and 8, respectively
2 and 8, respectively
8 and 6, respectively
8 and 10, respectively
A
Save & Next حفظ والتالي
٨:٤١ ص

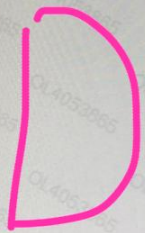
MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 1

What is the oxidation number of manganese in MnO_2?

- +4
- +1
- 0
- +2



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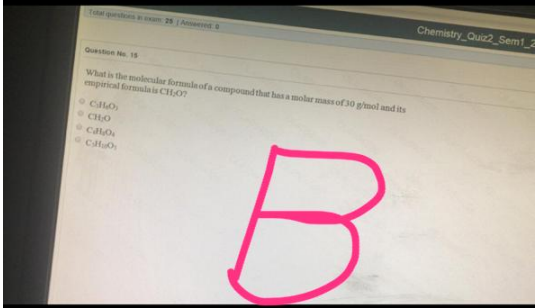
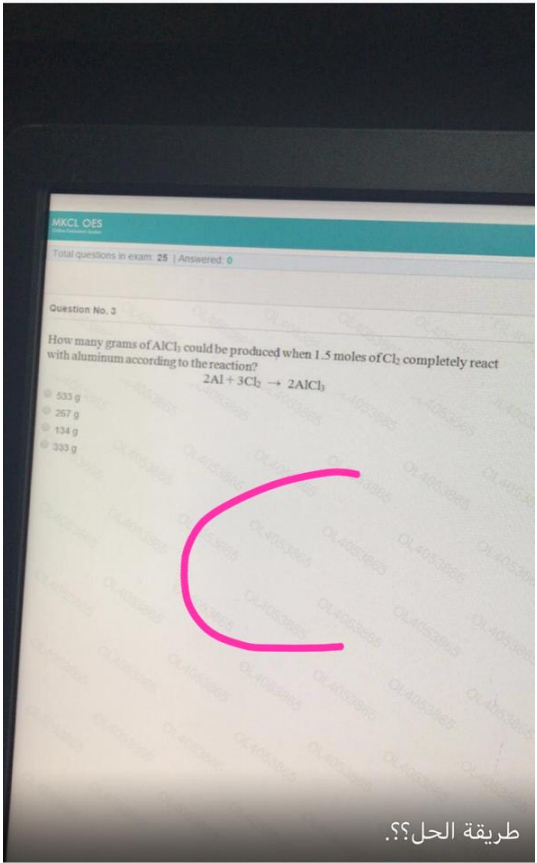
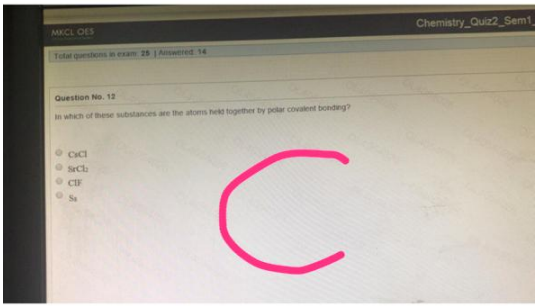
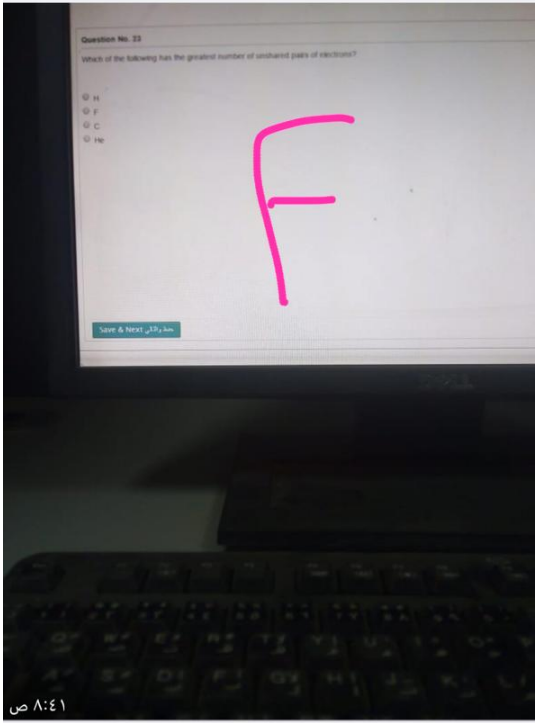
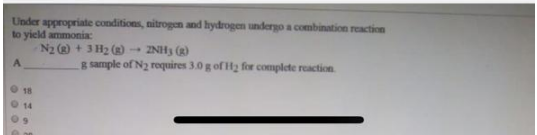

Total questions in exam: 25 | Answered: 0
Chemistry_Quiz2_Sem1_
Question No. 15
What is the molecular formula of a compound that has a molar mass of 30 g/mol and its empirical formula is CH_2O?
$C_2H_4O_2$
CH_2O
$C_4H_8O_4$
$C_3H_{12}O_3$
B


MKCL OES
Total questions in exam: 25 | Answered: 0
Question No. 3
How many grams of $AlCl_3$ could be produced when 1.5 moles of Cl_2 completely react with aluminum according to the reaction?
$2Al + 3Cl_2 \rightarrow 2AlCl_3$
533 g
267 g
134 g
333 g
طريقة الحل؟؟.


MKCL OES
Chemistry_Quiz2_Sem1
Total questions in exam: 25 | Answered: 14
Question No. 12
In which of these substances are the atoms held together by polar covalent bonding?
CsCl
SrCl2
ClF
Sn

٨:٤١ ص


Question No. 33
Which of the following has the greatest number of unshared pairs of electrons?
H
F
C
He
F
Save & Next حفظ والتالي
٨:٤١ ص


Under appropriate conditions, nitrogen and hydrogen undergo a combination reaction to yield ammonia:
$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$
A _____ g sample of N_2 requires 3.0 g of H_2 for complete reaction.
18
14
9

MKCL OES

Question No. 6

Which molecule contains the weakest carbon-carbon bond?

- $H_3C{-}CH_3$
- $HC{\equiv}CH$
- $F_3C{-}CF_3$
- $H_3C{-}CH_3$

Total questions in exam: 25 | Answered: 0

Question No. 7

Calculate the molar mass for $Mg(ClO_4)_2$.

- 123 g/mol
- 223 g/mol
- 119 g/mol
- 247 g/mol

Save & Next حفظ والتالي

Question No. 12

Calcium carbide (CaC_2) reacts with water to produce acetylene (C_2H_2):

$$CaC_2\,(s) + 2\,H_2O\,(g) \rightarrow Ca(OH)_2\,(s) + C_2H_2\,(g)$$

Production of 13 g of C_2H_2 requires consumption of _____ g of H_2O.

- 4.5
- 18
- 4.8×10^2
- 9.0

Save & Next

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 1

What is the oxidation number of manganese in MnO_2?

- +4
- +1
- 0
- +2

Save &

uestion No. 15

redict which of the following compounds has an ionic bond.

CCl_4

HI

LiH

IBr

ler appropriate conditions, nitrogen and hydrogen undergo a combination reaction
ield ammonia:

$N_2\,(g) + 3\,H_2\,(g) \rightarrow 2NH_3\,(g)$

_____ g sample of N_2 requires 3.0 g of H_2 for complete reaction.

Save & Next حفظ والتالي

Question No. 22

What is the term for a bond in which a pair of electrons is shared equally?

- electrovalent bond
- polar covalent bond
- ionic bond
- nonpolar covalent bond

Save & Next حفظ والتالي

Question No. 11

Which of the following is the electron dot formula (Lewis structure) for an atom of boron?

(a) B· (b) ·B· (c) ·B: (d) :B:

- (b)
- (a)
- (c)
- (d)

Which of these compounds is a

- ionic salt
- strong base
- strong acid
- weak base

Save & Next حفظ والتالي

HP L1710

Question No. 24

The oxidation number of iodine in KIO_4 is

- +7
- -7
- +1
- -1

Save & Next حفظ والتالي

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HP L1710

Total questions in exam: **25** | Answered: **25**

Question No. 22

Analysis of an unknown substance showed that it has a high boiling point and is brittle. It is an insulator as a solid but conducts electricity when melted. Which of the following substances would have those characteristics?

- (•) KBr
- ( ) CO
- ( ) HCl
- ( ) Al

Save & Next حفظ والتالي

AA4028829

CL OES

tal questions in exam: 25 | Answered: 0

estion No. 10

ich of the following compounds gives a strong electrolyte aqueous solution?

CH_4

H_2SO_4

$C_6H_{12}O_6$

CH_3OH

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 8

How many valence electrons are in an oxygen atom and an oxide ion?

- 6 and 8, respectively
- 2 and 8, respectively
- 8 and 6, respectively
- 8 and 10, respectively

Save & Next

Total questions in exam: **25** | Answered: **0**

Question No. 6

Which of the following statements is FALSE regarding an ionic bond?

- An ionic bond is the attraction between a metal ion and a nonmetal ion.
- Metal atoms lose electrons and nonmetal atoms gain electrons.
- Breaking an ionic bond between two ions requires energy.
- The ionic bond is formed between two nonmetal atoms.

Save & Next

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 4

How many grams of NaF are there in 120 mL of a 1.8 M aqueous solution of NaF?

- 3.12 g
- 1.34 g
- 9.07 g
- 13.43 g

Save & Next جतन व पुढचे

Question No. 1

When the following equation is balanced, the coefficient of H_2O is _____.

$$Ca\ (s) + H_2O\ (l) \rightarrow Ca(OH)_2\ (aq) + H_2\ (g)$$

- 1
- 4
- 2
- 3

Question No. 23

Which of the following has the greatest number of unshared pairs of electrons?

- H
- F
- C
- He

Save & Next حفظ والتالي

Total questions in exam: **25** | Answered: **0**

Question No. 5

The name of the chemical compound Cu_2O is

- Copper(II) oxide
- Copper(I) oxide
- Copper(III) oxide
- Copper oxide

For a given reaction

$$V_2O_5(s) + 5\,Ca(l) \rightarrow 2\,V(l) + 5\,CaO(s)$$

When 10 moles of V_2O_5 are mixed with 10 moles of Ca, which is the limiting reactant acco to the above equation?

- V_2O_5
- Ca
- V
- CaO

Question No. 31

Molarity is defined as

- moles of solute per liters of solution
- moles of solute per kilograms of solute
- grams of solute per moles of solvent
- grams of solute per liters of solution

Save & Next اگلا سوال

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 3

How many grams of $AlCl_3$ could be produced when 1.5 moles of Cl_2 completely react with aluminum according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 533 g
- 267 g
- 134 g
- 333 g

Question No. 11

What is the term for a bond composed of three electron pairs shared between two atoms?

- ionic bond
- double bond
- single bond
- triple bond

Save & Next حفظ و التالي

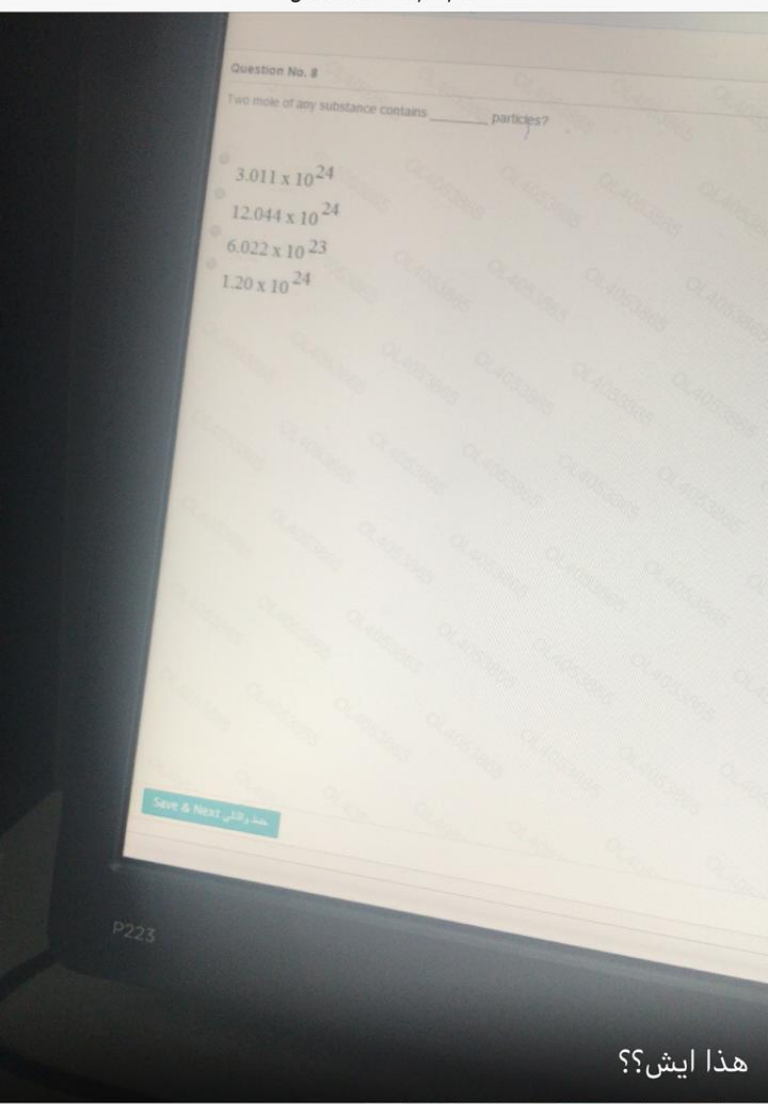
P223

Question No. 2

Which element is *reduced* in the following reaction?

$$Cu + 2H_2SO_4 \rightarrow CuSO_4 + SO_2 + 2H_2O$$

- H
- S
- O
- Cu


Question No. 8
Two mole of any substance contains _____ particles?
3.011×10^{24}
12.044×10^{24}
6.022×10^{23}
1.20×10^{24}
Save & Next
P223

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MKCL OES

Total questions in exam: 25 | Answered: 1

Question No. 2

What is the reducing agent in the following reaction?

$$Mg + NiO_2 + 2H_2O \rightarrow Mg(OH)_2 + Ni(OH)_2$$

- O
- Ni
- Mg
- H

Question No. 19

The name of the chemical compound $MgSO_4$ is:

- magnesium sulfite
- magnesium sulfate
- magnesium(II) sulfate
- magnesium(II) sulfite

B

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Chemistry-Q2(2).pdf
chapter 3 1-معدل.pdf
Chapter 4 (1).docx

MKCL OES

Total questions in exam: 25 | Answered: 0

Question No. 3

How many grams of $AlCl_3$ could be produced when 1.5 moles of Cl_2 completely react with aluminum according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 533 g
- 267 g
- 134 g
- 333 g

Question No. 3

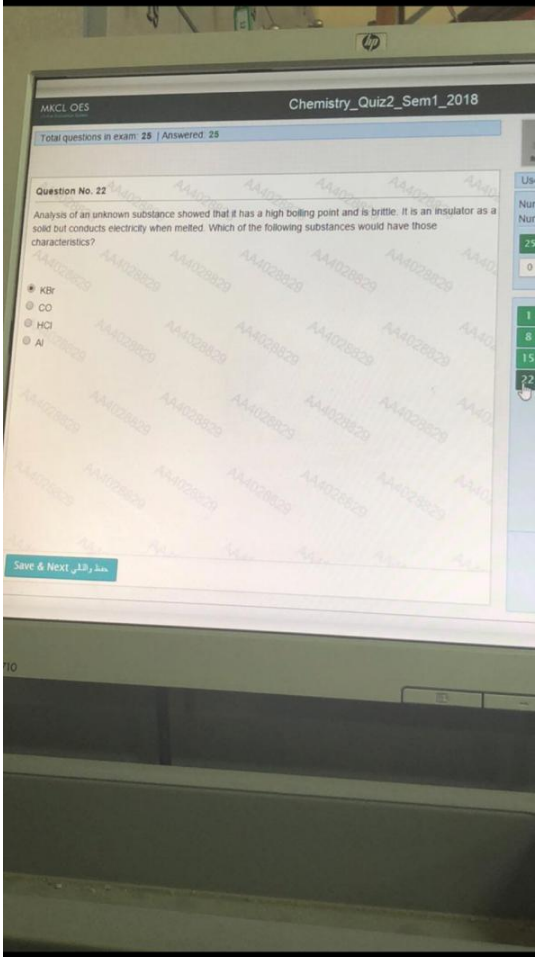
Calcium carbide (CaC_2) reacts with water to produce acetylene (C_2H_2):

$$CaC_2\,(s) + 2\,H_2O\,(g) \rightarrow Ca(OH)_2\,(s) + C_2H_2\,(g)$$

The complete reaction of 57.4 g of CaC_2 requires consumption of _____ g of H_2O.

- 1.79
- 32.3
- 0.895
- 64.1

Save & Next

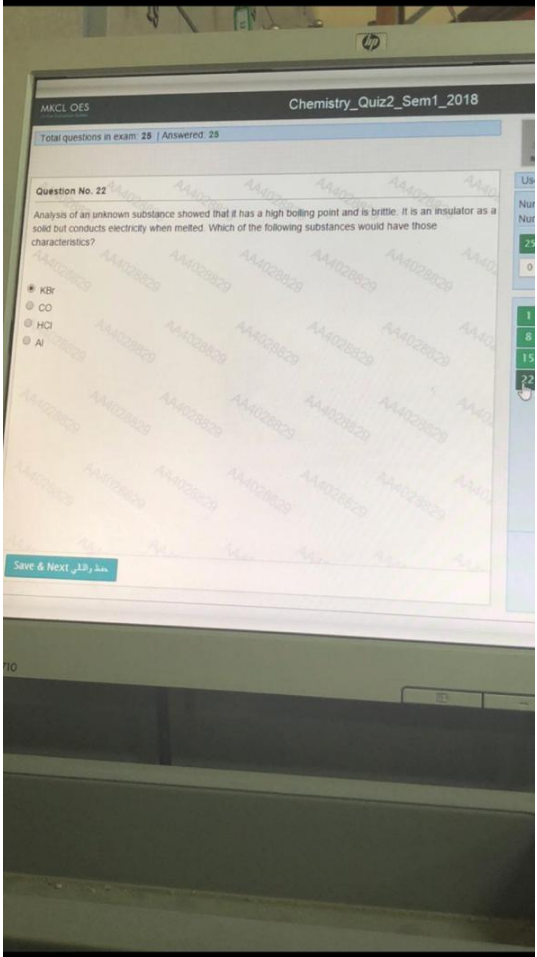

MKCL OES
Chemistry_Quiz2_Sem1_2018
Total questions in exam: 25 | Answered: 25
Question No. 22
Analysis of an unknown substance showed that it has a high boiling point and is brittle. It is an insulator as a solid but conducts electricity when melted. Which of the following substances would have those characteristics?
KBr
CO
HCl
Al
Save & Next

Question No. 11

Which of the following is the electron dot formula (Lewis structure) for an atom of boron?

(a) B· (b) ·B· (c) ·B: (d) :B:

- (b)
- (a)
- (c)
- (d)


MKCL OES
Chemistry_Quiz2_Sem1_2018
Total questions in exam: 25 | Answered: 25
Question No. 22
Analysis of an unknown substance showed that it has a high boiling point and is brittle. It is an insulator as a solid but conducts electricity when melted. Which of the following substances would have those characteristics?
KBr
CO
HCl
Al
Save & Next

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **0**

Question No. 5

What is the term for a bond composed of three electron pairs shared between two atoms?

- Ionic bond
- double bond
- triple bond
- single bond

Save & Next حفظ واقلي

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MKCL OES

Question No. 6

Which molecule contains the weakest carbon-carbon bond?

- $H_3C{-}CH_3$
- $HC{\equiv}CH$
- $F_3C{-}CF_3$
- $H_3C{-}CH_3$

MKCL OES

Chen

Question No. 8

Nitrogen and oxygen form an extensive series of oxides with the general formula N_xO_y. What is the empirical formula for an oxide that contains 63.65% nitrogen?

- NO_2
- N_2O_3
- NO
- N_2O

Question No. 16

What is the oxidation number of iron in Fe_2O_3?

- -3
- +3
- +6
- -6

Total questions in exam: 25 | Answered:

Question No. 1

How many moles of cesium (Cs) are contained in 595 kg of cesium?

- 7.91×10^4 moles Cs
- 2.23×10^2 moles Cs
- 4.48×10^3 moles Cs
- 1.26×10^3 moles Cs

Save & Next

BD4005891

Question No. 1

What is the final molarity of H_2SO_4 solution, if 240mL of 4M H_2SO_4 was diluted to a final volume of 0.5 L?

- 0.96 M
- 1.60 M
- 1.34 M
- 1.92 M

Save & Next حفظ والتالي

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Question No. 4

How many moles of H_2O could be produced when 3 moles of C_2H_6O completely r with oxygen gas according to the reaction?

$$C_2H_6O + 3\ O_2 \rightarrow 2\ CO_2 + 3\ H_2O$$

- 9 mol
- 6 mol
- 3 mol
- 12 mol

Save & Next حفظ و التالي

BD4005891

Question No. 2

What is the coefficient of oxygen gas after balancing the following equation?

$$\_\_H_2O_2(l) \xrightarrow{\Delta} \_\_H_2O(l) + \_\_O_2(g)$$

- 4
- 2
- 3
- 1

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 2

What is the coefficient of oxygen gas after balancing the following equation?

$$\_\_H_2O_2(l) \xrightarrow{\Delta} \_\_H_2O(l) + \_\_O_2(g)$$

- 4
- 2
- 3
- 1

Save & Next حفظ والتالي

OL4028305

HP Compaq LE1711

Total questions in exam: **25** | Answered: **0**

Question No. 3

What is the percent yield for a reaction if its theoretical yield is 122 g and its actual yield is 32 g?

- 26%
- 60%
- 53%
- 40%

Save & Next حفظ و التالي

HP Compaq LE1711

Total questions in exam: 25 | Answered: 0

Question No. 4

How many grams of $AlCl_3$ could be produced when 3 mole of Al completely react with Cl_2 according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 400 g
- 133 g
- 266 g
- 67 g

Save & Next جمع و التالي

HP Compaq LE1711

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **0**

Question No. 5

What is the term for a bond composed of three electron pairs shared between two atoms?

- Ionic bond
- double bond
- triple bond
- single bond

Save & Next حفظ واقبلي

OL4028305

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Total questions in exam: 25 | Answered: 0

Question No. 6

How many grams of nitrogen are needed to produce 325 grams of ammonia?

$$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$$

- 178 g
- 651 g
- 163 g
- 267 g

Save & Next حفظ والتالي

OL4028305

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Question No. 7

Calculate the molar mass of potassium chloride, KCl.

- 67.4 g/mol
- 54.5 g/mol
- 6.74 g/mol
- 74.6 g/mol

Save & Next حفظ والتالي

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Question No. 8

Lithium and nitrogen react in a combination reaction to produce lithium nitride:

$$6Li\,(s) + N_2\,(g) \rightarrow 2Li_3N\,(s)$$

How many moles of lithium are needed to produce 0.20 mol of Li_3N when the reaction is carried out in the presence of excess nitrogen?

- 0.13
- 0.6
- 0.067
- 0.1

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 9

Calculate the mass percent composition of sulfur in $Ga_2(SO_4)_3$.

- 32.6 %
- 9.4 %
- 44.9 %
- 22.5%

Save & Next حفظ والتالي

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Question No. 10

In the reaction below, what is the theoretical yield in moles for NO when 6 moles of NH_3 react with 7 moles of O_2?

$$4\ NH_3 + 5\ O_2 \rightarrow 4\ NO + 6\ H_2O$$

- 7.0 mol
- 5.6 mol
- 6.8 mol
- 4.6 mol

Save & Next حفظ والتالي

OL4028305

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Question No. 11

How many valence electrons are in a chlorine atom and a chloride ion?

- 17 and 18, respectively
- 7 and 8, respectively
- 8 and 7, respectively
- 1 and 8, respectively

Save & Next حفظ والتالي

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Question No. 12

Oxidation cannot occur without _____

- reduction
- air
- water
- acid

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 13

The name of the chemical compound $NaNO_3$ is:

- sodium(I) nitrite
- sodium(I) nitrate
- sodium nitrite
- sodium nitrate

Save & Next حفظ والتالي

OL4028305

Question No. 14

The coefficients (a,b,c) needed to balance the equation below are:

$$a\,NaNO_3 \rightarrow b\,NaNO_2 + c\,O_2$$

- (1,1,2)
- (1,2,1)
- (1,2,2)
- (2,2,1)

Save & Next حفظ والتالي

OL4028305

Question No. 15

The empirical formula of the molecular compound C_6H_{12} is:

- C_3H_8
- C_6H_{14}
- C_2H_7
- CH_2

Save & Next حفظ والتالي

OL4028305

Question No. 16

The elemental mass percent composition of succinic acid is: C 40.68% , H 5.12% , and O 54.19%.

- C_2H_4O
- $C_2H_3O_2$
- C_2H_2O
- CH_2O

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 17

How many grams of NaCl are there in 55.0 mL of a 1.90 M aqueous solution of NaCl?

- 12.2
- 6.11
- 3.21
- 0.105

Save & Next حفظ والتالي

HP Compaq LE1711

Total questions in exam: 25 | Answered: 0

Question No. 18

The most correct name for the compound CS_2 is:

- carbon disulfide
- carbon trisulfide
- carbon sulfide
- monocarbon disulfide

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 19

In an oxidation-reduction reaction, the reduced substance always _____

- shows gain of electrons.
- shows loss of electrons.
- gives up hydrogen atoms.
- takes on oxygen atoms.

Save & Next حفظ والتالي

OL4028305

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Total questions in

Question No. 20

The systematic name for the chemical compound Na_2S is:

- sodium sulfide
- sodium monosulfide
- sodium(II) sulfide
- sodium(I) sulfide

Save & Next حفظ والتالي

OL4028305

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Question No. 21

Which of the following is an ionic compound?

- CCl_4
- CO_2
- PCl_3
- LiBr

Save & Next حفظ والتالي

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Question No. 22

What volume of 12 M HCl solution must be diluted with distilled water to prepare 5.0 L o

- 60.2 mL
- 6.3 mL
- 41.7 mL
- 0.41 mL

Save & Next حفظ والتالي

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Which of the following is the electron
fluorine?

(a) F·

(b) ·F·

(c) ·F·

(d) :F:

(d)
(b)
(c)
(a)

Save & Next حفظ وتالي

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_____ gives a non-electrolyte when dissolved in water.

- $C_{12}H_{22}O_{11}$
- HNO_3
- $CaCl_2$
- weak base

Save & Next حفظ والتالي

OL4028305

HP Compaq LE1711

How many moles of cesium (Cs) are contained in 595 kg of

- 1.26×10^3 moles Cs
- 2.23×10^2 moles Cs
- 7.91×10^4 moles Cs
- 4.48×10^3 moles Cs

Save & Next حفظ والتالي

OL4028305

HP Compaq LE1711

Total questions in exam: 25 | Answered:

Question No. 1

How many moles of cesium (Cs) are contained in 595 kg of cesium?

- 7.91×10^4 moles Cs
- 2.23×10^2 moles Cs
- 4.48×10^3 moles Cs
- 1.26×10^3 moles Cs

Save & Next

BD4005891

s in exam: 25 | Answered: 2

2

llowing statements is FALSE regarding an ionic bond?

ionic bond between two ions requires energy.
d is the attraction between a metal ion and a nonmetal ion.
lose electrons and nonmetal atoms gain electrons.
nd is formed between two nonmetal atoms.

Question No. 4

How many moles of H_2O could be produced when 3 moles of C_2H_6O completely r with oxygen gas according to the reaction?

$$C_2H_6O + 3\ O_2 \rightarrow 2\ CO_2 + 3\ H_2O$$

- 9 mol
- 6 mol
- 3 mol
- 12 mol

Save & Next जतन व पुढील

BD4005891

Question No. 3

Calcium carbide (CaC_2) reacts with water to produce acetylene (C_2H_2):

$$CaC_2\,(s) + 2\,H_2O\,(g) \rightarrow Ca(OH)_2\,(s) + C_2H_2\,(g)$$

The complete reaction of 57.4 g of CaC_2 requires consumption of _____ g of H_2O.

- 1.79
- 32.3
- 0.895
- 64.1

Save & Next

Question No. 5

In the reaction below, what is the theoretical yield in moles for LiOH when 6 grams of Li_2O react with 7 grams of H_2O?

$$Li_2O + H_2O \rightarrow 2LiOH$$

- 1.0 mol
- 0.8 mol
- 0.4 mol
- 0.6 mol

Save & Next جفظ و تالي

Question No. 19

Choose the best classification of the reaction represented by the following equation:

$$2Fe^{3+}(aq) + Fe(s) \rightarrow 3Fe^{2+}(aq)$$

- Oxidation-reduction
- combustion
- precipitation
- acid-base

Save & Next حفظ والتالي

Total questions in exam: 25 | Answered: 0

Question No. 7

Calculate the molar mass for $Mg(ClO_4)_2$.

- 123 g/mol
- 223 g/mol
- 119 g/mol
- 247 g/mol

Save & Next حفظ والتالي

Question No. 1

When the following equation is balanced, the coefficient of wate

$$N_2O_5\,(g) + H_2O\,(l) \rightarrow HNO_3\,(aq)$$

- 1
- 5
- 3
- 2

11 Photos

Question No. 1

What is the oxidation number of manganese in MnO_2?

- +4
- +1
- 0
- +2

OL4053865

_____ gives a non-electrolyte when dissolved in water.

- $C_{12}H_{22}O_{11}$
- HNO_3
- $CaCl_2$
- weak base

Save & Next حفظ والتالي

OL4028305

HP Compaq LE1711

Question No. 1

Which of the following statements about atoms is FALSE?

- Atoms compose all matter.
- Atoms are the basic building block of nature.
- An atom is the smallest identifiable unit of an element.
- Atoms is composed from two or more elements.

Question No. 20

In an oxidation-reduction reaction, the oxidized substance always _____

- shows loss of electrons.
- shows gain of electrons.
- shows gain of neutrons.
- gives up hydrogen atoms.

حفظ وتالي ext

Question No. 16

The coefficients (a,b,c) needed to balance the equation below are:

$$a\,Al + b\,O_2 \rightarrow c\,Al_2O_3$$

- (3,2,4)
- (4,2,2)
- (4,3,2)
- (4,2,3)

Save & Next

Total questions in exam: 25 | Answered:

Question No. 1

How many moles of cesium (Cs) are contained in 595 kg of cesium?

- 7.91×10^4 moles Cs
- 2.23×10^2 moles Cs
- 4.48×10^3 moles Cs
- 1.26×10^3 moles Cs

Save & Next

BD4005891

Question No. 4

How many moles of H_2O could be produced when 3 moles of C_2H_6O completely r with oxygen gas according to the reaction?

$$C_2H_6O + 3\ O_2 \rightarrow 2\ CO_2 + 3\ H_2O$$

- 9 mol
- 6 mol
- 3 mol
- 12 mol

Save & Next حفظ و التالي

BD4005891

Question No. 1

What is the final molarity of H_2SO_4 solution, if 240mL of 4M H_2SO_4 was diluted to a final volume of 0.5 L?

- 0.96 M
- 1.60 M
- 1.34 M
- 1.92 M

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 2

What is the coefficient of oxygen gas after balancing the following equation?

$$\_\_H_2O_2(l) \xrightarrow{\Delta} \_\_H_2O(l) + \_\_O_2(g)$$

- 4
- 2
- 3
- 1

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 2

What is the coefficient of oxygen gas after balancing the following equation?

$$\_\_H_2O_2(l) \xrightarrow{\Delta} \_\_H_2O(l) + \_\_O_2(g)$$

- 4
- 2
- 3
- 1

Save & Next حفظ والتالي

OL4028305

HP Compaq LE1711

Total questions in exam: **25** | Answered: **0**

Question No. 3

What is the percent yield for a reaction if its theoretical yield is 122 g and its actual yield is 32 g?

- 26%
- 60%
- 53%
- 40%

Save & Next حفظ و التالي

HP Compaq LE1711

Total questions in exam: 25 | Answered: 0

Question No. 4

How many grams of $AlCl_3$ could be produced when 3 mole of Al completely react with Cl_2 according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 400 g
- 133 g
- 266 g
- 67 g

Save & Next جمع و التالي

HP Compaq LE1711

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **0**

Question No. 5

What is the term for a bond composed of three electron pairs shared between two atoms?

- Ionic bond
- double bond
- triple bond
- single bond

Save & Next حفظ واقلي

HP Compaq LE1711

Total questions in exam: 25 | Answered: 0

Question No. 6

How many grams of nitrogen are needed to produce 325 grams of ammonia?

$$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$$

- 178 g
- 651 g
- 163 g
- 267 g

Save & Next حفظ والتالي

OL4028305

HP Compaq LE1711

Question No. 7

Calculate the molar mass of potassium chloride, KCl.

- 67.4 g/mol
- 54.5 g/mol
- 6.74 g/mol
- 74.6 g/mol

Save & Next حفظ والتالي

OL4028305

HP Compaq LE1711

Question No. 8

Lithium and nitrogen react in a combination reaction to produce lithium nitride:

$$6Li\,(s) + N_2\,(g) \rightarrow 2Li_3N\,(s)$$

How many moles of lithium are needed to produce 0.20 mol of Li_3N when the reaction is carried out in the presence of excess nitrogen?

- 0.13
- 0.6
- 0.067
- 0.1

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 9

Calculate the mass percent composition of sulfur in $Ga_2(SO_4)_3$.

- 32.6 %
- 9.4 %
- 44.9 %
- 22.5%

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 10

In the reaction below, what is the theoretical yield in moles for NO when 6 moles of NH_3 react with 7 moles of O_2?

$$4\ NH_3 + 5\ O_2 \rightarrow 4\ NO + 6\ H_2O$$

- 7.0 mol
- 5.6 mol
- 6.8 mol
- 4.6 mol

Save & Next حفظ والتالي

OL4028305

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Question No. 11

How many valence electrons are in a chlorine atom and a chloride ion?

- 17 and 18, respectively
- 7 and 8, respectively
- 8 and 7, respectively
- 1 and 8, respectively

Save & Next حفظ والتالي

OL4028305

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Question No. 12

Oxidation cannot occur without _____

- reduction
- air
- water
- acid

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Question No. 13

The name of the chemical compound $NaNO_3$ is:

- sodium(I) nitrite
- sodium(I) nitrate
- sodium nitrite
- sodium nitrate

Save & Next حفظ والتالي

OL4028305

Question No. 14

The coefficients (a,b,c) needed to balance the equation below are:

$$a\,NaNO_3 \rightarrow b\,NaNO_2 + c\,O_2$$

- (1,1,2)
- (1,2,1)
- (1,2,2)
- (2,2,1)

Save & Next حفظ والتالي

OL4028305

Question No. 15

The empirical formula of the molecular compound C_6H_{12} is:

- C_3H_8
- C_6H_{14}
- C_2H_7
- CH_2

Save & Next حفظ والتالي

OL4028305

Question No. 16

The elemental mass percent composition of succinic acid is: C 40.68% , H 5.12% , and O 54.19%.

- C_2H_4O
- $C_2H_3O_2$
- C_2H_2O
- CH_2O

Save & Next حفظ والتالي

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Question No. 17

How many grams of NaCl are there in 55.0 mL of a 1.90 M aqueous solution of NaCl?

- 12.2
- 6.11
- 3.21
- 0.105

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HP Compaq LE1711

Total questions in exam: 25 | Answered: 0

Question No. 18

The most correct name for the compound CS_2 is:

- carbon disulfide
- carbon trisulfide
- carbon sulfide
- monocarbon disulfide

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 19

In an oxidation-reduction reaction, the reduced substance always _____

- shows gain of electrons.
- shows loss of electrons.
- gives up hydrogen atoms.
- takes on oxygen atoms.

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Total questions in

Question No. 20

The systematic name for the chemical compound Na_2S is:

- sodium sulfide
- sodium monosulfide
- sodium(II) sulfide
- sodium(I) sulfide

Save & Next حفظ والتالي

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Question No. 21

Which of the following is an ionic compound?

- CCl_4
- CO_2
- PCl_3
- LiBr

Save & Next حفظ والتالي

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Question No. 22

What volume of 12 M HCl solution must be diluted with distilled water to prepare 5.0 L o

- 60.2 mL
- 6.3 mL
- 41.7 mL
- 0.41 mL

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HP Compaq LE1711

Question No. 2

What is the coefficient of oxygen gas after balancing the following equation?

$$\_\_H_2O_2(l) \xrightarrow{\Delta} \_\_H_2O(l) + \_\_O_2(g)$$

- 4
- 2
- 3
- 1

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 24

The molarity (M) of an aqueous solution containing 52.5 g of sucrose ($C_{12}H_{22}O_{11}$) in 35.5 mL of solution is _____.

- 1.48
- 0.104
- 5.46
- 4.32

Question No. 20

A chemical reaction produces 7.25 moles of barium sulfate, $BaSO_4$. What mass (in grams) of barium sulfate is produced?

- 253 g
- 1.69×10^3 g
- 185 g
- 31.1 g

Save & Next حفظ والتالي

Question No. 3

In the reaction below, what is the theoretical yield in grams for B_2H_6 when 6 moles of BF_3 react with 5 moles of NaH?

$$8\ BF_3 + 6\ NaH \rightarrow 6\ NaBF_4 + B_2H_6$$

- 47 g
- 21 g
- 39 g
- 32 g

Save & Next حفظ والتالي

P223

Question No. 9

Calculate the mass percent composition of sulfur in $Ga_2(SO_4)_3$.

- 32.6 %
- 9.4 %
- 44.9 %
- 22.5%

Save & Next حفظ والتالي

HP Compaq LE1711

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HP Compaq LE1711

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HP Compaq LE1711

Question No. 10

In the reaction below, what is the theoretical yield in moles for NO when 6 moles of NH_3 react with 7 moles of O_2?

$$4\,NH_3 + 5\,O_2 \rightarrow 4\,NO + 6\,H_2O$$

- 7.0 mol
- 5.6 mol
- 6.8 mol
- 4.6 mol

Save & Next

OL4028305

HP Compaq LE1711

Total questions in exam: **25** | Answered: **0**

Question No. 5

What is the term for a bond composed of three electron pairs shared between two atoms?

- Ionic bond
- double bond
- triple bond
- single bond

Save & Next जतन व पुढील

HP Compaq LE1711

Question No. 8

Lithium and nitrogen react in a combination reaction to produce lithium nitride:

$$6Li\,(s) + N_2\,(g) \rightarrow 2Li_3N\,(s)$$

How many moles of lithium are needed to produce 0.20 mol of Li_3N when the reaction is carried out in the presence of excess nitrogen?

- 0.13
- 0.6
- 0.067
- 0.1

Save & Next حفظ والتالي

HP Compaq LE1711

Question No. 7

Calculate the molar mass of potassium chloride, KCl.

- 67.4 g/mol
- 54.5 g/mol
- 6.74 g/mol
- 74.6 g/mol

Save & Next حفظ والتالي

OL4028305

HP Compaq LE1711

Total questions in exam: 25 | Answered: 0

Question No. 6

How many grams of nitrogen are needed to produce 325 grams of ammonia?

$$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$$

- 178 g
- 651 g
- 163 g
- 267 g

Save & Next حفظ والتالي

OL4028305

HP Compaq LE1711

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **1**

Question No. 5

The number of lone electron pairs in the N_2 molecule is _

- 1 pair
- 5 pairs
- 3 pairs
- 2 pairs

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **1**

Question No. 6

Which one of the following contains 92.3% carbon by mass?

- C_2H_2
- CO_2
- CH_3F
- CH_4

Question No. 19

Choose the best classification of the reaction represented by the following equation:

$$2Fe^{3+}(aq) + Fe(s) \rightarrow 3Fe^{2+}(aq)$$

- Oxidation-reduction
- combustion
- precipitation
- acid-base

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Save & Next حفظ والتالي

Question No. 10

In the reaction below, what is the theoretical yield in moles for NO when 6 moles of NH_3 react with 7 moles of O_2?

$$4\,NH_3 + 5\,O_2 \rightarrow 4\,NO + 6\,H_2O$$

- 7.0 mol
- 5.6 mol
- 6.8 mol
- 4.6 mol

Save & Next

OL4028305

HP Compaq LE1711

Total questions in exam: 25 | Answered: 6

Question No. 5

In the reaction below, what is the theoretical yield in moles for LiOH when 6 grams of Li_2O react with 7 grams of H_2O?

$$Li_2O + H_2O \rightarrow 2LiOH$$

- 1.0 mol
- 0.8 mol
- 0.4 mol
- 0.6 mol

Save & Next

MKCL OES

Total questions in exam: 25 | Answered: 5

Question No. 8

Which one of the following is oxidized in the following equation?

$$2Al(s) + 3H_2SO_4(aq) \rightarrow Al_2(SO_4)_3(aq) + 3H_2(g)$$

- $Al_2(SO_4)_3(aq)$
- $H_2(g)$
- $Al(s)$
- $H_2SO_4(aq)$

Question No. 3

Calcium carbide (CaC_2) reacts with water to produce acetylene (C_2H_2):

$$CaC_2\,(s) + 2\,H_2O\,(g) \rightarrow Ca(OH)_2\,(s) + C_2H_2\,(g)$$

The complete reaction of 57.4 g of CaC_2 requires consumption of _____ g of H_2O.

- 1.79
- 32.3
- 0.895
- 64.1

Save & Next

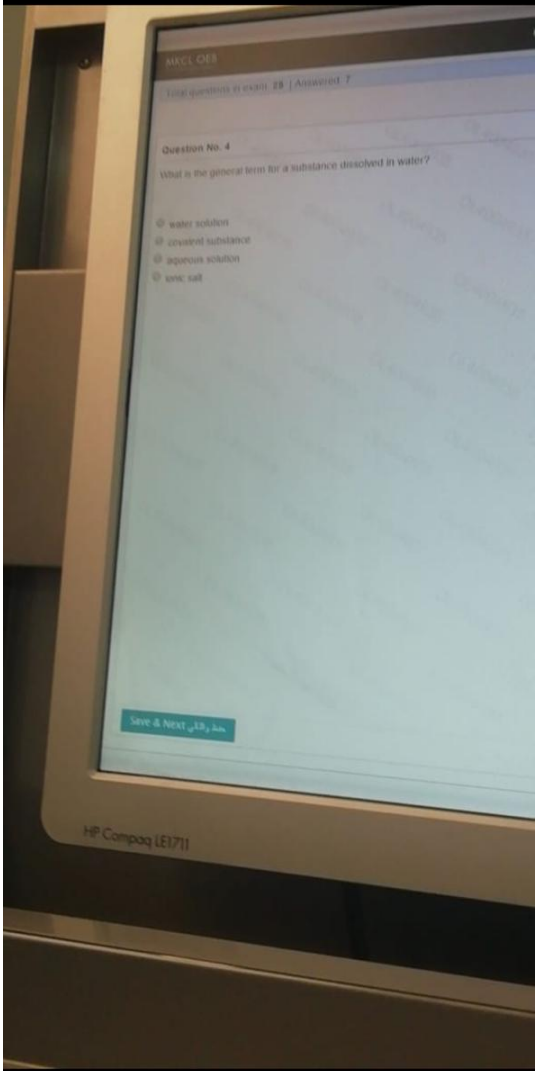
carbon

phosphorus

calcium

hydrogen

Save & Next حفظ والتالي


MKCL OES
Total questions in exam: 25 | Answered: 7
Question No. 4
What is the general term for a substance dissolved in water?
water solution
covalent substance
aqueous solution
ionic salt
Save & Next جतन व पुढचे
HP Compaq LE1711

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **12**

Question No. 18

How many moles of $AlCl_3$ could be produced when 284 grams of Cl_2 c
with aluminum according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 3.92 mol
- 3.33 mol
- 4.60 mol
- 2.67 mol

Question No. 16

The elemental mass percent composition of succinic acid is: C 40.68% , H 5.12% , and O 54.19%.

- C_2H_4O
- $C_2H_3O_2$
- C_2H_2O
- CH_2O

Save & Next حفظ والتالي

HP Compaq LE1711

What is the oxidation number of nitrogen in NO_3^{-1}?

- 0
- -5
- -3
- +5

Question No. 9

Calculate the mass percent composition of sulfur in $Ga_2(SO_4)_3$.

- 32.6 %
- 9.4 %
- 44.9 %
- 22.5%

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **12**

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$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 3.92 mol
- 3.33 mol
- 4.60 mol
- 2.67 mol

In a

- must be water
- substance present in a small amount
- the substance present in the greatest amount
- always a solid

MKCL OES

Total questions in exam: 25 | Answered: 21

Question No. 15

What is the final molarity of HNO_3 solution, if 95 mL of 2M HNO_3 was diluted to a final volume of 0.3 L?

- 0.43 M
- 0.53 M
- 0.42 M
- 0.63 M

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **12**

Question No. 14

The number of lone pairs of electrons in HCl is ____ pairs.

- 2
- 1
- 3
- 0

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EN English

Question No. 2

What is the coefficient of oxygen gas after balancing the following equation?

$$\_\,H_2O_2(l) \xrightarrow{\Delta} \_\,H_2O(l) + \_\,O_2(g)$$

- 4
- 2
- 3
- 1

Save & Next حفظ والتالي

HP Compaq LE1711

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The elemental mass percent composition of succinic acid is: C 40.68% , H 5.12% , and O 54.19%.

- C_2H_4O
- $C_2H_3O_2$
- C_2H_2O
- CH_2O

Save & Next احفظ والتالي

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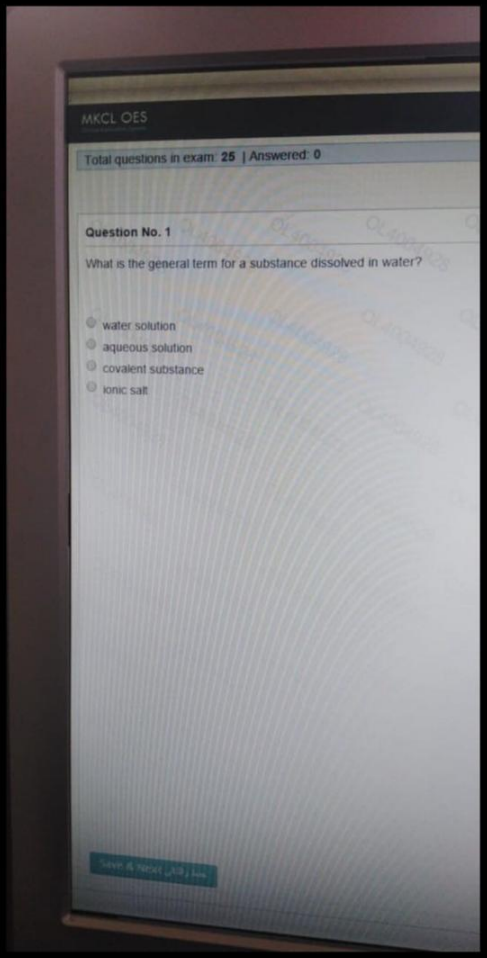

MKCL OES
Total questions in exam: 25 | Answered: 0
Question No. 1
What is the general term for a substance dissolved in water?
water solution
aqueous solution
covalent substance
ionic salt

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **1**

Question No. 6

Which one of the following contains 92.3% carbon by mass?

- C_2H_2
- CO_2
- CH_3F
- CH_4

ical reaction produces 7.25 moles of barium sulfate, $BaSO_4$. What mass (in grams) of barium su

g

9×10^3 g

1 g

3 g

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **12**

Question No. 18

How many moles of $AlCl_3$ could be produced when 284 grams of Cl_2 c
with aluminum according to the reaction?

$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

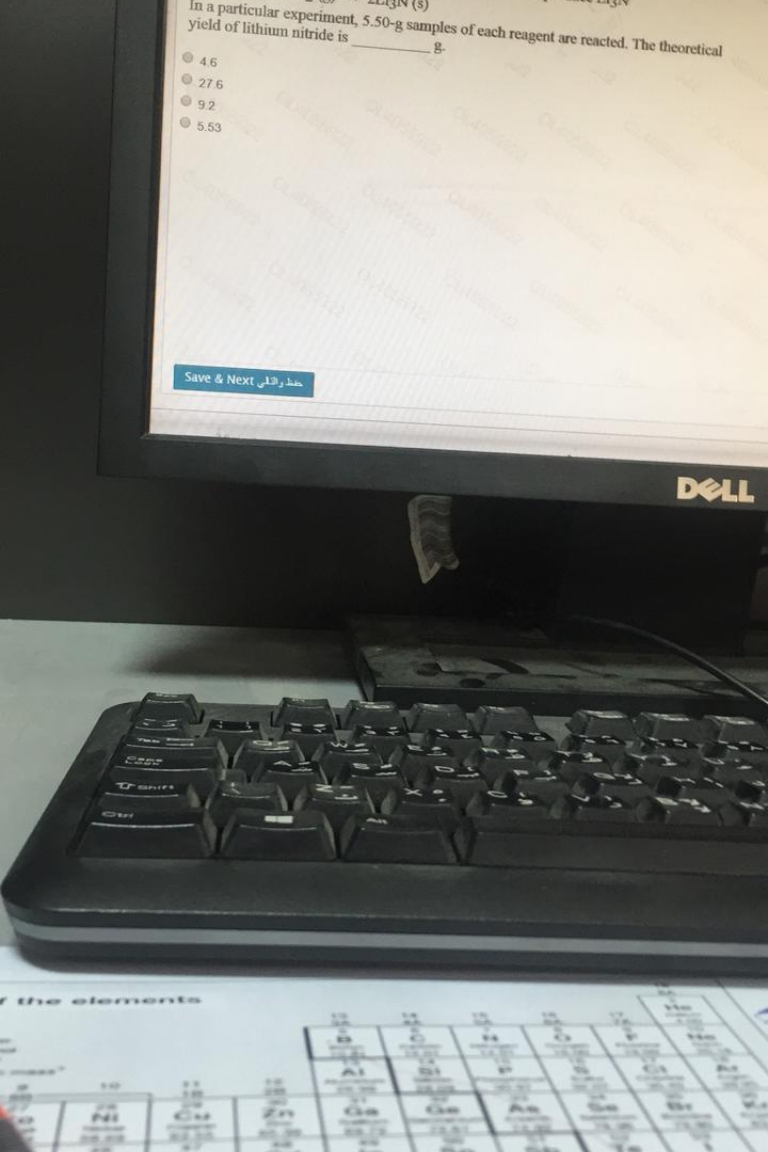
- 3.92 mol
- 3.33 mol
- 4.60 mol
- 2.67 mol

$2Li_3N(s)$

In a particular experiment, 5.50-g samples of each reagent are reacted. The theoretical yield of lithium nitride is _____ g.

- 4.6
- 27.6
- 9.2
- 5.53

Save & Next حفظ والتالي



When 22.0 g NaCl and 21.0 g H_2SO_4 are mixed and react according to the equation below, which is the limiting reagent?

$$2NaCl + H_2SO_4 \rightarrow Na_2SO_4 + 2HCl$$

- H_2SO_4
- Na_2SO_4
- NaCl
- HCl

MKCL OES
Online Evaluation System

Total questions in exam: **25** | Answered: **12**

Question No. 18

How many moles of $AlCl_3$ could be produced when 284 grams of Cl_2 c
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$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

- 3.92 mol
- 3.33 mol
- 4.60 mol
- 2.67 mol

he molar mass of $Mg(OH)_2$ is equ

30 g/mol

41 g/mol

58 g/mol

[illegible] g/mol

The empirical formula of the molecular compound C_4H_{10} is:

- C_2H_5
- CH
- C_4He
- C_2H_4

Question No. 11

For a given reaction

$$V_2O_5(s) + 5\,Ca(l) \rightarrow 2\,V(l) + 5\,CaO(s)$$

When 10 moles of V_2O_5 are mixed with 10 moles of Ca, which is the limiting to the above equation?

- V_2O_5
- V
- Ca
- CaO

Save & Next حفظ والتالي

Household sugar, sucrose, has the molecular formula $C_{12}H_{22}O_{11}$. What is the mass percent of carbon in sucro

- 0.333
- 0.421
- 0.267
- 0.528

In any balanced chemical equation, the number of atoms of each element on both sides of the equation is

- decreased by one
- always the same
- doubled
- increased by one

Choose the best classification of the reaction represented by the following equation

$$2Fe^{3+}(aq) + Fe(s) \rightarrow 3Fe^{2+}(aq)$$

- acid-base
- precipitation
- combustion
- Oxidation-reduction

What is the oxidation number of carbon in CO_3^{2-} ?

- +2
- +4
- -2
- +6

The most correct name for the compound SCl_8 is:

- sulfur octachloride
- sulfur chloride
- monosulfur octachloride
- sulfur nonachloride

The empirical formula of the molecular compound C_8H_{20} is:

- C_2H_5
- CH
- C_4H_{10}
- C_2H_4

The number of grams of NaCl (molar mass = 58.5 g/mol) that are required to make 250 mL of a 2 M solution is:

- 58.5 g
- 14.6 g
- 29.2 g
- 20.5 g

How many moles of CO_2 could be produced when 2 moles of C_2H_6O completely react with oxygen gas according to the reaction?

$$C_2H_6O + 3\ O_2 \rightarrow 2\ CO_2 + 3\ H_2O$$

- 4 mol
- 6 mol
- 2 mol
- 8 mol

The empirical formula of the molecular compound C_8H_{20} is:

- C_2H_5
- CH
- C_4H_{10}
- C_2H_4

The molar mass of $Mg(OH)_2$ is equal to

- 30 g/mol
- 41 g/mol
- 58 g/mol
- 50 g/mol

Which of the following polyatomic ions has a positive charge?

- sulfate ion
- carbonate ion
- hydroxide ion
- ammonium ion

uestion No. 7

diffusion .....

- Both solvents and solutes moves
- None of them
- Only solutes moves
- Only solvents moves

MKCL OES

Chemistry_

Total questions in exam: 25 | Answered: 0

Question No. 16

The elemental mass percent composition of succinic acid is: C 40.68% , H 5.12% ; and O 54.19%

- C_2H_4O
- $C_2H_3O_2$
- C_2H_2O
- CH_2O

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Question No. 3

Calcium carbide (CaC_2) reacts with water to produce acetylene (C_2H_2):

$$CaC_2\,(s) + 2\,H_2O\,(g) \rightarrow Ca(OH)_2\,(s) + C_2H_2\,(g)$$

The complete reaction of 57.4 g of CaC_2 requires consumption of _____ g of H_2O.

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- 0.895
- 64.1

Save & Next

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- doubled
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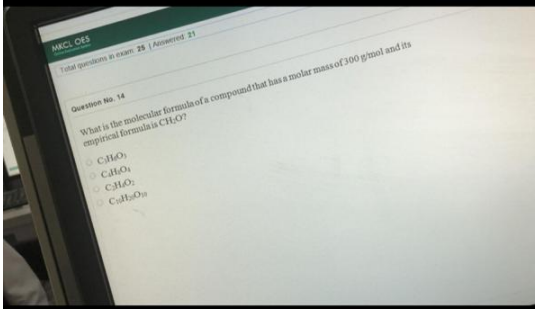
$$2Fe^{3+}(aq) + Fe(s) \rightarrow 3Fe^{2+}(aq)$$

- acid-base
- precipitation
- combustion
- Oxidation-reduction

What is the oxidation number of carbon in CO_3^{2-} ?

- +2
- +4
- -2
- +6


MKCL OES
Total questions in exam: 25 | Answered: 21
Question No. 14
What is the molecular formula of a compound that has a molar mass of 300 g/mol and its empirical formula is CH_2O?
$C_3H_6O_3$
$C_5H_{10}O_5$
$C_2H_4O_2$
$C_{10}H_{20}O_{10}$

Which of the following statements is FALSE regarding an ionic bond?

- An ionic bond is the attraction between a metal ion and a nonmetal ion.
- Breaking an ionic bond between two ions requires energy.
- Metal atoms lose electrons and nonmetal atoms gain electrons.
- The ionic bond is formed between two nonmetal atoms.

The number which is located on the LEFT of a chemical formula that helps to balance a chemical equation is called ____

- exponent
- subscript
- superscript
- coefficient

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- +4
- -2
- +6

Question No. 20

In an oxidation-reduction reaction, the oxidized substance always _____

- shows loss of electrons.
- shows gain of electrons.
- shows gain of neutrons.
- gives up hydrogen atoms.

A chemical reaction produces 7.25 moles of barium sulfate, $BaSO_4$. What mass (in grams) of barium sulfate is p

- 185 g
- 1.69×10^3 g
- 31.1 g
- 233 g

Total questions in exam: **25** | Answered: **14**

Question No. 12

In which of these substances are the atoms held together by polar covalent bonding?

- CsCl
- $SrCl_2$
- ClF
- S_8

Question No. 2

What is the coefficient of oxygen gas after balancing the following equation?

$$\_\_H_2O_2(l) \xrightarrow{\Delta} \_\_H_2O(l) + \_\_O_2(g)$$

- 4
- 2
- 3
- 1

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- $C_2H_3O_2$
- C_2H_2O
- CH_2O

Save & Next حفظ والتالي

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Question No. 8

Lithium and nitrogen react in a combination reaction to produce lithium nitride:

$$6Li\,(s) + N_2\,(g) \rightarrow 2Li_3N\,(s)$$

How many moles of lithium are needed to produce 0.20 mol of Li_3N when the reaction is carried out in the presence of excess nitrogen?

- 0.13
- 0.6
- 0.067
- 0.1

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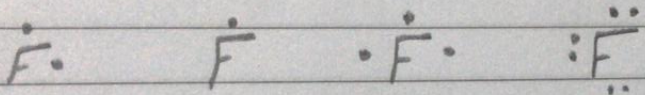
HP Compaq LE1711

What is the Lewis structure of an atom of fluorine?



Question No. 7

What is the molecular formula of a compound that has a molar mass of 180 g/mol and its empirical formula is CH_2O?

- $C_6H_{12}O_6$
- $C_5H_{10}O_5$
- $C_2H_2O_2$
- $C_4H_4O_4$

Save & Next

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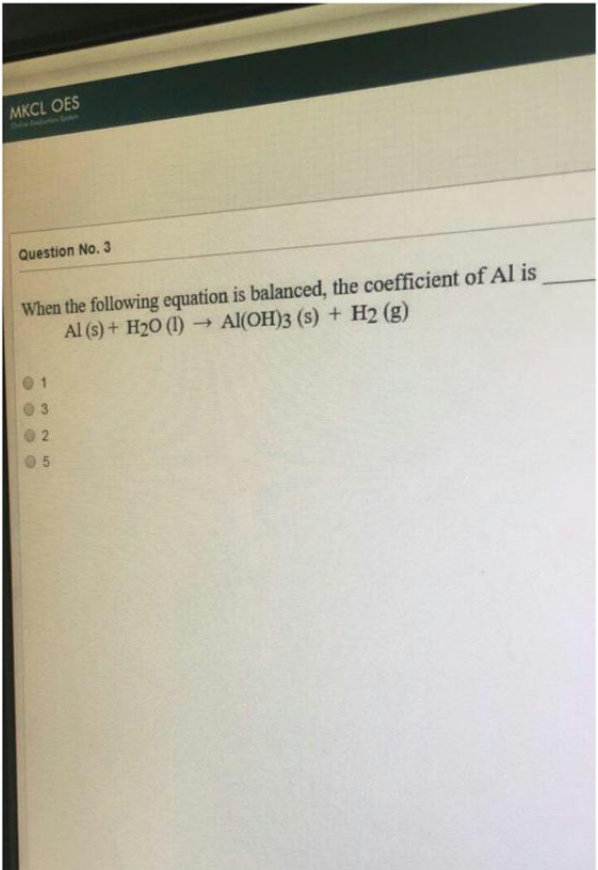
Select the element whose Lewis dot symbol is correct.

- He•
- •Fr•
- :Te:
- •Ra•

Save & Next


MKCL OES
Question No. 3
When the following equation is balanced, the coefficient of Al is ____
Al (s) + H2O (l) → Al(OH)3 (s) + H2 (g)
1
3
2
5

24 Photos

Question No. 1

What is the oxidation number of manganese in MnO_2?

- +4
- +1
- 0
- +2

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Total questions in exam: **25** | Answered: **0**

Question No. 6

Which of the following statements is FALSE regarding an ionic bond?

- An ionic bond is the attraction between a metal ion and a nonmetal ion.
- Metal atoms lose electrons and nonmetal atoms gain electrons.
- Breaking an ionic bond between two ions requires energy.
- The ionic bond is formed between two nonmetal atoms.

Save & Next

Total questions in exam: 25 | Answered: 6

Question No. 5

In the reaction below, what is the theoretical yield in moles for LiOH when 6 grams of Li_2O react with 7 grams of H_2O?

$$Li_2O + H_2O \rightarrow 2LiOH$$

- 1.0 mol
- 0.8 mol
- 0.4 mol
- 0.6 mol

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Question No. 2

What is the coefficient of oxygen gas after balancing the following equation?

$$\_\,H_2O_2(l) \xrightarrow{\Delta} \_\,H_2O(l) + \_\,O_2(g)$$

- 4
- 2
- 3
- 1

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Which of the following is the electron
fluorine?

(a) $F\cdot$ (b) $\cdot\dot{F}\cdot$ (c) $\cdot\ddot{F}\cdot$ (d) $:\ddot{F}:$

(d)
(b)
(c)
(a)

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Question No. 1

When the following equation is balanced, the coefficient of wate

$$N_2O_5\,(g) + H_2O\,(l) \rightarrow HNO_3\,(aq)$$

- 1
- 5
- 3
- 2

Question No. 22

What is the term for a bond in which a pair of electrons is shared equally?

- electrovalent bond
- polar covalent bond
- ionic bond
- nonpolar covalent bond

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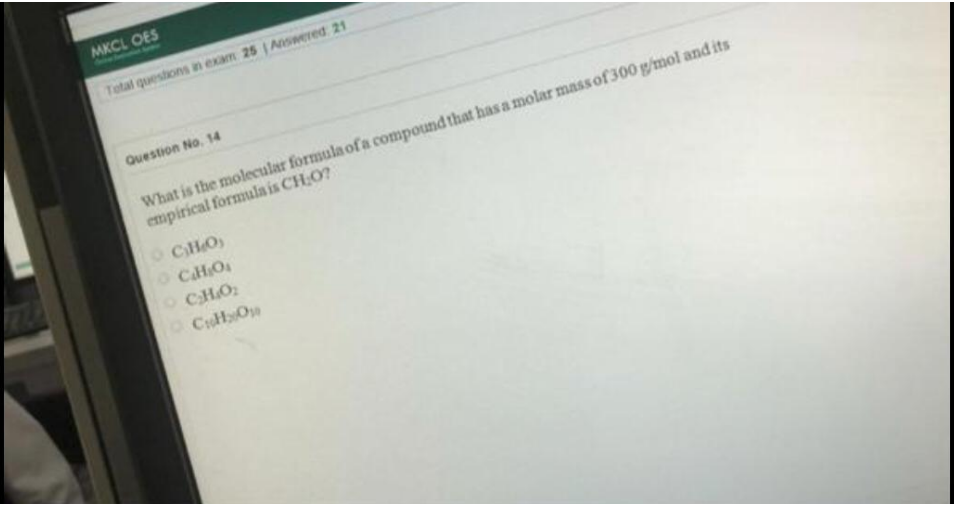
Which of the following statements is FALSE regarding an ionic bond?

- An ionic bond is the attraction between a metal ion and a nonmetal ion.
- Metal atoms lose electrons and nonmetal atoms gain electrons.
- Breaking an ionic bond between two ions requires energy.
- The ionic bond is formed between two nonmetal atoms.


MKCL OES
Total questions in exam: 25 | Answered: 21
Question No. 14
What is the molecular formula of a compound that has a molar mass of 300 g/mol and its empirical formula is CH_2O?
$C_3H_6O_3$
$C_4H_8O_4$
$C_2H_4O_2$
$C_{10}H_{20}O_{10}$

Question No. 8

How many lone pairs of electrons are on the P atom in PF_3?

- 1 pair
- 0 pairs
- 3 pairs
- 2 pairs